Supporting Information

Electrolytic approach towards the controllable synthesis of NiO nanocrystalline and self-assembly mechanism of Ni(OH)₂ precursor under the electric, temperature and magnetic field

Jian Wang,*^a Yuli Zhou,^a Mingyang He,^a Peihua Wangyang,^b Yangfan Lu*^c and Lin Gu^a

^a School of Materials Science and Engineering, Xihua University, Chengdu 610039,
People's Republic of China. E-mail: wangjianxhu@163.com; Fax: +86 28 8772 8746;
Tel: +86 28 8772 8746

^b Information Materials and Devices Application Key Laboratory of Sichuan Provincial Universities, School of Optoelectronic Technology, Chengdu University of Information Technology, Chengdu 610225, People's Republic of China

^c Materials Research Center for Element Strategy, Tokyo Institute of Technology, Yokohama 2268503, Japan. E-mail: yf_lu@mces.titech.ac.jp



Element	Point 1	Point 2	Point 3	Point 4	Point 5	Point 6	Point 7	Point 8	Point 9
Ni(k)	48.71	47.53	49.35	49.27	47.28	46.38	49.87	48.54	47.26
O(k)	51.29	52.47	50.65	50.73	52.72	53.62	50.13	51.46	52.74

Fig. S1 EDS spectral analysis from nine locations of hexagonal NiO nanosheets, showing the elemental compositions of NiO nanosheets.



Element	Point 1	Point 2	Point 3	Point 4	Point 5	Point 6	Point 7	Point 8	Point 9
Ni(k)	49.24	49.68	47.35	46.59	48.71	47.62	48.65	49.18	48.39
O(k)	50.76	50.32	52.65	53.41	51.29	52.38	51.35	50.82	51.61

Fig. S2 EDS spectral analysis from nine locations of hexagonal NiO nanoflowers, showing the elemental compositions of NiO nanoflowers.