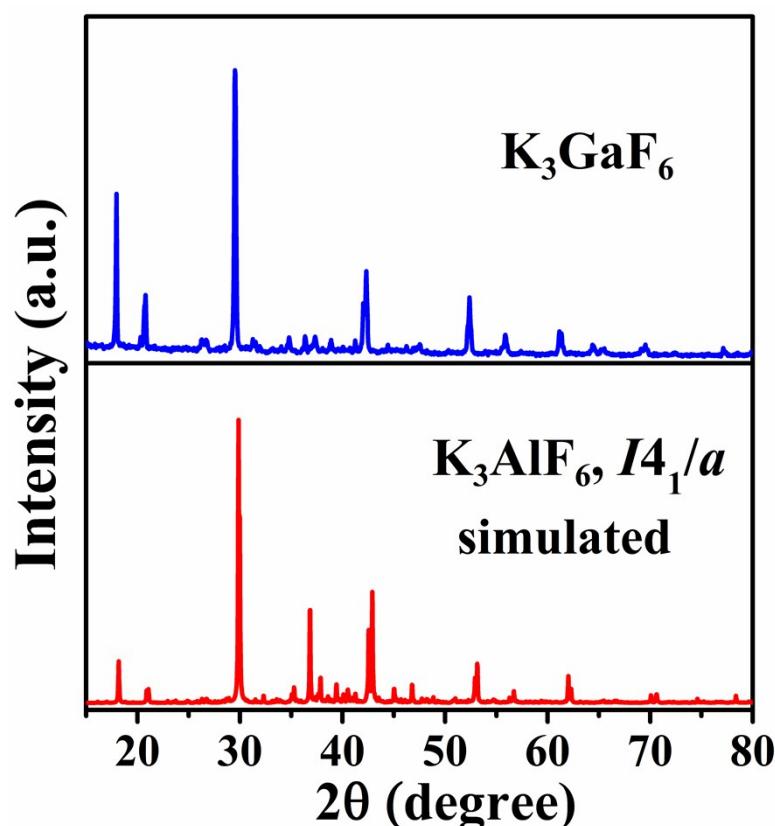


## Supporting Information

### Stable narrowband red phosphor $K_3GaF_6:Mn^{4+}$ derived from hydrous $K_2GaF_5(H_2O)$ and $K_2MnF_6$

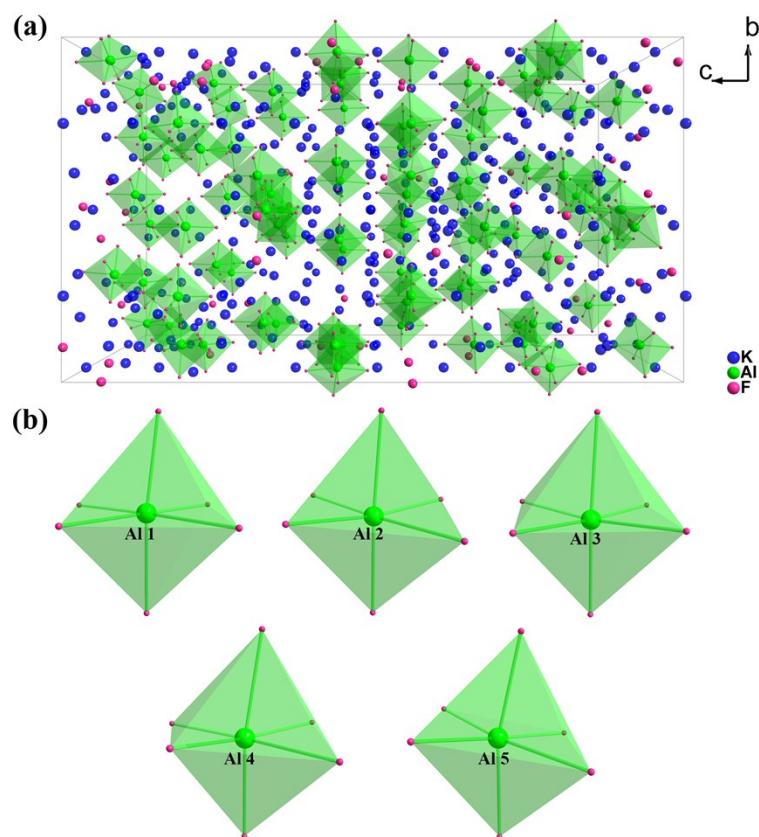
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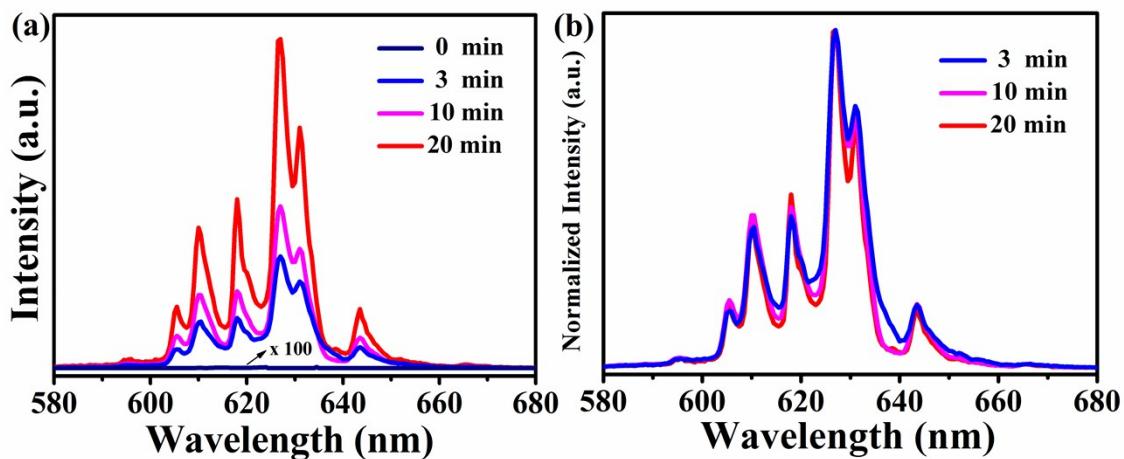


**Fig. S1** XRD patterns of  $K_3AlF_6$  and  $K_3GaF_6$  in  $I4_1/a$  space group.

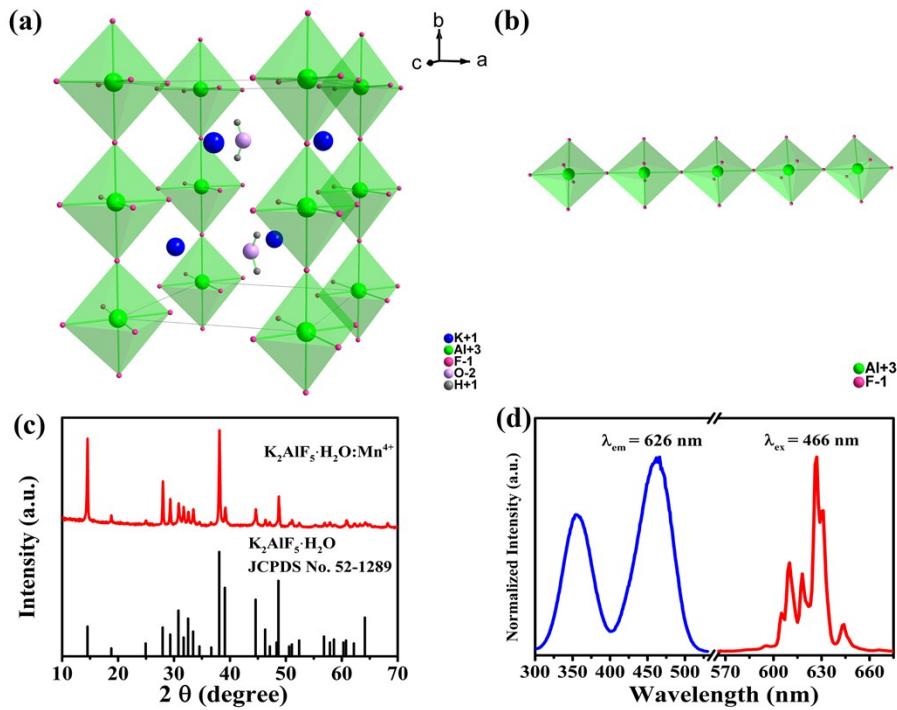
<sup>1</sup> Author to whom correspondence should be addressed; Email:[gyzhang@scut.edu.cn](mailto:gyzhang@scut.edu.cn)



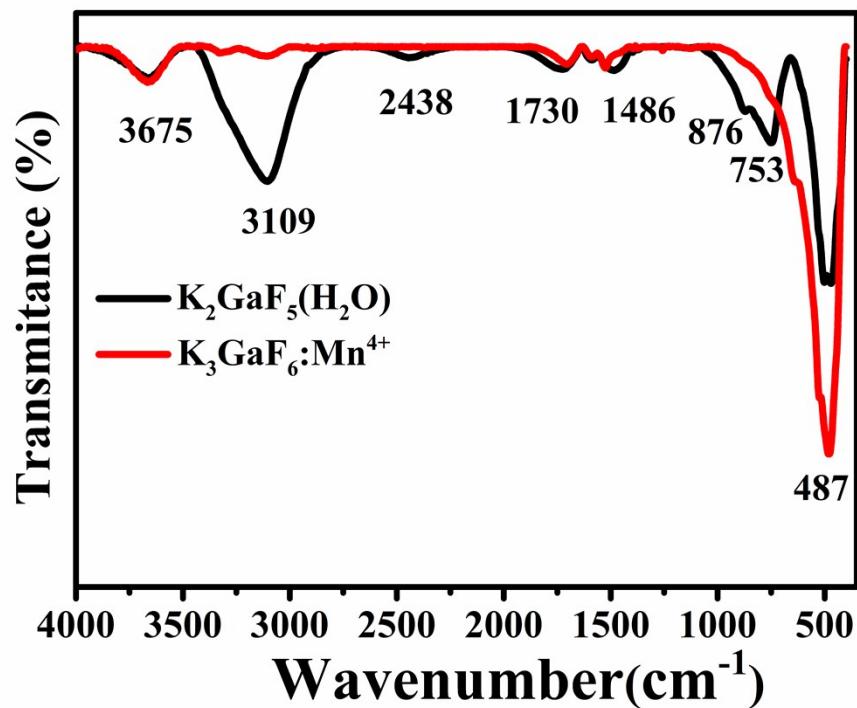
**Fig. S2** Crystal structure of  $\text{K}_3\text{AlF}_6$  in  $I4_1/a$  space group: (a) the unit cell and (b) five types of distorted  $[\text{AlF}_6]^{3-}$  octahedrons.



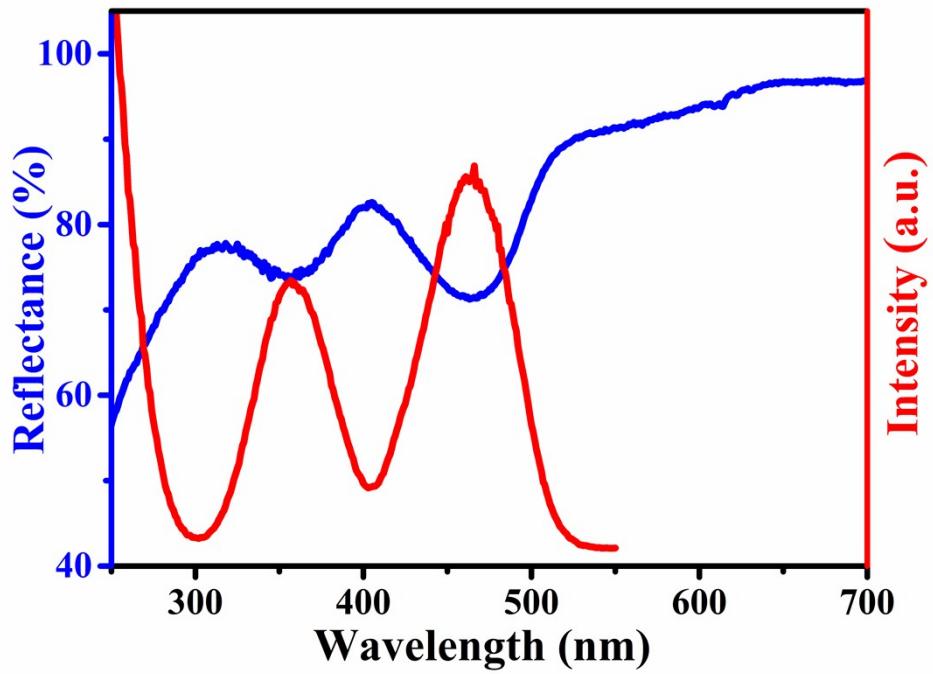
**Fig. S3** The emission spectrum (a) and normalized emission spectrum (b) of time-dependent products from reaction of  $\text{K}_3\text{GaF}_6:\text{Mn}^{4+}$ : 0, 3, 10, and 20 min, respectively.



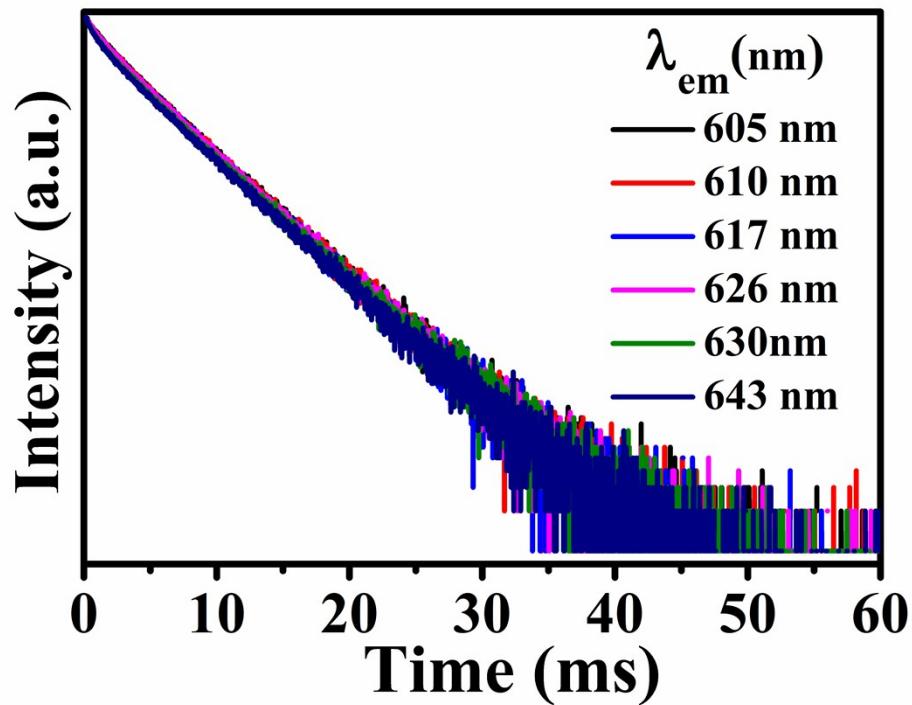
**Fig. S4** The crystal structure of K<sub>2</sub>AlF<sub>5</sub>·H<sub>2</sub>O in P2<sub>1</sub>/m space group: the unit cell (a) and Al-F chain composed of corner-sharing [AlF<sub>6</sub>] octahedrons (b). XRD pattern (c) and excitation, emission spectrum (d) of K<sub>2</sub>AlF<sub>5</sub>·H<sub>2</sub>O:Mn<sup>4+</sup>.



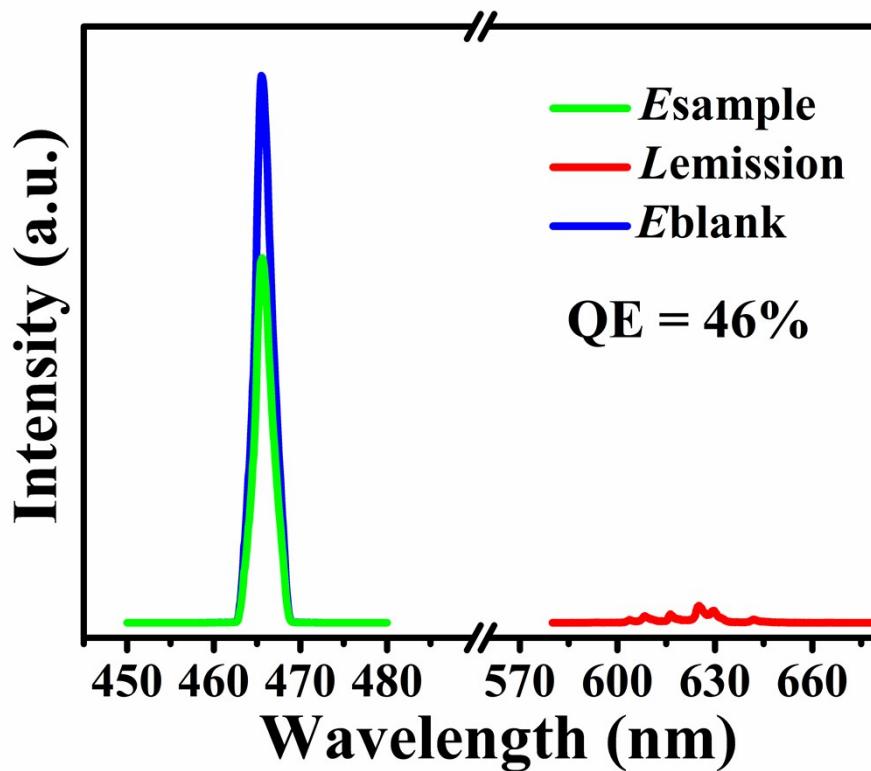
**Fig. S5** IR spectra of K<sub>2</sub>GaF<sub>5</sub>(H<sub>2</sub>O) and K<sub>3</sub>GaF<sub>6</sub>:Mn<sup>4+</sup>.



**Fig. S6** The excitation spectra and diffuse reflection spectrum of  $\text{K}_3\text{GaF}_6:\text{Mn}^{4+}$ .



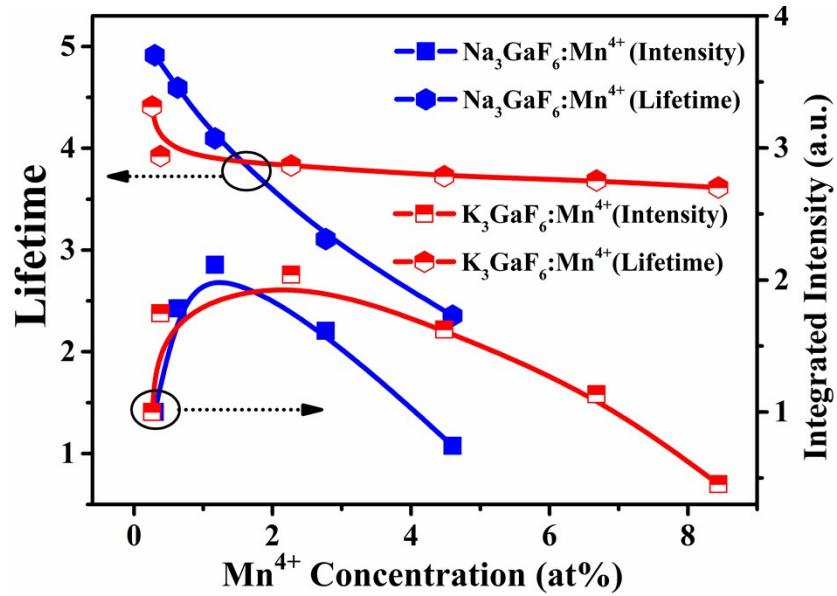
**Fig. S7** Luminescence decay curves of different emission peaks in  $\text{K}_3\text{GaF}_6:\text{Mn}^{4+}$  (2.27 at%) upon 467 nm excitation.



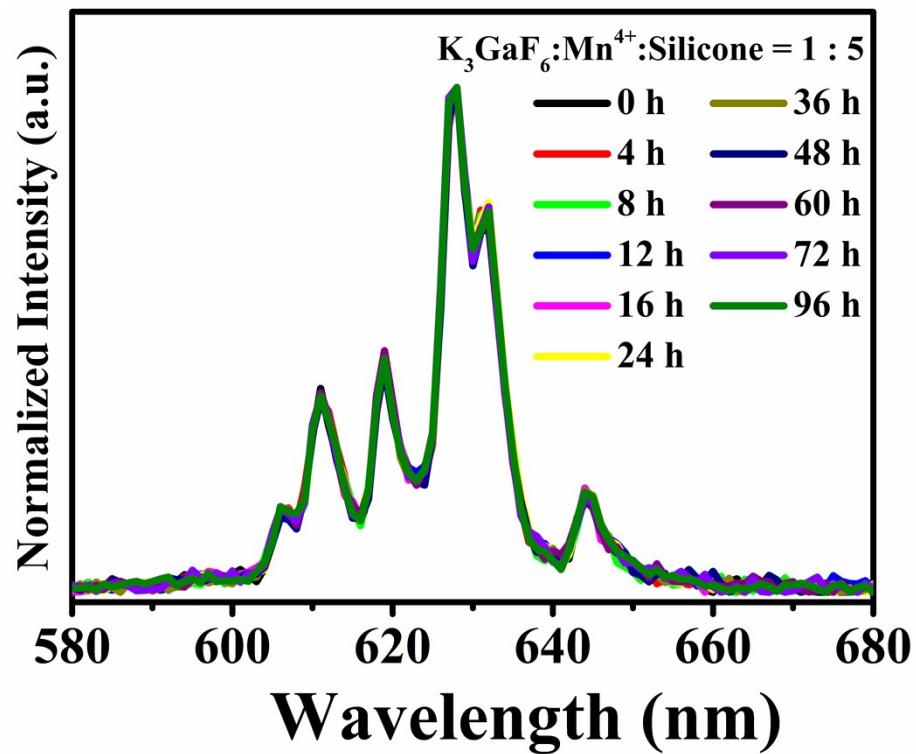
**Fig. S8** Excitation line of blank ( $\text{BaSO}_4$ ) ( $E_{\text{blank}}$ ) and excitation ( $E_{\text{sample}}$ ) and emission ( $L_{\text{emission}}$ ) spectra of  $\text{K}_3\text{GaF}_6:\text{Mn}^{4+}$  (2.27 at%). The quantum efficiency (QE) can be calculated as follows:

$$QE = \frac{\int L_{\text{emission}}}{\int E_{\text{blank}} - \int E_{\text{sample}}} \quad (1)$$

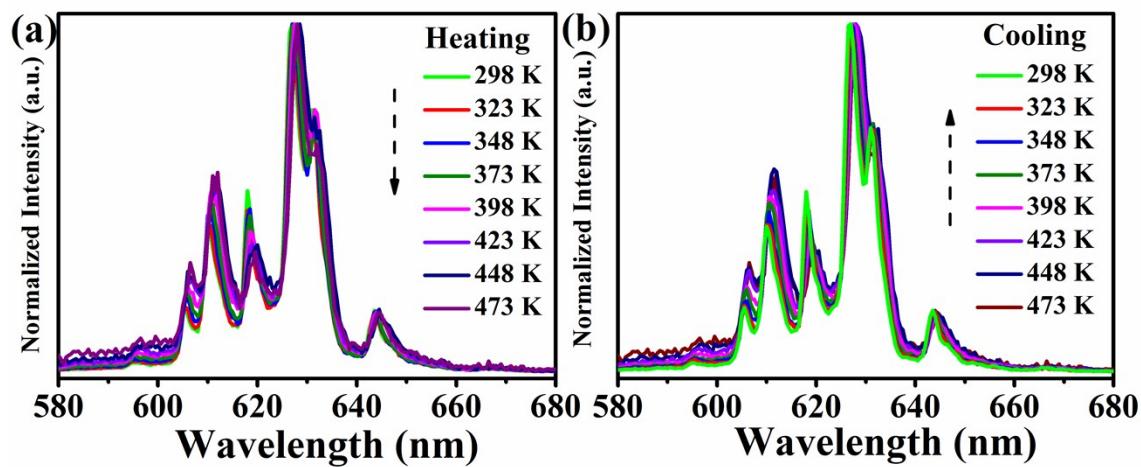
where  $L_{\text{emission}}$  and  $E_{\text{sample}}$  are the emission and excitation spectra of objective samples;  $E_{\text{blank}}$  is the spectrum of the excitation light without samples in integrating sphere. All of spectra were recorded by an optical integrating sphere coupled to the FLS920 fluorescence spectrophotometer at room temperature.



**Fig. S9** Concentration dependent integrated intensities and decay lifetimes of  $\text{A}_3\text{GaF}_6:\text{Mn}^{4+}$  (A = Na, K) with various  $\text{Mn}^{4+}$  contents.



**Fig. S10 (a)** Time-dependent normalized luminescence spectra of  $\text{K}_3\text{GaF}_6:\text{Mn}^{4+}$  (2.27 at%) loaded in the 85 % humidity and 85 °C environment.



**Fig. S11** Normalized temperature-dependent luminescence spectra of  $\text{K}_3\text{GaF}_6:\text{Mn}^{4+}$  (2.27 at%) in the cyclic process of heating and cooling between 298 to 473 K.

**Table S1** Synthesis strategies of  $A_3GaF_6:Mn^{4+}$  ( $A = Na, K$ ) by co-precipitation method.

Compound	Ga source	A source	Mn precursor	Molar ratio of Ga atom to A atom	Amount of HF solution (49 wt.%)/mL
$Na_3GaF_6:Mn^{4+}$	$Ga_2O_3$	$NaF$	$K_2MnF_6$	5:10	10
$K_3GaF_6:Mn^{4+}$	$Ga_2O_3$	$K_2CO_3$		5:16	8

**Table S2** Selected bond lengths ( $\text{\AA}$ ) and bond valence sum (BVS) calculations for  $K_2GaF_5(H_2O)$ .<sup>a</sup>

Bond	Distance	BVS	BVS(Sum)
$Ga(1)-F(1)$	1.8681(19)	0.511	
$Ga(1)-F(1)\#A$	1.8681(19)	0.511	
$Ga(1)-F(2)$	1.8897(18)	0.482	
$Ga(1)-F(2)\#A$	1.8897(18)	0.482	2.989
$Ga(1)-F(3)$	1.876(3)	0.501	
$Ga(1)-O(1)$	1.986(4)	0.501	
$K(1)-F(1)$	2.696(2)	0.123	
$K(1)-F(1)\#B$	2.687(2)	0.152	
$K(1)-F(1)\#C$	2.792(2)	0.114	
$K(1)-F(2)\#D$	2.674(2)	0.157	
$K(1)-F(2)$	3.400(2)	0.058	
$K(1)-F(2)\#E$	2.788(2)	0.116	
$K(1)-F(3)$	2.7805(11)	0.111	1.262
$K(1)-F(3)\#A$	2.7805(11)	0.111	
$K(1)-F(3)\#F$	2.8040(12)	0.111	
$K(1)-F(3)\#D$	2.8040(12)	0.111	
$K(1)-O(1)\#E$	3.2454(13)	0.049	
$K(1)-O(1)\#C$	3.2454(13)	0.049	
$O(1)-Ga(1)$	1.986(4)	0.501	
$O(1)-K(1)\#C$	3.2454(13)	0.049	0.598
$O(1)-K(1)\#E$	3.2454(13)	0.049	

<sup>a</sup>Symmetry codes for  $K_2GaF_5(H_2O)$ : (A)  $-x, y, -z+1/2$ ; (B)  $-x+1/2, -y+1/2, -z$ ; (C)  $-x+1/2, y-1/2, -z+1/2$ ; (D)  $-x, -y+1, -z$ ; (E)  $x-1/2, y-1/2, z$ ; (F)  $x, -y+1, z+1/2$ .

**Table S3** ICP results of  $K_3GaF_6:Mn^{4+}$  phosphors prepared with different mole ratios of  $Ga_2O_3$  to  $K_2MnF_6$ .

Samples	Experimental molar ratio of	Actual doping concentration
	Ga <sub>2</sub> O <sub>3</sub> to K <sub>2</sub> MnF <sub>6</sub>	of Mn <sup>4+</sup> (mol%)
G1	100:0.5	100:0.26
G2	100:1	100:0.38
G3	100:3	100:2.27
G4	100:5	100:4.48
G5	100:7	100:6.68
G6	100:10	100:8.44

**Table S4** Photoelectric parameters of the three fabricated warm WLEDs.

Device	Chromaticity coordinate (x, y)	CCT(K)	Ra	R9	Efficacy (lm/W)
LED 1	(0.3822, 0.3902)	4048	83.9	34.7	102.7
LED 2	(0.3979, 0.3939)	3691	87.2	50.2	92.1
LED 3	(0.4476, 0.4253)	2994	88.4	55.0	71.0

**Table S5** Photoelectric parameters of warm WLEDs denoted as LED I and LED II by using Na<sub>3</sub>GaF<sub>6</sub>:Mn<sup>4+</sup> and K<sub>3</sub>GaF<sub>6</sub>:Mn<sup>4+</sup> as red light component, respectively.

No.	CIE coordinates	CCT (K)	Ra	R9	Efficacy (lmW <sup>-1</sup> )
LED I	(0.4434 0.4131)	2966	81	31	56.73
LED II	(0.4476, 0.4253)	2994	88.4	55.0	71.0