

**Preparation of High Performance Supercapacitor Material by Fast Pyrolysis of
Corn Gluten Meal Waste**

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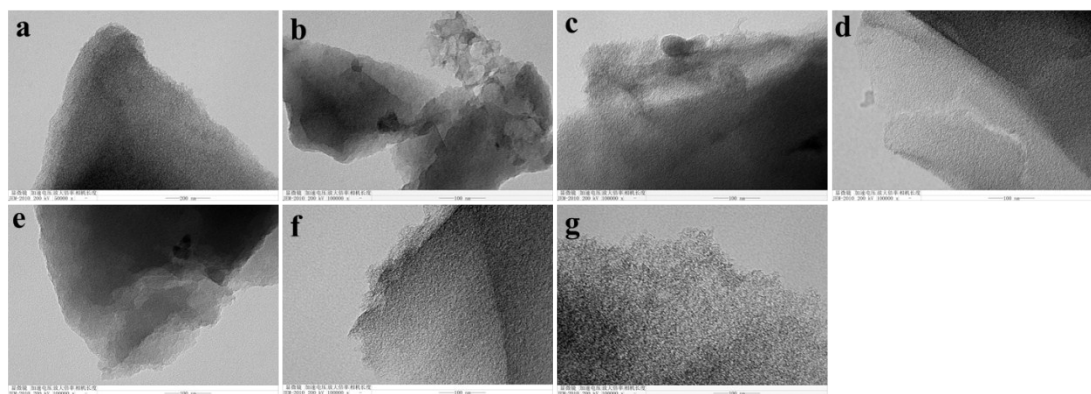


Fig. S1 TEM images of PCMs synthesized under different conditions. a-d) TEM images of PCM_{H600-900}; e-g) TEM images of PCM_{P300-500}.

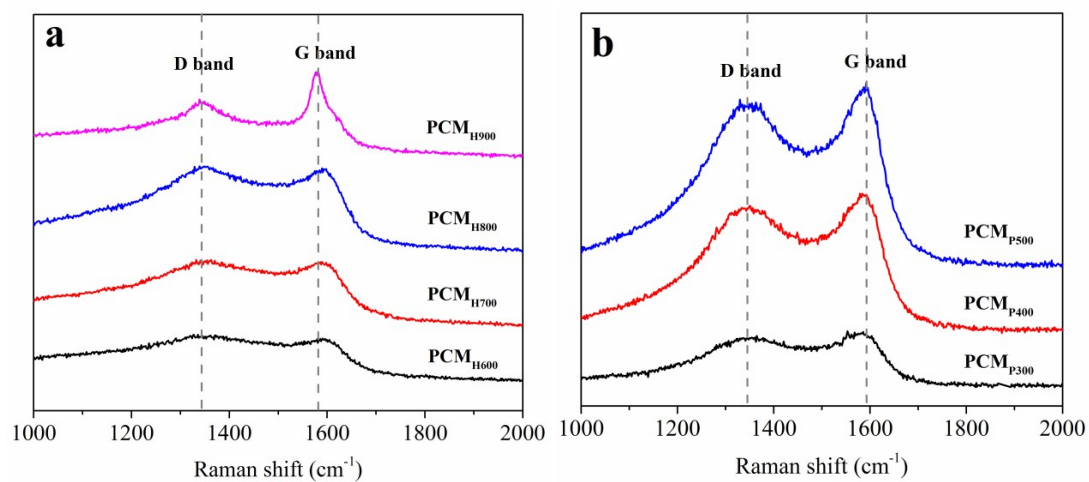


Fig. S2 Raman spectrum of a) PCM_{H600-900}, and b) PCM_{P300-500}.

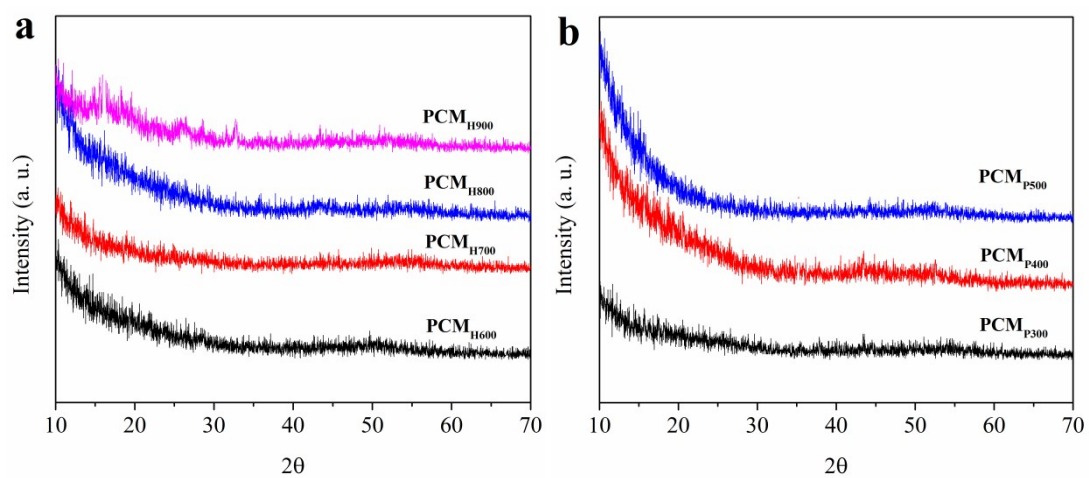


Fig. S3 XRD pattern of a) PCM_{H600-900}, and b) PCM_{P300-500}.

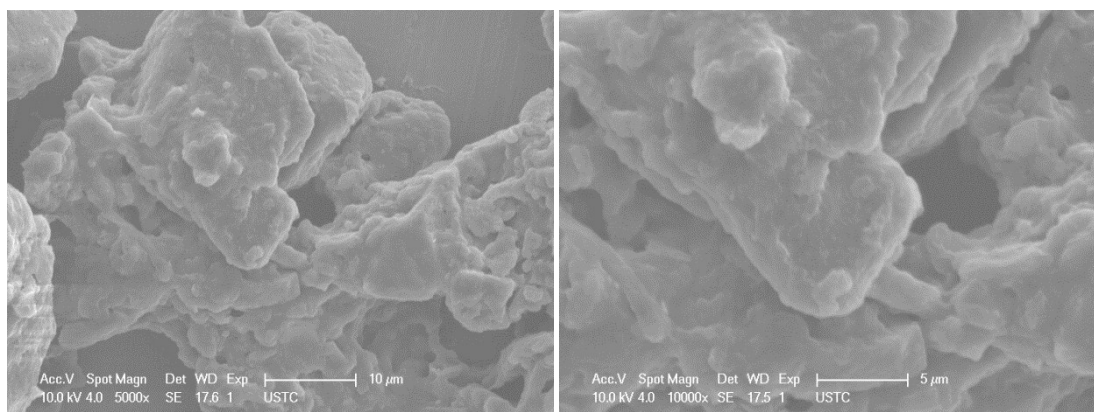


Fig. S4 SEM image of the CGM.

The elemental analysis shows the C, H, N and O contents in PCMs (Table 1). The contents of heteroatoms (N, O) in the bulk materials are 8.5~29.6% (N), 0.7~1.5% (O), and gradually decrease with increasing of pyrolysis temperature. The X-ray photoelectron spectrometer (XPS) was used to identify the surface elemental composition of the porous carbons obtained under different process. The high-resolution XPS spectra of N1s and O1s were collected to understand the formed N–C and O–C bonding. Nitrogen atoms were found in four different contributions in the carbon matrix: ¹pyridinic-N (N1, 398.0 eV), pyrrolic-N (N2, 399.6 eV), quaternary-N (N3, 400.8 eV) and oxidized-N (N4, 402.5 eV) (Figure S1b). Oxygen atoms were found in three different contributions in the carbon matrix: ¹carbonyl oxygen of Keto and quinone (O1, 530.8 eV), noncarbonyl (ether-type) oxygen atoms in esters and anhydrides (O2, 531.8 eV), and oxygen atoms in carboxylic groups (O3, 533.0 eV) (Figure S1c). These nitrogen and oxygen functional groups on the surface of PCMs coupled with high specific surface area offer a strong tendency to deliver exciting electrochemical performance. Therefore, these PCMs derived from CGM could be promising electrode materials for high power density supercapacitors.

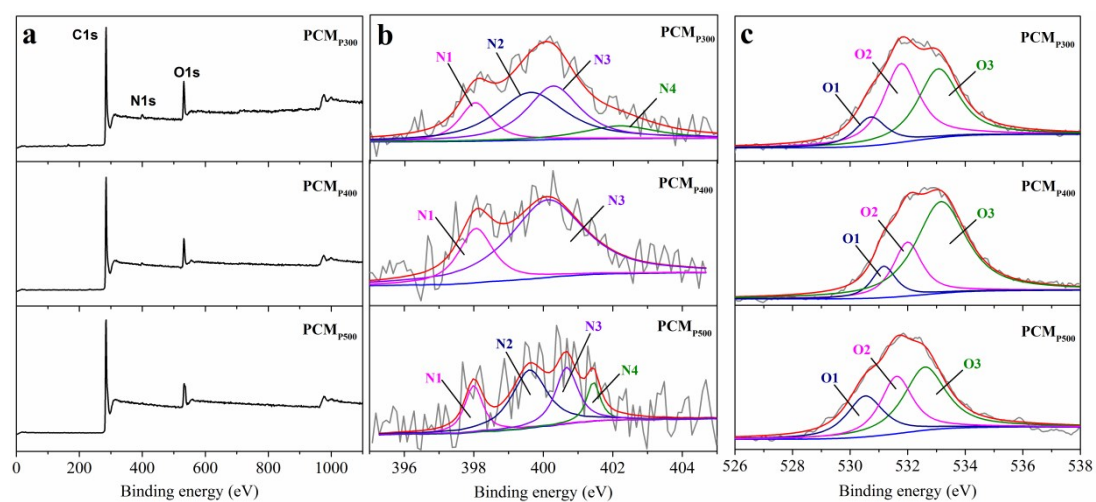


Figure S5. a) XPS survey spectra of the PCMs; b) XPS N1s spectra of the PCMs; c)

XPS O1s spectra of the PCMs.

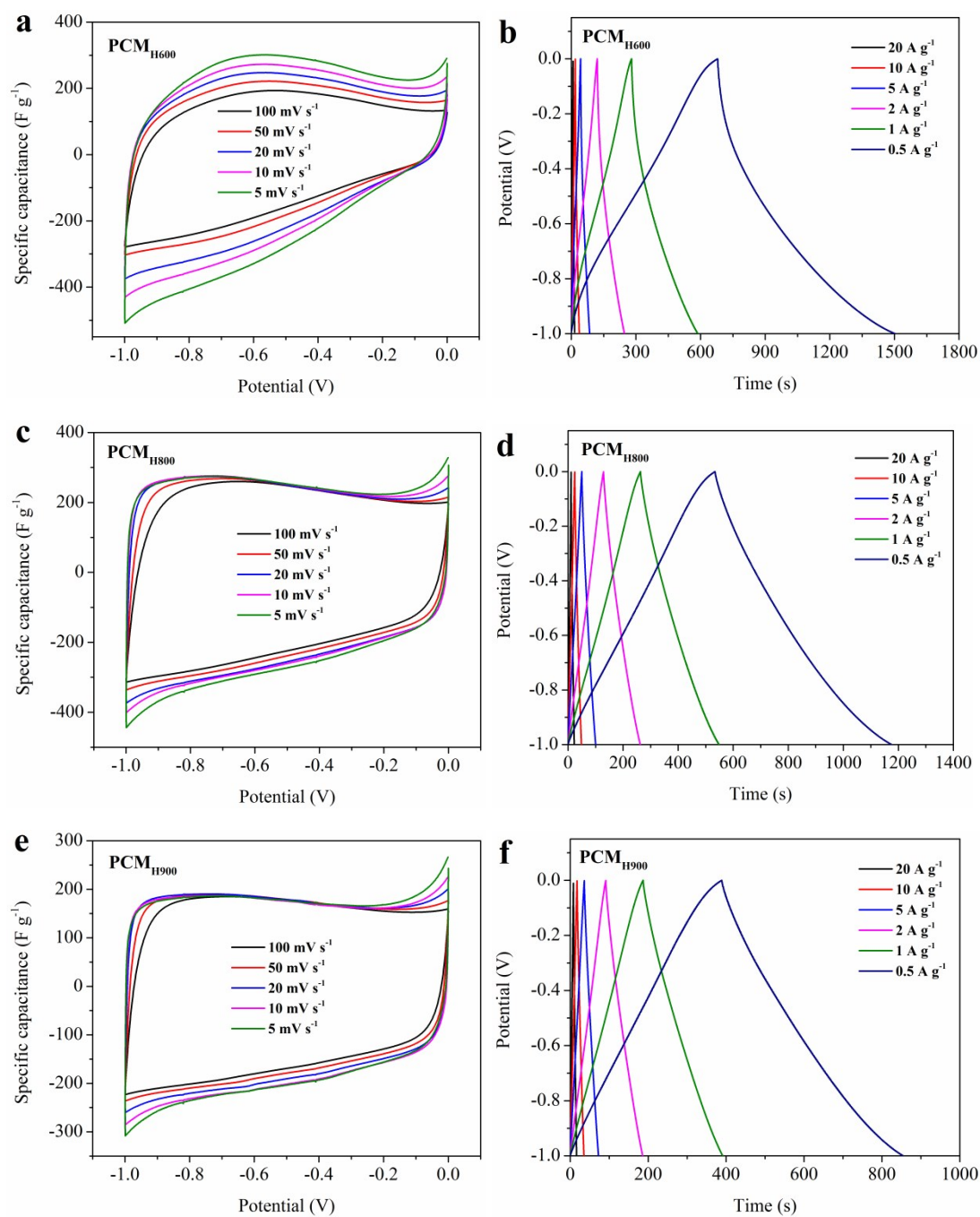


Figure S6. CV curves at different scan rates of a) PCM_{H600}, c) PCM_{H800}, e) PCM_{H900}; GCD profiles under different current densities of b) PCM_{H600}, d) PCM_{H800}, f) PCM_{H900}.

Table S1.Comparison of specific capacitance between the PCM_{P500} and other EDLCs.

Supercapacitor materials	Specific surface area (m ² g ⁻¹)	Pore volume (cm ³ g ⁻¹)	Max. capacitance (F g ⁻¹)	Scan rate or current density	Electrolyte	Cycle number	Stability	Ref.
Nitrogen-Containing Hydrothermal Carbons	598	0.34	220	0.1 A g ⁻¹	6 M KOH	-	-	²
HTC of Natural Organic Chemicals	2967	1.35	236	1 mV s ⁻¹	TEABF ₄ / AN	-	-	³
Activation of Graphene a-MEGO	~3100	2.14	165	1.4 A g ⁻¹	TEABF ₄ / AN	10 000	97%	⁴
Nitrogen-Doped Porous Carbon Nanofibers	562	0.51	202	1 A g ⁻¹	6 M KOH	3 000	97%	⁵
Nitrogen-Doped Carbon Monolith	679	0.46	246	1 mV s ⁻¹	6 M KOH	-	-	⁶
Hierarchical Porous Graphene-Like Networks	1810	1.22	305	0.5 A g ⁻¹	6 M KOH	15 000	~100%	⁷
Functionalized 3D Hierarchical Porous Carbon	2870	2.19	318	0.5 A g ⁻¹	6 M KOH	10 000	95.8%	⁸
3D Micro-porous Conducting Carbon Beehive	1327	-	254	0.5 A g ⁻¹	1 M H ₂ SO ₄	5 000	90%	⁹
Highly Porous Interconnected Carbon Nanosheets	~2200	1.30	150	1 mV s ⁻¹	1 M TEABF ₄ / AN	10 000	90-94%	¹⁰
Oxygen- and Nitrogen-Enriched 3D Porous Carbon	1003	0.62	440	0.5 A g ⁻¹	6 M KOH	10 000	92.3%	¹
Microporous Doped Carbon	1680	0.86	340	2 mV s ⁻¹	1 M H ₂ SO ₄	-	-	¹¹
Carbon Materials by Direct Pyrolysis of Seaweeds	1300	-	264	2 mV s ⁻¹	1 M H ₂ SO ₄	-	-	¹²
Functional MicroporousCarbon from Dead Leaves	1230		400	0.5 A g ⁻¹	1 M H ₂ SO ₄	-	-	¹³
Carbon Materials from High Ash Biochar	3310	1.85	260	0.6 A g ⁻¹	6 M KOH	2 000	99%	¹⁴
Porous Graphitic Carbon NanosheetfromCornstalk Biomass	540	0.48	213	1 A g ⁻¹	6 M KOH	6 000	98%	¹⁵
Porous 3D Carbon from Rice Bran	2475	1.21	265	10 A g ⁻¹	6 M KOH	10 000	-	¹⁶
Human Hair-Derived Carbon Flakes	1306	0.90	340	1 A g ⁻¹	6 M KOH	20 000	98%	¹⁷
Porous Carbon from BiowasteCorncob Residue	1210	0.67	314	5 mV s ⁻¹	6 M KOH	-	-	¹⁸
PCM _{P500}	3485	2.03	465	0.5 A g ⁻¹	6 M KOH	10000	93%	This work

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