

Supporting Information (SI)

Cooperativity Between Hydrogen- and Halogen Bonds: the Case of Selenourea

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Titration studies.

Benzoate + **I1**

Table S1. ^{19}F chemical shift of the αF of **I1** (22.6 mM) at different concentrations of [**TBABzO**].

[TBABzO] (M)	$\delta_{\alpha\text{F}}(\text{I1})$ (ppm)
0	-65,1614
0,0134	-70,5373
0,0176	-71,8447
0,0192	-72,159
0,0247	-73,4476
0,03	-74,361
0,036	-75,1553
0,047	-76,0039
0,0747	-77,0226
0,00062	-65,3094
0,0021	-66,206
0,0046	-67,0994
0,0092	-69,0357

Table S2. ^1H chemical shift of the NH (a and b) of **SeU** (64 mM) at different concentrations of [**TBABzO**].

[TBABzO] (M)	$\delta_{\text{NH}a}(\text{SeU})$ (ppm)	$\delta_{\text{NH}b}(\text{SeU})$ (ppm)	averaged $\delta_{\text{NH}a,b}(\text{SeU})$ (ppm)
0	a	a	7,2633
0,0011	a	a	7,301
0,00185	a	a	7,31
0,00232	7,382	7,27	7,326
0,00366	7,498	7,198	7,348
0,00529	7,6154	7,1621	7,38875
0,00706	7,7309	7,1264	7,42865
0,00919	7,8361	7,1165	7,4763
0,01123	7,973	7,0851	7,52905
0,01194	8,003	7,0728	7,5379
0,01393	8,1369	7,0454	7,59115
0,01736	8,3132	7,0239	7,66855
0,0198	8,4553	6,992	7,72365
0,02198	8,5753	6,971	7,77315
0,02343	8,6519	6,9546	7,80325
0,02815	8,9529	6,8914	7,92215
0,03337	9,2298	6,8538	8,0418
0,04514	9,9484	6,727	8,3377
0,04954	10,2051	6,6848	8,44495
0,06337	10,9646	6,5673	8,76595
0,08337	11,3006	6,529	8,9148
0,1078	11,3122	6,5491	8,93065

^aThe two peaks overlap, showing the same δ (the averaged one).

Diffusion NMR studies.

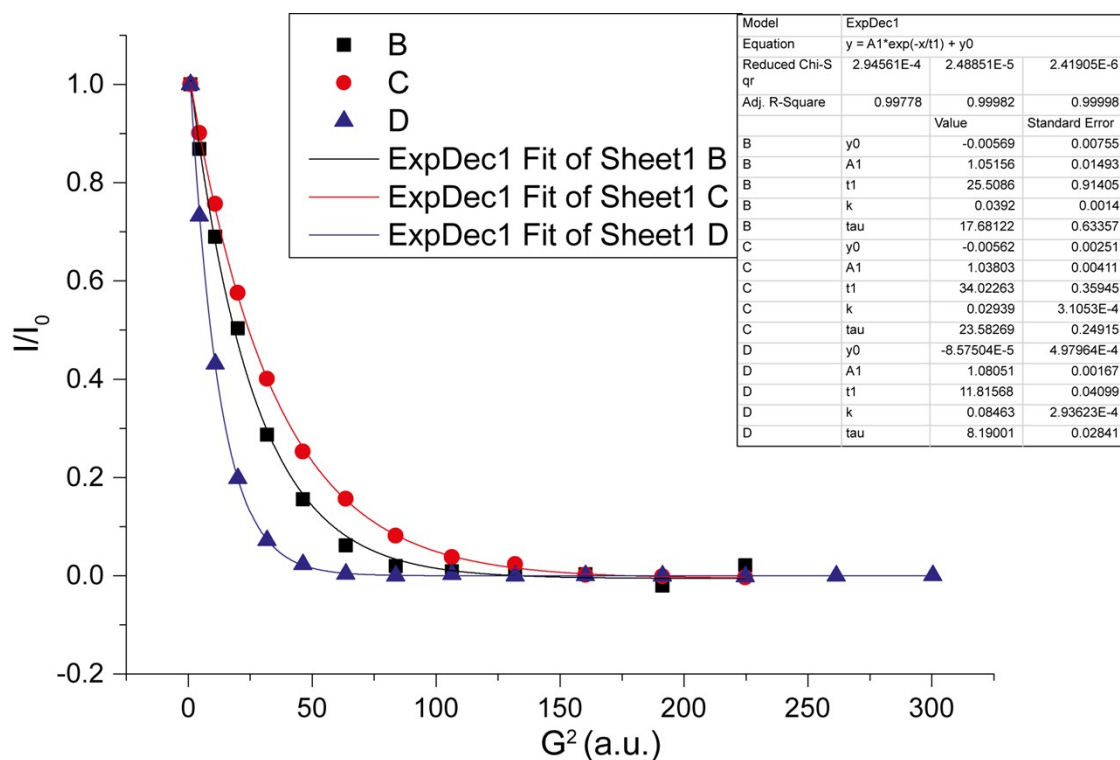


Figure S1. I/I_0 versus G^2 for TBABzO in acetone- d_6 ($C = 2.5$ mM). B = benzoate; C = Tetrabutylammonium; D = solvent.

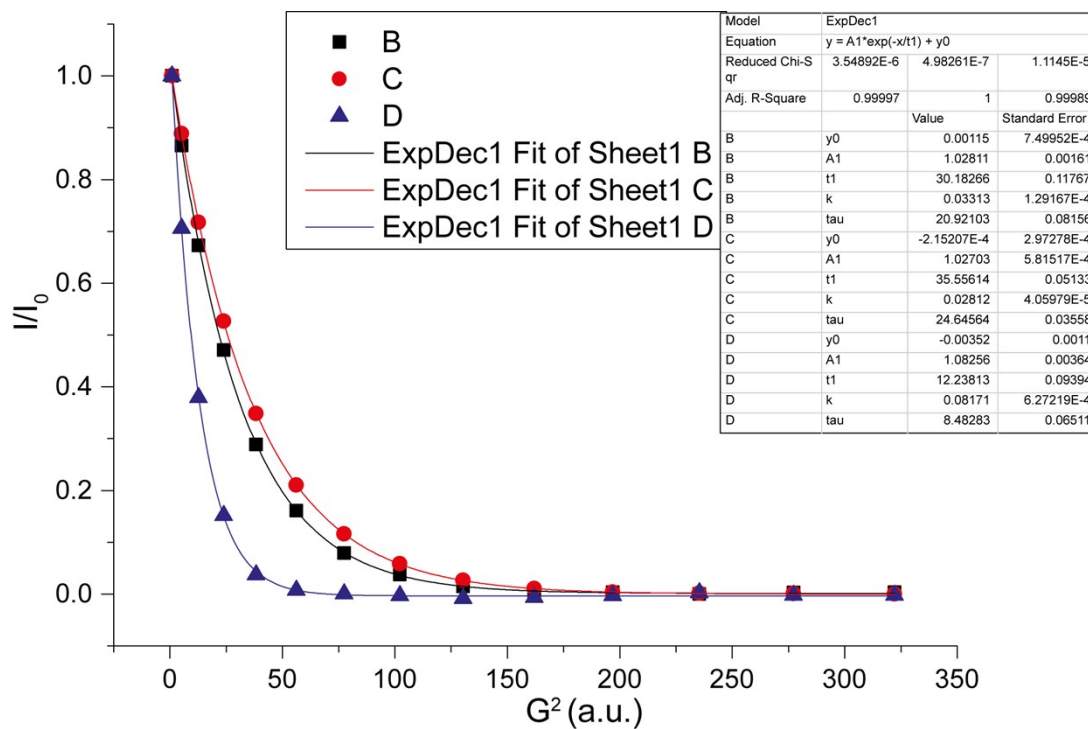


Figure S2 I/I_0 versus G^2 for TBABzO in acetone- d_6 ($C = 20.5$ mM). B = benzoate; C = Tetrabutylammonium; D = solvent.

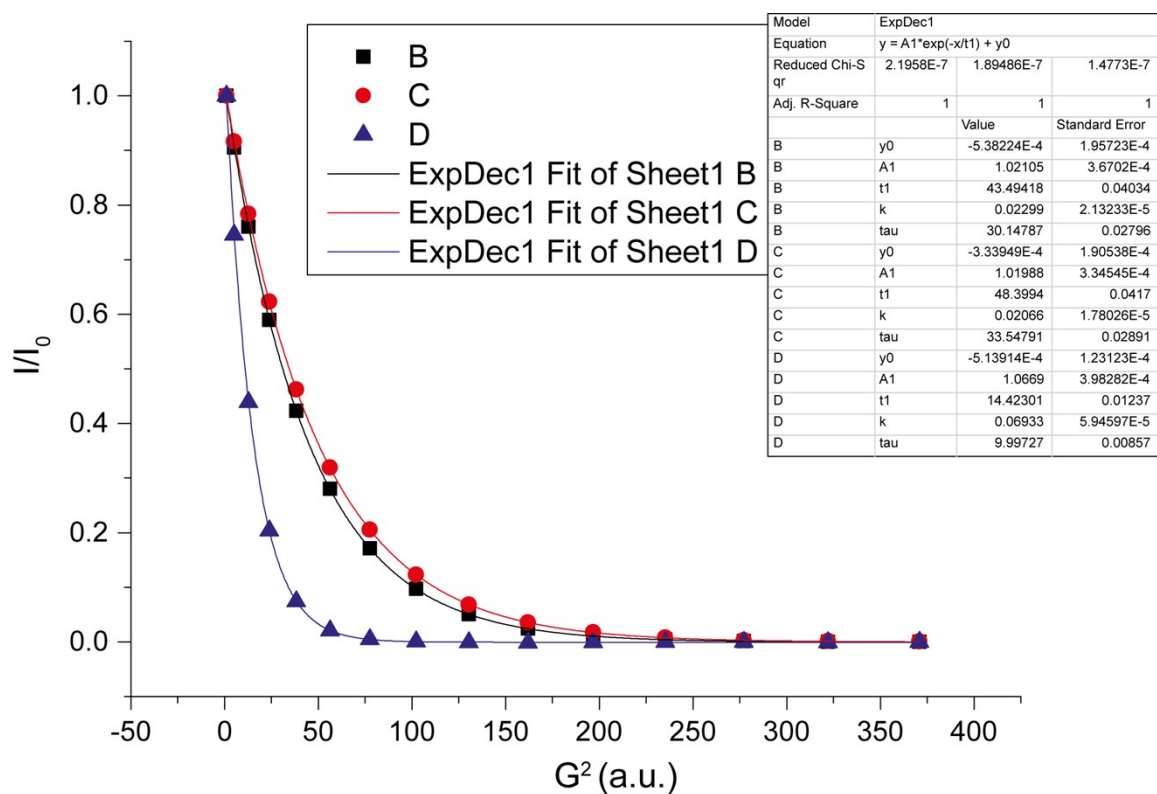


Figure S3 I/I_0 versus G^2 for TBABzO in acetone- d_6 ($C = 185$ mM). B = benzoate; C = Tetrabutylammonium; D = solvent.

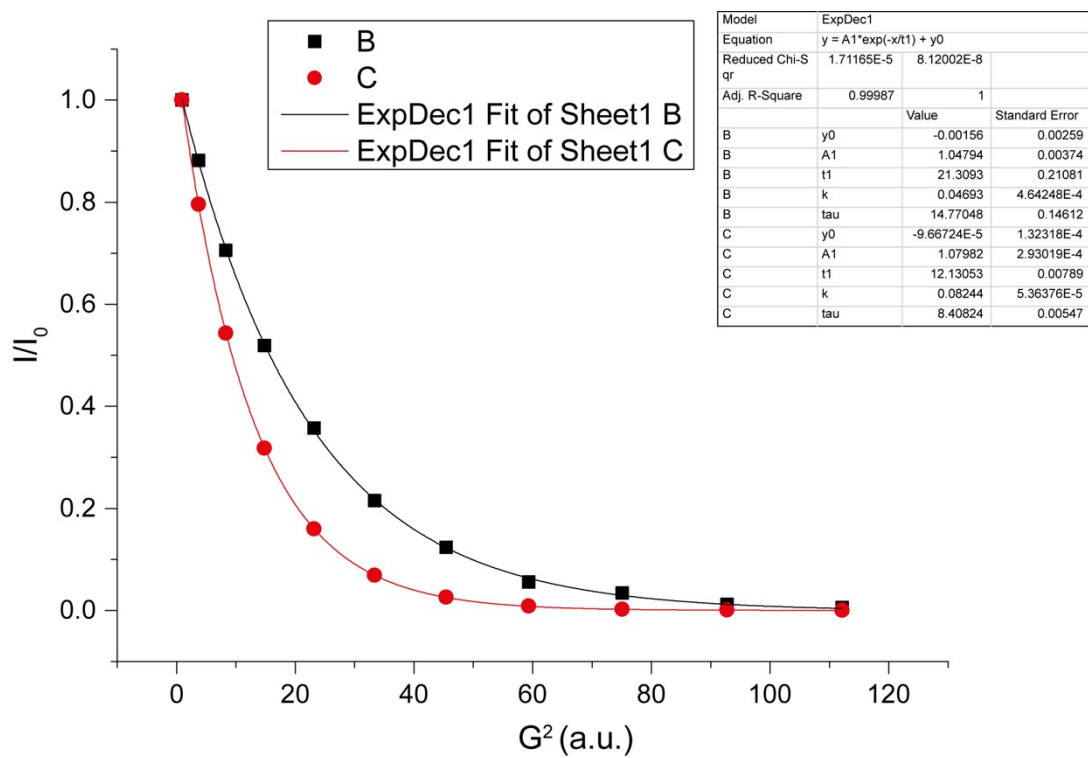


Figure S4 I/I_0 versus G^2 for SeU in acetone- d_6 ($C = 16$ mM). B = SeU; C = solvent.

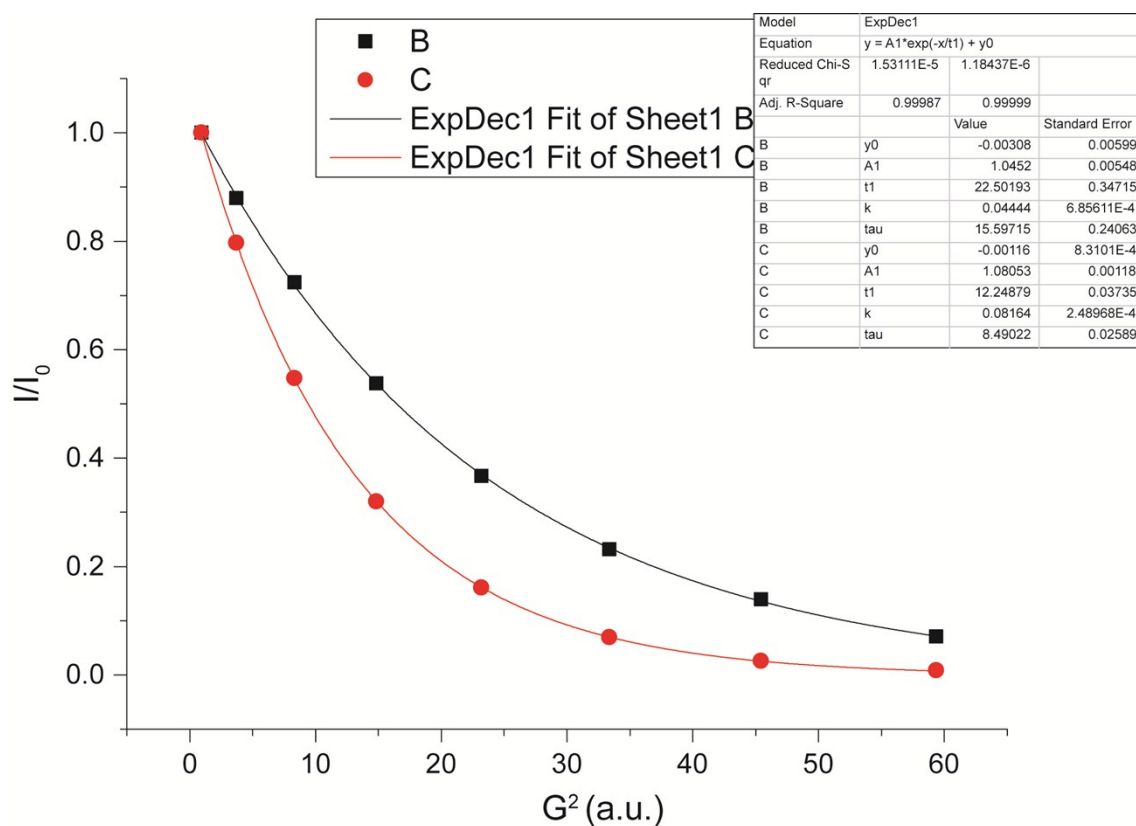


Figure S5 I/I_0 versus G^2 for **SeU** in acetone- d_6 ($C = 16$ mM) + **I1** (19 mM). $B = \text{SeU}$; $C = \text{solvent}$.

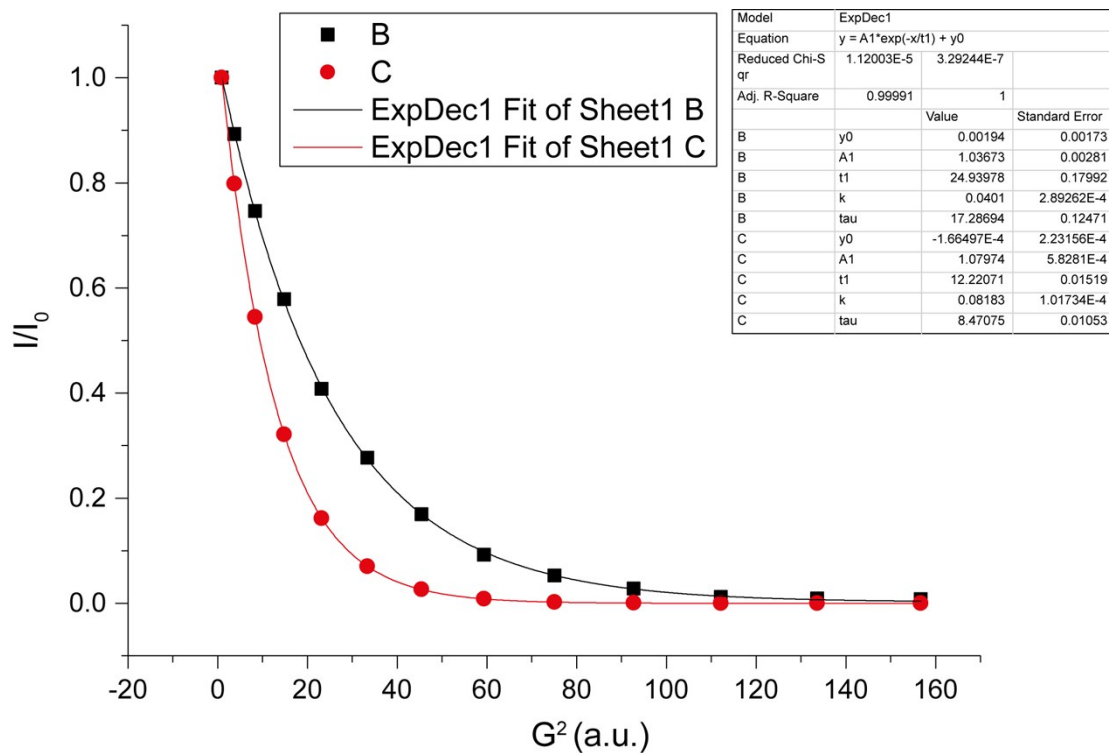


Figure S6 I/I_0 versus G^2 for **SeU** in acetone- d_6 ($C = 16$ mM) + **I1** (82 mM). $B = \text{SeU}$; $C = \text{solvent}$.

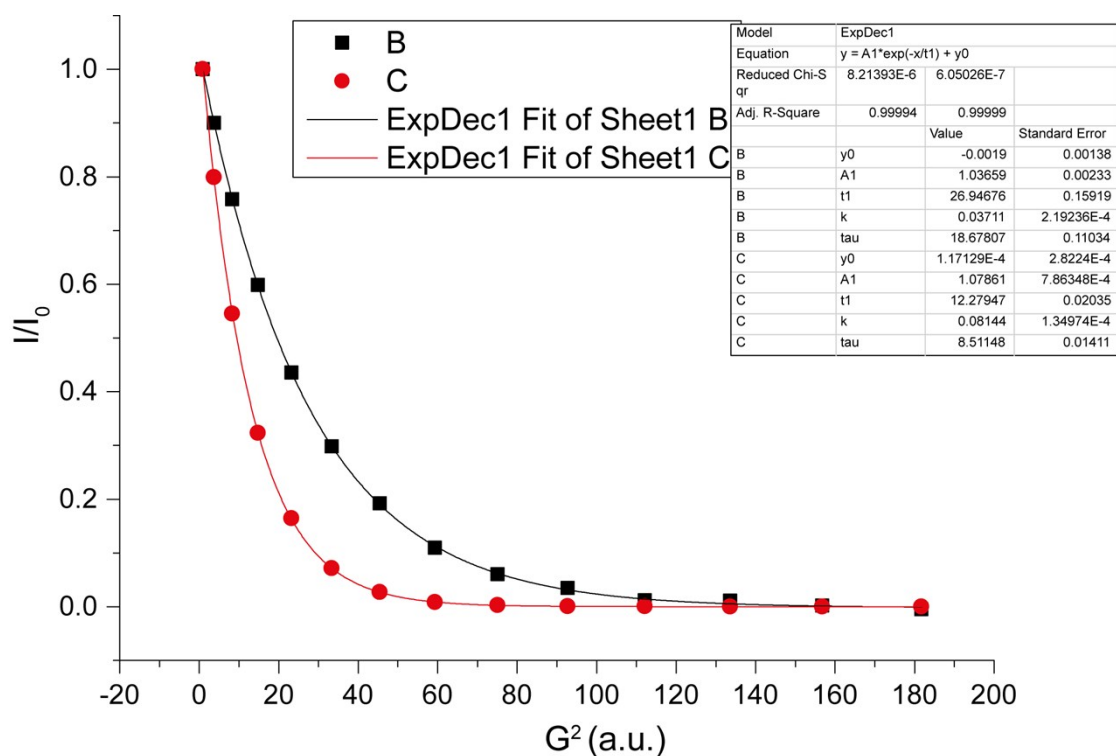


Figure S7 I/I_0 versus G^2 for SeU in acetone- d_6 (C = 16 mM) + I1 (209 mM). B = SeU; C = solvent.

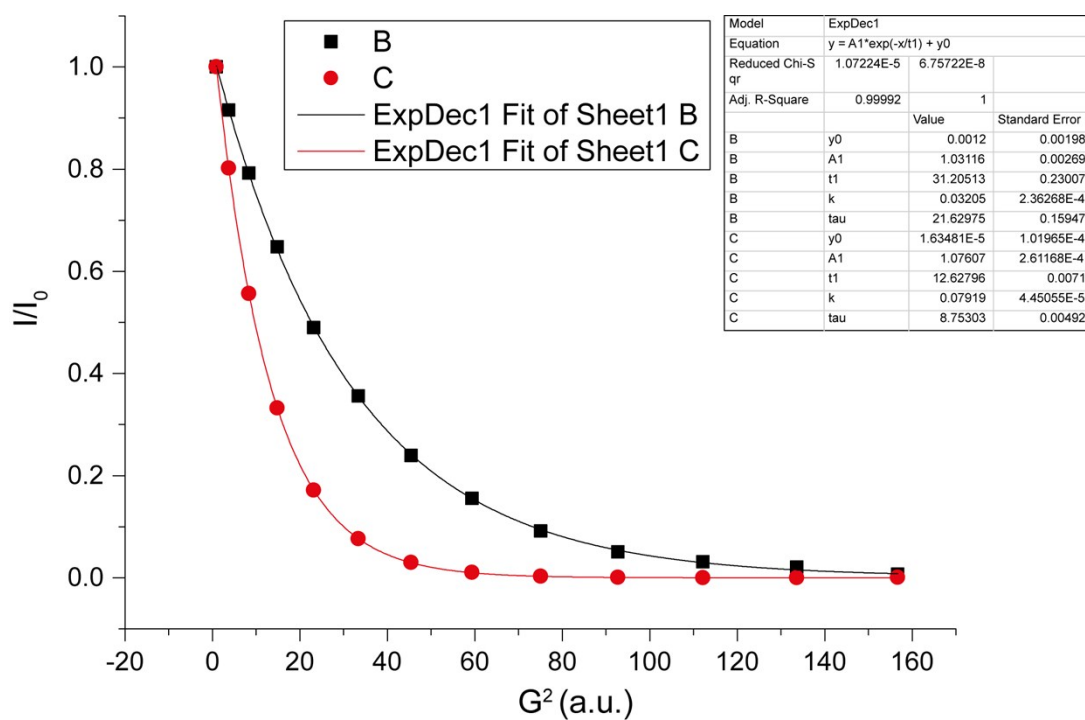


Figure S8 I/I_0 versus G^2 for SeU in acetone- d_6 (C = 16 mM) + I1 (460 mM). B = SeU; C = solvent.

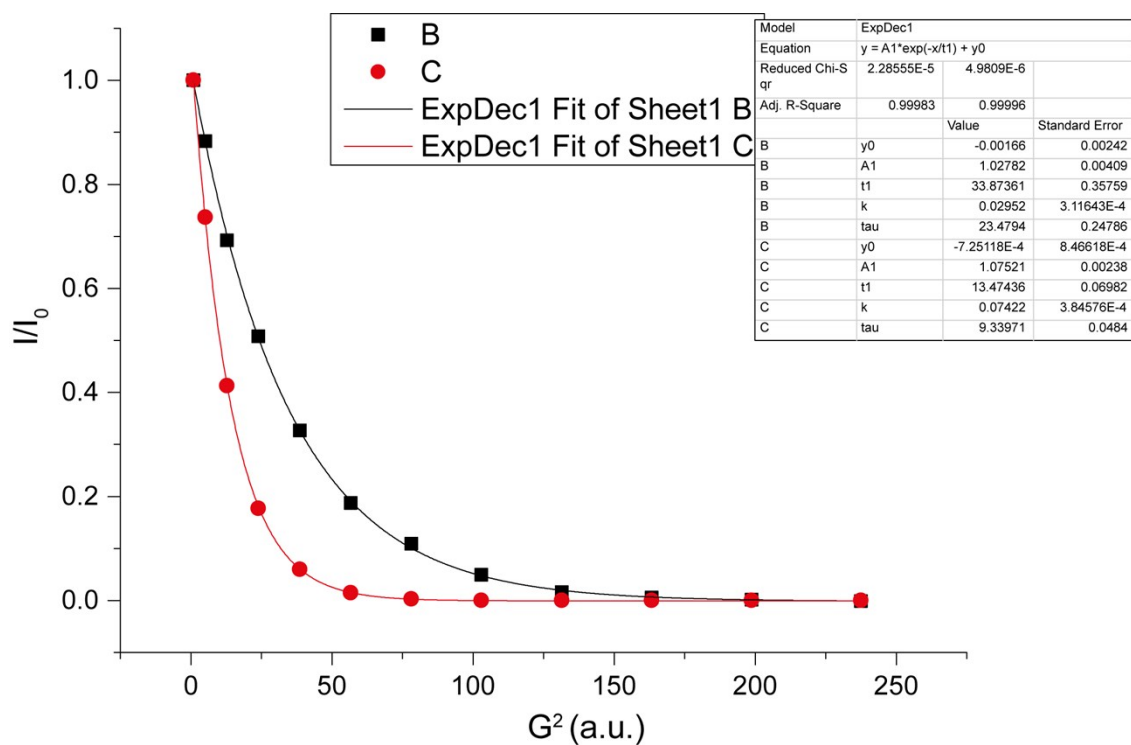


Figure S9 I/I_0 versus G^2 for SeU in acetone- d_6 (C = 16 mM) + **I1** (931 mM). B = SeU; C = solvent.

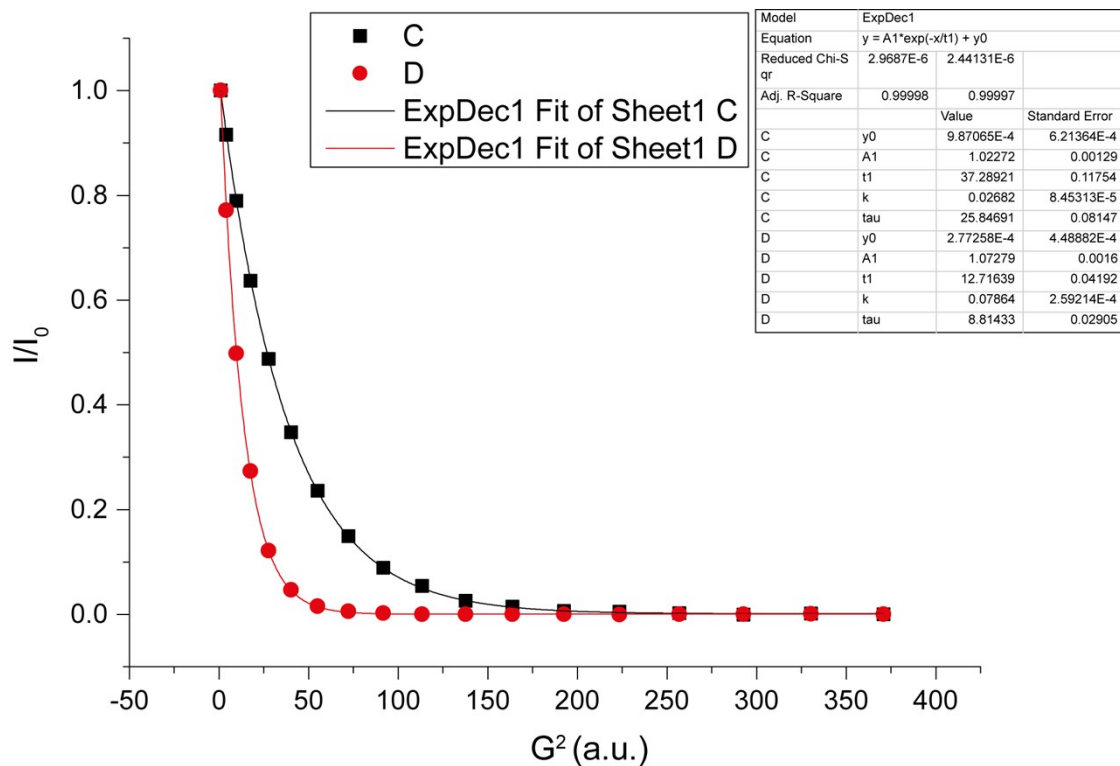


Figure S10. I/I_0 versus G^2 for SeU in acetone- d_6 (C = 32 mM) + **TBABzO** (74 mM). B = SeU; C = solvent.

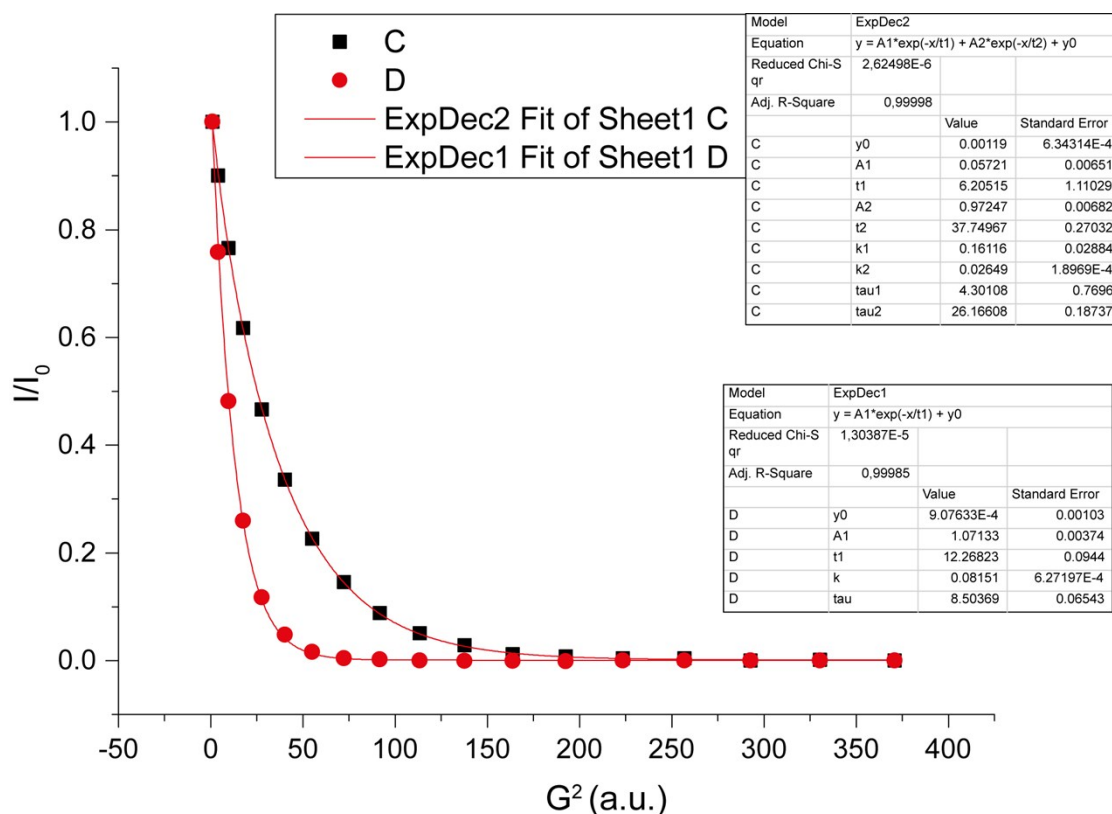


Figure S11. I/I_0 versus G^2 for SeU in acetone- d_6 (C = 32 mM) + TBABzO (74 mM) + I1 (16 mM). B = SeU; C = solvent.

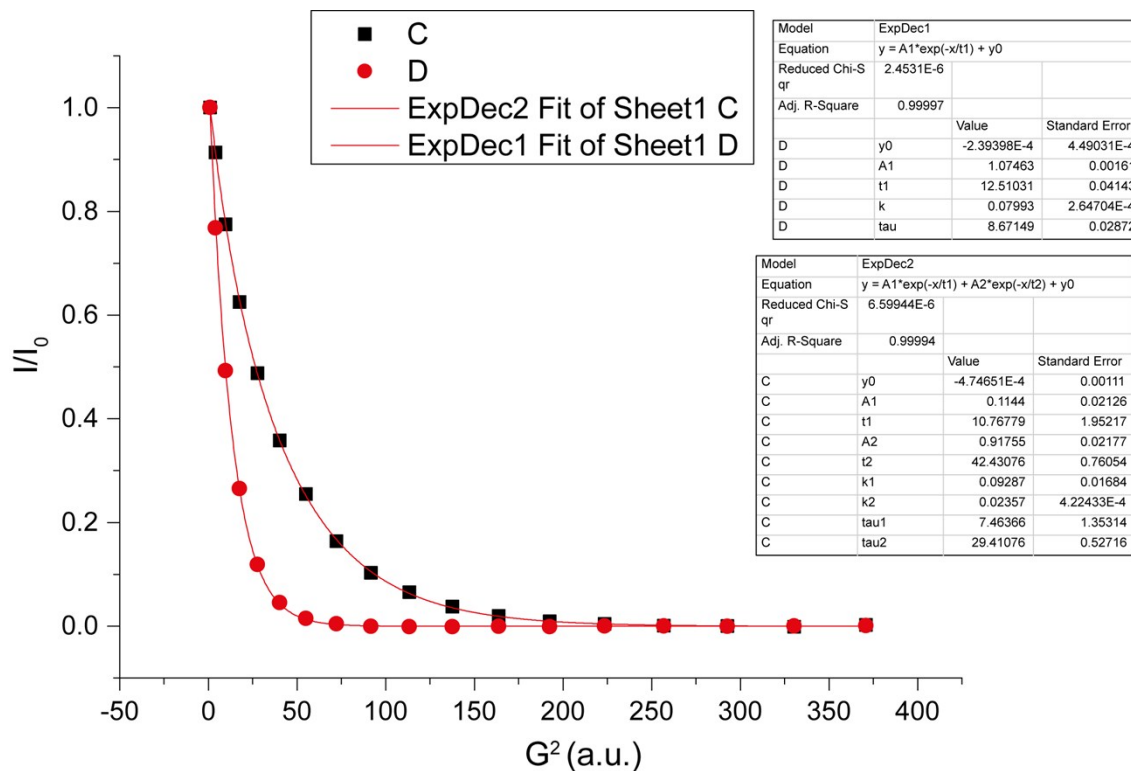


Figure S12. I/I_0 versus G^2 for SeU in acetone- d_6 (C = 32 mM) + TBABzO (74 mM) + I1 (30 mM). B = SeU; C = solvent.

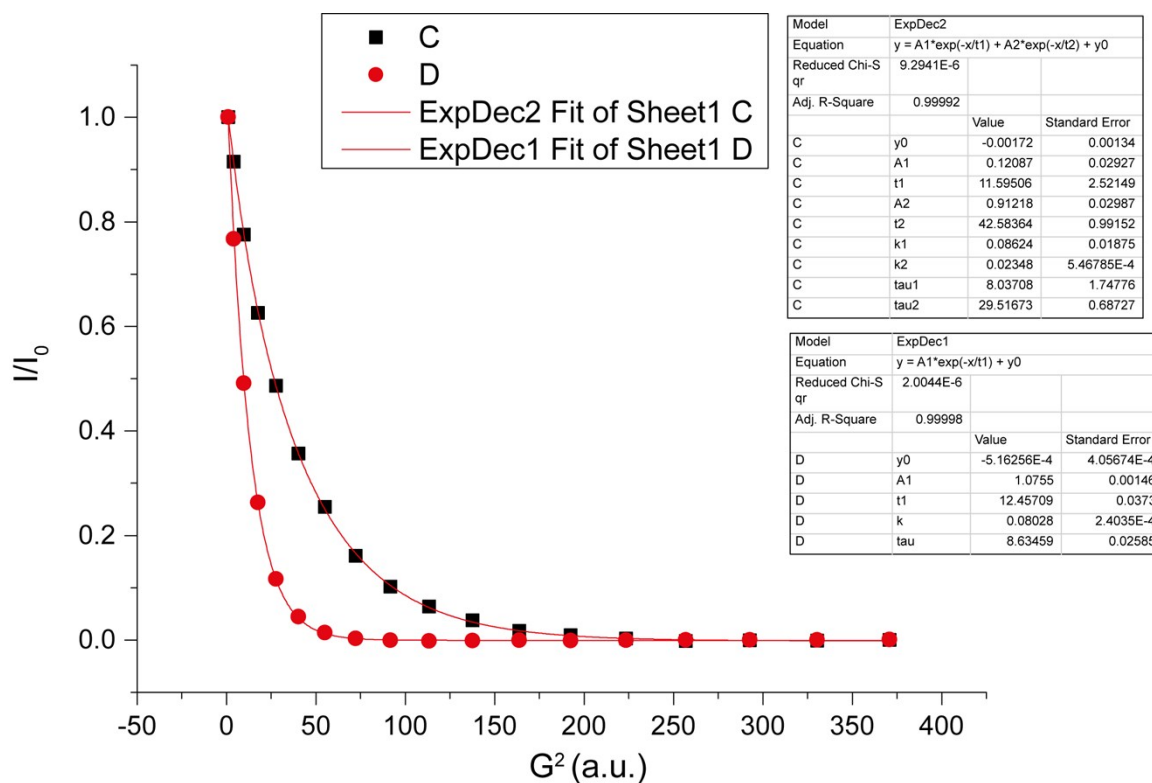


Figure S13 I/I_0 versus G^2 for **SeU** in acetone- d_6 (C = 32 mM) + **TBABzO** (74 mM) + **I1** (52.8 mM). B = **SeU**; C = solvent.

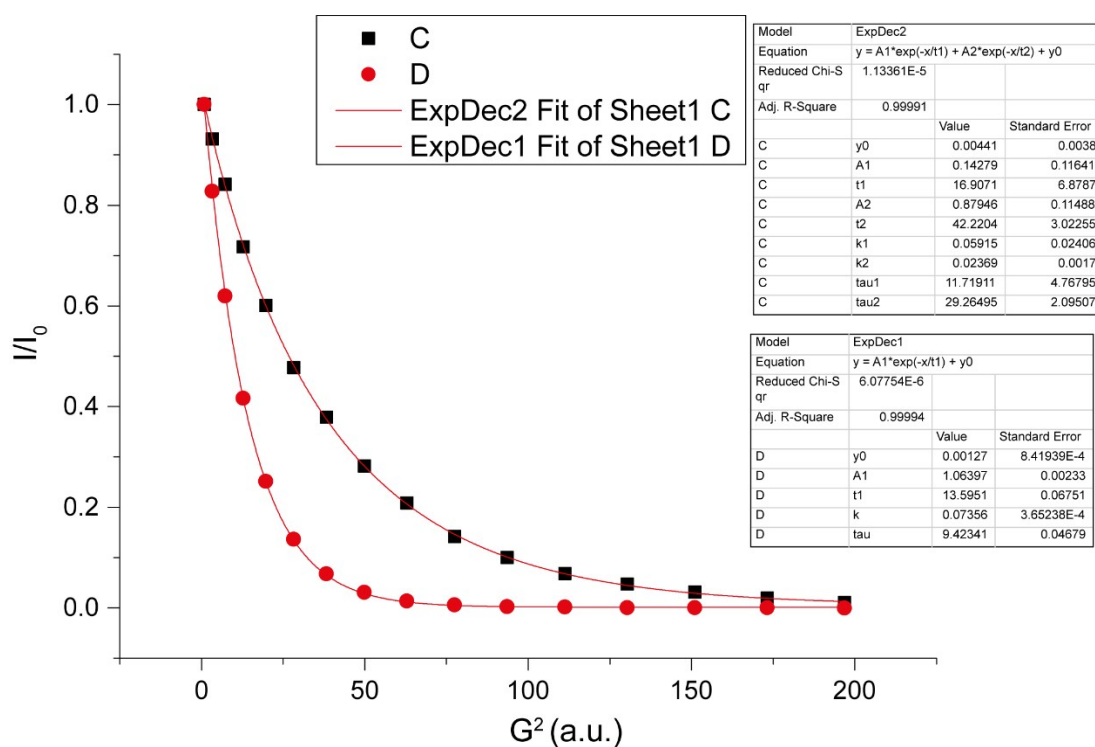


Figure S14 I/I_0 versus G^2 for **SeU** in acetone- d_6 (C = 32 mM) + **TBABzO** (74 mM) + **I1** (248 mM). B = **SeU**; C = solvent.

Table S3. Interaction energy (in kcal/mol), thermodynamic parameters and bond lengths (in Å) for the different adducts at M06-2X/aug-TZVP//BP86-D3/aug-TZVP level (gas phase). The parenthesis indicate that the interaction energy is referred to the pre-formed binary adduct rather than the three isolated components.

Adduct	ΔE	ΔH^a	ΔS^a	ΔG^a	C-Se	C-N
SeU	-	-	-	-	1.830	1.366
S-I_m	-6.2	-5.9	-10.3	4.4	1.855	1.351 ^b
B-S	-39.4	-41.3	-14.1	-25.5	1.881	1.344
B-I_m	-21.6	-21.4	-13.2	-8.1	-	-
B-S-I_m	-52.4	-54.3	-29.3	-25.0	1.903	1.334 ^b
(B-S)-I_m	-13.0	-12.9	-13.5	0.6	-	-
B-(S-I_m)	-46.2	-48.4	-19.0	-29.3	-	-

^a Calculated at 298 K; ^b Averaged between the two CN bonds

Table S4. Relevant donor – acceptor (D → A) natural bond orbital interactions and their second-order perturbation stabilization energies ($E^{(2)}$, in kcal/mol) for halogen- and hydrogen bonds in **S-I_m**, **B-S** and **B-S-I_m** adducts.

Adduct	D → A interaction	$E^{(2)}$
S-I_m	LP(1)Se2 → BD*(1)I9C10	1.37
	LP(2)Se2 → BD*(1)I9C10	25.04
	LP(1)N3 → BD*(1)C1Se2	82.66
	LP(1)N6 → BD*(1)C1Se2	72.89
	LP(3)I9 → BD*(1)N3H4	4.51
	BD(1)C1Se2 → RY(2)I9	0.23
B-S	LP(2)O21 → BD*(1)N6H8	13.49
	LP(2)O22 → BD*(1)N3H4	13.46
	LP(2)Se2 → LV(1)C1	349.15
	LP(1)N3 → LV(1)C1	161.99
	LP(1)N6 → LV(1)C1	163.03
B-S-I_m	LP(1)Se2 → BD*(1)I23C24	2.16
	LP(2)Se2 → BD*(1)I23C24	50.87
	LP(2)O21 → BD*(1)N6H8	17.81
	LP(2)O22 → BD*(1)N3H4	20.84
	LP(3)I23 → BD*(1)N6H7	1.80

LP(2)Se2 $\rightarrow$ LV(1)C1	178.79
LP(1)N6 $\rightarrow$ LV(1)C1	185.41
LP(1)N3 $\rightarrow$ LV(1)C1	177.85
