## **Supporting Information**

## Potassium-rich iron hexacyanoferrate/dipotassium terephthalate-

## carbon nanotube composite for K-ion batteries with optimized

## electrolyte

Jiaying Liao<sup>a</sup>, Qiao Hu<sup>a</sup>, Yingtao Yu<sup>a</sup>, Heyang Wang<sup>a</sup>, Zhongfeng Tang<sup>a</sup>, Zhaoyin Wen<sup>b</sup>, Chunhua Chen<sup>a,\*</sup>

 <sup>a</sup> CAS Key Laboratory of Materials for Energy Conversions, Department of Materials Science and Engineering & Collaborative Innovation Center of Suzhou Nano Science and Technology, University of Science and Technology of China, Hefei 230026, Anhui, China
<sup>b</sup> Shanghai Institute of Ceramics, Chinese Academy of Sciences, Shanghai 200050, China



**Fig. S1** The characterizations of KFeHCF: (a) TG curve, (b) FTIR spectrum, (c) EDS results.



Fig. S2 Charge-discharge curves of the initial 3 cycles of KFeHCF in different electrolytes at a current density of 26 mA  $g^{-1}$  (0.2C).



Fig. S3 Charge-discharge curves of KFeHCF in 0.8 M KPF<sub>6</sub> EC/DEC. It was first charged and discharged in 0.8 M KPF<sub>6</sub> EC/DEC+1 vol% FEC, washed by DEC and then reassembled.



**Fig. S4** (a) Initial three-cycle CV curves of KFeHCF in 0.05 M KClO<sub>4</sub>/PC from 2.0-4.4V at a scan rate of 0.1 mV s<sup>-1</sup>; (b) The comparison of EIS spectra of KFeHCF in 0.05 M KClO<sub>4</sub>/PC at different cycles. (The inset graph is the enlarged region at the high frequency end).



Fig. S5 Linear sweep voltammetry of potassium half cells using aluminum foil as counter electrode in different electrolytes from OCV to 5.0 V at a scan rate of 0.1 mV s<sup>-1</sup>.



Fig. S6 (a) Different states of KFeHCF during the first charge-discharge process. It was tested in 0.8 M KPF<sub>6</sub> EC/DEC+1%FEC at 0.2C. (b) Ex-situ XRD patterns at these states.



**Fig. S7** The comparison of the second charge-discharge curves of KFeHCF in sodium (black line) and potassium (red line) half cells.

The voltage difference resulted from  $K^+$  concentration in sodium-ion battery is calculated as :

$$K_{2}FeFe(CN)_{6} + xNa^{+} \leftrightarrow K_{2-x}FeFe(CN)_{6} + xK^{+} + xNa$$
$$E = E^{\Theta} - \frac{RT}{xF} \ln \frac{\left[K^{+}\right]^{x}}{\left[Na^{+}\right]^{x}}$$
$$\Delta E = -\frac{RT}{F} \ln \frac{\left[K^{+}\right]}{\left[Na^{+}\right]} = 0.11 \text{ V}$$



Fig. S8 The SEM image of bulk K<sub>2</sub>TP.



Fig. S9 The charge-discharge curves of  $K_2$ TP-60 (a),  $K_2$ TP-120 (b), bulk  $K_2$ TP (c) and CNT (d) in 1 M KPF<sub>6</sub>/DME. (e) Charge-discharge curves of CNT at different current densities.



**Fig. S10** (a) The charge-discharge curves of the selected cycles of  $K_2TP-75$  in 1 M KPF<sub>6</sub>/DME at a current density of 5C (1250 mA g<sup>-1</sup>). (b) EIS spectrum of  $K_2TP-75$  after 400 cycles.



**Fig. S11** (a) Initial three-cycle CV curves of the KFeHCF/K<sub>2</sub>TP full cell from 1.5 to 3.8V at a scan rate of 0.1 mV s<sup>-1</sup>. (b) EIS spectra of the KFeHCF/K<sub>2</sub>TP full cell after 3 and 60 cycles.