

Design considerations for chiral frustrated Lewis pairs: B/N FLPs derived from 3,5-bicyclic aryl piperidines

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Supporting Information

Table S5. Optimized Cartesian coordinates of 10-F	21
Table S6. Optimized Cartesian coordinates of 11	22
Table S7. Optimized Cartesian coordinates of 11-F	23
Table S8. Optimized Cartesian coordinates of BPh_3	24
Table S9. Optimized Cartesian coordinates of $BPh_3\text{-F}$	25
Figure S34. Depictions of HOMO (left) and LUMO (right) of 10	26
Figure S35. Depictions of HOMO (left) and LUMO (right) of 11	26
References	27

General Considerations

All manipulations were carried out under dry, O₂-free N₂ using an MBraun glovebox and a Schlenk vacuum-line. Pentane and dichloromethane were collected from a Grubbs-type column system manufactured by Innovative Technology and into thick-walled glass Schlenk bombs with Young-type Teflon valve stopcocks. Chloroform-d was obtained from Cambridge Isotope Laboratories, dried over CaH₂, and vacuum-transferred into Young bombs, Toluene-d₈ and benzene-d₆ was obtained from Sigma-Aldrich, dried over Na/benzophenone and vacuum-transferred into Young bombs. All solvents were degassed after purification and stored over 4 Å molecular sieves. B(C₆F₅)₃ was purchased from Boulder Scientific and used without further purification for the synthesis of bis(perfluorophenyl)borane. (E)-N-tert-butyl-1-phenylmethanimine and triethylsilane were purchased from Sigma-Aldrich, and dried and stored over 4 Å molecular sieves. Commercial reagents were purchased from Sigma-Aldrich, TCI Chemicals, Strem Chemicals or Alfa Aesar, and used without further purification unless indicated otherwise. Hydrogen gas (Grade 5.0) was obtained from Linde and purified through a Matheson Nanochem WeldAssure™ gas purifier column prior to use. Deuterium hydride (extent of labeling: 96 mol% HD, 98 atom % D) and carbon ¹³C dioxide (99 atom % ¹³C, <3 atom % ¹⁸O) were purchased from Sigma Aldrich. (1R,3s,5S)-7,8-difluoro-3-isopropyl-2,3,4,5-tetrahydro-1H-1,5-methanobenzo[d]azepin-3-iium chloride was received from Pfizer and used without further purification. (E)-N-(1-phenylethylidene)aniline,¹ tris(2,6-difluorophenyl)borane,² (bis(perfluorophenyl)borane,³ were prepared according to literature methods.

NMR spectra were recorded on a Bruker Avance 400 MHz spectrometer or a Varian Mercury Plus 400 MHz spectrometer at 25 °C. Chemical shifts are given relative to SiMe₄ and referenced to the residual solvent signal (¹H, ¹³C) or relative to an external standard (¹¹B: 15% (Et₂O)BF₃, ³¹P: 85% H₃PO₄, ¹⁹F: CFCl₃). Chemical shifts are reported in ppm and coupling constants as scalar values in Hz. Mass spectrometry was carried out using an AB/Sciex QStar mass spectrometer with an ESI source or on a Hewlett-Packard GC/MS 6890 N that works with the EI technique (70 eV). Elemental analyses (C, H, N) were performed in-house with a Perkin Elmer 2400 Series II CHNS Analyzer

Spectra

Figure S1. ^1H NMR spectrum of **1**

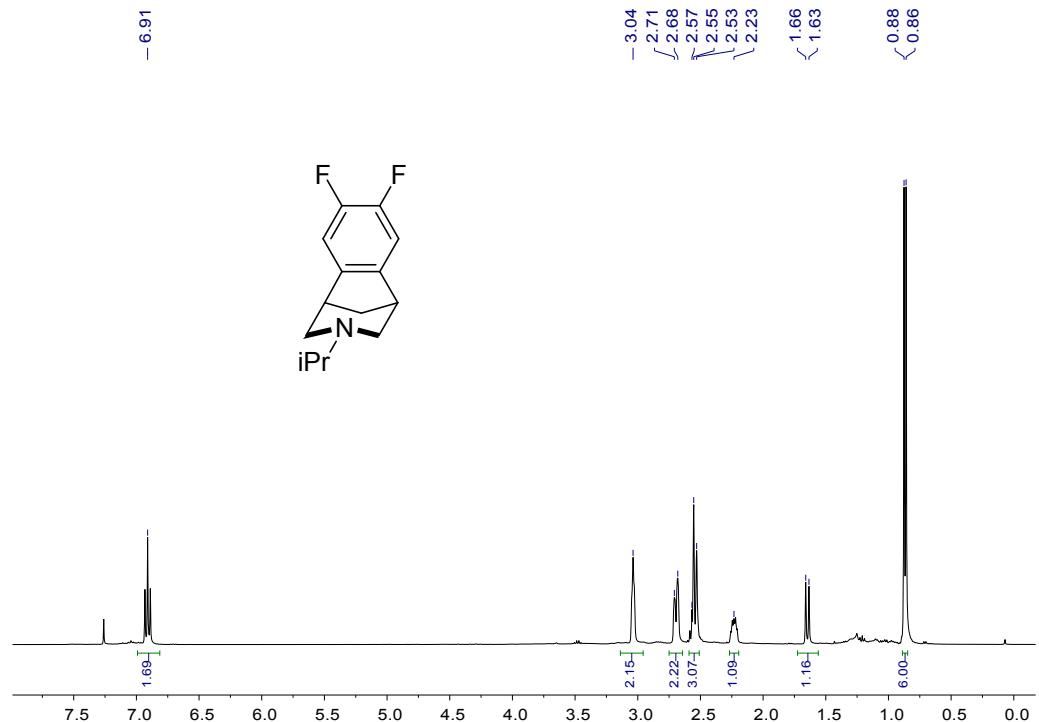


Figure S2. $^{13}\text{C}\{\text{H}\}$ NMR spectrum of **1**

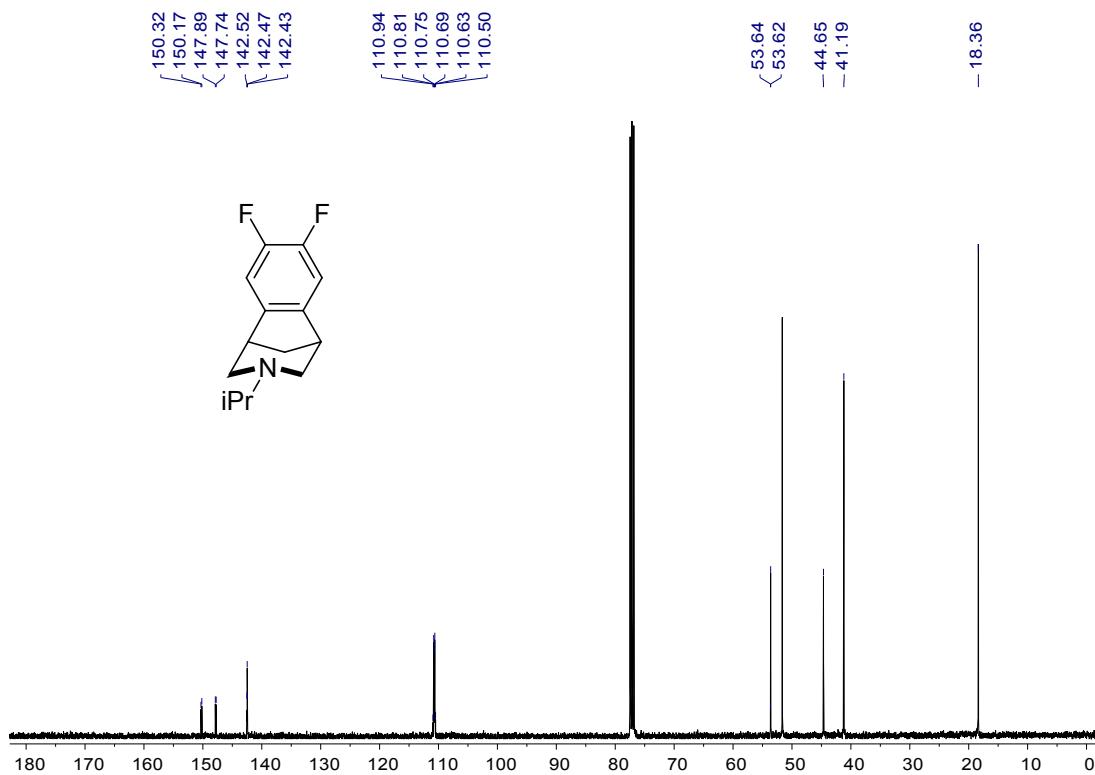


Figure S3. ^{19}F NMR spectrum of **1**

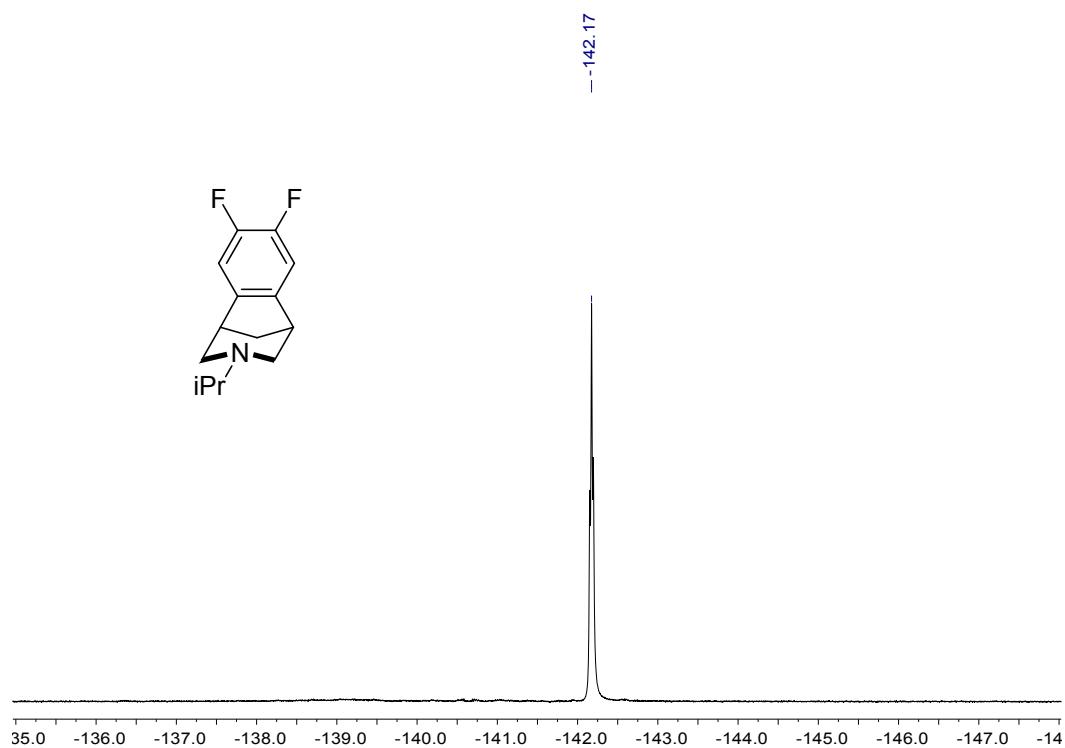


Figure S4. ^1H NMR spectrum of **2**

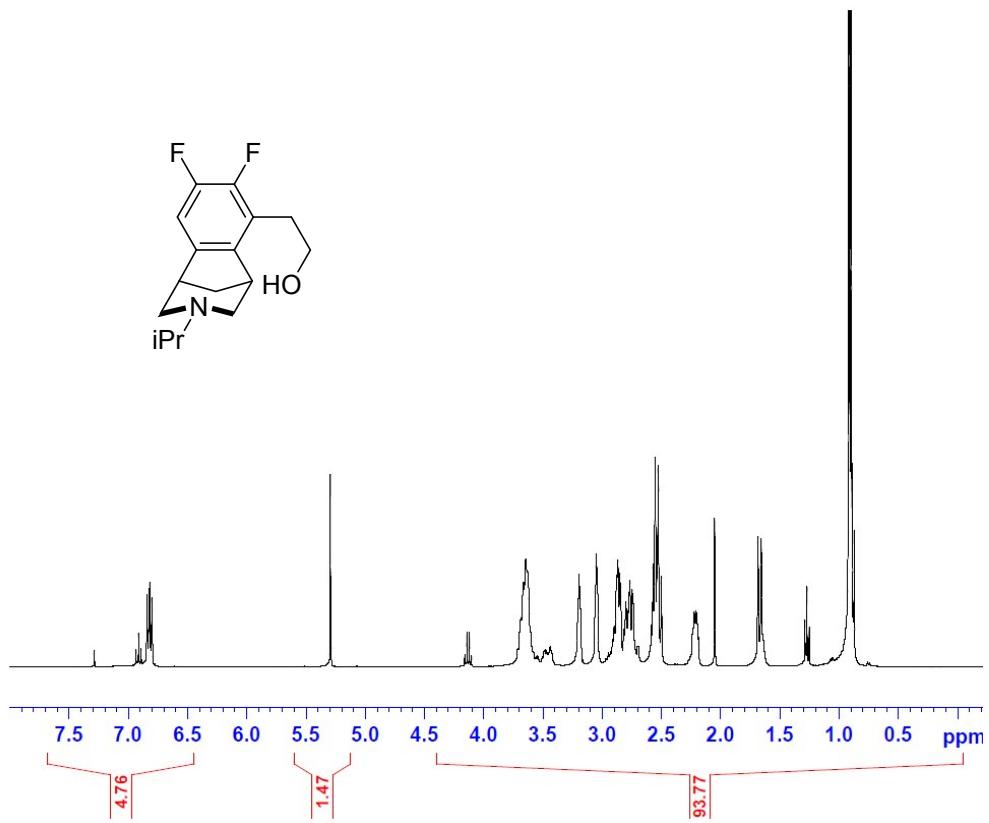


Figure S5. ^{13}C NMR spectrum of **2**

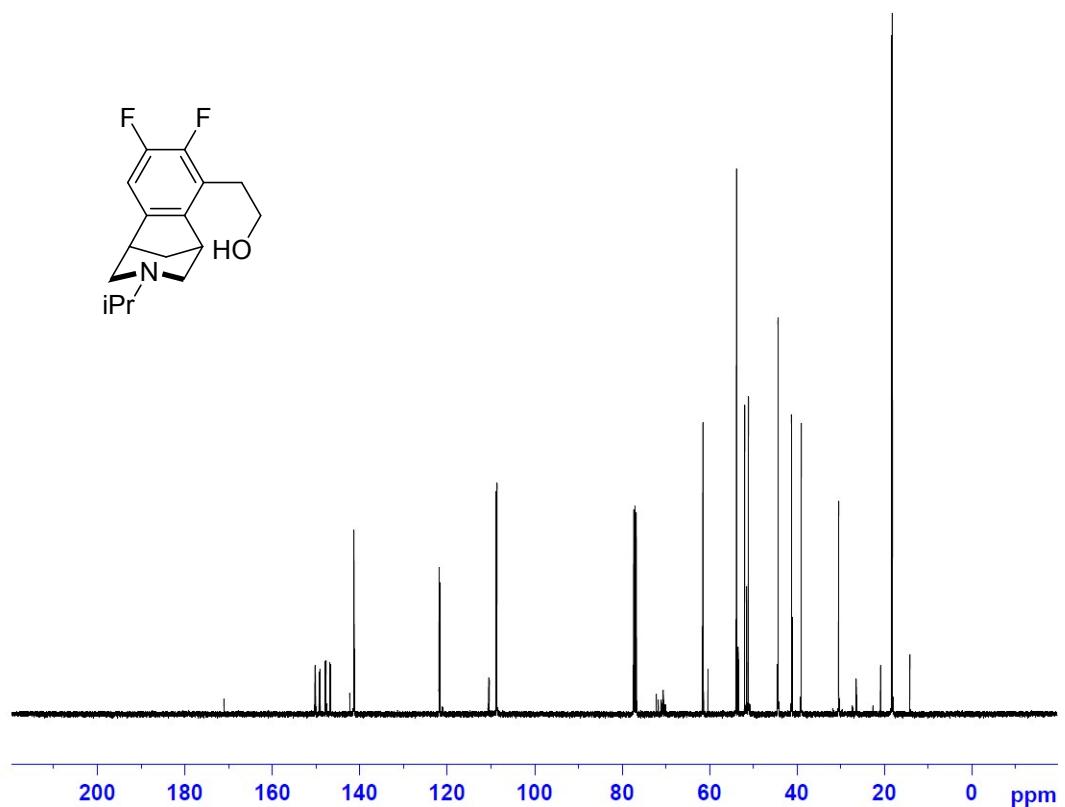


Figure S6. ^{19}F NMR spectrum of **2**

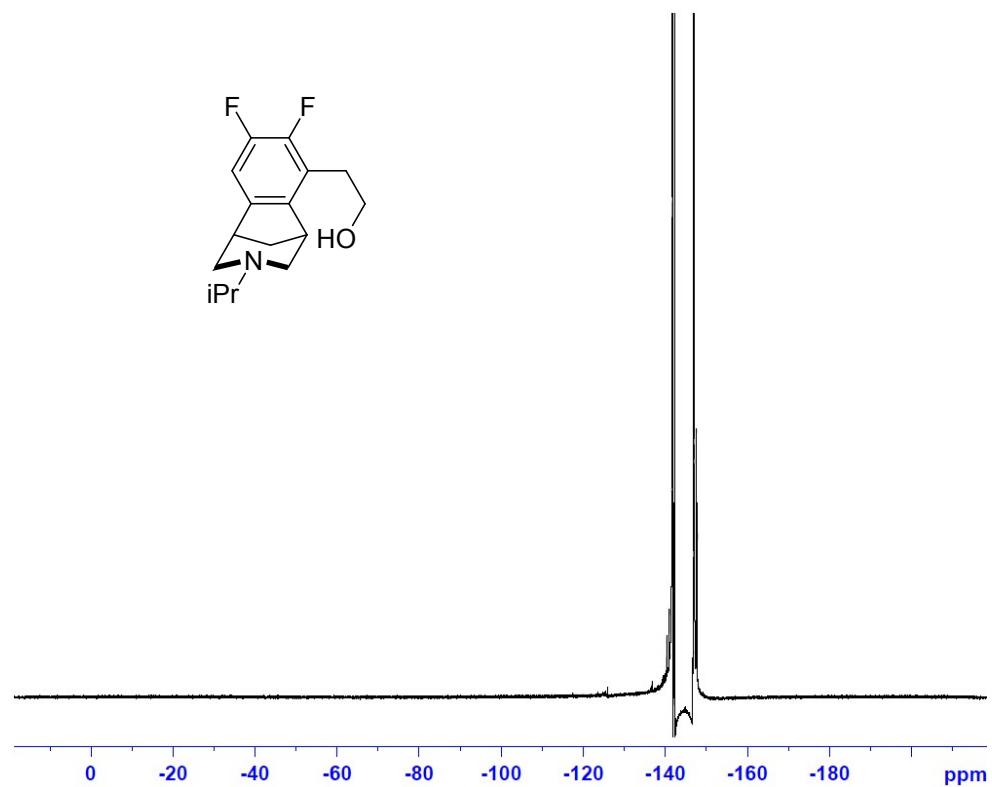


Figure S7. ^1H NMR spectrum of **3**

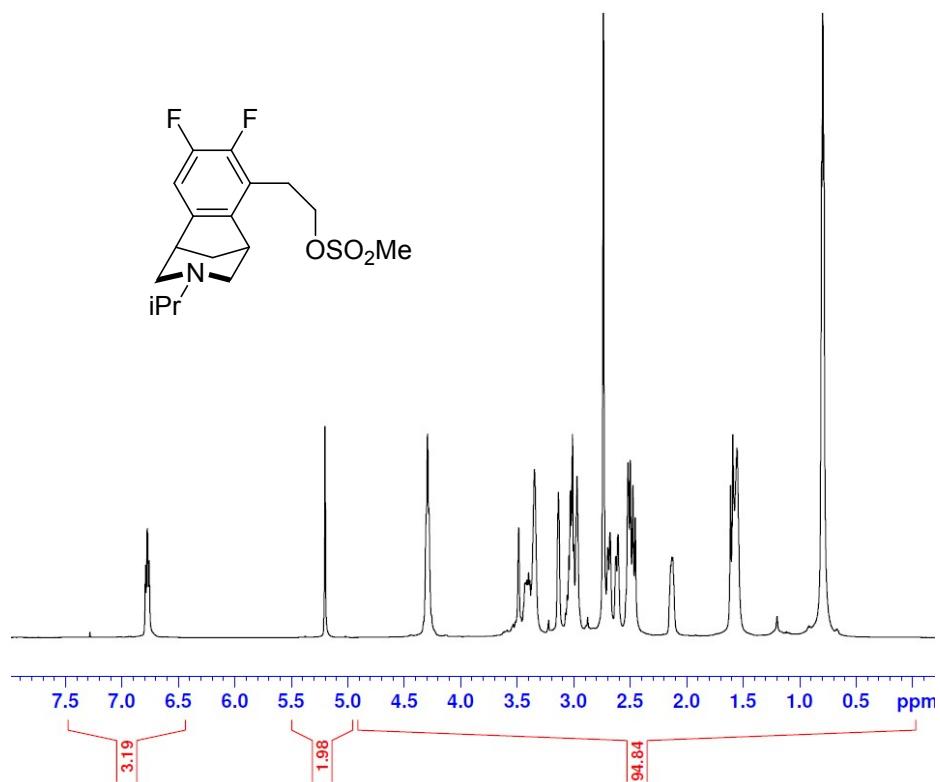


Figure S8. ^{13}C NMR spectrum of **3**

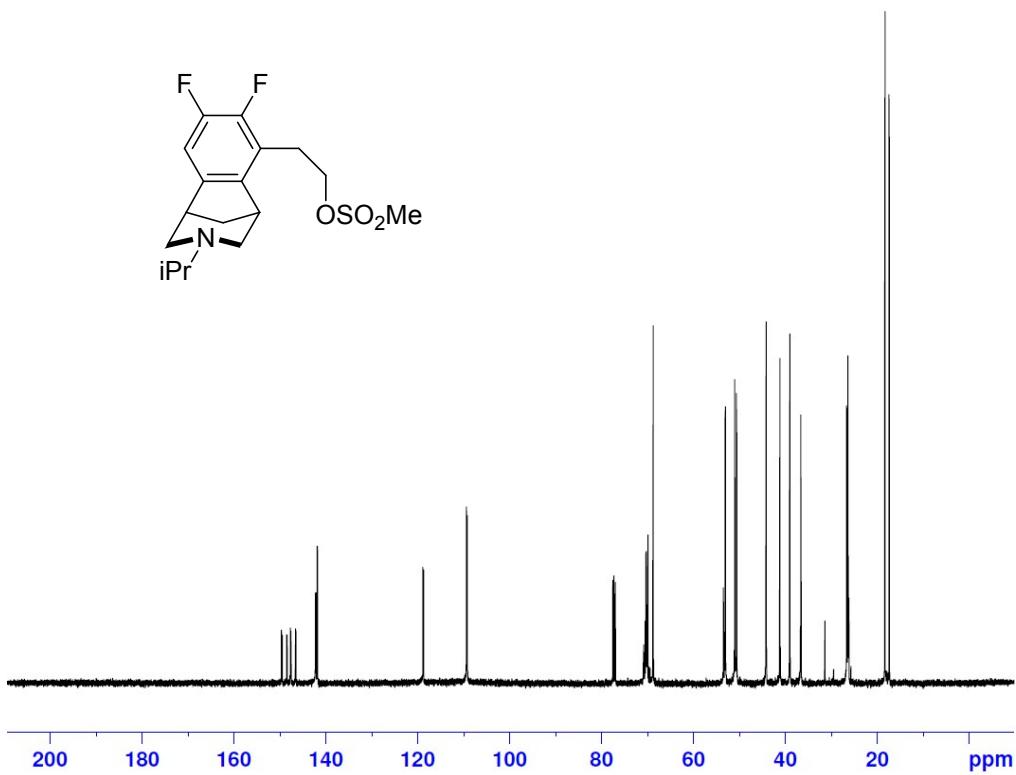


Figure S9. ^1H NMR spectrum of **4**

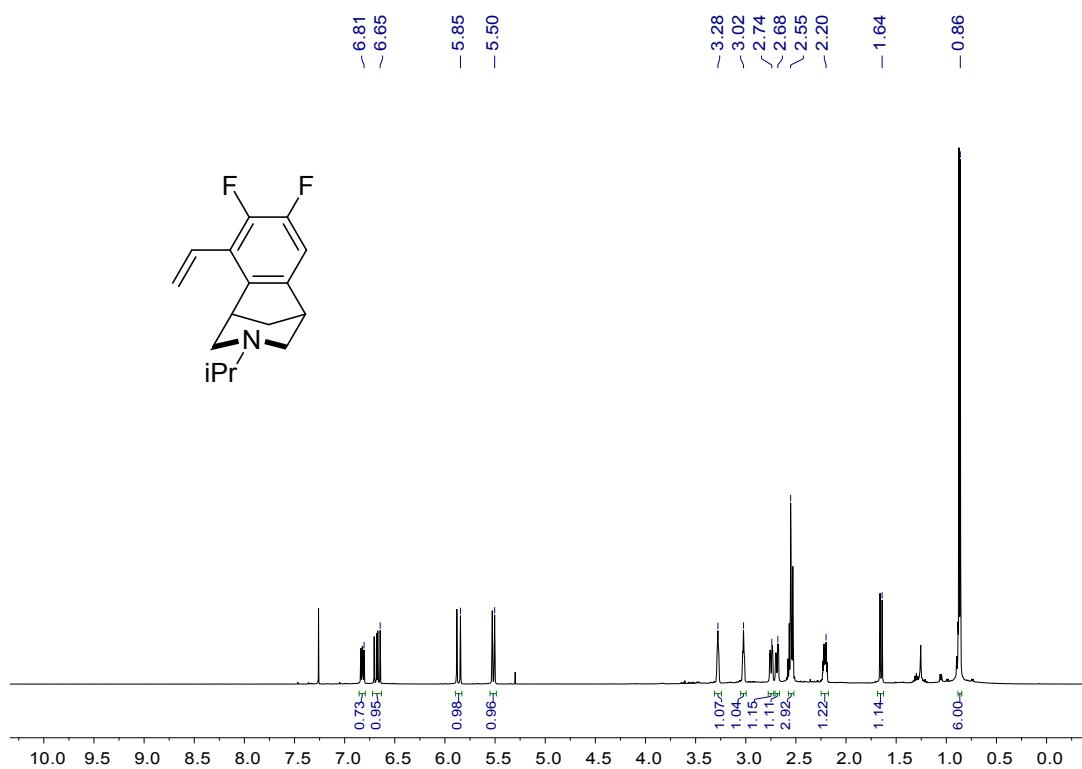


Figure S10. $^{13}\text{C}\{\text{H}\}$ NMR spectrum of **4**

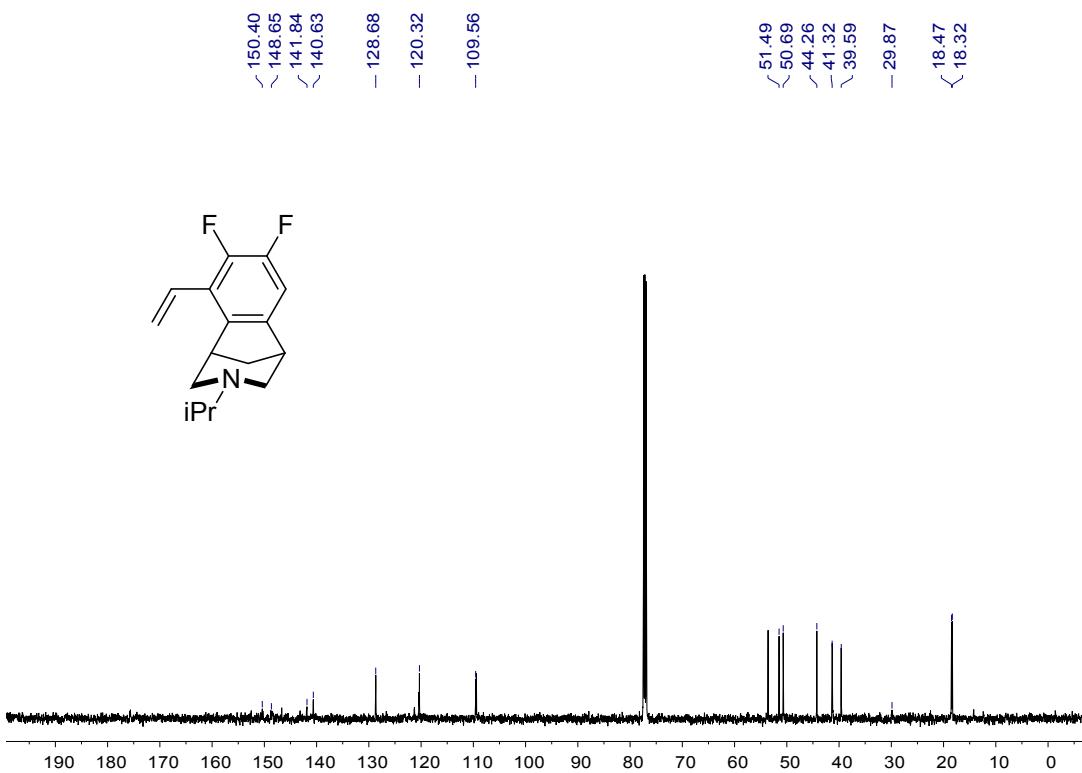


Figure S11. ^{19}F NMR spectrum of **4**

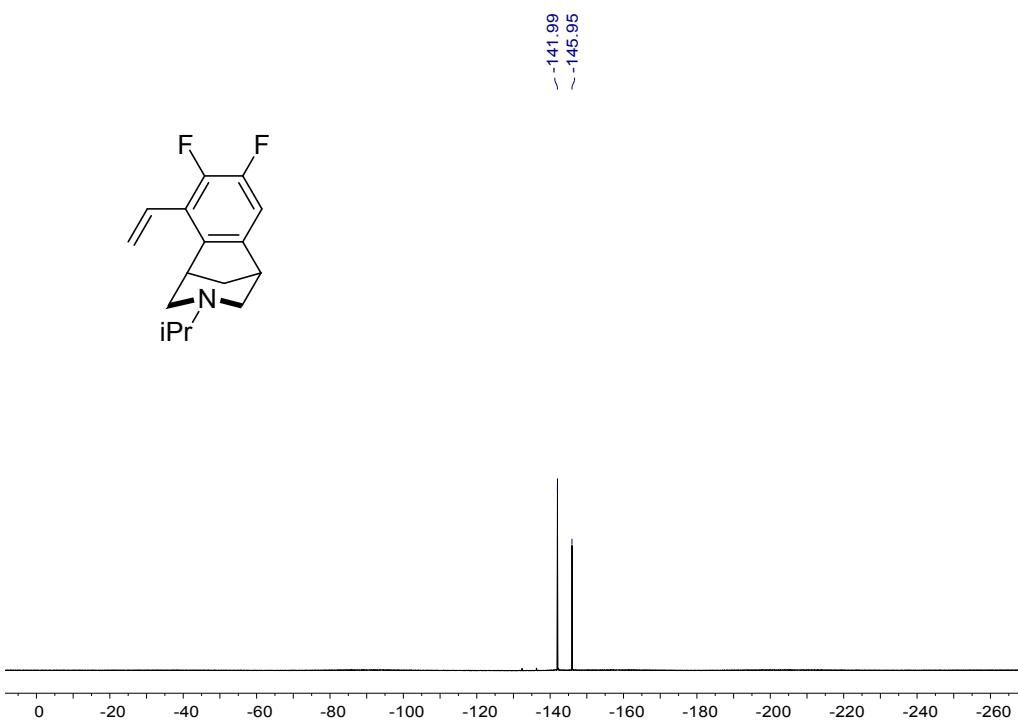


Figure S12. ^1H NMR spectrum of **6**

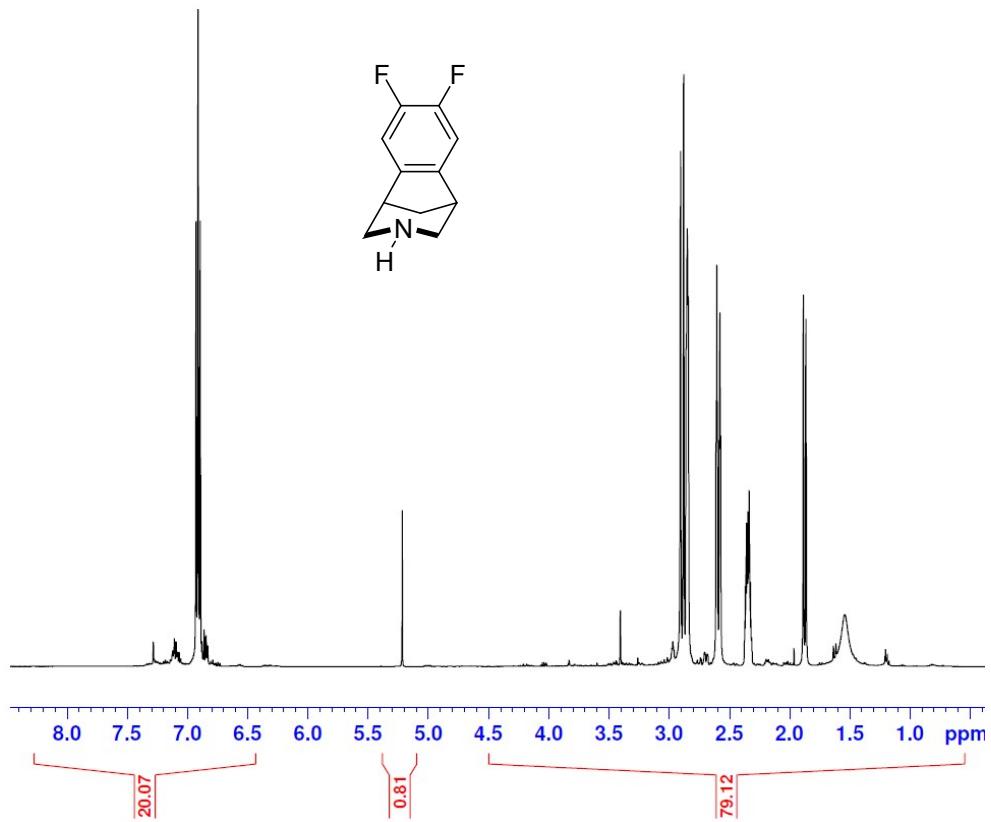


Figure S13. ^1H NMR spectrum of **7**

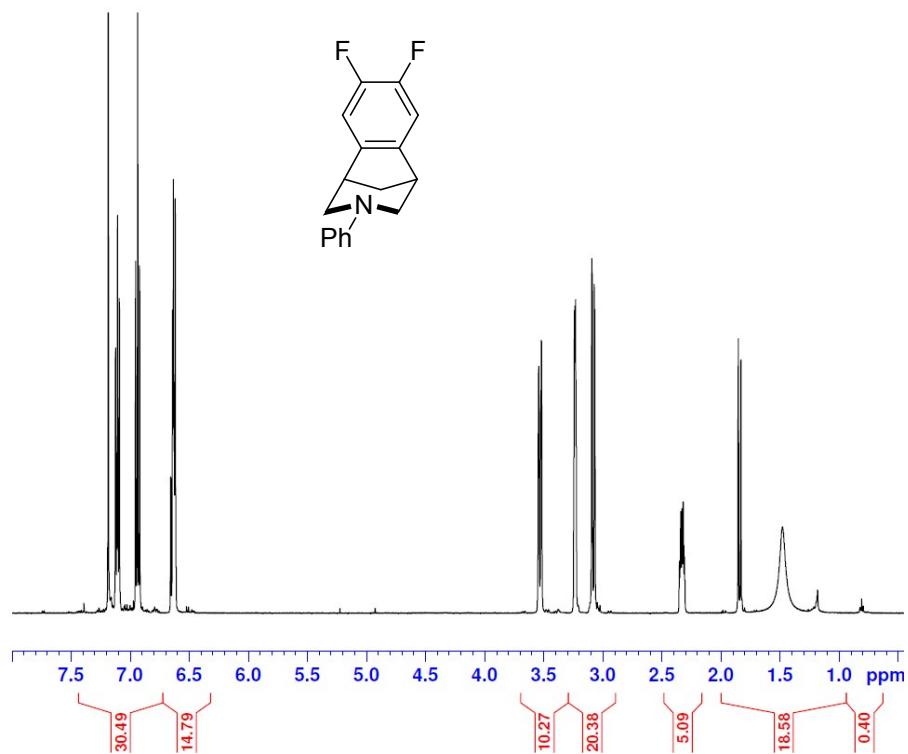


Figure S14. ^{13}C NMR spectrum of **7**

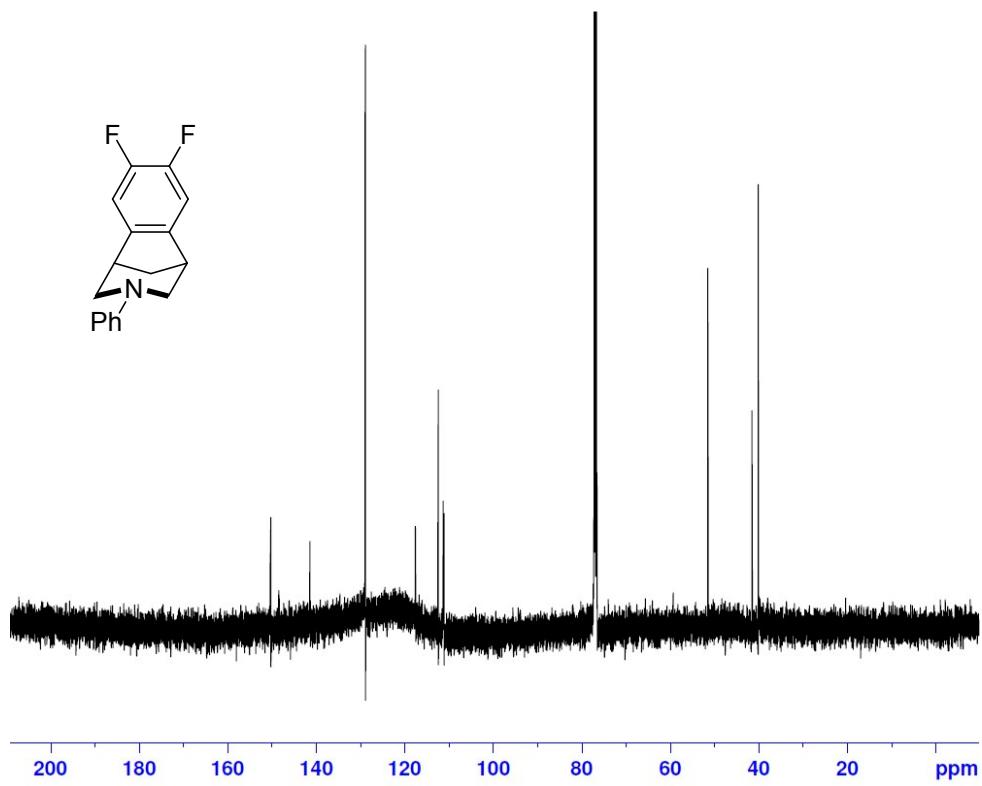


Figure S15. ^1H NMR spectrum of **8**

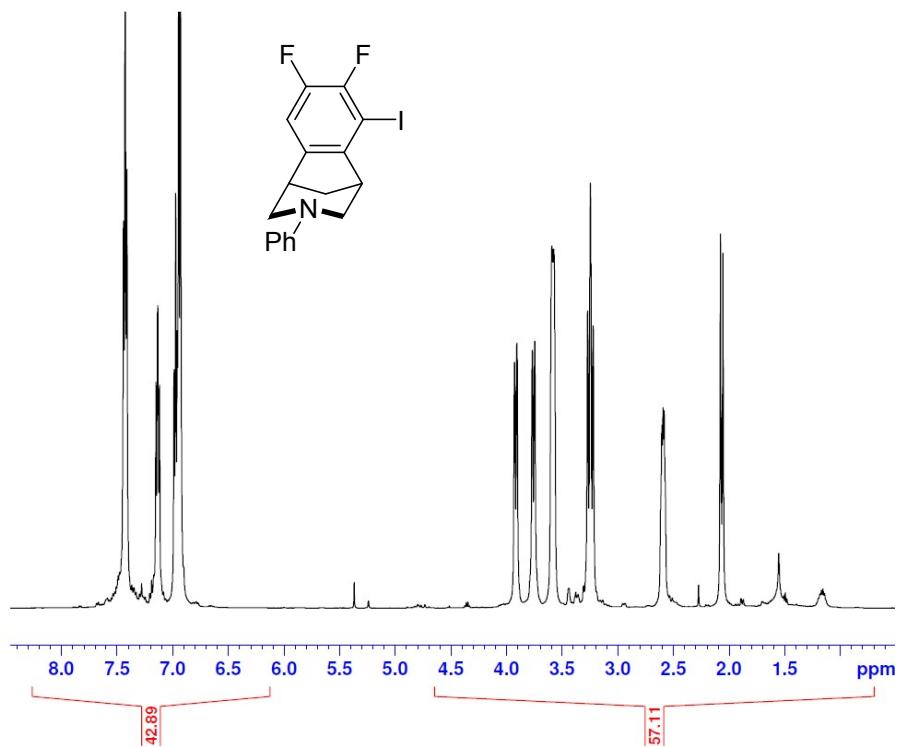


Figure S16. ^{13}C NMR spectrum of **8**

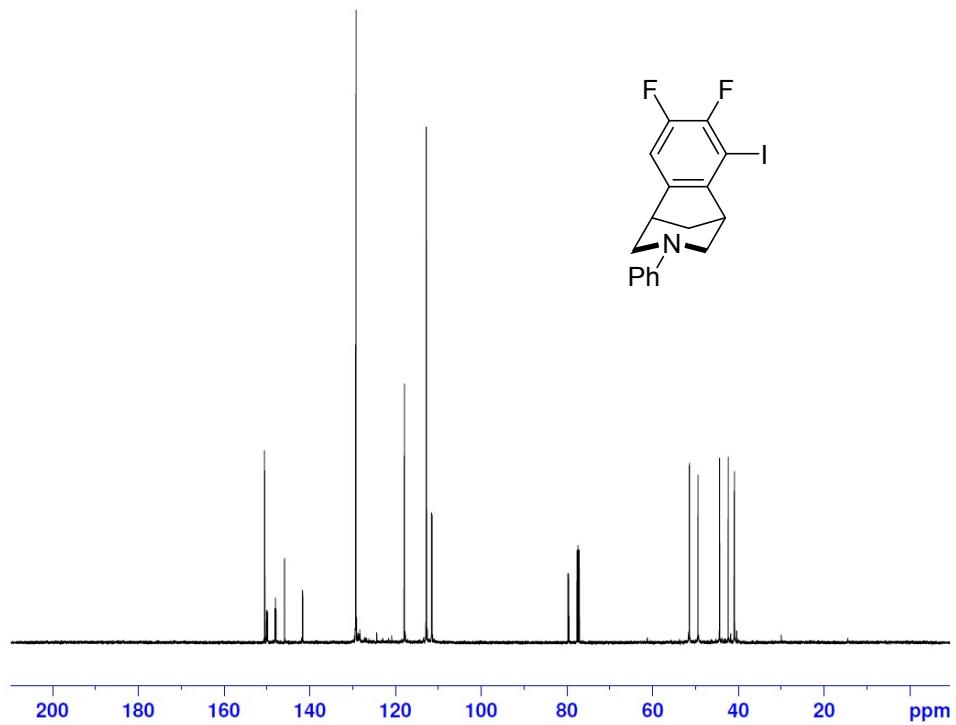


Figure S17. ^1H NMR spectrum of **9**

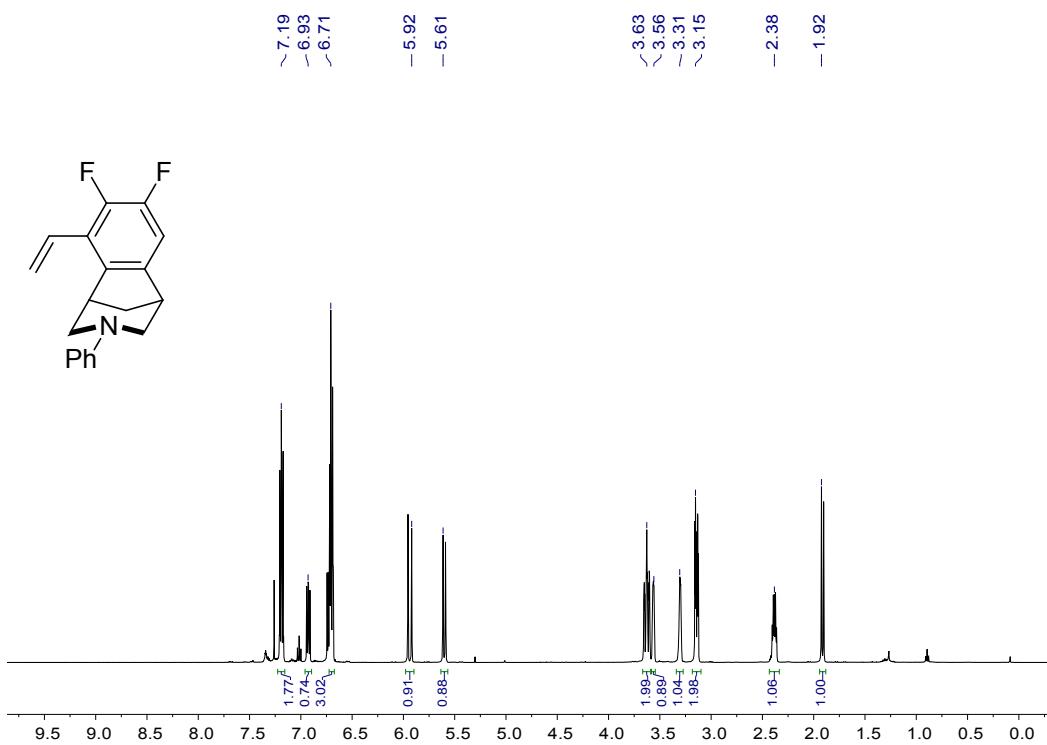


Figure S18. $^{13}\text{C}\{\text{H}\}$ NMR spectrum of **9**

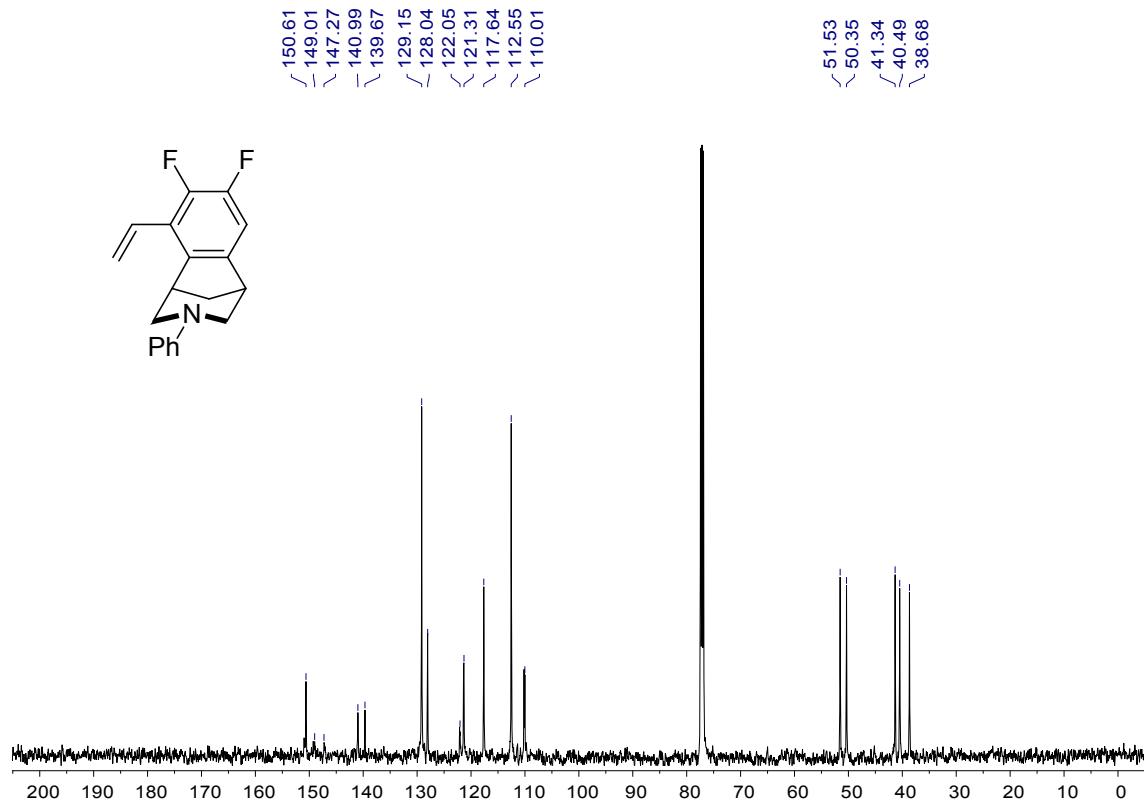


Figure S19. ^{19}F NMR spectrum of **9**

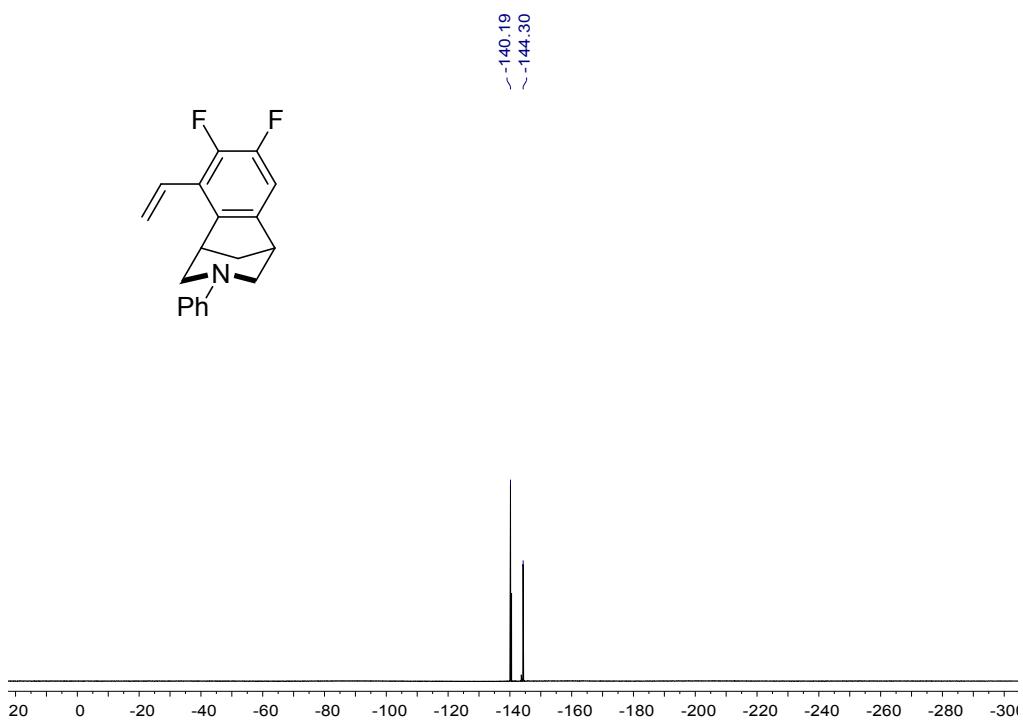


Figure S20. ^1H NMR spectrum of **10**

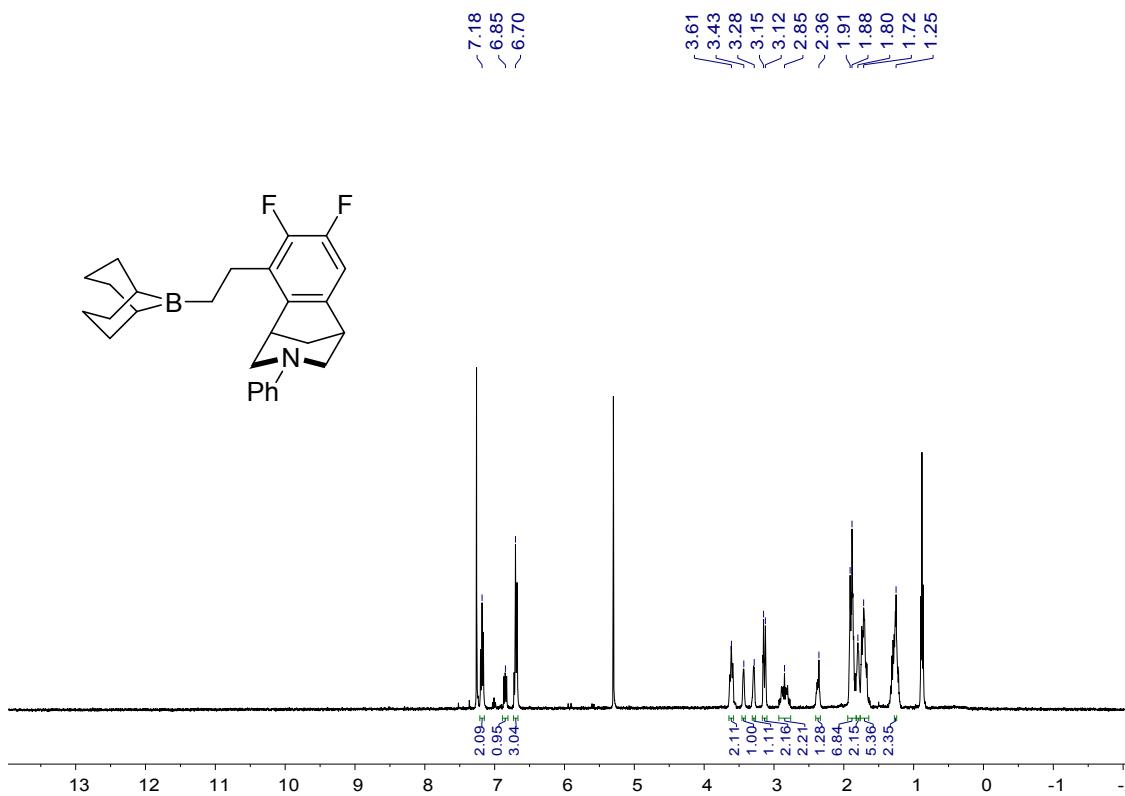


Figure S21. $^{13}\text{C}\{\text{H}\}$ NMR spectrum of **10**

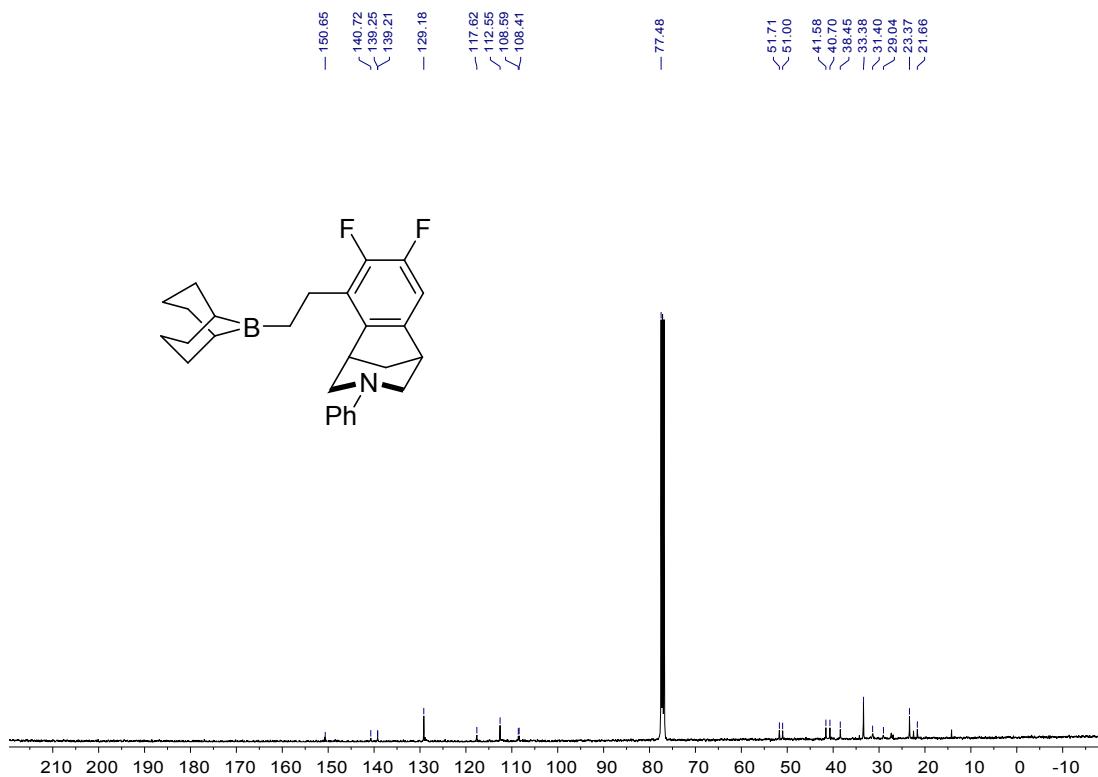


Figure S22. ^{11}B NMR spectrum of **10**

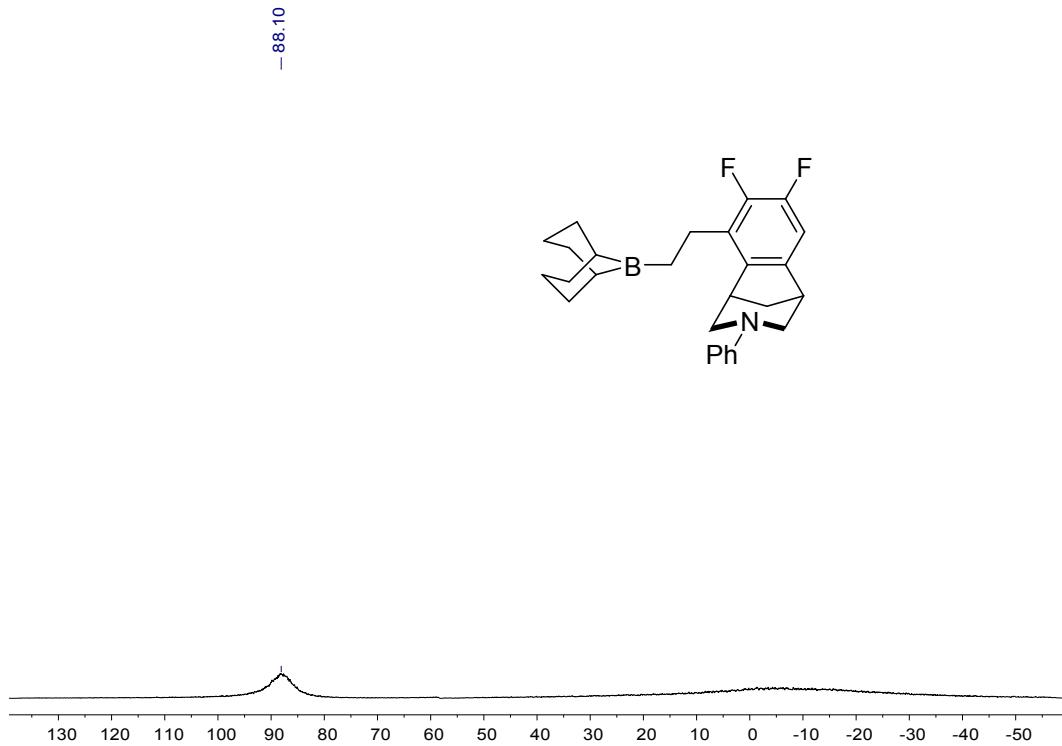


Figure S23. ^{19}F NMR spectrum of **10**



Figure S24. ^1H NMR spectrum of **11**

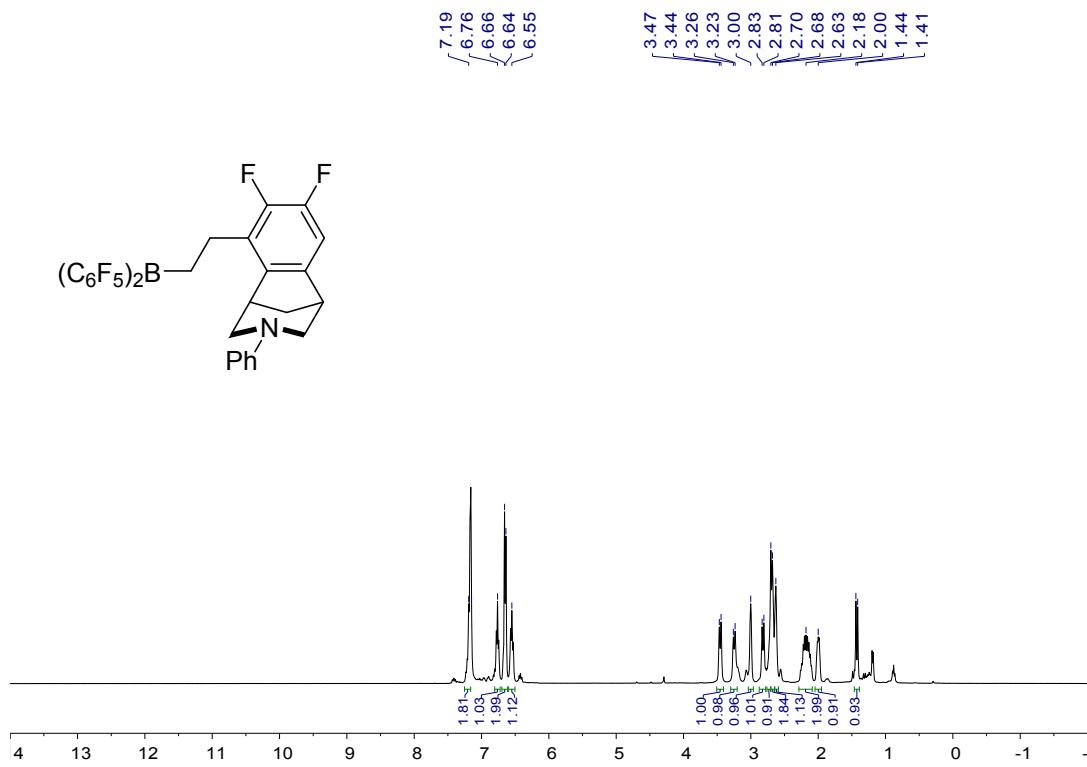


Figure S25. $^{13}\text{C}\{\text{H}\}$ NMR spectrum of **11**

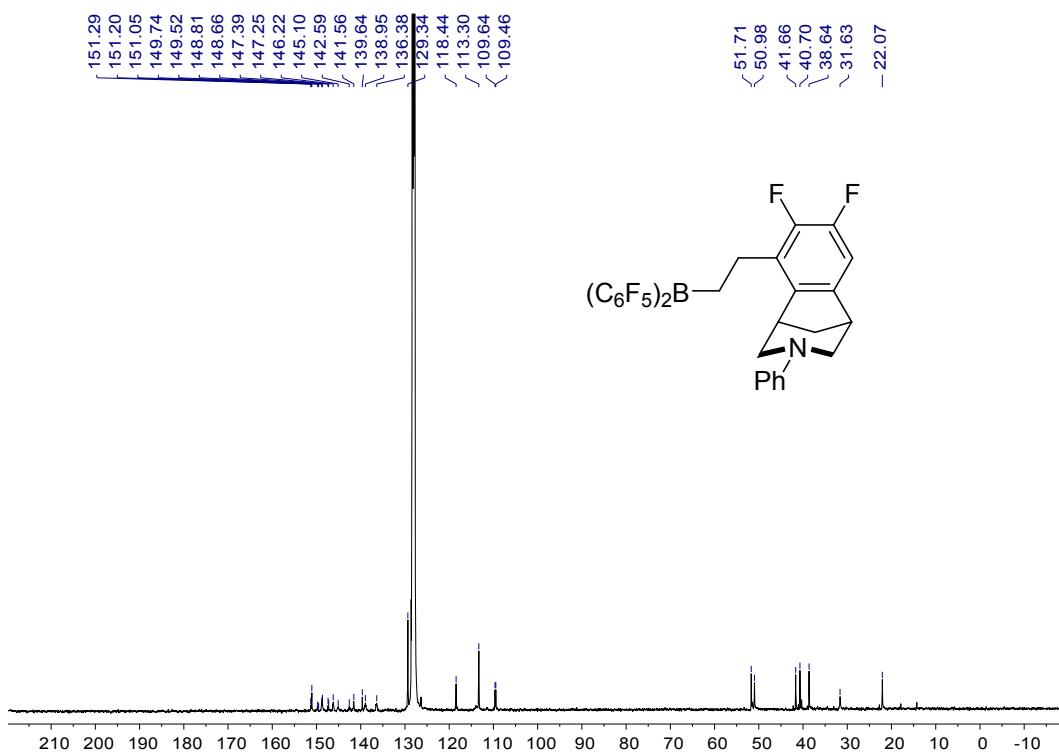


Figure S26. ^{11}B NMR spectrum of **11**

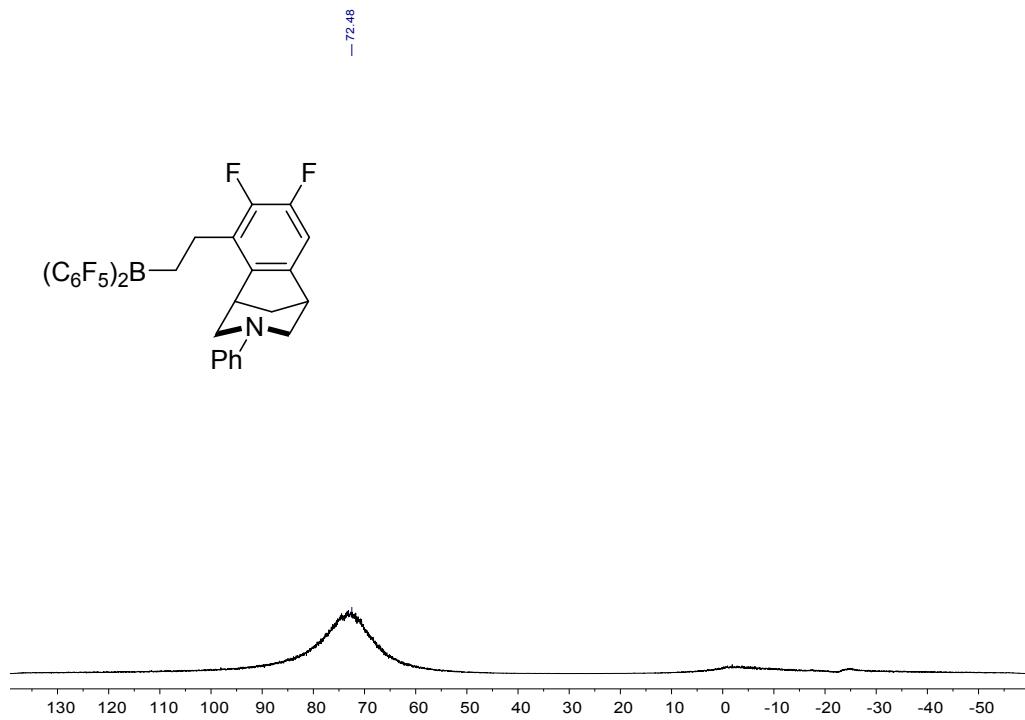


Figure S27. ^{19}F NMR spectrum of **11**

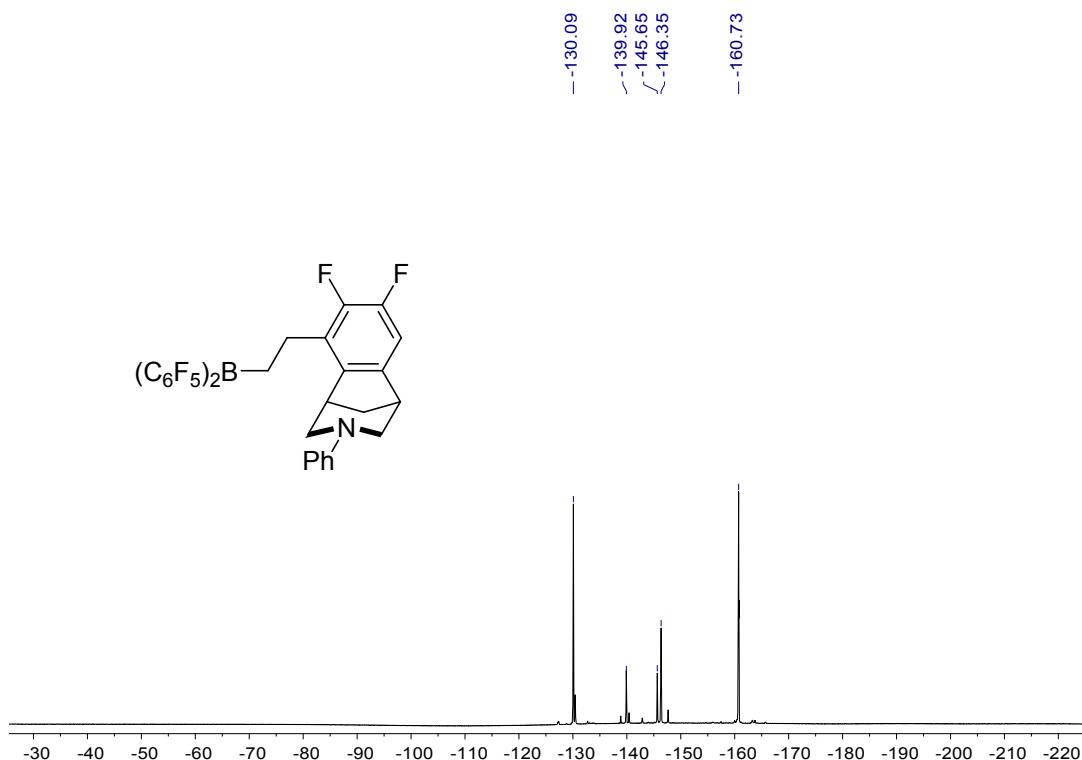


Figure S28. ^1H NMR spectrum of **12**

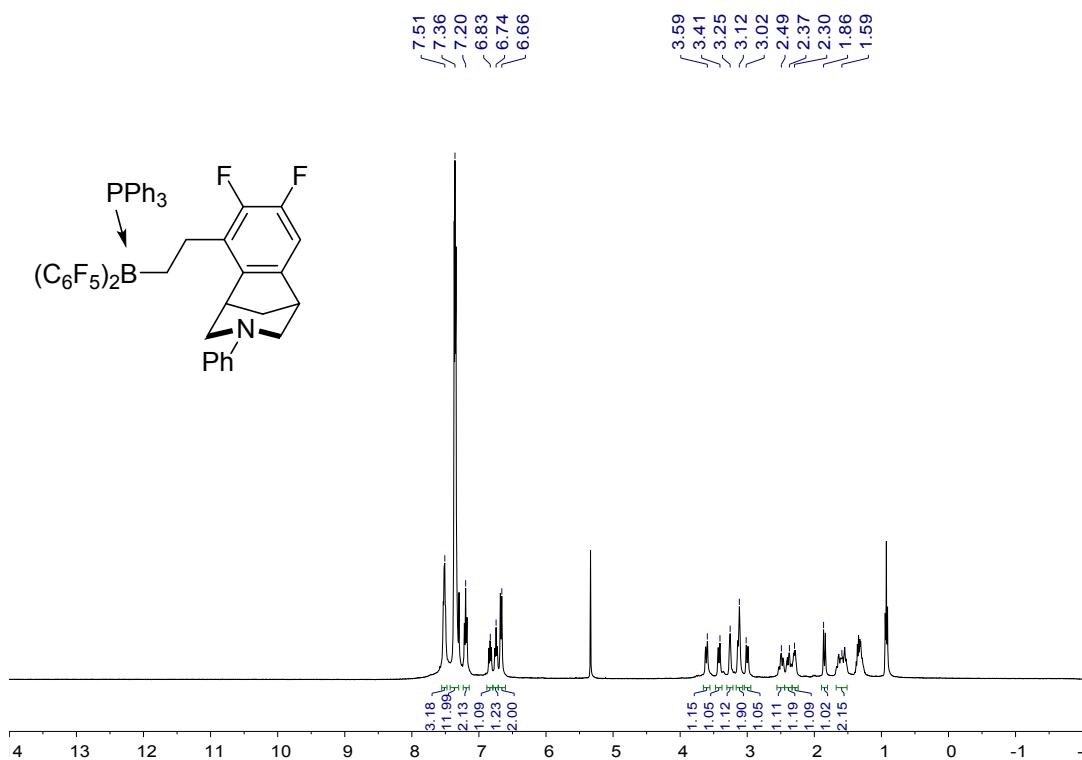


Figure S29. $^{13}\text{C}\{\text{H}\}$ NMR spectrum of **12**

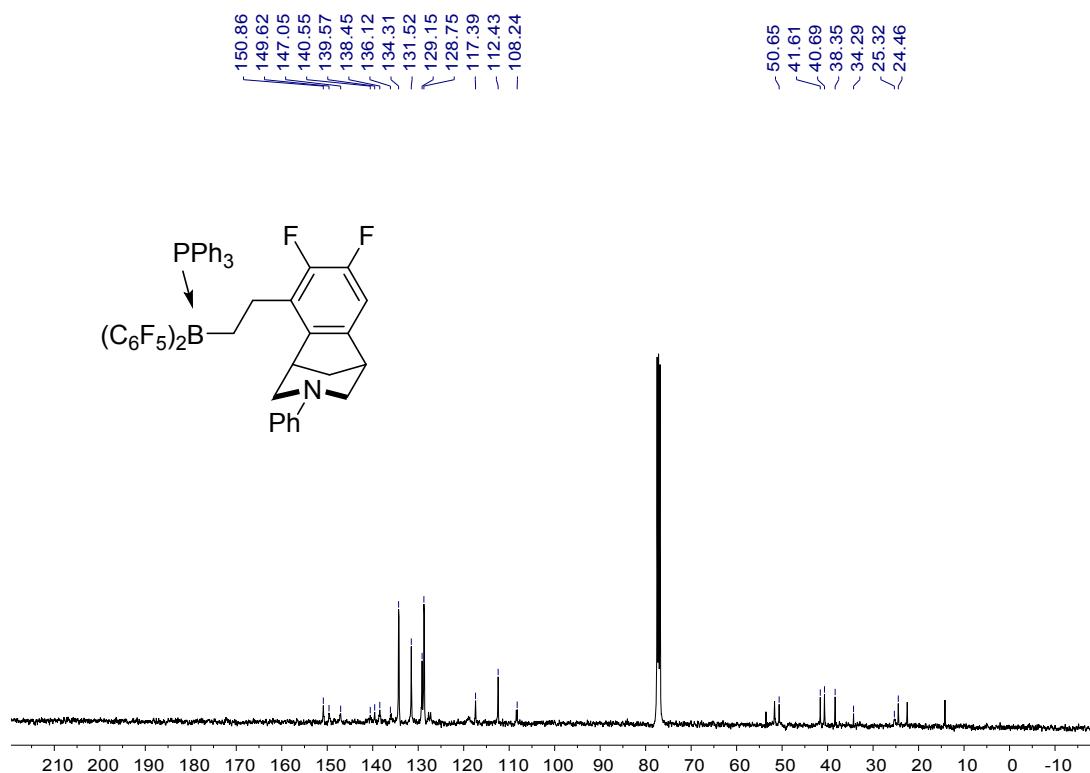


Figure S30. ^{31}P NMR spectrum of **12**

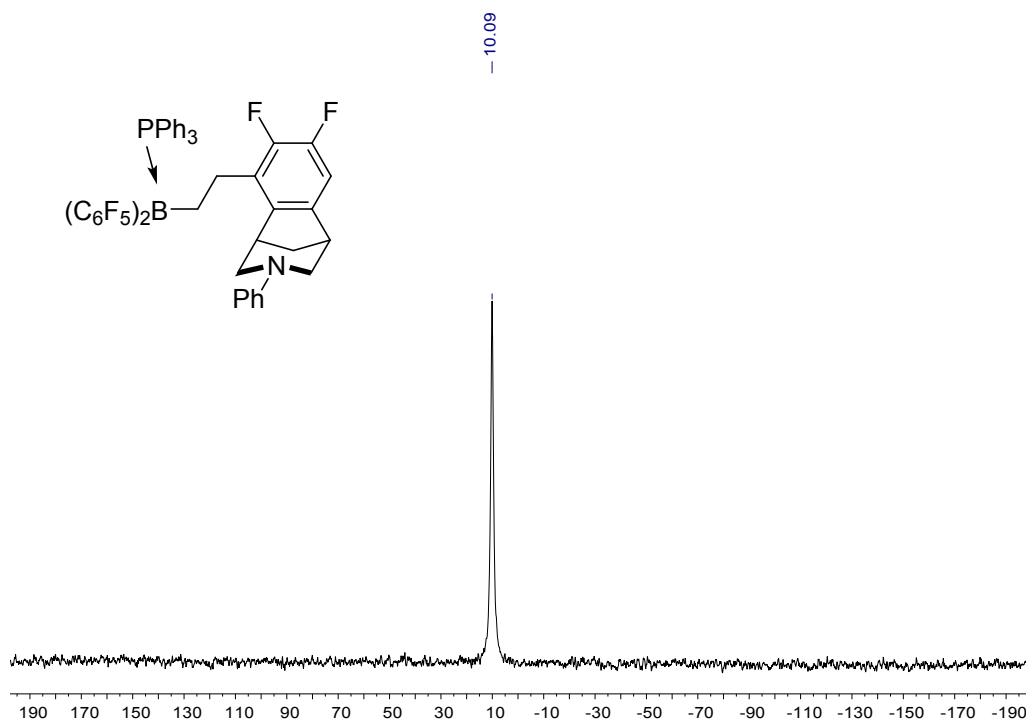


Figure S31. ^{11}B NMR spectrum of **12**

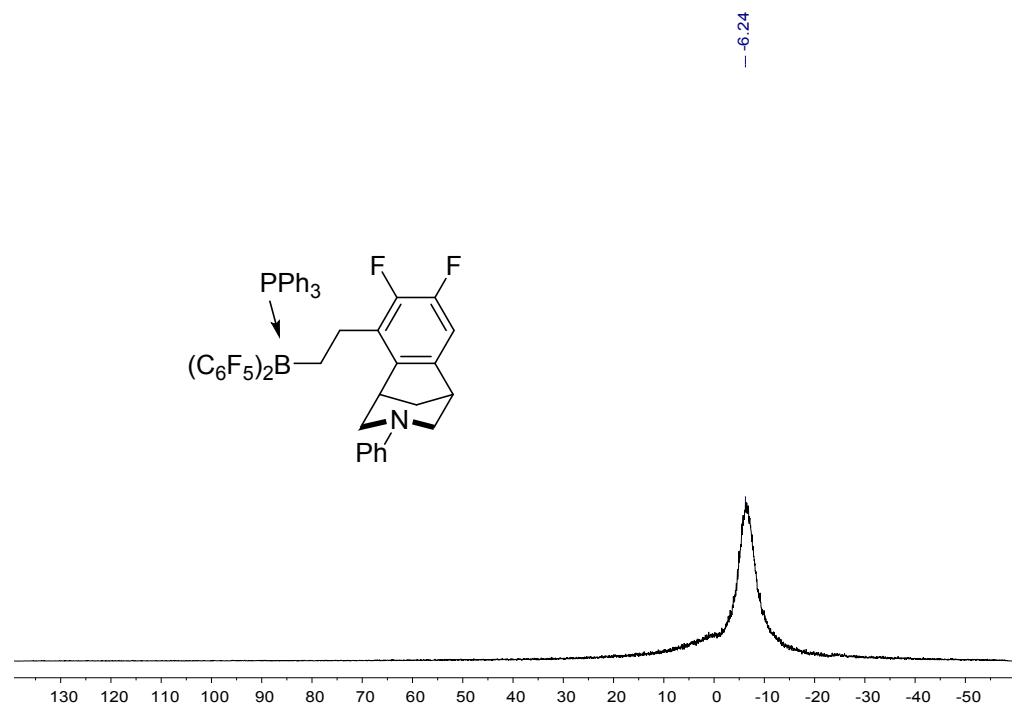
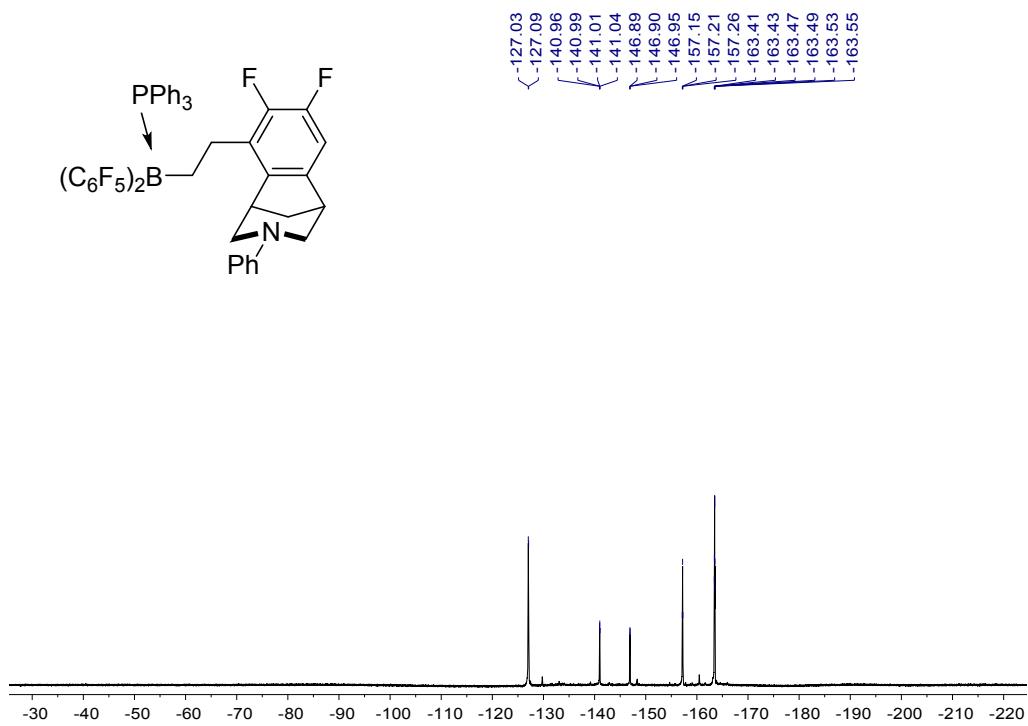


Figure S32. ^{19}F NMR spectrum of **12**



BPh ₃	-1885741.0	-2147911.5	328.1
B(C ₆ H ₃ F ₂) ₃	-3447380.9	-3709585.2	362.3
11	-6473827.8	-6736111.6	441.4
B(C ₆ F ₅) ₃	-5789595.6	-6051890.2	452.6

Table S3. GEI data for boranes

Compound	Energy of HOMO (eV)	Energy of LUMO (eV)	GEI (eV)
10	-5.445	-0.713	1.00
BPh ₃	-6.888	-2.256	2.04
B(C ₆ H ₃ F ₂) ₃	-6.993	-2.547	2.56
11	-5.538	-2.938	3.45
B(C ₆ F ₅) ₃	-7.651	-3.523	3.78

Figure S34. Depictions of HOMO (left) and LUMO (right) of **10**

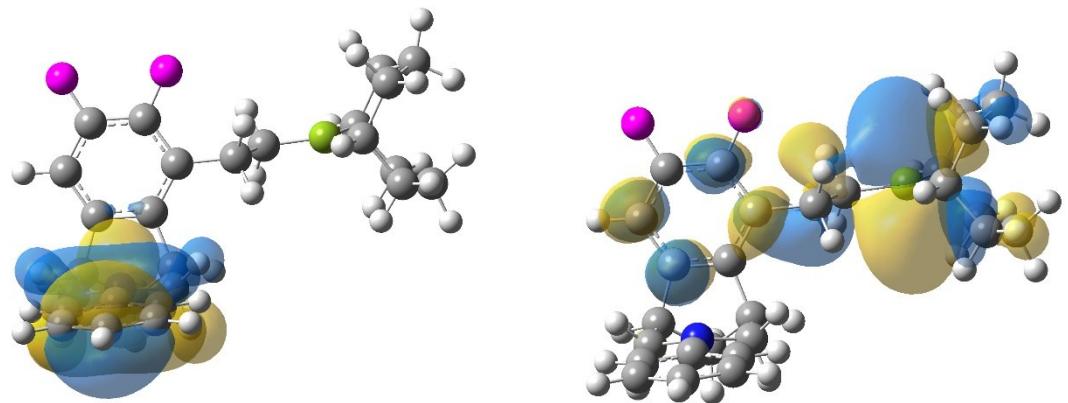
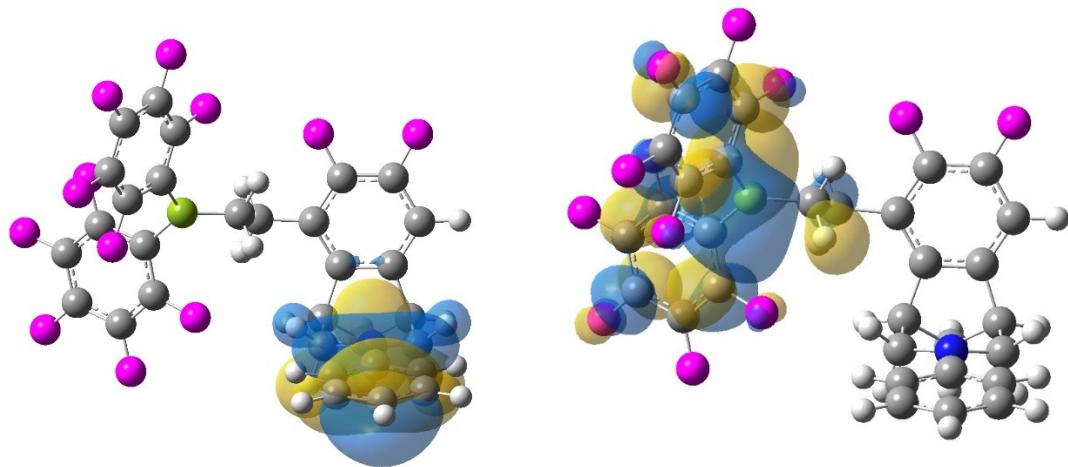


Figure S35. Depictions of HOMO (left) and LUMO (right) of **11**



References

1. Hansen, M. C.; Buchwald, S. L., A Method for the Asymmetric Hydrosilylation of N-Aryl Imines. *Organic Letters* **2000**, 2 (5), 713-715.
2. Nicasio, J. A.; Steinberg, S.; Ines, B.; Alcarazo, M., Tuning the Lewis acidity of boranes in frustrated Lewis pair chemistry: implications for the hydrogenation of electron-poor alkenes. *Chemistry – A European Journal* **2013**, 19 (33), 11016-20.
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4. Beckett, M. A.; Strickland, G. C.; Holland, J. R.; Sukumar Varma, K., A convenient n.m.r. method for the measurement of Lewis acidity at boron centres: correlation of reaction rates of Lewis acid initiated epoxide polymerizations with Lewis acidity. *Polymer* **1996**, 37 (20), 4629-4631.