Supporting Information



Figure S1. The oxygen concentration effect on the formation of concave Cu_2O truncated cubes with velocity of 10 mL/min and etching time of 0.5 h, (a) air, (b) 50%, (c) 80%, (d) pure oxygen.



Figure S2. The flow velocity effect on the formation of concave Cu_2O truncated cubes with air and etching time of 1 h, (a) 5 mL/min, (b) 10 mL/min, (c) 20 mL/min.



Figure S3. Selective etching of $\{100\}$ facets for Cu₂O hexapod microcrystals with the assistance of PVP at air velocity of 10 mL/min, (a) 0 h, (b) 1 h, (c) 2 h, (d) 2.5 h.

The experimental details for the synthesis of hexapod Cu₂O microcrystals:

In a typical procedure, an aqueous solution was first prepared by mixing 34 mL of deionized water (18 MΩ) and 2 mL of 0.68 M CuSO₄·5H₂O in a 100 mL flat bottomed flask. After the mixture was stirred with a magnetic blender for about 15 min, 2 mL mixed solution of 0.74 M Na₃C₆H₅O₇·2H₂O and 1.2 M Na₂CO₃ were dropped into the solution. Stirred for another 10 min, then, 0.5 g NaCl was added. When the NaCl powder was dissolved, 2 mL of 1.4 M glucose was slowly added in a dropwise manner. Kept the solution in water bath at temperature of 80 °C for 15 min, then, introduced the air into the reaction system, at flow velocity of 10 mL/min. Controlling the speed of magnetic stirring to make sure the gas was uniformly diffused in the solution. 2 h later, the brick red precipitate was centrifuged and washed with distilled water and absolute alcohol for several times, dried in a vacuum oven at 60 °C for further characterization.



Figure S4. Absorption spectra of Cu_2O cuboctahedrons and Cu_2O cuboctahedrons etched.



Figure S5. Absorption spectra of CR solution in the presence of the as-prepared Cu_2O , (a) basic structure, (b) {100} facets etching, (c) {111} facets etching.