

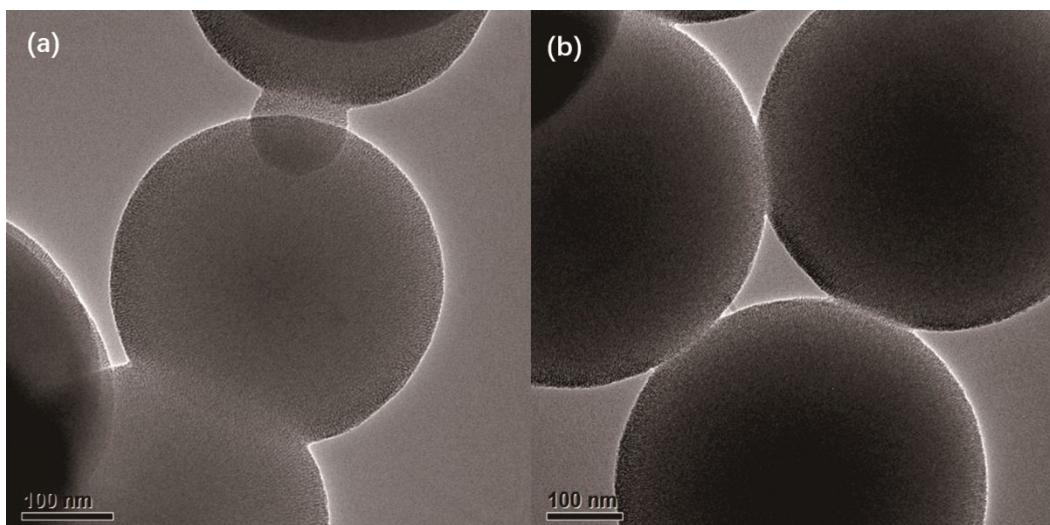
## Supporting Information

### Heterogeneous hybrid of propyl amino functionalized MCM-41 and 1H-1,2,4-triazole for high efficient intermediate temperature proton conductor

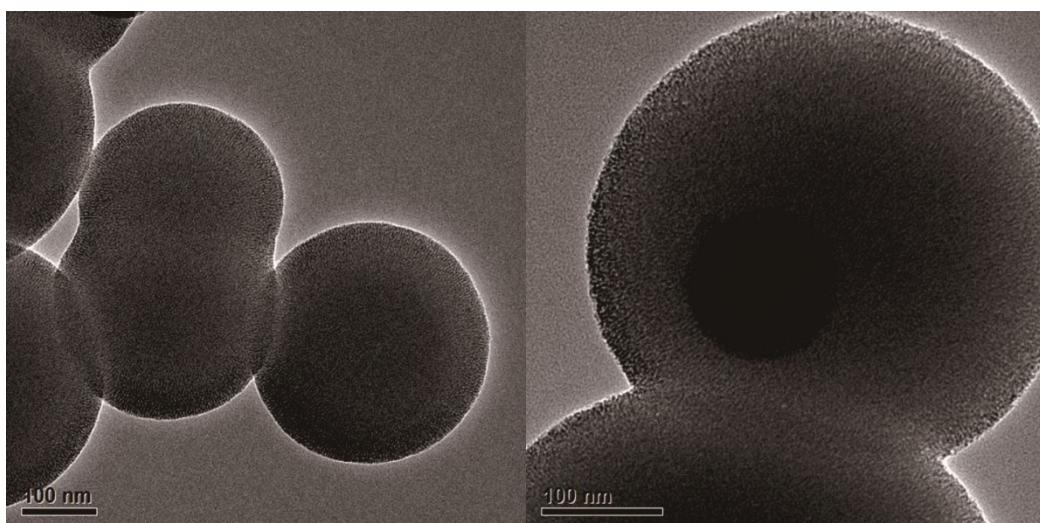
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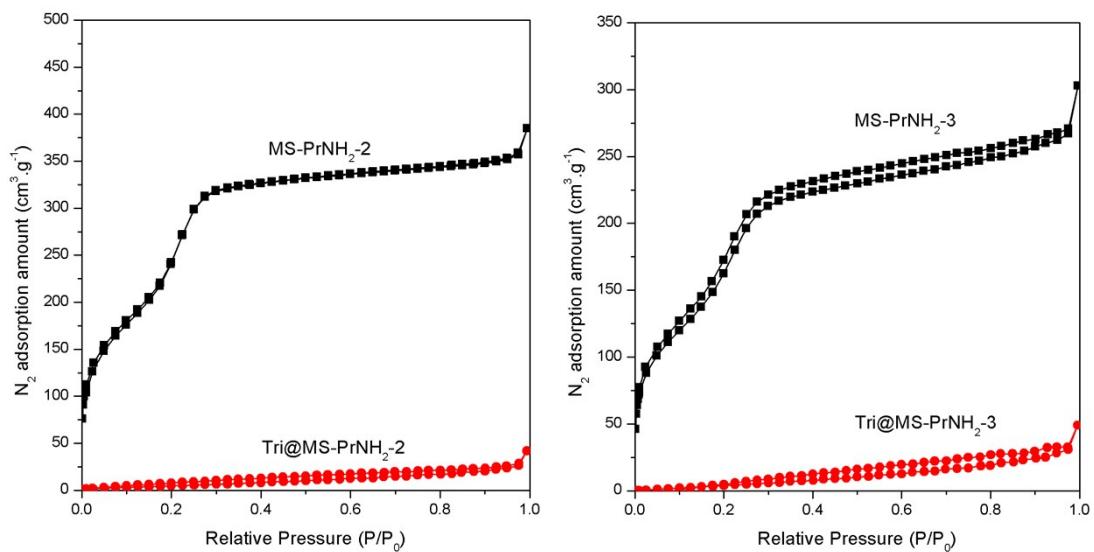
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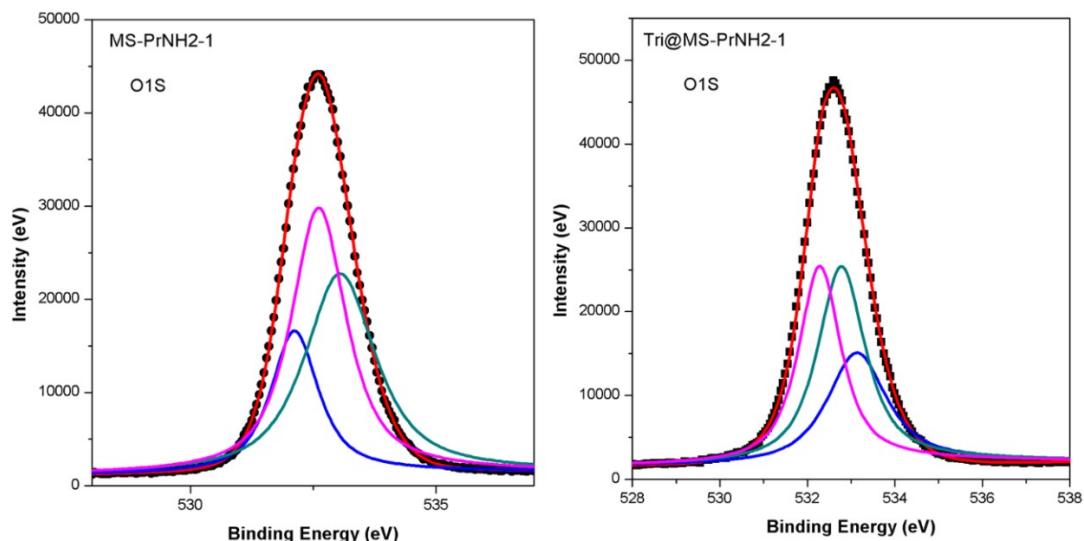
**Figure S1.** TEM images of MS-PrNH<sub>2</sub>-2 (a) and Tri@MS-PrNH<sub>2</sub>-2



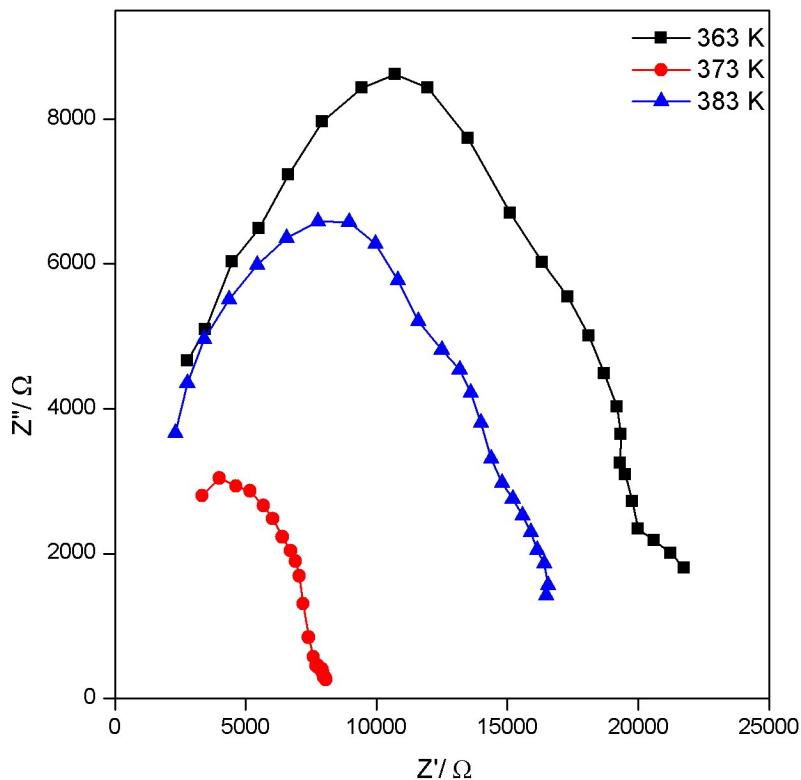
**Figure S2.** TEM images of MS-PrNH<sub>2</sub>-3 and Tri@MS-PrNH<sub>2</sub>-3



**Figure S3.** N<sub>2</sub> adsorption isotherms of MS-PrNH<sub>2</sub>-2, Tri@MS-PrNH<sub>2</sub>-2, MS-PrNH<sub>2</sub>-3 and Tri@MS-PrNH<sub>2</sub>-3



**Figure S4.** O 1s X-ray photoelectron spectroscopy of MS-PrNH<sub>2</sub>-1 and Tri@MS-PrNH<sub>2</sub>-1.



**Figure S5.** The Nyquist plots of MCM-41 loaded with 1H-1,2,4-triazole

**Table S1.** Comparison of proton conductivity of Tri@MS-PrNH<sub>2</sub>-1 with the reported porous materials.

Materials	Conductivity (S/cm)	Ea (eV)	Measurement Conditions	Ref
His@[Al(OH)(ndc)] <sub>n</sub>	1.7×10 <sup>-3</sup>	0.25	Anhydrous, 150°C	<i>Angew. Chem. Int. Ed.</i> 2011, 50, 11706
β-PCMOF-2(Tz) <sub>0.45</sub>	5.0×10 <sup>-4</sup>	0.34	Anhydrous, 150°C	<i>Nat. Chem.</i> 2009, 1, 75
In-IA-2D-2	1.2×10 <sup>-5</sup>	0.48	Anhydrous, 90°C	<i>Chem. Commun.</i> 2013, 49, 6197
Im@[Al(μ2-OH)(1,4-bdc)] <sub>n</sub>	2.2×10 <sup>-5</sup>	0.90	Anhydrous, 120°C	<i>Nat. Mater.</i> 2009, 8, 831
[ImH <sub>2</sub> ][Cu(H <sub>2</sub> PO <sub>4</sub> ) <sub>1.5</sub> (HPO <sub>4</sub> ) <sub>0.5</sub> •Cl <sub>0.5</sub> ]	2.0×10 <sup>-2</sup>	1.10	Anhydrous, 130°C	<i>Chem. Commun.</i> . 2014, 50, 10241
PA@Tp-Azo	6.7×10 <sup>-5</sup>	0.11	Anhydrous, 67°C	<i>J. Am. Chem. Soc.</i> 2014, 136, 6570
[Zn <sub>3</sub> (H <sub>2</sub> PO <sub>4</sub> ) <sub>6</sub> ](Hbim)	1.3×10 <sup>-3</sup>	0.50	Anhydrous, 120°C	<i>J. Am. Chem. Soc.</i> 2013, 135, 11345
Im@Td-PPI 2	3.49×10 <sup>-4</sup>	0.30	Anhydrous,	<i>J. Am. Chem. Soc.</i>

			90°C	2015,137,913
[Zn(HPO <sub>4</sub> )(H <sub>2</sub> PO <sub>4</sub> ) <sub>2</sub> ](I mH <sub>2</sub> ) <sub>2</sub>	$2.5 \times 10^{-4}$	0.47	Anhydrous, 130°C	<i>J. Am. Chem. Soc.</i> 2012,134,7612
<i>Tri@MS-NH<sub>2</sub>-1</i>	$8.34 \times 10^{-3}$	$0.55(< 80^{\circ}\text{C})$ $1.303(80\text{--}120^{\circ}\text{C})$	<i>Anhydrous,</i> $120^{\circ}\text{C}$	<i>This work</i>
<i>Tri@MS-NH<sub>2</sub>-2</i>	$2.68 \times 10^{-3}$	$0.48(< 80^{\circ}\text{C})$ $1.159(80\text{--}120^{\circ}\text{C})$	<i>Anhydrous,</i> $120^{\circ}\text{C}$	<i>This work</i>
<i>Tri@MS-NH<sub>2</sub>-3</i>	$7.29 \times 10^{-4}$	$0.40(< 80^{\circ}\text{C})$ $0.937(80\text{--}120^{\circ}\text{C})$	<i>Anhydrous,</i> $120^{\circ}\text{C}$	<i>This work</i>