

## *Supplementary Information*

### **Symmetric synergy of hybrid CoS<sub>2</sub>-WS<sub>2</sub> electrocatalysts for hydrogen evolution reaction**

Xiaofeng Zhou,<sup>ab</sup> Xiulin Yang,<sup>b</sup> Henan Li,<sup>b</sup> Mohamed Nejib Hedhili,<sup>b</sup> Kuo-Wei Huang,<sup>b</sup> Lain-Jong Li<sup>b\*</sup> and Wenjing Zhang<sup>a\*</sup>

<sup>a</sup> SZU-NUS Collaborative Innovation Center for Optoelectronic Science & Technology, Key Laboratory of Optoelectronic Devices and Systems of Ministry of Education and Guangdong Province, College of Optoelectronic Engineering, Shenzhen University, Shenzhen 518060, China

<sup>b</sup> KAUST Catalysis Center, Physical Sciences and Engineering, King Abdullah University of Science and Technology, Thuwal, 23955-6900, Kingdom of Saudi Arabia

\*E-mail: lance.li@kaust.edu.sa & wjzhang@szu.edu.cn

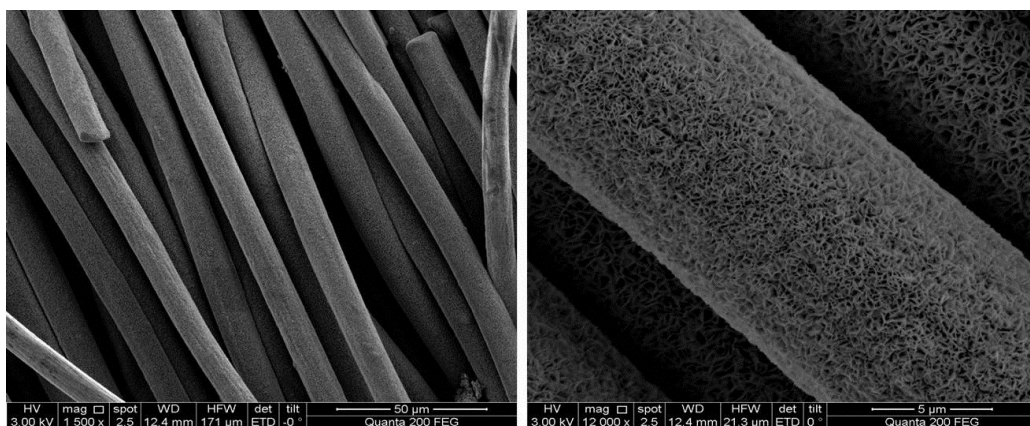


Figure S1. SEM images of Co(OH)<sub>2</sub>/CC with different magnifications.

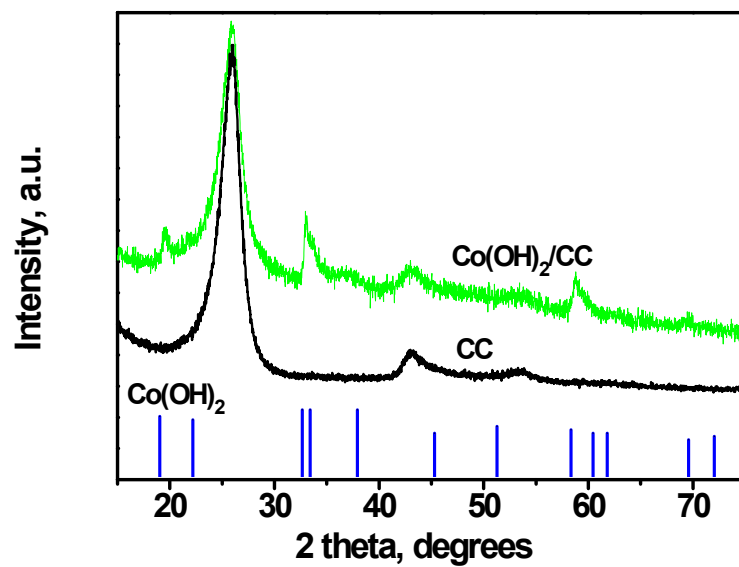


Figure S2. XRD patterns of the electrochemical deposited  $\text{Co(OH)}_2$  on carbon cloth (CC) and bare CC.

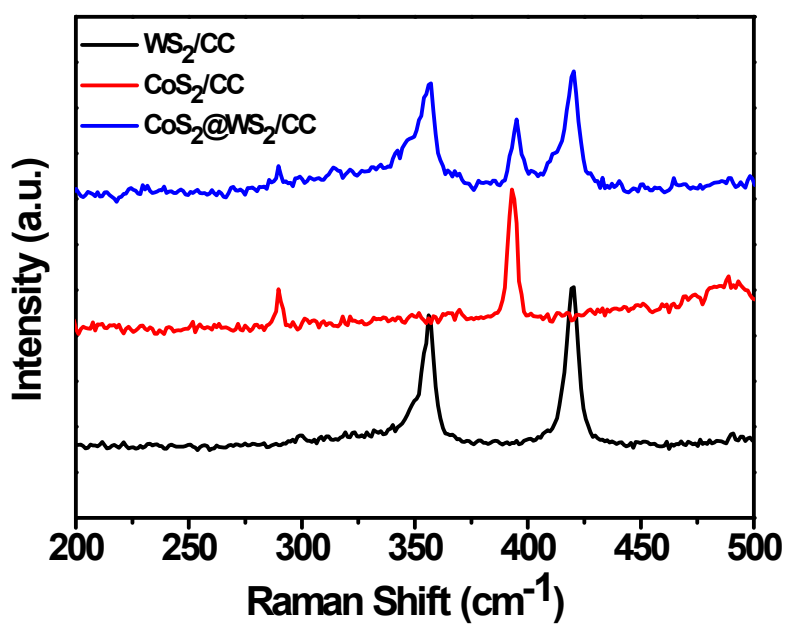


Figure S3. Raman spectrum of the samples WS<sub>2</sub>/CC, CoS<sub>2</sub>/CC, and CoS<sub>2</sub>@WS<sub>2</sub>/CC.

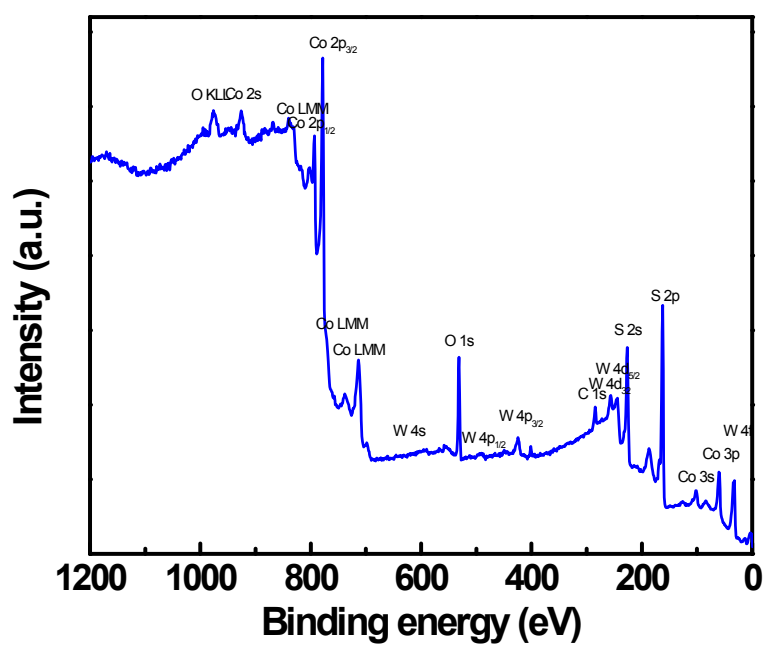


Figure S4. The XPS survey spectra of CoS<sub>2</sub>@WS<sub>2</sub>/CC sample.

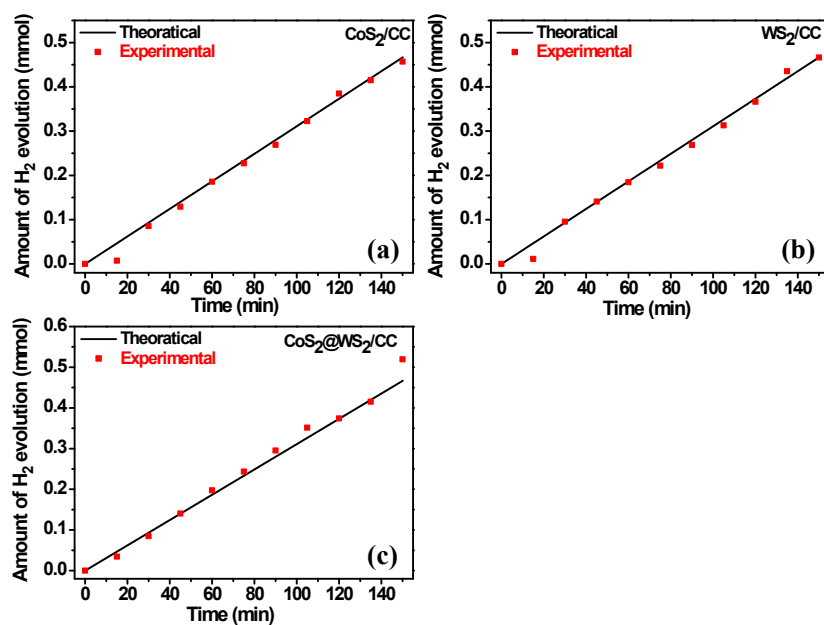


Figure S5. Faradic efficiency of HER in 0.5 M H<sub>2</sub>SO<sub>4</sub> solution with CoS<sub>2</sub>/CC, WS<sub>2</sub>/CC, and CoS<sub>2</sub>@WS<sub>2</sub>/CC as the cathode, respectively.

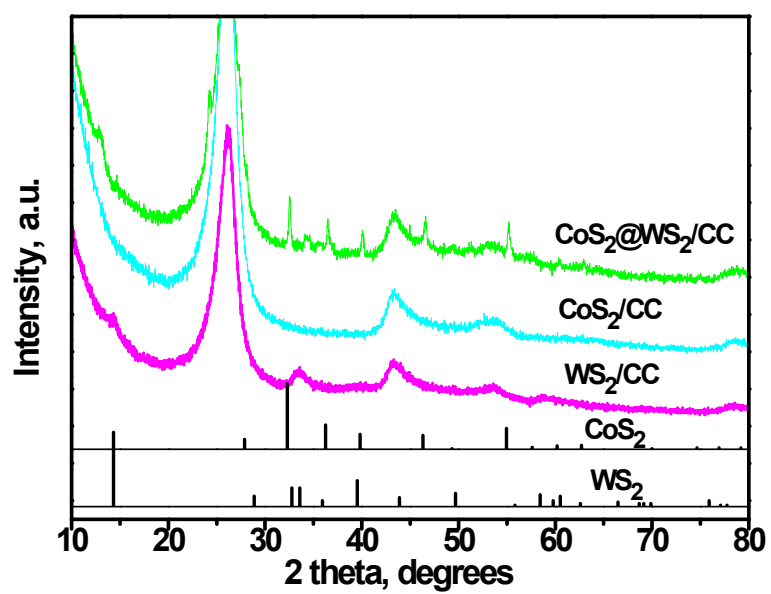


Figure S6. XRD patterns of samples  $\text{CoS}_2/\text{CC}$ ,  $\text{WS}_2/\text{CC}$ , and  $\text{CoS}_2@\text{WS}_2/\text{CC}$  after the stability evaluation.

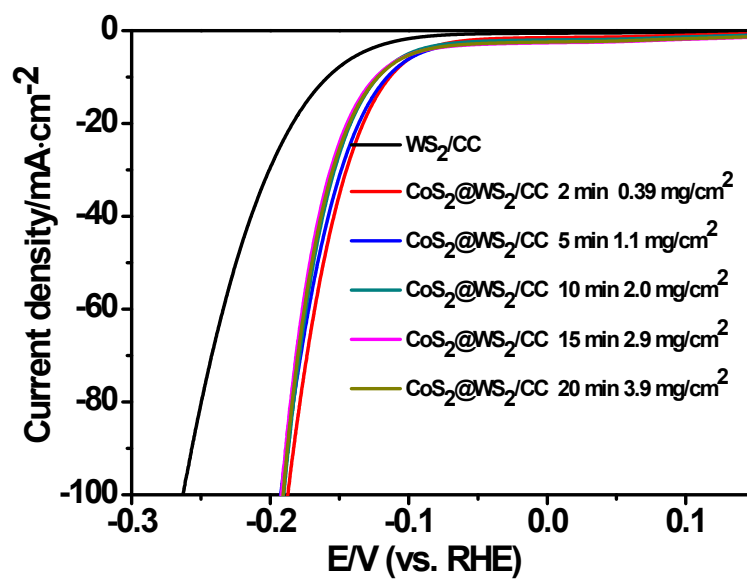


Figure S7. Polarization curves of WS<sub>2</sub> (loading of 4.0 mg/cm<sup>2</sup>) and a series of CoS<sub>2</sub>@WS<sub>2</sub>/CC samples with fixed amount of WS<sub>2</sub> (loading of 4.0 mg/cm<sup>2</sup>) and different electrodeposition amounts of CoS<sub>2</sub>, where the current is normalized by the geometrical area of carbon cloth and the potential is after internal resistance correction.



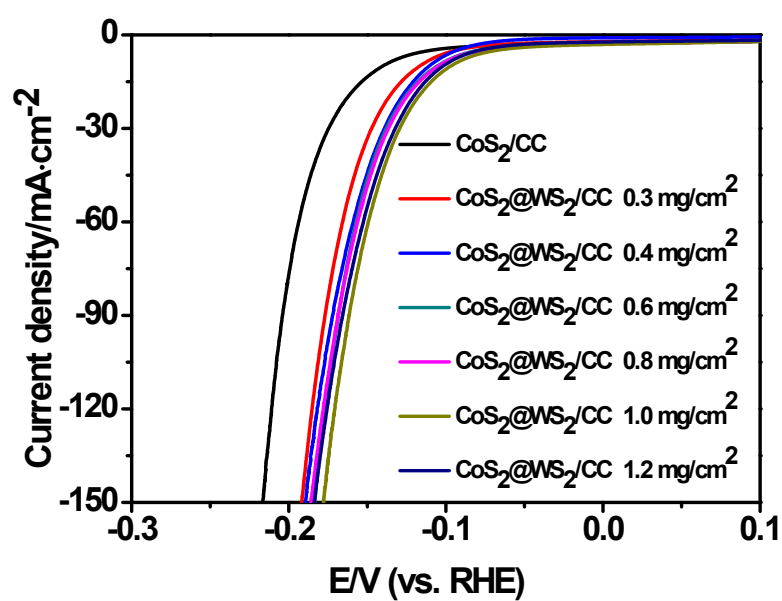


Figure S8. Polarization curves of CoS<sub>2</sub> (loading of 2.0 mg/cm<sup>2</sup>) and a series of CoS<sub>2</sub>@WS<sub>2</sub>/CC samples with different dip-coating content of WS<sub>2</sub>, where the current is normalized by the geometrical area of carbon cloth and the potential is after internal resistance correction.

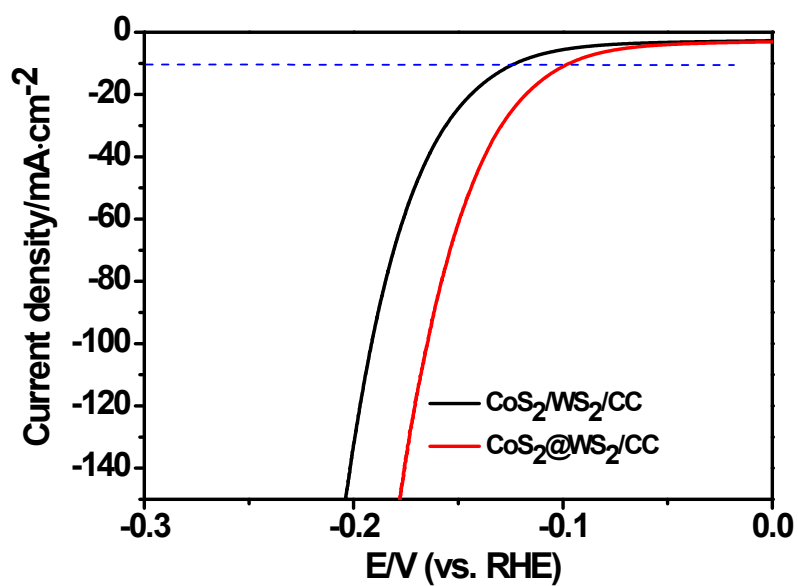


Figure S9. Polarization curves of CoS<sub>2</sub>/WS<sub>2</sub>/CC (preparing by a two-step of thermolysis) and CoS<sub>2</sub>@WS<sub>2</sub>/CC (preparing by a one-step of thermolysis) samples, where the current is normalized by the geometrical area of carbon cloth and the potential is after internal resistance correction.

**Table S1.** Comparison of HER performance for CoS<sub>2</sub>-based and WS<sub>2</sub>-based sulfides catalysts published recently.

Catalyst	substrate	Loading amount (mg/cm <sup>2</sup> )	Current density ( <i>j</i> , mA/cm <sup>2</sup> )	Overpotential at the corresponding <i>j</i> (mV)	Tafel slope (mV/dec)	Electrolyte Conc. (H <sub>2</sub> SO <sub>4</sub> )
CoS <sub>2</sub> (this work)	CC	2.0±0.1	-10	-140.3	71.5	0.5 M
WS <sub>2</sub> (this work)	CC	1.0±0.1	-10	-151.7	79.6	0.5 M
CoS <sub>2</sub> @WS <sub>2</sub> (this work)	CC	2.9±0.1	-10	-97.2	66.0	0.5 M
CoWS <sup>[1]</sup>	glassy carbon	5.11	-10	-236	74	0.5 M
CoS <sub>2</sub> @MoS <sub>2</sub> nanoarray <sup>[2]</sup>	Ti foil	---	-10	-110.5	57.3	0.5 M
MoS <sub>2</sub> /CoS <sub>2</sub> nanorods <sup>[3]</sup>	CC	18.6	-10	-87	73.4	0.5 M
CoS <sub>2</sub> <sup>[3]</sup>	CC	16.5	-10	-288	210.7	0.5 M
CoS <sub>2</sub> thin film <sup>[4]</sup>	graphite disk	---	-10	-192	52	0.5 M
CoS <sub>2</sub> nanowire <sup>[5]</sup>	graphite disk	1.7±0.3	-10	-145	51.6	0.5 M
N-doped CoS <sub>2</sub> <sup>[6]</sup>	carbon					
	black and glass	0.32	-10	-57	48	0.5 M
WS <sub>2</sub> nanoparticles <sup>[7]</sup>	CC	14	-42	-300	78	0.5 M
WS <sub>2</sub> nanosheets <sup>[8]</sup>	graphite disk	1.0±0.2	-10	-142	70	0.5 M
WS <sub>2</sub> nanoribbons <sup>[9]</sup>	glassy carbon	---	-10	-225	68	0.5 M
WS <sub>2</sub> nanoflakes <sup>[10]</sup>	FTO	1.0	-2.1	-200	200	0.5 M
WS <sub>2</sub> @WS <sub>2</sub> nanorattle <sup>[11]</sup>	glassy carbon	0.35	-71	-300	68	0.5 M

## References

- [1] L. Yang, X. L. Wu, X. S. Zhu, C. Y. He, M. Meng, Z. X. Gan and P. K. Chu, *Appl. Surf. Sci.*, 2015, **341**, 149-156.
- [2] H. C. Zhang, Y. J. Li, T. H. Xu, J. B. Wang, Z. Y. Huo, P. B. Wan and X. M. Sun, *J. Mater. Chem A*, 2015, **3**, 15020.
- [3] J. L. Huang, D. M. Hou, Y. C. Zhou, W. J. Zhou, G. Q. Li, Z. H. Tang, L. G. Li and S. W. Chen, *J. Mater. Chem. A*, 2015, **3**, 22886.

- [4] M. S. Faber, M. A. Lukowski, Q. Ding, N. S. Kaiser and S. Jin, *J. Phys. Chem. C*, 2014, **118**, 21347.
- [5] M. S. Faber, R. Dziedzic, M. A. Lukowski, N. S. Kaiser, Q. Ding and S. Jin, *J. Am. Chem. Soc.*, 2014, **136**, 10053.
- [6] J. Y. Zhang, W. Xiao, P. X. Xi, S. B. Xi, Y. H. Du, D. Q. Gao and J. Ding, *ACS Energy Lett.*, 2017, **2**, 1022.
- [7] T. Chen, Y. H. Chang, C. L. Hsu, K. H. Wei, C. Y. Chiang and L. J. Li, *Int. J. Hydrogen. Energy*, 2013, **38**, 12302.
- [8] M. A. Lukowski, A. S. Daniel, C. R. English, F. Meng, A. Forticaux, R. J. Hamers and S. Jin, *Energy Environ. Sci.*, 2014, **7**, 2608.
- [9] J. Lin, Z. W. Peng, G. Wang, D. Zakhidov, E. Larios, M. J. Yacaman and J. M. Tour, *Adv. Energy Mater.*, 2014, **4**, 1301875.
- [10] C. L. Choi, J. Feng, Y. G. Li, J. Wu, A. Zak, R. Tenne and H.J. Dai, *Nano Res.*, 2013, **6**, 921.
- [11] Y. Wen, Y. Xia and S. W. Zhang, *J. Power Sources*, 2016, **307**, 593.