

Supplementary Information for

**Methoxy-substituted tetrakisquinoline analogs of EGTA and BAPTA for  
fluorescent detection of Cd<sup>2+</sup>**

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**Table S1.** Crystallographic Data for BAPTO

BAPTO	
Formula	C <sub>54</sub> H <sub>44</sub> N <sub>6</sub> O <sub>2</sub>
FW	808.98
Crystal system	triclinic
Space group	P-1
<i>a</i> , Å	10.7273(18)
<i>b</i> , Å	11.8900(18)
<i>c</i> , Å	17.402(3)
$\alpha$ , deg	77.787(5)
$\beta$ , deg	87.943(6)
$\gamma$ , deg	73.552(4)
<i>V</i> , Å <sup>3</sup>	2079.9(6)
<i>Z</i>	2
<i>D</i> <sub>calc</sub> , g cm <sup>-3</sup>	1.292
$\mu$ , mm <sup>-1</sup>	0.0799
2 $\theta$ <sub>max</sub> , deg	55
temp, K	153
no. reflns collected	16366
no. reflns used	9080
no. of params	735
<i>R</i> <sub>int</sub>	0.0176
Final <i>R</i> 1 ( <i>I</i> > 2 $\sigma$ ( <i>I</i> )) <sup>a</sup>	0.0470
<i>wR</i> 2 (all data) <sup>b</sup>	0.1160
GOF	1.115

<sup>a</sup> $R1 = \sum | |F_o| - |F_c| | / \sum |F_o|$ .    <sup>b</sup> $wR2 = [\sum w[(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2]]^{1/2}$ .

**Table S2.** Crystallographic Data for [Cd(EGTQ)](ClO<sub>4</sub>)<sub>2</sub>·CH<sub>3</sub>OH and [Cd<sub>2</sub>(μ-OH)(EGTQ)(H<sub>2</sub>O)<sub>2</sub>](ClO<sub>4</sub>)<sub>3</sub>·2CH<sub>3</sub>CN

	[Cd(EGTQ)](ClO <sub>4</sub> ) <sub>2</sub> ·CH <sub>3</sub> OH	[Cd <sub>2</sub> (μ-OH)(EGTQ)(H <sub>2</sub> O) <sub>2</sub> ](ClO <sub>4</sub> ) <sub>3</sub> ·2CH <sub>3</sub> CN
Formula	C <sub>47</sub> H <sub>48</sub> CdCl <sub>2</sub> N <sub>6</sub> O <sub>11</sub>	C <sub>50</sub> H <sub>55</sub> Cd <sub>2</sub> Cl <sub>3</sub> N <sub>8</sub> O <sub>17</sub>
FW	1056.25	1371.21
Crystal system	triclinic	monoclinic
Space group	P-1	P2 <sub>1</sub> /c
a, Å	11.0758(14)	22.6992(5)
b, Å	13.7776(16)	11.29840(10)
c, Å	14.882(2)	23.9600(5)
α, deg	87.257(3)	90
β, deg	89.363(4)	114.7039(9)
γ, deg	75.001(5)	90
V, Å <sup>3</sup>	2191.1(5)	5582.51(18)
Z	2	4
D <sub>calc</sub> , g cm <sup>-3</sup>	1.601	1.631
μ, mm <sup>-1</sup>	0.6925	0.9835
2θ <sub>max</sub> , deg	55	55
temp, K	153	153
no. reflns collected	21773	42577
no. reflns used	9852	12743
no. of params	609	723
R <sub>int</sub>	0.0136	0.0174
Final R1 (I > 2σ(I)) <sup>a</sup>	0.0370	0.0387
wR2 (all data) <sup>b</sup>	0.0940	0.1040
GOF	1.025	1.042

<sup>a</sup>R1 = Σ ||F<sub>o</sub>| - |F<sub>c</sub>| | / Σ |F<sub>o</sub>|.    <sup>b</sup>wR2 = [Σw[(F<sub>o</sub><sup>2</sup> - F<sub>c</sub><sup>2</sup>)<sup>2</sup>] / Σ[w(F<sub>o</sub><sup>2</sup>)<sup>2</sup>]]<sup>1/2</sup>.

**Table S3.** Crystallographic Data for  $[\text{Zn}_2(\mu\text{-OH})(\text{EGTQ})](\text{ClO}_4)_3\cdot\text{CH}_3\text{OH}$  and  $[\text{Zn}_4(\mu\text{-CO}_3)_2(\text{EGTQ})_2](\text{ClO}_4)_4\cdot\text{CH}_3\text{CN}\cdot 2\text{CHCl}_3\cdot\text{CH}_3\text{OH}\cdot\text{H}_2\text{O}$ 

	$[\text{Zn}_2(\mu\text{-OH})(\text{EGTQ})]\cdot(\text{ClO}_4)_3\cdot\text{CH}_3\text{OH}$	$[\text{Zn}_4(\mu\text{-CO}_3)_2(\text{EGTQ})_2](\text{ClO}_4)_4\cdot\text{CH}_3\text{CN}\cdot 2\text{CHCl}_3\cdot\text{CH}_3\text{OH}\cdot\text{H}_2\text{O}$
Formula	$\text{C}_{47}\text{H}_{49}\text{Cl}_3\text{N}_6\text{O}_{16}\text{Zn}_2$	$\text{C}_{99}\text{H}_{99}\text{Cl}_{10}\text{N}_{13}\text{O}_{28}\text{Zn}_4$
FW	1191.05	2534.99
Crystal system	triclinic	monoclinic
Space group	$P\bar{1}$	$P2_1/n$
$a$ , Å	11.91350(10)	23.6390(13)
$b$ , Å	12.44190(10)	15.9622(9)
$c$ , Å	17.5021(2)	30.6105(18)
$\alpha$ , deg	78.577(3)	90
$\beta$ , deg	79.314(3)	112.100(2)
$\gamma$ , deg	78.273(4)	90
$V$ , Å <sup>3</sup>	2461.18(6)	10701.6(11)
Z	2	4
$D_{\text{calc}}$ , g cm <sup>-3</sup>	1.607	1.573
$\mu$ , mm <sup>-1</sup>	1.2159	1.2178
$2\theta_{\text{max}}$ , deg	55	55
temp, K	153	153
no. reflns collected	24466	105545
no. reflns used	11062	24506
no. of params	673	1408
$R_{\text{int}}$	0.0153	0.0517
Final $R1$ ( $I > 2\sigma(I)$ ) <sup>a</sup>	0.0405	0.0673
$wR2$ (all data) <sup>b</sup>	0.1108	0.1818
GOF	1.028	1.056

<sup>a</sup> $R1 = \sum ||F_o| - |F_c|| / \sum |F_o|$ .    <sup>b</sup> $wR2 = [\sum w[(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2]]^{1/2}$ .

**Table S4.** Crystallographic Data for [Cd(BAPTQ)](ClO<sub>4</sub>)<sub>2</sub>·H<sub>2</sub>O and [Cd<sub>2</sub>(BAPTQ)(CH<sub>3</sub>CN)<sub>4</sub>(H<sub>2</sub>O)<sub>2</sub>](ClO<sub>4</sub>)<sub>4</sub>

	[Cd(BAPTQ)]-(ClO <sub>4</sub> ) <sub>2</sub> ·H <sub>2</sub> O	[Cd <sub>2</sub> (BAPTQ)(CH <sub>3</sub> CN) <sub>4</sub> (H <sub>2</sub> O) <sub>2</sub> ](ClO <sub>4</sub> ) <sub>4</sub>
Formula	C <sub>54</sub> H <sub>46</sub> CdCl <sub>2</sub> N <sub>6</sub> O <sub>11</sub>	C <sub>62</sub> H <sub>60</sub> Cd <sub>2</sub> Cl <sub>4</sub> N <sub>10</sub> O <sub>20</sub>
FW	1138.31	1631.84
Crystal system	monoclinic	triclinic
Space group	P2 <sub>1</sub> /n	P-1
a, Å	15.4421(8)	12.3738(8)
b, Å	15.5215(8)	17.9552(14)
c, Å	20.5748(10)	18.3514(13)
α, deg	90	103.789(3)
β, deg	92.922(3)	90.8895(10)
γ, deg	90	108.016(3)
V, Å <sup>3</sup>	4925.1(4)	3747.9(5)
Z	4	2
D <sub>calc</sub> , g cm <sup>-3</sup>	1.535	1.446
μ, mm <sup>-1</sup>	0.6226	0.7834
2θ <sub>max</sub> , deg	55	54.9
temp, K	153	153
no. reflns collected	50632	36976
no. reflns used	11303	16805
no. of params	703	887
R <sub>int</sub>	0.0224	0.0387
Final R1 (I > 2σ(I)) <sup>a</sup>	0.0508	0.0584
wR2 (all data) <sup>b</sup>	0.1562	0.1653
GOF	1.049	1.061

<sup>a</sup>R1 = Σ ||F<sub>o</sub>| - |F<sub>c</sub>| | / Σ |F<sub>o</sub>|.    <sup>b</sup>wR2 = [Σw[(F<sub>o</sub><sup>2</sup> - F<sub>c</sub><sup>2</sup>)<sup>2</sup>] / Σ[w(F<sub>o</sub><sup>2</sup>)<sup>2</sup>]]<sup>1/2</sup>.

**Table S5.** Crystallographic Data for  $[Zn_2(\mu\text{-OH})(BAPTQ)][Zn_2(\mu\text{-OH})(BAPTQ)(CH_3CN)](ClO_4)_6 \cdot 2CH_3CN \cdot C_2H_5OC_2H_5$  and  $[Zn_2(\mu\text{-OH})(6\text{-MeOEGTQ})](ClO_4)_3$

	$[Zn_2(\mu\text{-OH})(BAPTQ)][Zn_2(\mu\text{-OH})(BAPTQ)(CH_3CN)](ClO_4)_6 \cdot 2CH_3CN \cdot C_2H_5OC_2H_5$	$[Zn_2(\mu\text{-OH})(6\text{-MeOEGTQ})](ClO_4)_3$
Formula	$C_{126}H_{121}Cl_6N_{19}O_{31}Zn_4$	$C_{50}H_{53}Cl_3N_6O_{19}Zn_2$
FW	2871.69	1279.12
Crystal system	monoclinic	triclinic
Space group	$P2_1/c$	$P-1$
$a$ , Å	23.1074(4)	13.244(3)
$b$ , Å	28.4612(6)	14.742(3)
$c$ , Å	19.2564(4)	17.352(3)
$\alpha$ , deg	90	91.076(3)
$\beta$ , deg	91.4750(11)	96.137(2)
$\gamma$ , deg	90	105.435(3)
$V$ , Å <sup>3</sup>	12660.0(4)	3242.9(11)
Z	4	2
$D_{\text{calc}}$ , g cm <sup>-3</sup>	1.507	1.310
$\mu$ , mm <sup>-1</sup>	0.9609	0.9310
$2\theta_{\text{max}}$ , deg	55	55
temp, K	153	153
no. reflns collected	128986	25309
no. reflns used	28808	14133
no. of params	1742	729
$R_{\text{int}}$	0.0294	0.0160
Final $R1$ ( $I > 2\sigma(I)$ ) <sup>a</sup>	0.0695	0.0516
$wR2$ (all data) <sup>b</sup>	0.2282	0.1499
GOF	1.050	1.065

<sup>a</sup> $R1 = \sum ||F_o| - |F_c|| / \sum |F_o|$ .    <sup>b</sup> $wR2 = [\sum w[(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2]]^{1/2}$ .

**Table S6.** Crystallographic Data for  $[\text{Cd}_2\text{Br}_2(\text{TriMeOBAPTQ})(\text{CH}_3\text{OH})(\text{H}_2\text{O})](\text{ClO}_4)_2$ 

$[\text{Cd}_2\text{Br}_2(\text{TriMeOBAPTQ})(\text{CH}_3\text{OH})(\text{H}_2\text{O})](\text{ClO}_4)_2$	
Formula	$\text{C}_{67}\text{H}_{74}\text{Br}_2\text{Cd}_2\text{Cl}_2\text{N}_6\text{O}_{24}$
FW	1802.88
Crystal system	triclinic
Space group	<i>P</i> -1
<i>a</i> , Å	13.614(3)
<i>b</i> , Å	16.137(3)
<i>c</i> , Å	20.661(4)
$\alpha$ , deg	72.534(5)
$\beta$ , deg	85.863(6)
$\gamma$ , deg	76.622(6)
<i>V</i> , Å <sup>3</sup>	4212.2(15)
<i>Z</i>	2
<i>D</i> <sub>calc</sub> , g cm <sup>-3</sup>	1.421
$\mu$ , mm <sup>-1</sup>	1.5908
$2\theta_{\max}$ , deg	55
temp, K	153
no. reflns collected	44019
no. reflns used	19115
no. of params	941
<i>R</i> <sub>int</sub>	0.0288
Final <i>R</i> 1 ( $I > 2\sigma(I)$ ) <sup>a</sup>	0.0555
<i>wR</i> 2 (all data) <sup>b</sup>	0.1824
GOF	1.071

<sup>a</sup> $R1 = \sum |F_o| - |F_c| | / \sum |F_o|$ .    <sup>b</sup> $wR2 = [\sum w[(F_o^2 - F_c^2)^2] / \sum [w(F_o^2)^2]]^{1/2}$ .

**Table S7.** Fluorescence Lifetimes for Cd<sup>2+</sup> and Zn<sup>2+</sup> Complexes of TQTACN, 6-MeOTQTACN and TriMeOTQTACN in DMF-H<sub>2</sub>O (1:1)<sup>a</sup>

Ligand	Metal	$\lambda_{\text{em}}$	BPF	$\tau$ (nsec) <sup>c</sup>
	Ion	(nm)	(nm) <sup>b</sup>	
TQTACN	Zn <sup>2+</sup>	393	400	1.46 (48%), 6.94 (17%)
	Cd <sup>2+</sup>	ND <sup>d</sup>	-	-
6-MeOTQTACN	Zn <sup>2+</sup>	420	430	1.02 (27%), 3.68 (38%), 9.18 (35%)
	Cd <sup>2+</sup>	420	430	0.99 (30%), 3.22 (49%), 8.62 (21%)
TriMeOTQTACN	Zn <sup>2+</sup>	498	490	3.07 (3%), 8.76 (86%), 23.34 (11%)
	Cd <sup>2+</sup>	489	490	2.72 (1%), 10.81 (87%), 22.78 (12%)

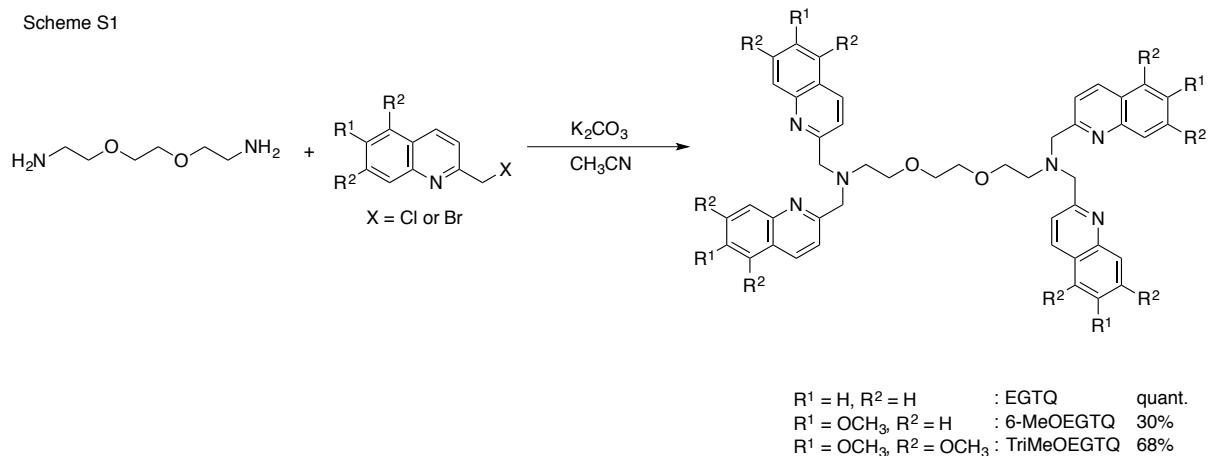
<sup>a</sup> Conditions: 34 μM solution in DMF-H<sub>2</sub>O (1:1) at 25 °C in the presence of 2 equiv. of metal ion ( $\lambda_{\text{ex}} = 331$  nm).

<sup>b</sup> Bandpath filter used (±10 nm).

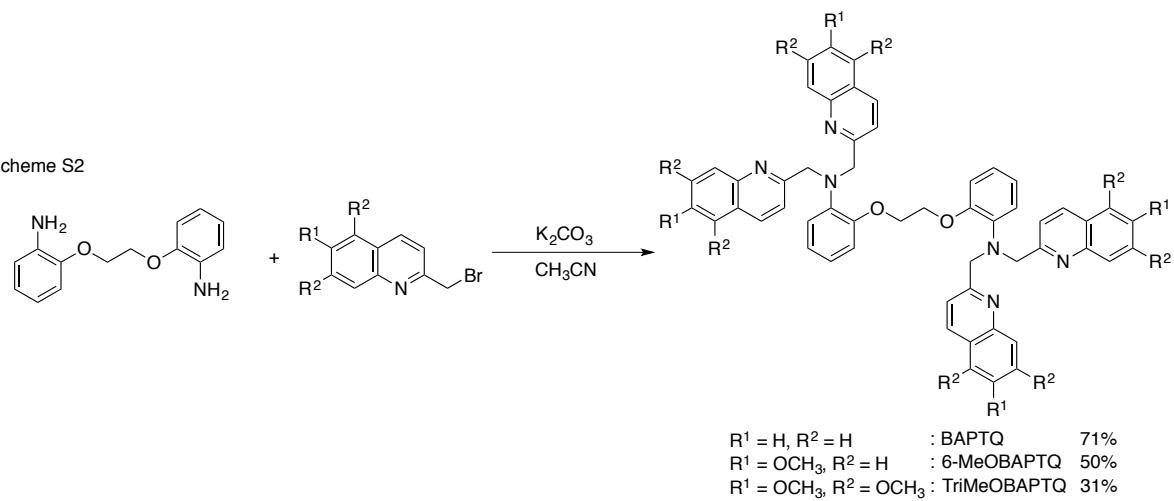
<sup>c</sup> Components with extremely short lifetime (< 1 nsec) were omitted.

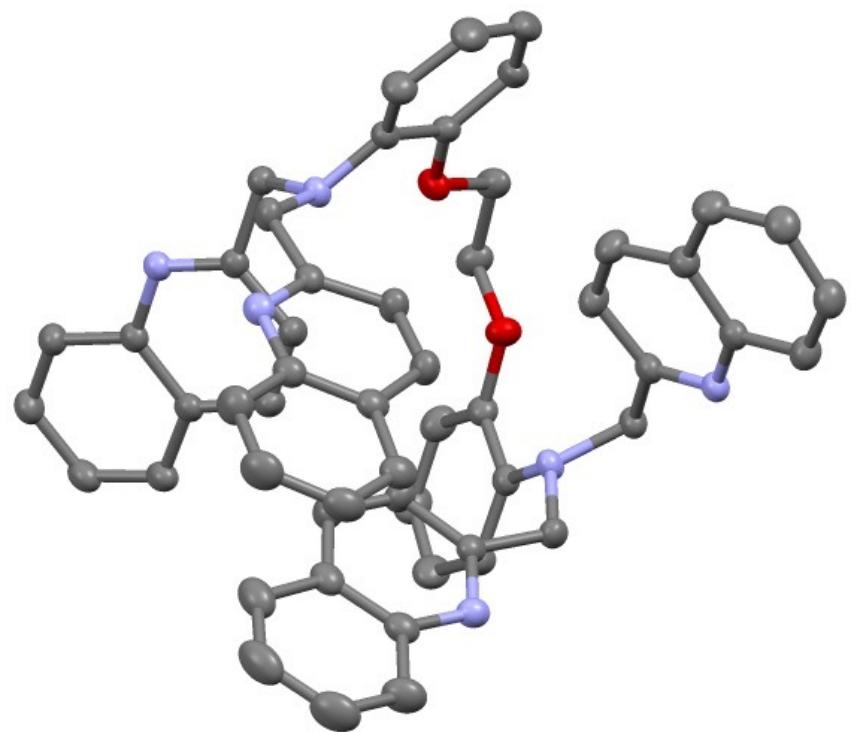
<sup>d</sup> No emission was observed.

Scheme S1

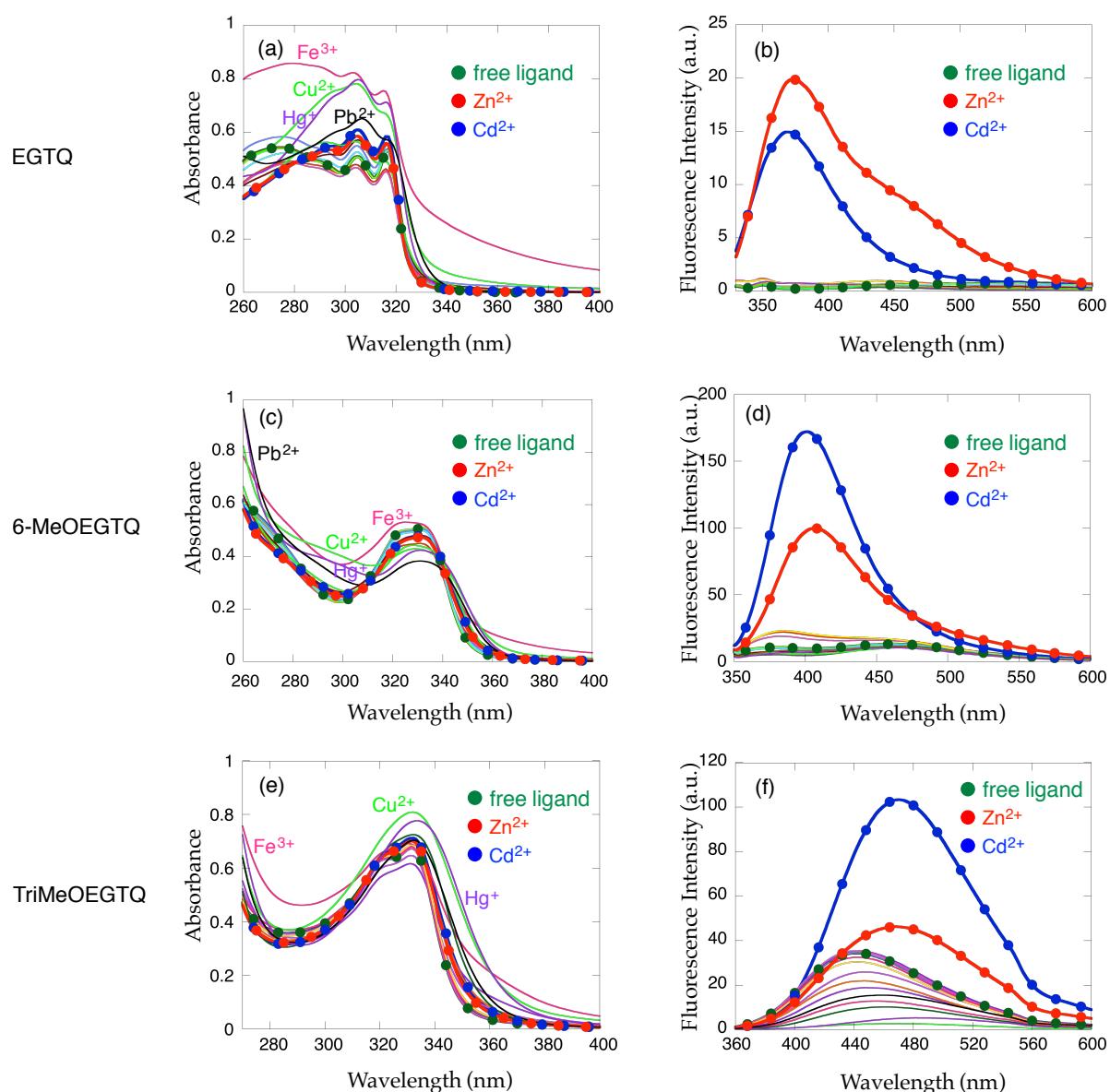


Scheme S2

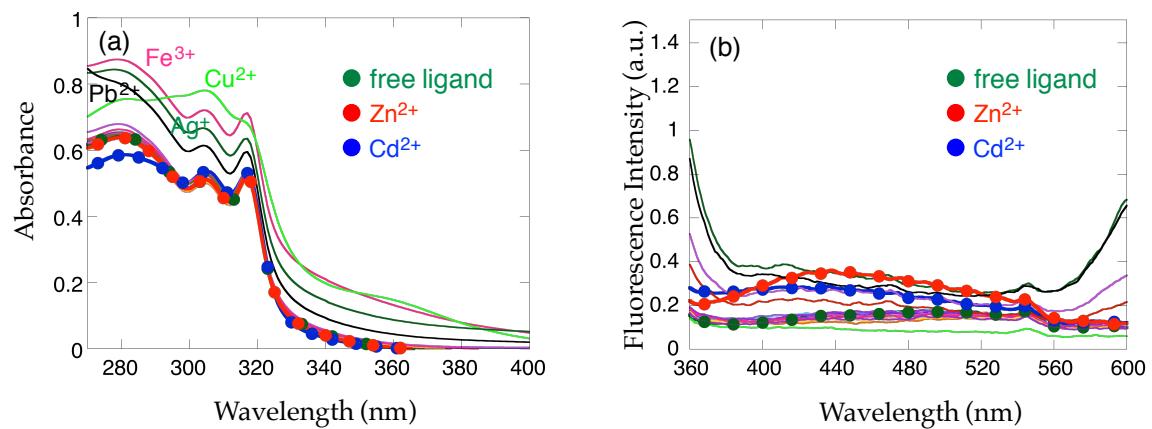




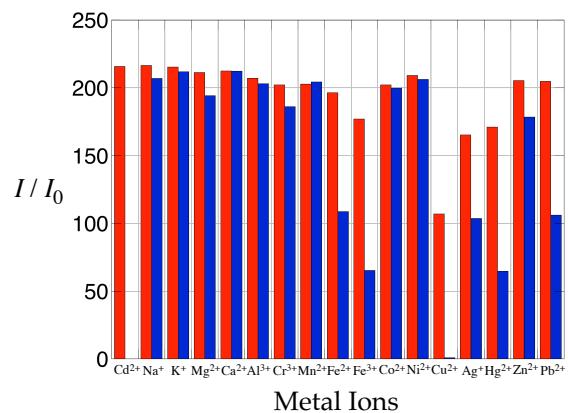
**Fig. S1.** Perspective view for BAPTQ in 50% probability. Hydrogen atoms are omitted for clarity.



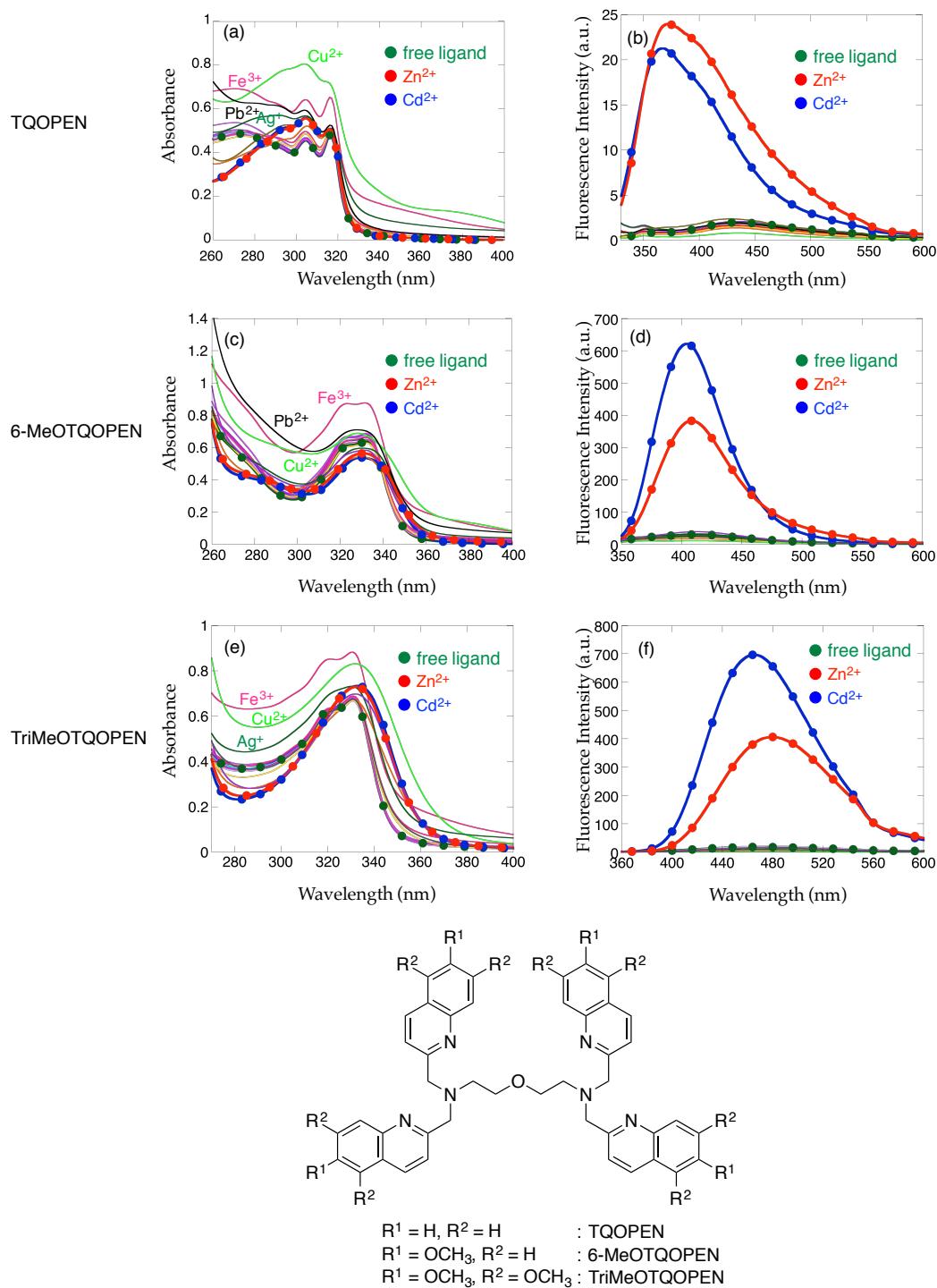
**Fig. S2.** (a, c, e) UV-vis absorption and (b, d, f) fluorescence spectral changes of 34  $\mu\text{M}$  (a, b) EGTQ ( $\lambda_{\text{ex}} = 317 \text{ nm}$ ), (c, d) 6-MeOEGTQ ( $\lambda_{\text{ex}} = 342 \text{ nm}$ ) and (e, f) TriMeOEGTQ ( $\lambda_{\text{ex}} = 342 \text{ nm}$ ) in DMF-H<sub>2</sub>O (1:1) at 25 °C in the presence of (a,b) 5 equiv. or (c-f) 2 equiv. of various metal ions.



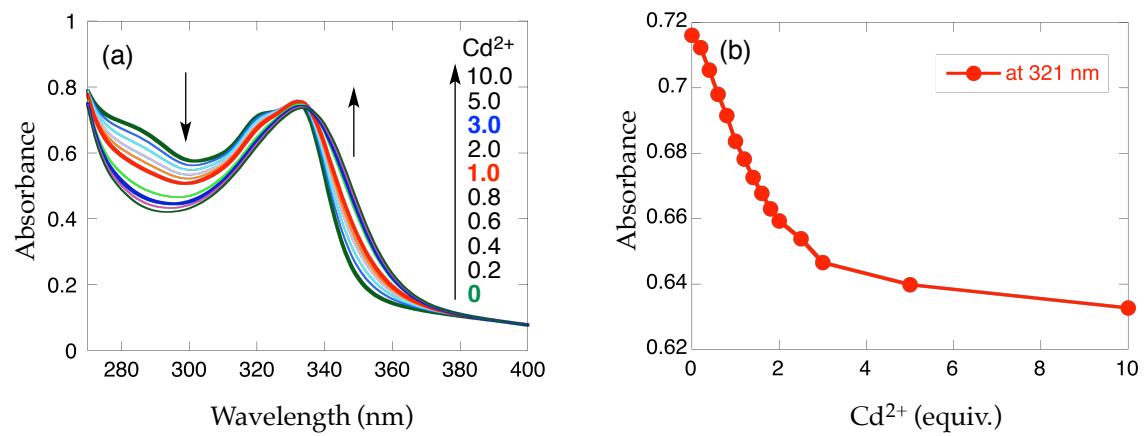
**Fig. S3.** (a) UV-vis absorption and (b) fluorescence spectral changes of 34  $\mu\text{M}$  BAPTO ( $\lambda_{\text{ex}} = 317 \text{ nm}$ ) in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25  $^{\circ}\text{C}$  in the presence of 3 equiv. of various metal ions.



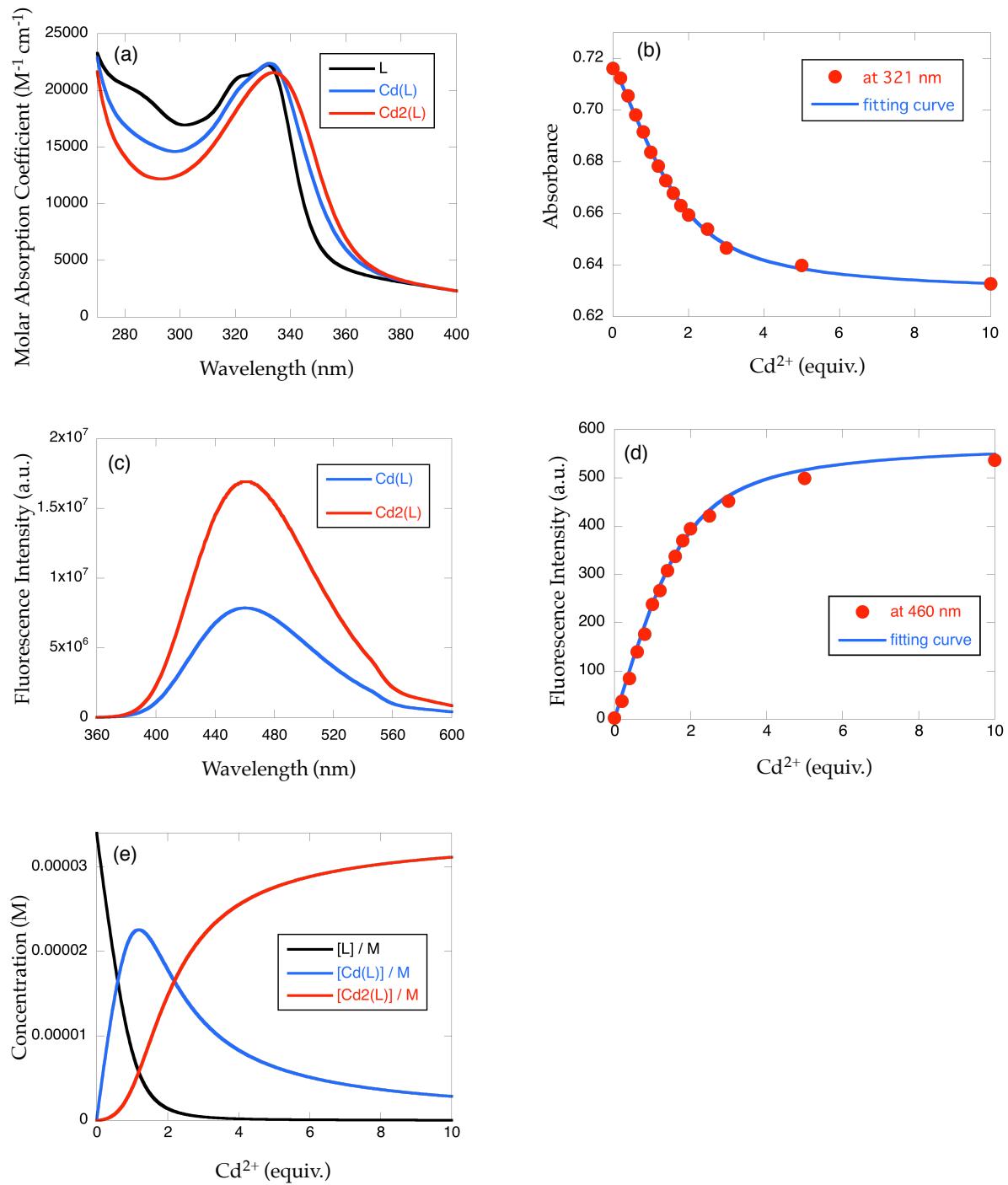
**Fig. S4.** Plot of fluorescence intensity of 34  $\mu\text{M}$  TriMeOBAPQ at 460 nm in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25 °C in the presence of 3 equiv. of  $\text{Cd}^{2+}$  + 3 equiv. of metal ions (red bars) and 3 equiv. of  $\text{Cd}^{2+}$  + 30 equiv. of metal ions (blue bars) ( $\lambda_{\text{ex}} = 347$  nm).



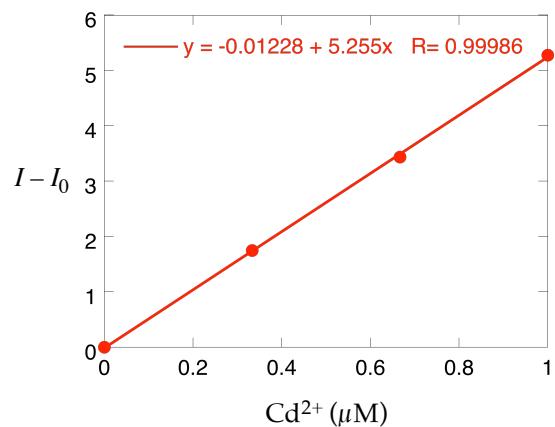
**Fig. S5.** (a, c, e) UV-vis absorption and (b, d, f) fluorescence spectral changes of 34  $\mu$ M (a, b) TQOPEN ( $\lambda_{\text{ex}} = 317$  nm), (c, d) 6-MeOTQOPEN ( $\lambda_{\text{ex}} = 342$  nm) and (e, f) TriMeOTQOPEN ( $\lambda_{\text{ex}} = 342$  nm) in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25 °C in the presence of 3 equiv. of various metal ions.



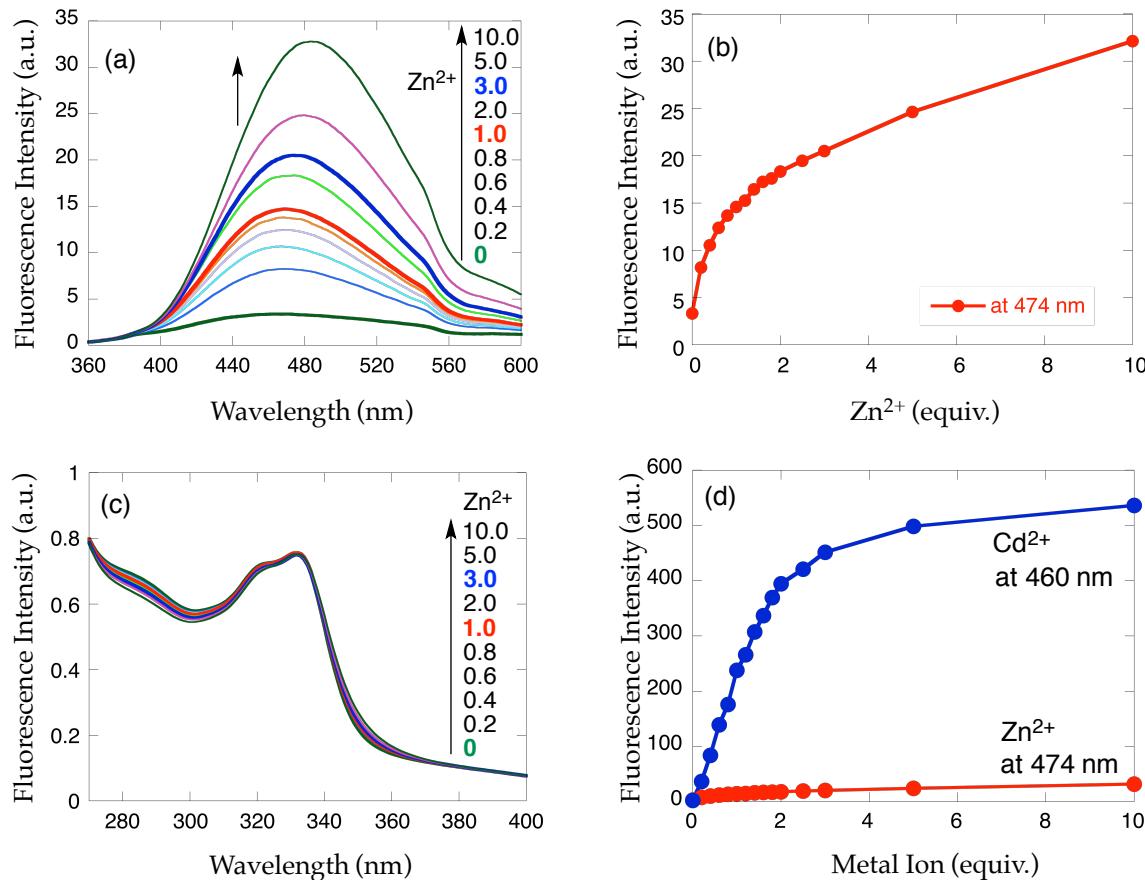
**Fig. S6.** (a) UV-vis absorption spectra and (b) plot of absorbance change at 321 nm of 34  $\mu\text{M}$  TriMeOBAPTQ in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25  $^{\circ}\text{C}$  in the presence of increasing amount of  $\text{Cd}^{2+}$ .



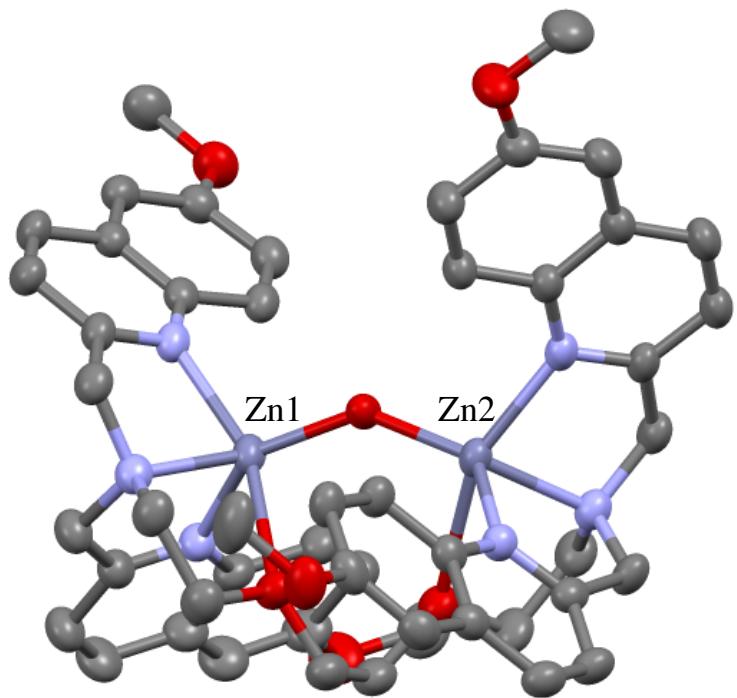
**Fig. S7.** (a) Separated UV-vis absorption spectra of free ligand (L), Cd(L) and Cd<sub>2</sub>(L), (b) simulated absorbance change at 321 nm, (c) separated fluorescence spectra of Cd(L) and Cd<sub>2</sub>(L), (d) simulated fluorescence intensity change at 460 nm and (e) distribution profile of L, Cd(L) and Cd<sub>2</sub>(L) during the titration of 34  $\mu\text{M}$  TriMeOBAPTQ in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25 °C in the presence of increasing amount of Cd<sup>2+</sup>.



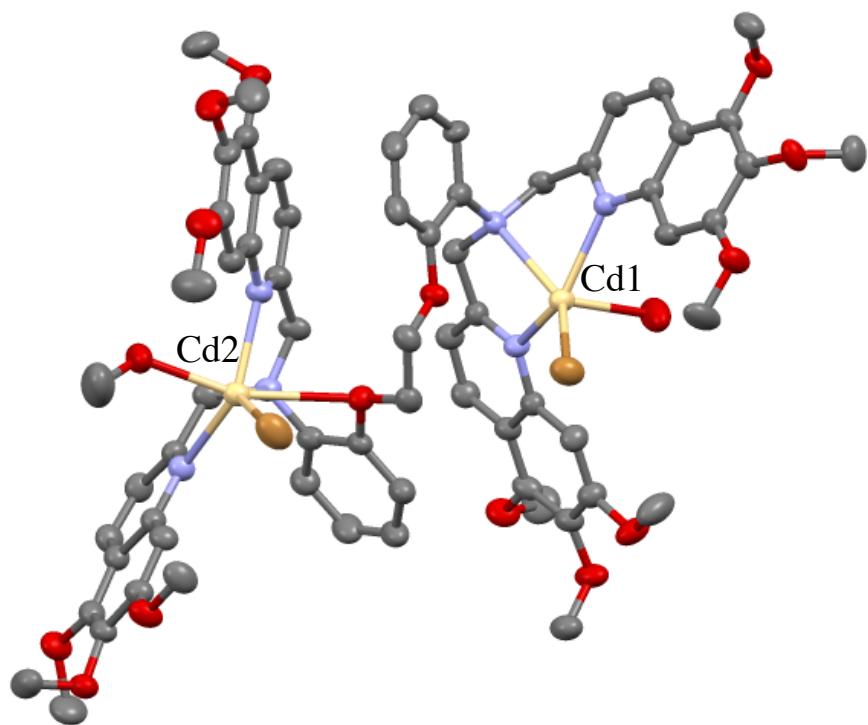
**Fig. S8.** Estimation of LOD (limit of detection) for  $\text{Cd}^{2+}$  with TriMeOBAPTQ in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25 °C. The  $3\sigma$  value ( $\sigma$  corresponds to standard deviation from 7 measurements) of blank solution (34 µM TriMeOBAPTQ) is 0.051831 in fluorescence intensity unit, which corresponds to 9.9 nM from the slope of the liner dynamic fluorescence intensity plot ( $k$ ) shown above ( $\text{LOD} = 3\sigma/k$ ).



**Fig. S9.** (a) Fluorescence spectra ( $\lambda_{\text{ex}} = 347 \text{ nm}$ ), (b) plot of fluorescence intensity change at 474 nm and (c) UV-vis spectra of 34  $\mu\text{M}$  TriMeOBAPTQ in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25 °C in the presence of increasing amount of  $\text{Zn}^{2+}$ . (d) Comparison of fluorescence intensity changes of 34  $\mu\text{M}$  TriMeOBAPTQ in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 25 °C in the presence of increasing amount of  $\text{Cd}^{2+}$  at 460 nm and  $\text{Zn}^{2+}$  at 474 nm ( $\lambda_{\text{ex}} = 347 \text{ nm}$ ).

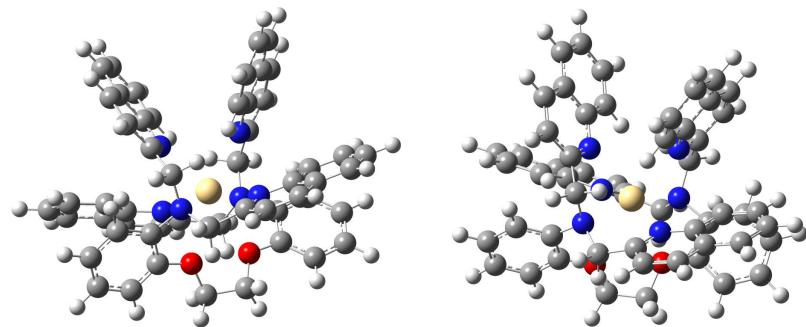


**Fig. S10.** Perspective view for  $[\text{Zn}_2(\mu\text{-OH})(6\text{-MeOEGTQ})](\text{ClO}_4)_3$  in 50% probability. Hydrogen atoms, counteranions and solvents are omitted for clarity.

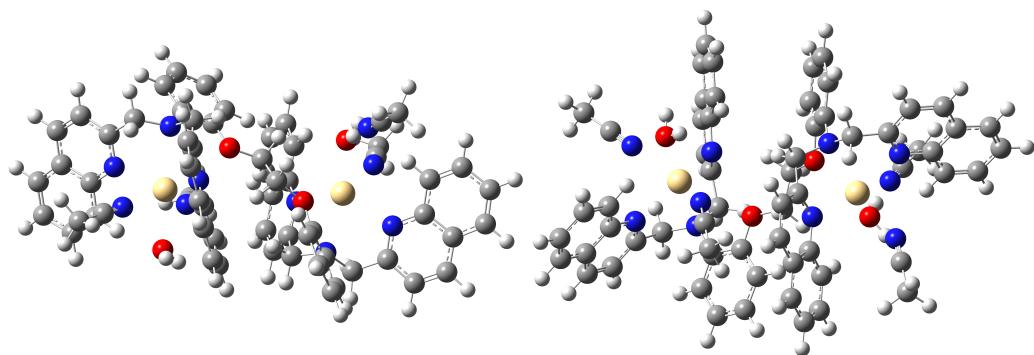


**Fig. S11.** Perspective view for  $[\text{Cd}_2\text{Br}_2(\text{TriMeOBAPTQ})(\text{CH}_3\text{OH})(\text{H}_2\text{O})](\text{ClO}_4)_2$  in 50% probability. Hydrogen atoms, counteranions and solvents are omitted for clarity.

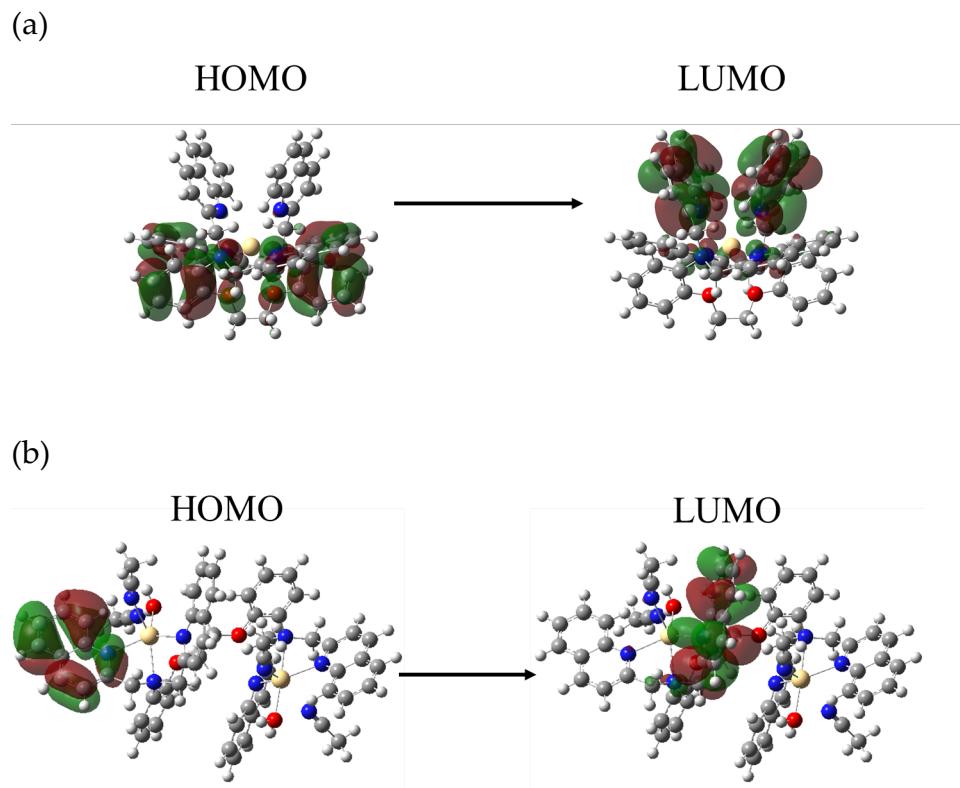
(a)



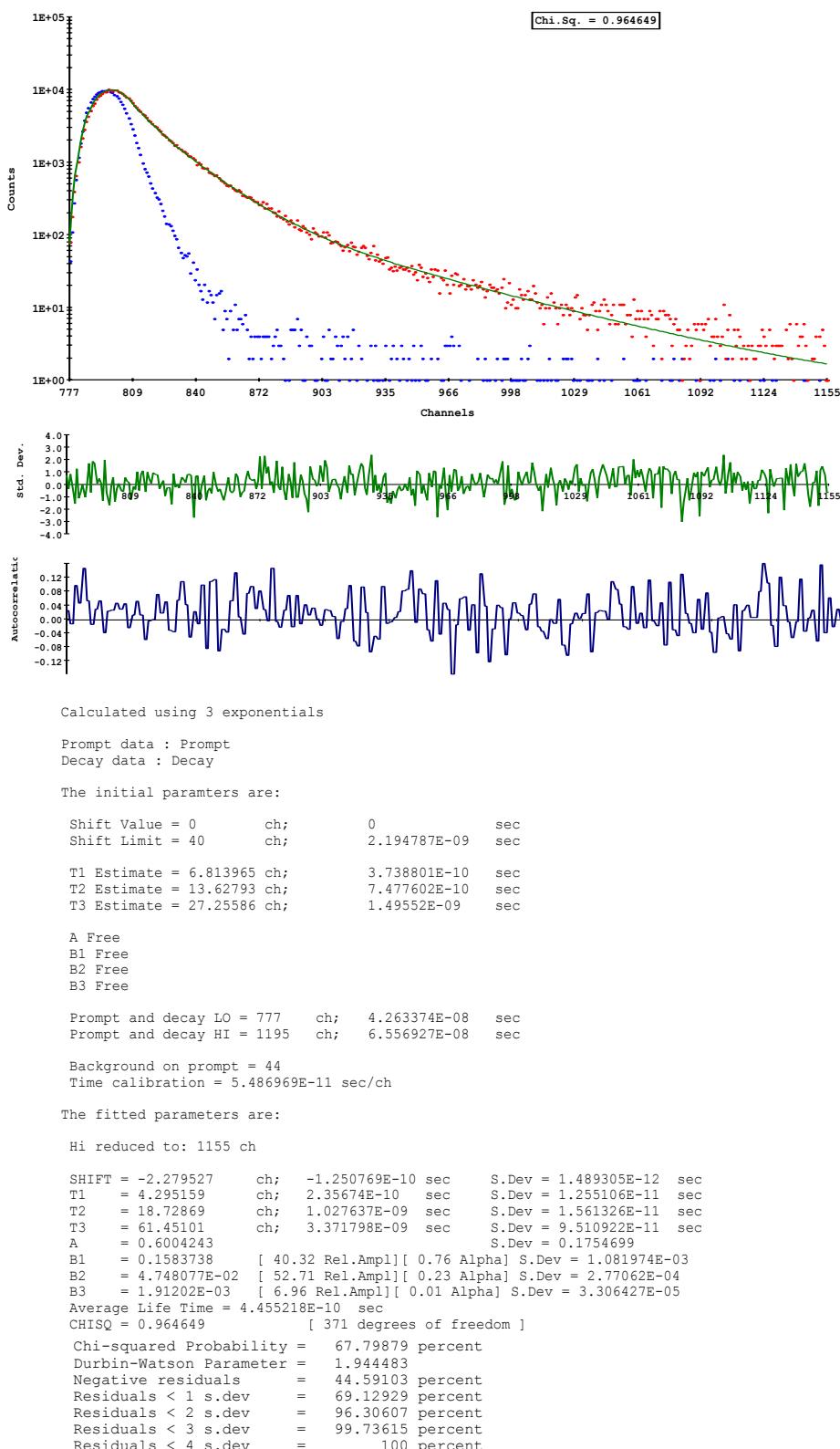
(b)



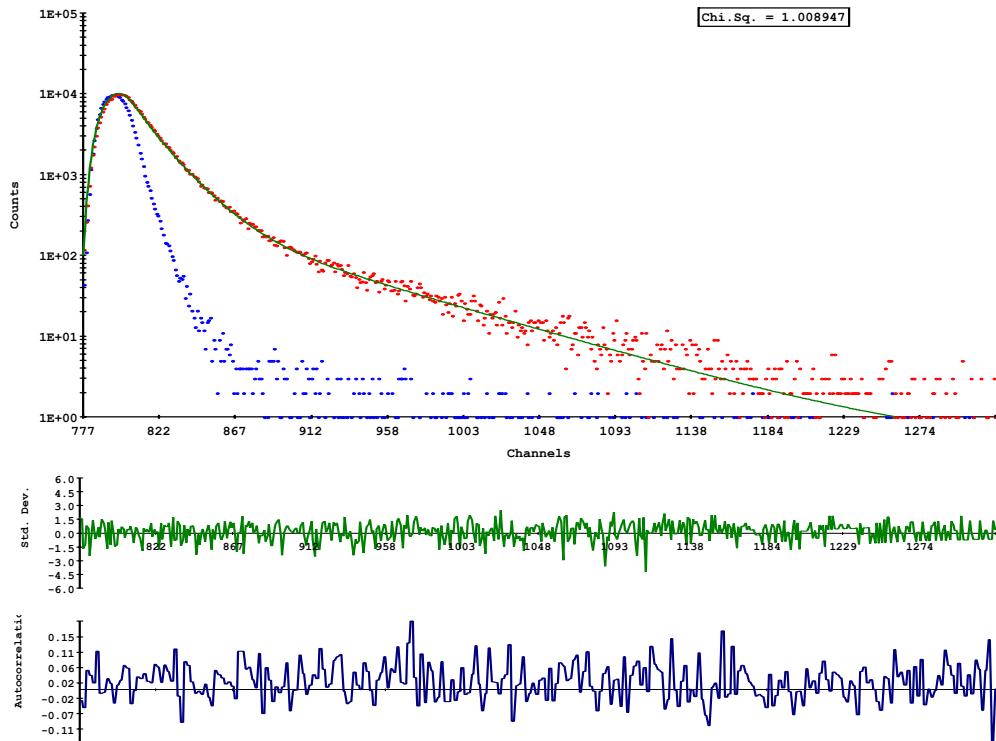
**Figure S12.** Optimized structures from different two directions for (a)  $[\text{Cd}(\text{BAPTQ})]^{2+}$  and (b)  $[\text{Cd}_2(\text{BAPTQ})(\text{H}_2\text{O})_2(\text{CH}_3\text{CN})_4]^{4+}$ .



**Figure S13.** HOMO and LUMO for (a)  $[\text{Cd}(\text{BAPTQ})]^{2+}$  and (b)  $[\text{Cd}_2(\text{BAPTQ})(\text{H}_2\text{O})_2(\text{CH}_3\text{CN})_4]^{4+}$ .



**Fig. S14.** Fluorescent lifetime measurement of 34  $\mu$ M EGTQ in the presence of 3 equiv. of Zn<sup>2+</sup> in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 370 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



Calculated using 3 exponentials

Prompt data : Prompt  
Decay data : Decay

The initial parameters are:

Shift Value = 0	ch;	0	sec
Shift Limit = 40	ch;	2.194787E-09	sec
T1 Estimate = 7.586029	ch;	4.16243E-10	sec
T2 Estimate = 15.17206	ch;	8.32486E-10	sec
T3 Estimate = 30.34412	ch;	1.664972E-09	sec

A Free  
B1 Free  
B2 Free  
B3 Free

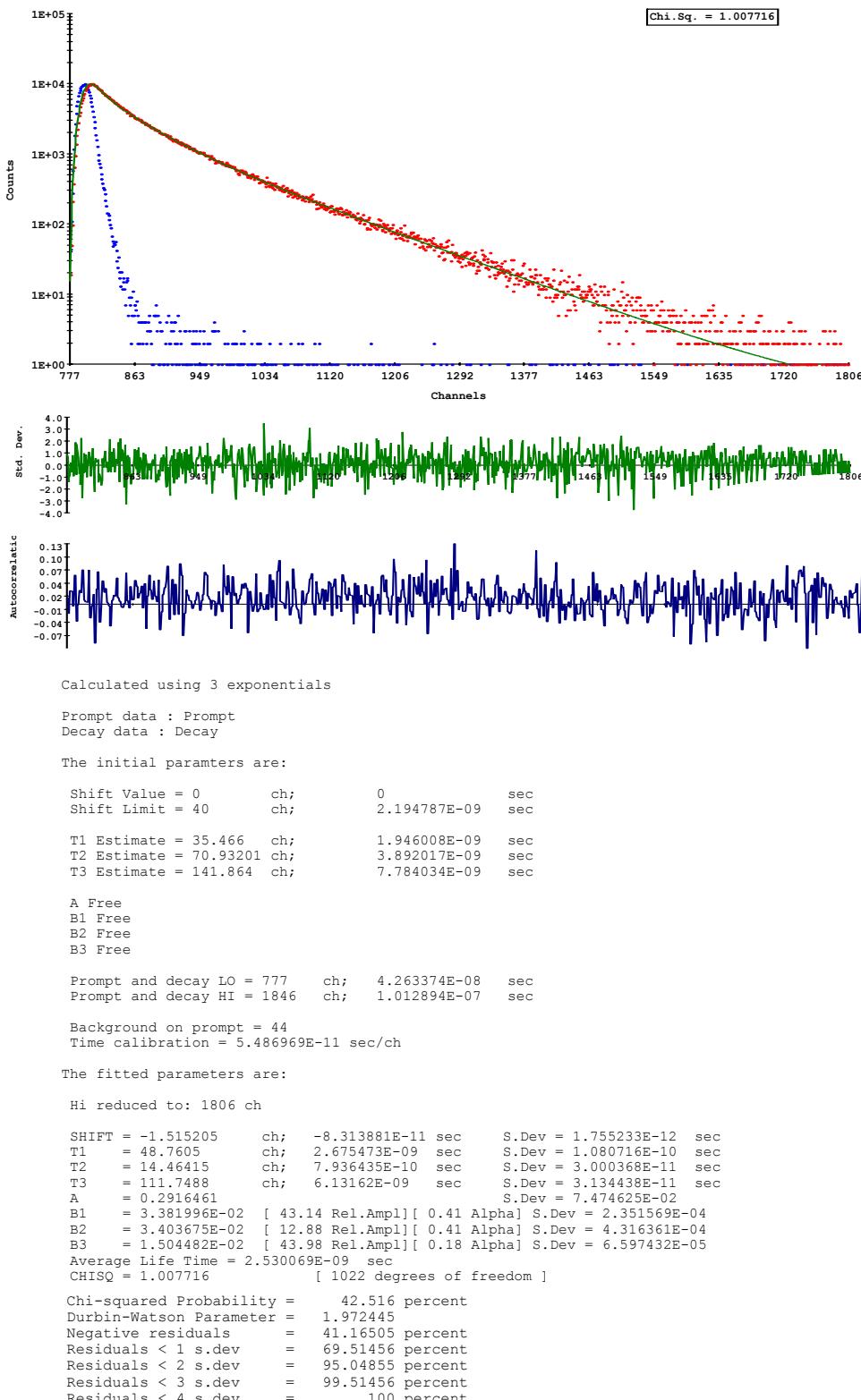
Prompt and decay LO = 777 ch; 4.263374E-08 sec  
Prompt and decay HI = 1359 ch; 7.45679E-08 sec

Background on prompt = 44  
Time calibration = 5.486969E-11 sec/ch

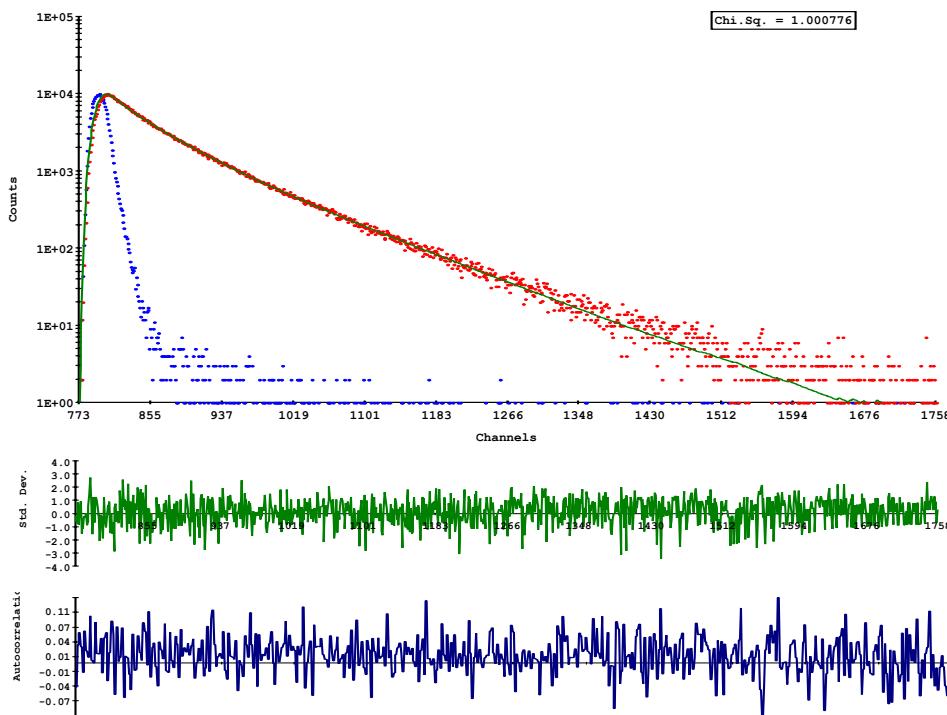
The fitted parameters are:

Chi-squared Probability =	43.40381 percent
H-Durbin-Watson Parameter =	2.053351
Negative residuals =	39.22652 percent
S1 Residuals < 1 s.dev =	67.9558 percent $\text{lev} = 1.846938E-12$ sec
T1 Residuals < 2 s.dev =	97.05341 percent $\text{lev} = 1.748121E-11$ sec
T2 Residuals < 3 s.dev =	99.63168 percent $\text{lev} = 1.031822E-11$ sec
A Residuals < 4 s.dev =	99.81584 percent $\text{lev} = 9.419094E-02$
B1 = 0.1266575	[ 31.52 Rel.Ampl] [ 0.67 Alpha] S.Dev = 1.130463E-03
B2 = 6.077619E-02	[ 59.97 Rel.Ampl] [ 0.32 Alpha] S.Dev = 3.14106E-04
B3 = 2.068207E-03	[ 8.51 Rel.Ampl] [ 0.01 Alpha] S.Dev = 2.423551E-05
Average Life Time =	5.053373E-10 sec
CHISQ = 1.008947	[ 535 degrees of freedom ]

**Fig. S15.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  EGTQ in the presence of 3 equiv. of  $\text{Cd}^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 370 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



**Fig. S16.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  6-MeOEGTQ in the presence of 3 equiv. of  $\text{Zn}^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 400 nm at 25 °C ( $\lambda_{\text{ex}} = 331 \text{ nm}$ ).



Calculated using 3 exponentials

Prompt data : Prompt  
Decay data : Decay

The initial parameters are:

Shift Value = 0	ch;	0	sec
Shift Limit = 40	ch;	2.194787E-09	sec
T1 Estimate = 34.15521	ch;	1.874086E-09	sec
T2 Estimate = 68.31042	ch;	3.748172E-09	sec
T3 Estimate = 136.6208	ch;	7.496343E-09	sec

A Free  
B1 Free  
B2 Free  
B3 Free

Prompt and decay LO = 773 ch; 4.241427E-08 sec  
Prompt and decay HI = 1798 ch; 9.86557E-08 sec

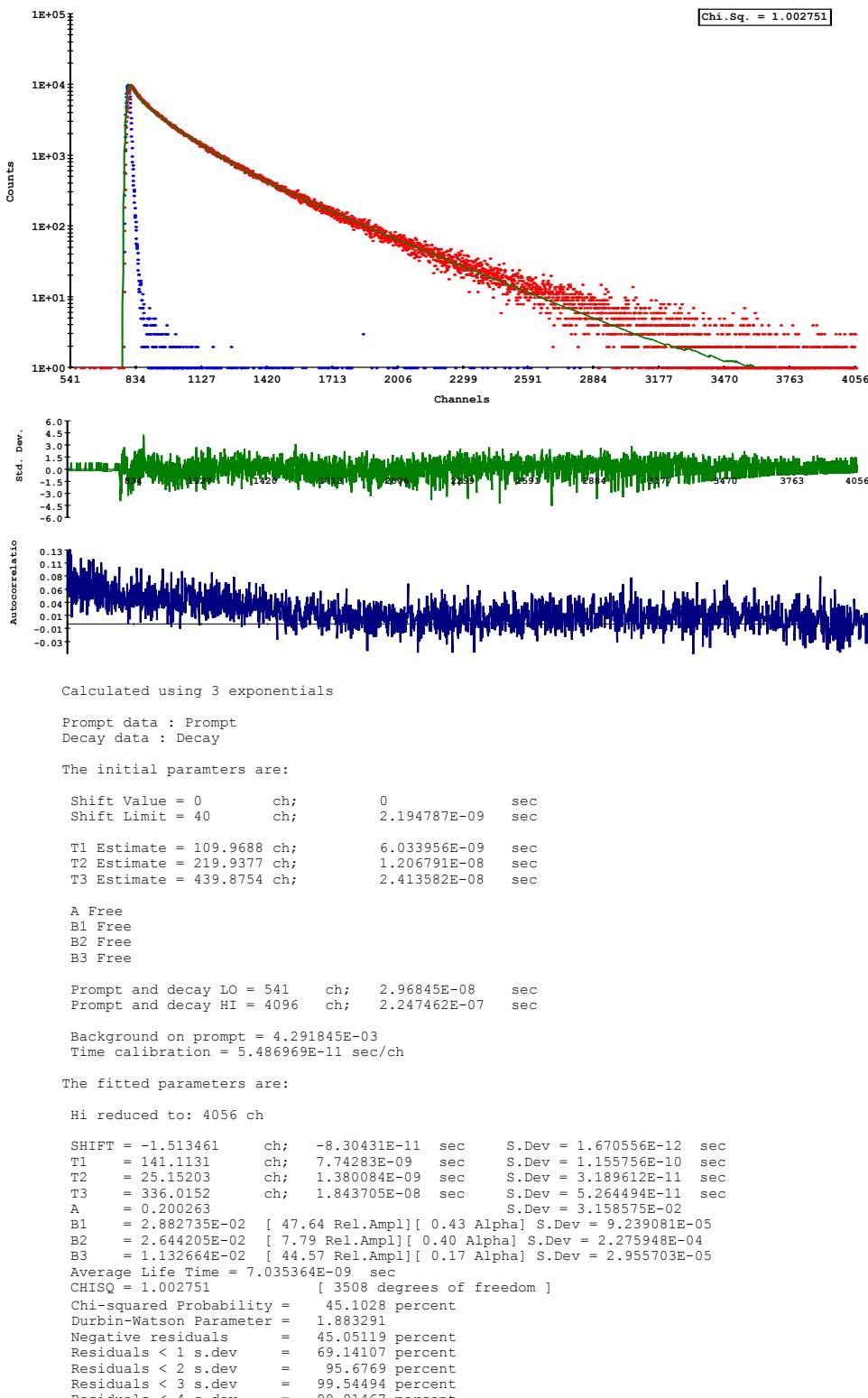
Background on prompt = 0  
Time calibration = 5.486969E-11 sec/ch

The fitted parameters are:

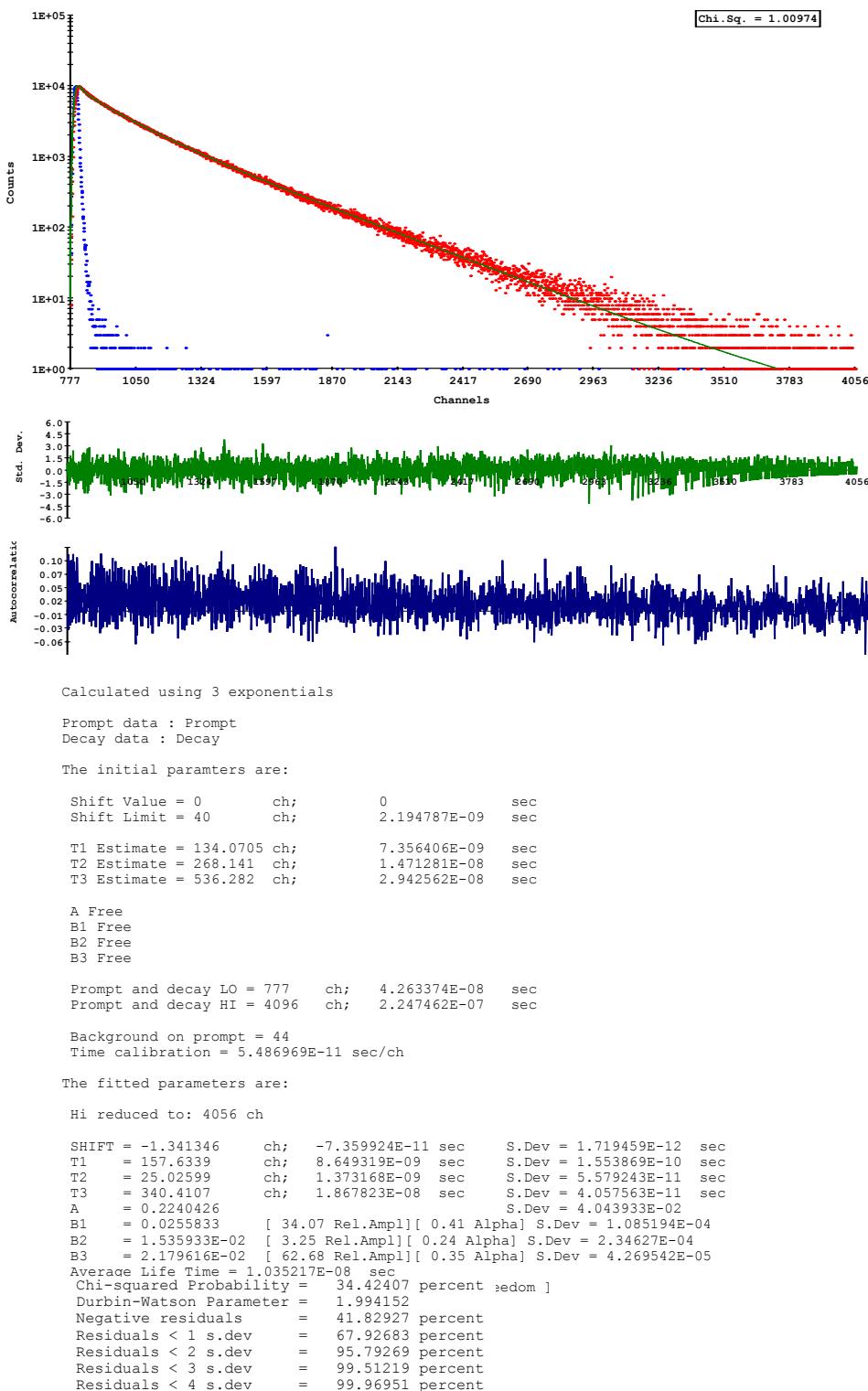
Hi reduced to: 1758 ch

SHIFT = -1.535257	ch;	-8.423909E-11	sec	S.Dev = 1.91863E-12	sec
T1 = 12.31128	ch;	6.755162E-10	sec	S.Dev = 8.275124E-11	sec
T2 = 46.84394	ch;	2.570312E-09	sec	S.Dev = 5.669799E-11	sec
T3 = 101.0398	ch;	5.544023E-09	sec	S.Dev = 2.748387E-11	sec
A = 0.1115768				S.Dev = 7.576162E-02	
B1 = 2.374297E-02	[ 7.48 Rel.Ampl][ 0.30 Alpha]	S.Dev = 4.676534E-04			
B2 = 3.720786E-02	[ 44.60 Rel.Ampl][ 0.47 Alpha]	S.Dev = 2.526488E-04			
B3 = 1.853837E-02	[ 47.93 Rel.Ampl][ 0.23 Alpha]	S.Dev = 8.037989E-05			
Average Life Time = 2.697873E-09 sec					
CHISQ = 1.000776 [ 978 degrees of freedom ]					
Chi-squared Probability = 48.71476 percent					
Durbin-Watson Parameter = 2.095582					
Negative residuals = 42.39351 percent					
Residuals < 1 s.dev = 67.13996 percent					
Residuals < 2 s.dev = 95.53753 percent					
Residuals < 3 s.dev = 99.69574 percent					
Residuals < 4 s.dev = 100 percent					

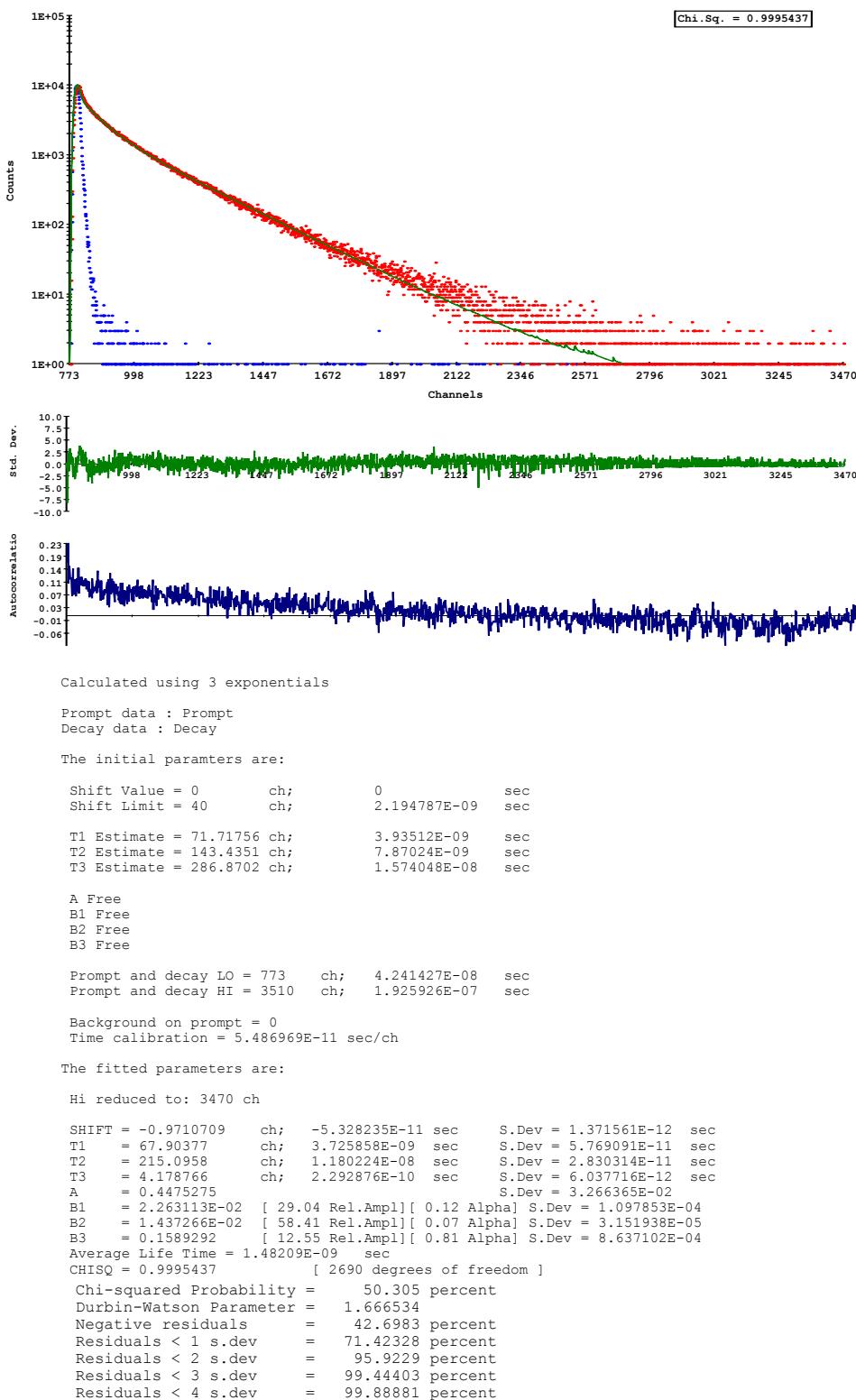
**Fig. S17.** Fluorescent lifetime measurement of 34  $\mu$ M 6-MeOEGTQ in the presence of 3 equiv. of Cd<sup>2+</sup> in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 400 nm at 25 °C ( $\lambda_{ex}$  = 331 nm).



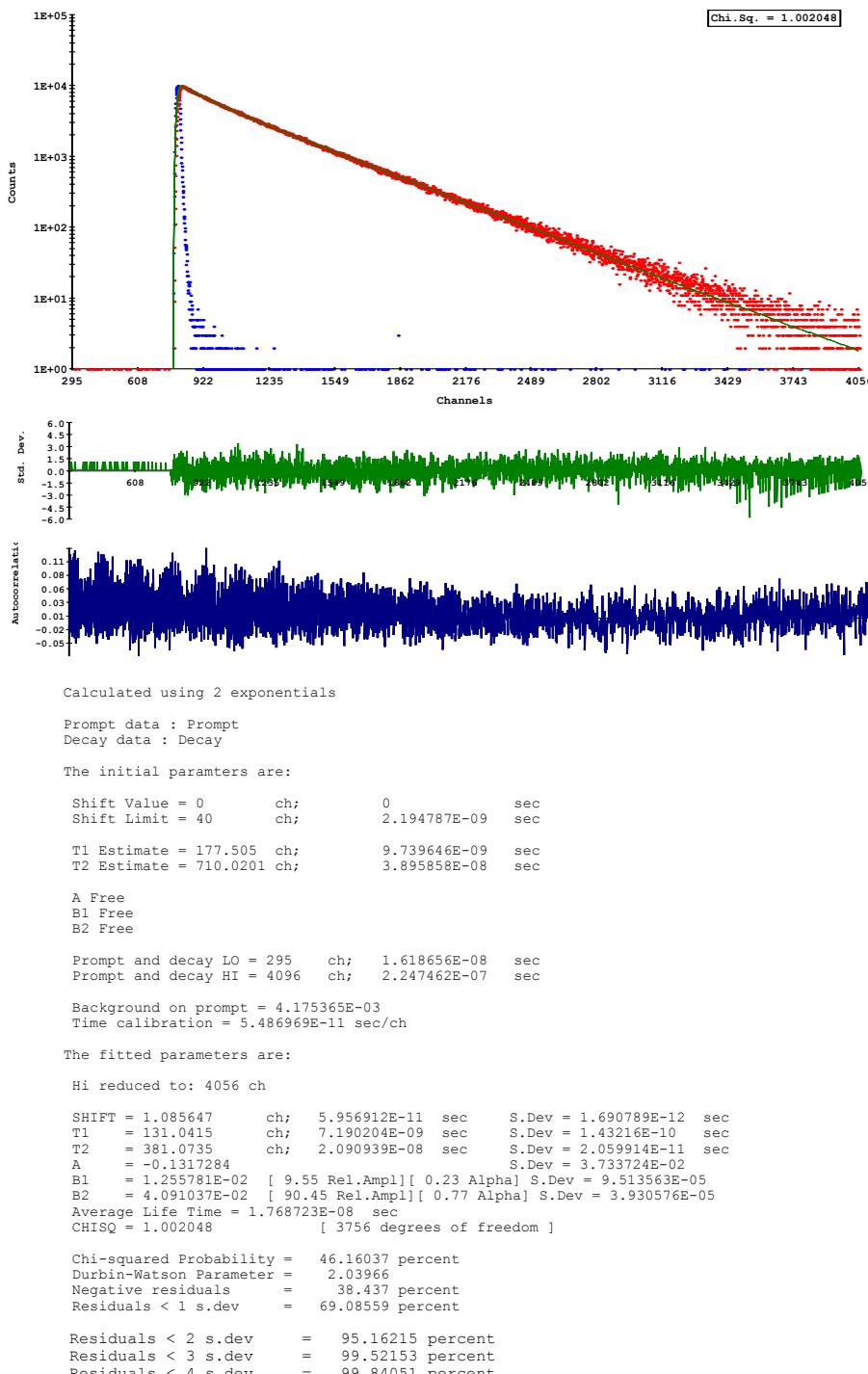
**Fig. S18.** Fluorescent lifetime measurement of 34  $\mu$ M TriMeOEGTQ in the presence of 3 equiv. of Zn<sup>2+</sup> in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 460 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



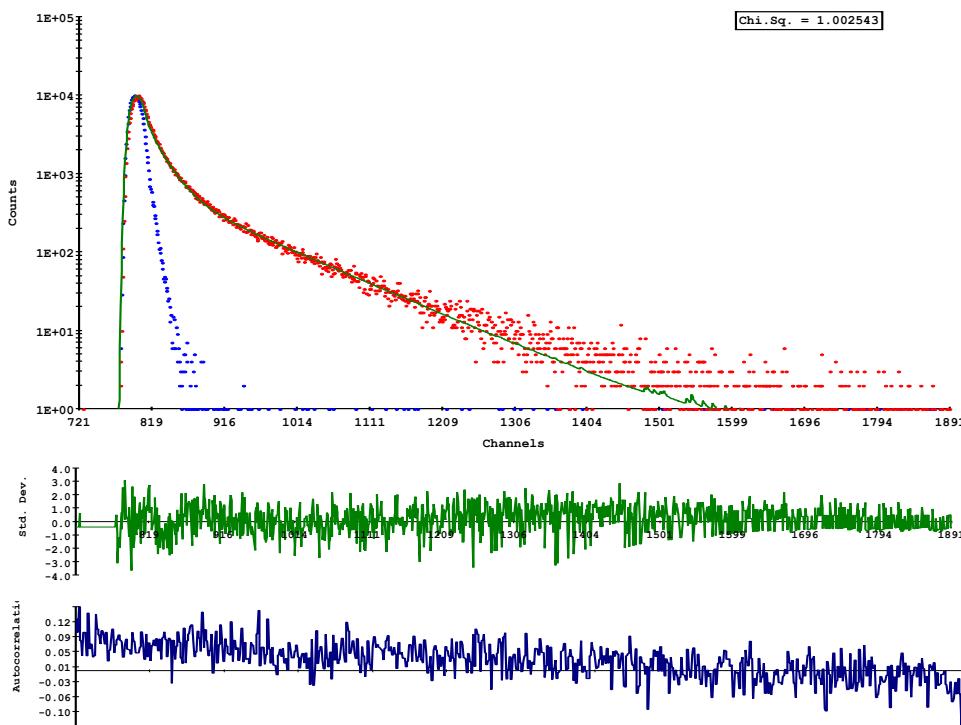
**Fig. S19.** Fluorescent lifetime measurement of 34  $\mu$ M TriMeOEGTQ in the presence of 3 equiv. of Cd<sup>2+</sup> in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 460 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



**Fig. S20.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  TriMeOBAPTQ in the presence of 3 equiv. of  $\text{Zn}^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 460 nm at 25 °C ( $\lambda_{\text{ex}} = 331 \text{ nm}$ ).



**Fig. S21.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  TriMeOBAPTO in the presence of 3 equiv. of  $\text{Cd}^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 460 nm at 25 °C ( $\lambda_{\text{ex}} = 331 \text{ nm}$ ).



Calculated using 3 exponentials

Prompt data : Prompt  
Decay data : Decay

The initial parameters are:

```

Shift Value = 0      ch;      0      sec
Shift Limit = 40     ch;     2.194787E-09  sec

T1 Estimate = 17.08746 ch;   9.375838E-10  sec
T2 Estimate = 34.17493 ch;   1.875168E-09  sec
T3 Estimate = 68.34985 ch;   3.750335E-09  sec

```

```

A Free
B1 Free
B2 Free
B3 Free

```

```

Prompt and decay LO = 721      ch;    3.956104E-08  sec
Prompt and decay HI = 1931     ch;    1.059534E-07  sec

```

```

Background on prompt = 0
Time calibration = 5.486969E-11 sec/ch

```

The fitted parameters are:

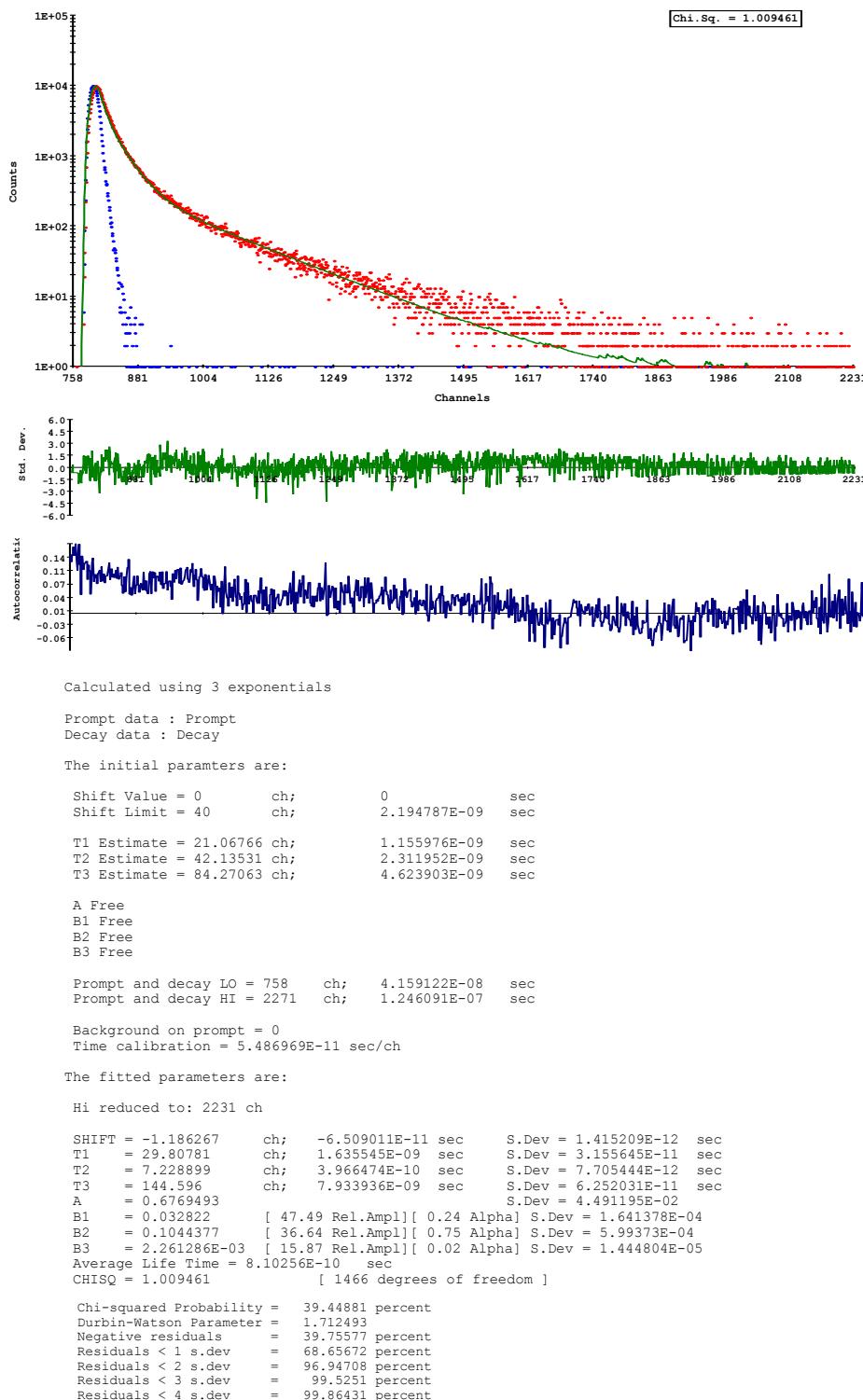
Hi reduced to: 1891 ch

```

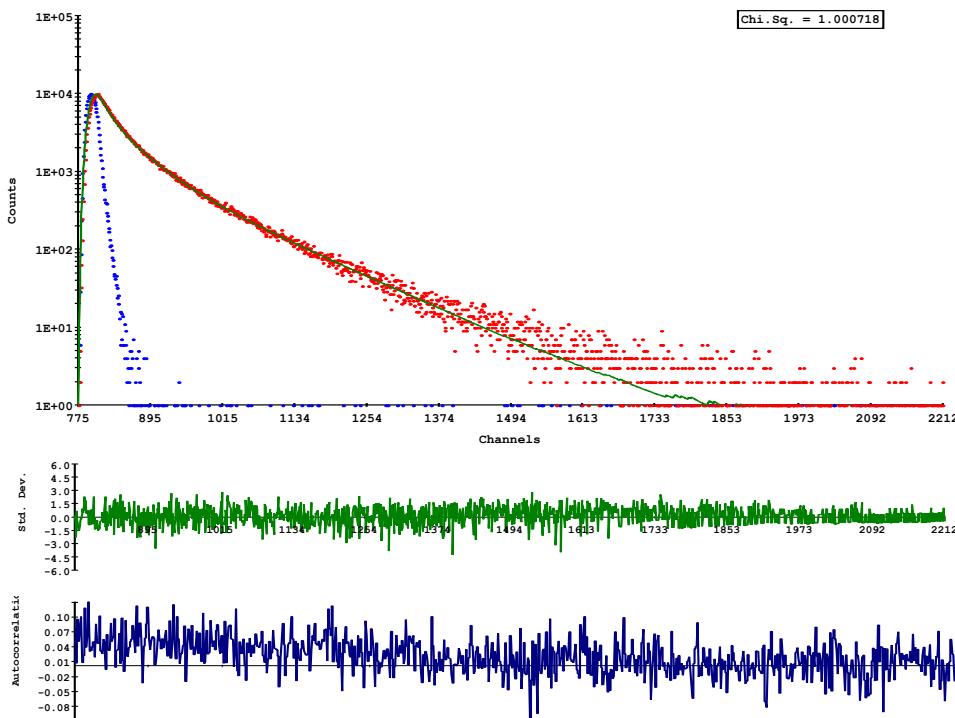
SHIFT = -0.6023685      ch;    -3.305177E-11  sec    S.Dev = 1.543106E-12  sec
T1    = 21.45963      ch;    1.177483E-09  sec    S.Dev = 2.105889E-11  sec
T2    = 2.673482      ch;    1.466931E-10  sec    S.Dev = 7.561361E-12  sec
T3    = 106.3772      ch;    5.836883E-09  sec    S.Dev = 3.693903E-11  sec
A     = 0.3993898
B1   = 0.033521      [ 37.34 Rel.Ampl][ 0.10 Alpha] S.Dev = 1.980716E-04
B2   = 0.3015539      [ 41.84 Rel.Ampl][ 0.89 Alpha] S.Dev = 1.562831E-03
B3   = 3.771347E-03      [ 20.82 Rel.Ampl][ 0.01 Alpha] S.Dev = 2.072202E-05
Average Life Time = 3.119976E-10  sec
CHISQ = 1.002543      [ 1163 degrees of freedom ]
Chi-squared Probability = 47.00817 percent
Durbin-Watson Parameter = 1.745352
Negative residuals = 44.1503 percent
Residuals < 1 s.dev = 70.70879 percent
Residuals < 2 s.dev = 95.13236 percent
Residuals < 3 s.dev = 99.40222 percent
Residuals < 4 s.dev = 100 percent

```

**Fig. S22.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  TQTACN in the presence of 2 equiv. of  $\text{Zn}^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 400 nm at 25 °C ( $\lambda_{\text{ex}} = 331 \text{ nm}$ ).



**Fig. S23.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  TQTACN in the presence of 2 equiv. of  $\text{Cd}^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 430 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



Calculated using 3 exponentials

Prompt data : Prompt  
Decay data : Decay

The initial parameters are:

Shift Value = 0	ch;	0	sec
Shift Limit = 40	ch;	2.194787E-09	sec
T1 Estimate = 32.99875	ch;	1.810631E-09	sec
T2 Estimate = 65.9975	ch;	3.621262E-09	sec
T3 Estimate = 131.995	ch;	7.242524E-09	sec

A Free  
B1 Free  
B2 Free  
B3 Free

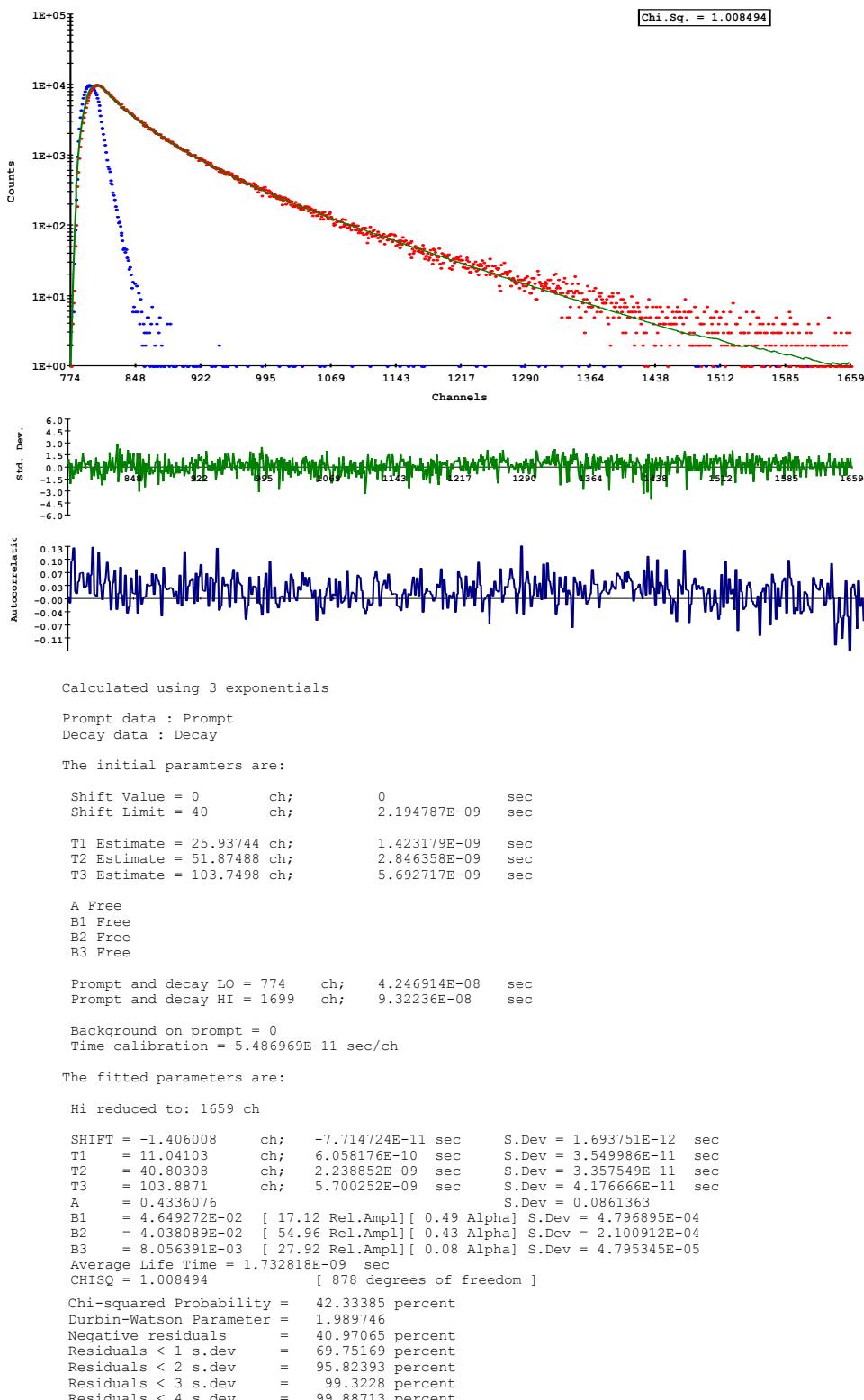
Prompt and decay LO = 775 ch; 4.252401E-08 sec  
Prompt and decay HI = 2252 ch; 1.235665E-07 sec

Background on prompt = 0  
Time calibration = 5.486969E-11 sec/ch

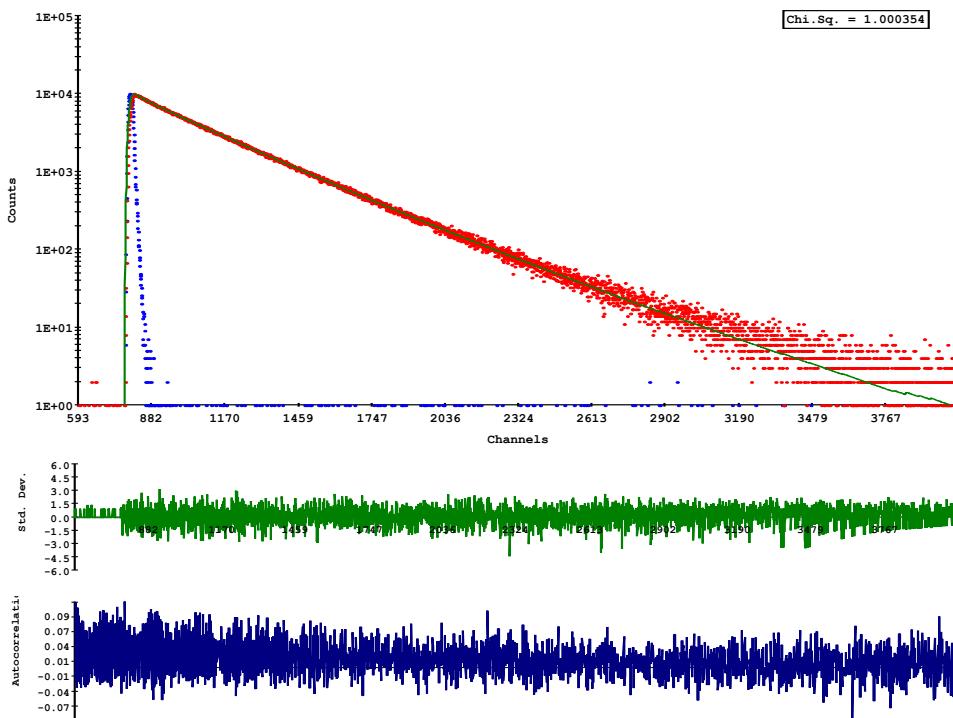
The fitted parameters are:

Hi reduced to: 2212 ch			
SHIFT = -0.9908015	ch;	-5.436497E-11 sec	S.Dev = 1.511775E-12 sec
T1 = 46.5347	ch;	2.553344E-09 sec	S.Dev = 1.03436E-10 sec
T2 = 15.67823	ch;	8.602596E-10 sec	S.Dev = 1.686203E-11 sec
T3 = 125.3808	ch;	6.879606E-09 sec	S.Dev = 3.646886E-11 sec
A = 0.3617765			S.Dev = 0.0455973
B1 = 2.860537E-02	[ 42.03 Rel.Ampl] [ 0.33 Alpha]	S.Dev = 2.005311E-04	
B2 = 4.895177E-02	[ 24.23 Rel.Ampl] [ 0.57 Alpha]	S.Dev = 3.972751E-04	
B3 = 8.52488E-03	[ 33.74 Rel.Ampl] [ 0.10 Alpha]	S.Dev = 4.077398E-05	
Average Life Time = 2.018986E-09 sec			
CHISQ = 1.000718 [ 1430 degrees of freedom ]			
Chi-squared Probability = 48.73662 percent			
Durbin-Watson Parameter = 1.857808			
Negative residuals = 43.0459 percent			
Residuals < 1 s.dev = 69.19332 percent			
Residuals < 2 s.dev = 96.17525 percent			
Residuals < 3 s.dev = 99.6523 percent			
Residuals < 4 s.dev = 99.93046 percent			

**Fig. S24.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  6-MeOTQTACN in the presence of 2 equiv. of  $\text{Zn}^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 430 nm at 25 °C ( $\lambda_{\text{ex}} = 331 \text{ nm}$ ).



**Fig. S25.** Fluorescent lifetime measurement of 34  $\mu$ M 6-MeOTQTACN in the presence of 2 equiv. of Cd<sup>2+</sup> in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 400 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



Calculated using 3 exponentials

Prompt data : Prompt  
Decay data : Decay

The initial parameters are:

Shift Value = 0	ch;	0	sec
Shift Limit = 40	ch;	2.194787E-09	sec
T1 Estimate = 148.943	ch;	8.172456E-09	sec
T2 Estimate = 297.886	ch;	1.634491E-08	sec
T3 Estimate = 595.7721	ch;	3.268983E-08	sec

A Free  
B1 Free  
B2 Free  
B3 Free

Prompt and decay LO = 593 ch; 3.253772E-08 sec  
Prompt and decay HI = 4096 ch; 2.247462E-07 sec

Background on prompt = 0  
Time calibration = 5.486969E-11 sec/ch

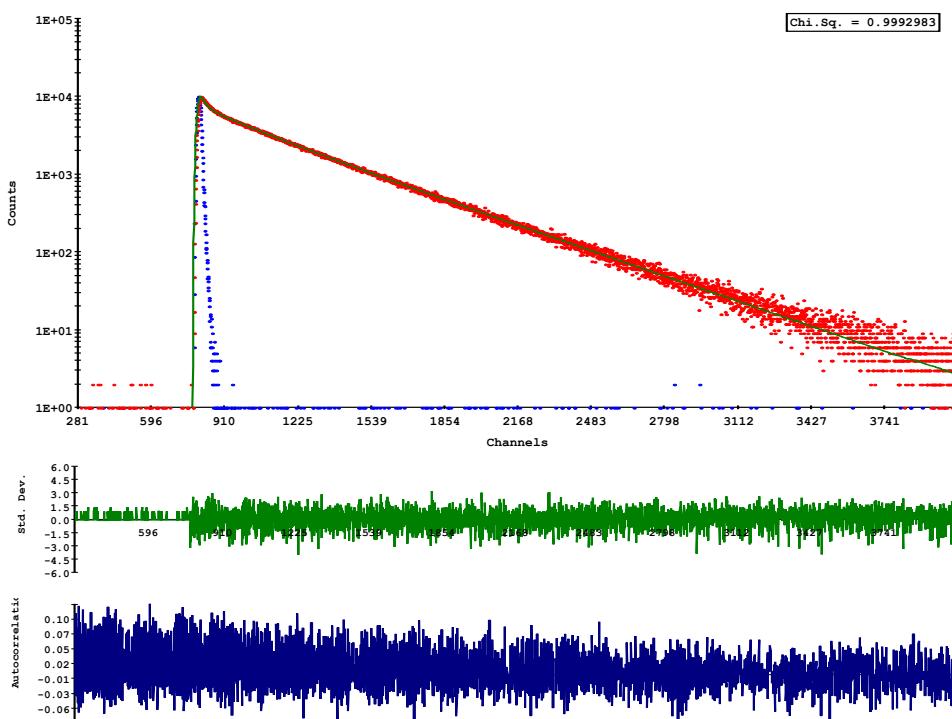
The fitted parameters are:

Hi reduced to: 4056 ch

SHIFT = -1.606446	ch;	-8.814519E-11	sec	S.Dev = 1.868659E-12	sec
T1 = 33.47426	ch;	1.836722E-09	sec	S.Dev = 2.101768E-10	sec
T2 = 262.092	ch;	1.43809E-08	sec	S.Dev = 2.093338E-10	sec
T3 = 389.0543	ch;	2.134728E-08	sec	S.Dev = 1.200582E-10	sec
A = 7.716215E-02				S.Dev = 4.246895E-02	
B1 = 5.068578E-03	[ 1.14 Rel.Ampl] [ 0.09 Alpha]	S.Dev = 1.818532E-04			
B2 = 3.668423E-02	[ 64.56 Rel.Ampl] [ 0.67 Alpha]	S.Dev = 1.115349E-04			
B3 = 1.313167E-02	[ 34.30 Rel.Ampl] [ 0.24 Alpha]	S.Dev = 6.582256E-05			
Average Life Time = 1.488923E-08 sec					
CHISQ = 1.000354 [ 3456 degrees of freedom ]					

Chi-squared Probability =	49.09356 percent
Durbin-Watson Parameter =	2.028572
Negative residuals =	44.97691 percent
Residuals < 1 s.dev =	67.20554 percent
Residuals < 2 s.dev =	95.95843 percent
Residuals < 3 s.dev =	99.62471 percent
Residuals < 4 s.dev =	99.94226 percent

**Fig. S26.** Fluorescent lifetime measurement of 34  $\mu$ M TriMeOTQTACN in the presence of 1 equiv. of  $Zn^{2+}$  in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 490 nm at 25 °C ( $\lambda_{ex}$  = 331 nm).



Calculated using 3 exponentials

Prompt data : Prompt  
Decay data : Decay

The initial parameters are:

Shift Value = 0	ch;	0	sec
Shift Limit = 40	ch;	2.194787E-09	sec
T1 Estimate = 185.0465	ch;	1.015344E-08	sec
T2 Estimate = 370.093	ch;	2.030689E-08	sec
T3 Estimate = 740.186	ch;	4.061377E-08	sec

A Free  
B1 Free  
B2 Free  
B3 Free

Prompt and decay LO = 281 ch; 1.541838E-08 sec  
Prompt and decay HI = 4096 ch; 2.247462E-07 sec

Background on prompt = 2.024292E-03  
Time calibration = 5.486969E-11 sec/ch

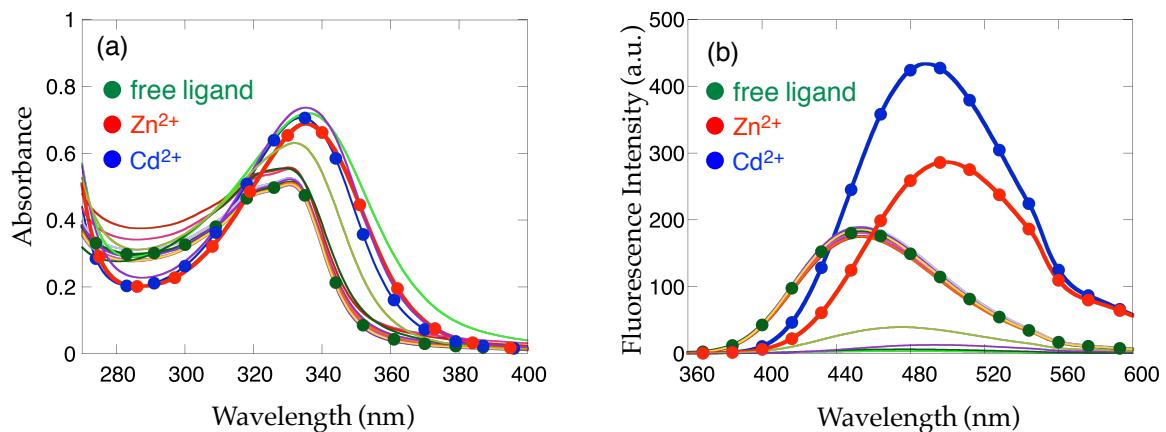
The fitted parameters are:

Hi reduced to: 4056 ch

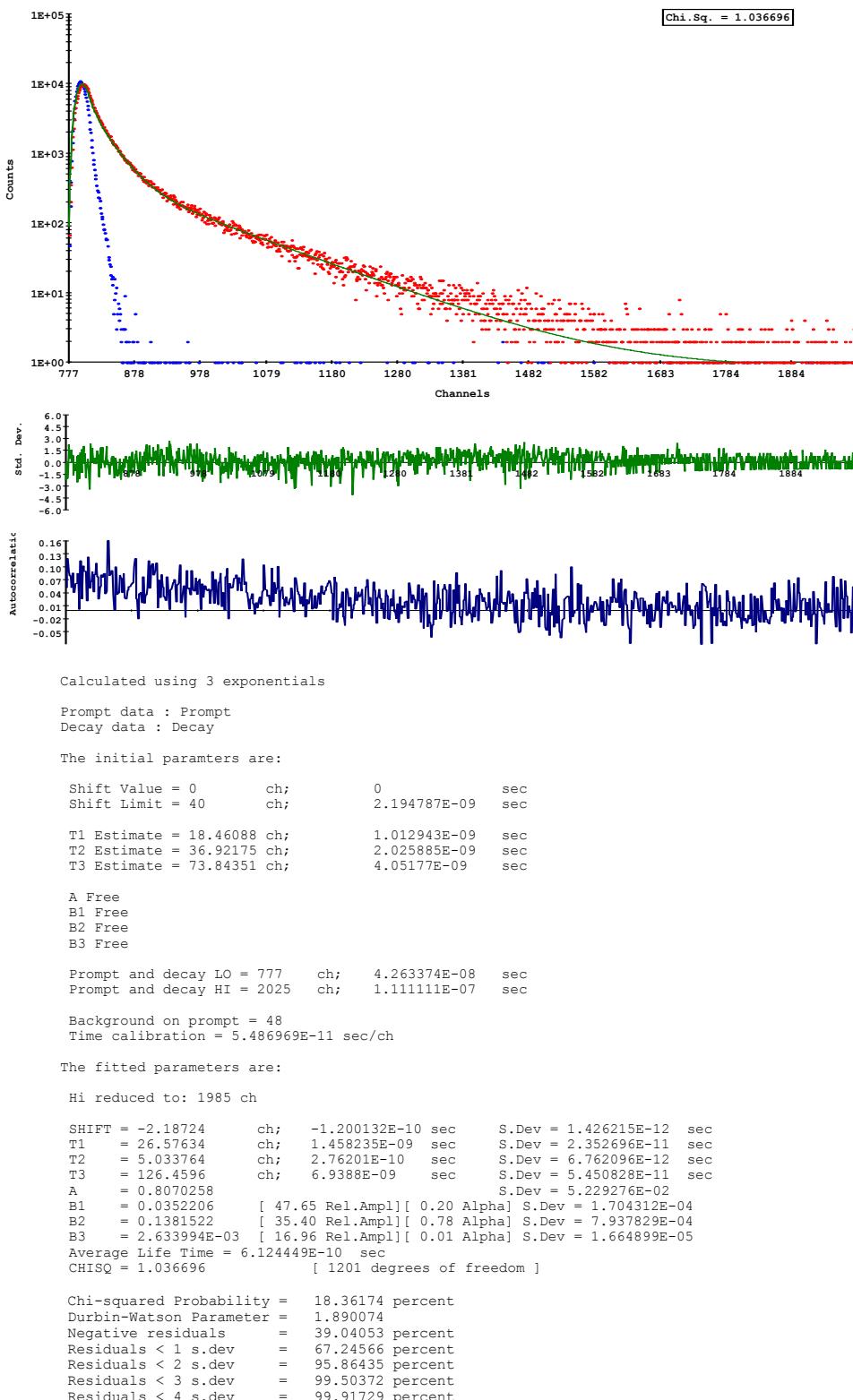
SHIFT = -2.160237	ch;	-1.185315E-10	sec	S.Dev = 1.586205E-12	sec
T1 = 201.0937	ch;	1.103395E-08	sec	S.Dev = 4.531502E-10	sec
T2 = 416.4765	ch;	2.285193E-08	sec	S.Dev = 3.739936E-11	sec
T3 = 22.18578	ch;	1.217327E-09	sec	S.Dev = 2.362652E-11	sec
A = 8.421974E-02				S.Dev = 3.938364E-02	
B1 = 8.697198E-03	[ 12.05 Rel.Ampl]	[ 0.13 Alpha]	S.Dev = 9.170287E-05		
B2 = 2.913807E-02	[ 83.58 Rel.Ampl]	[ 0.44 Alpha]	S.Dev = 4.240498E-05		
B3 = 2.864082E-02	[ 4.38 Rel.Ampl]	[ 0.43 Alpha]	S.Dev = 2.284693E-04		
Average Life Time = 1.198462E-08 sec					
CHISQ = 0.9992983 [ 3768 degrees of freedom ]					

Chi-squared Probability =	50.9089 percent
Durbin-Watson Parameter =	2.066888
Negative residuals =	48.99364 percent
Residuals < 1 s.dev =	70.23305 percent
Residuals < 2 s.dev =	94.86229 percent
Residuals < 3 s.dev =	99.44386 percent
Residuals < 4 s.dev =	99.97352 percent

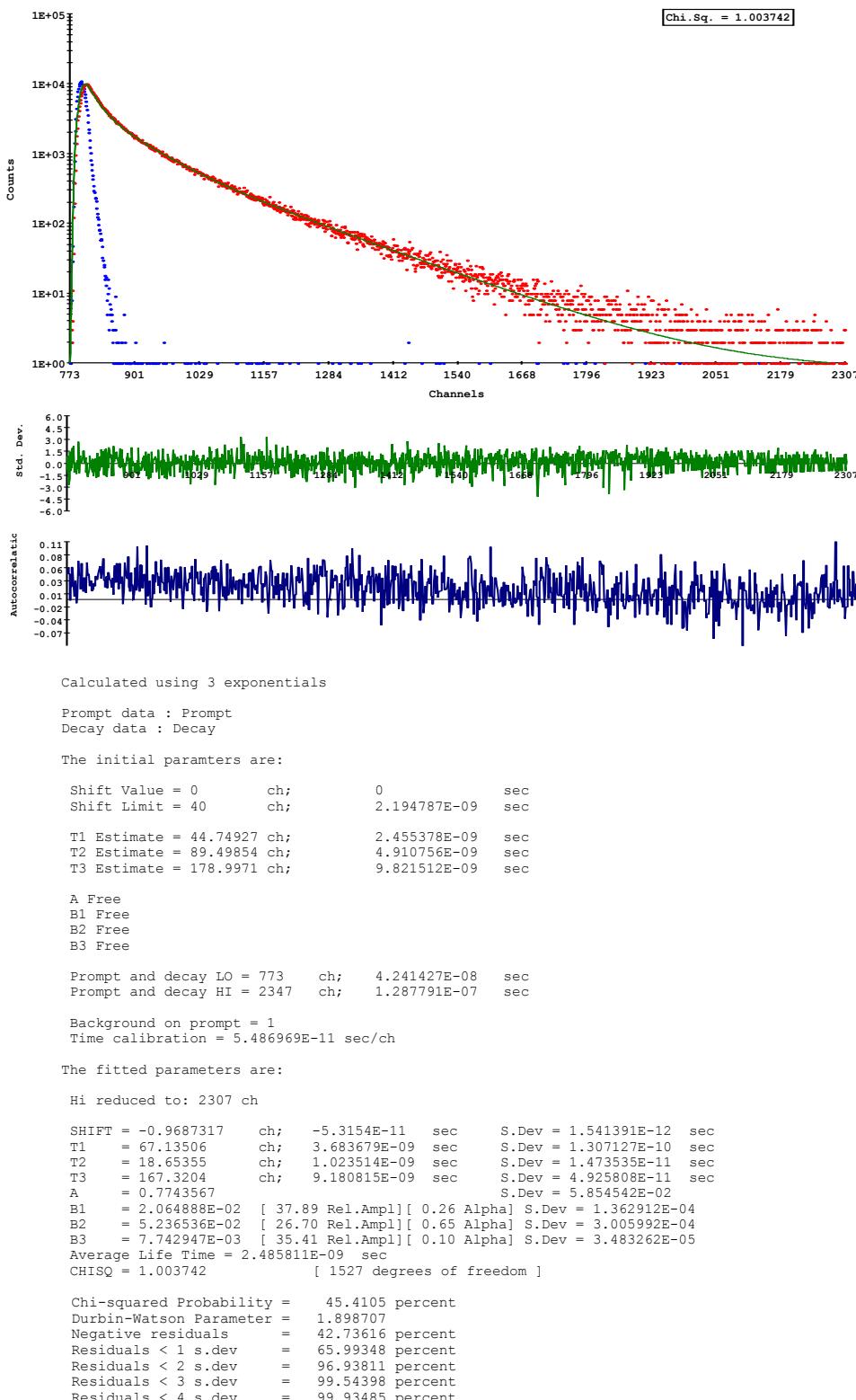
**Fig. S27.** Fluorescent lifetime measurement of 34  $\mu$ M TriMeOTQTACN in the presence of 1 equiv. of Cd<sup>2+</sup> in methanol-HEPES buffer (9:1, 50 mM HEPES, 0.1 M KCl, pH = 7.5) at 490 nm at 25 °C ( $\lambda_{ex}$  = 331 nm).



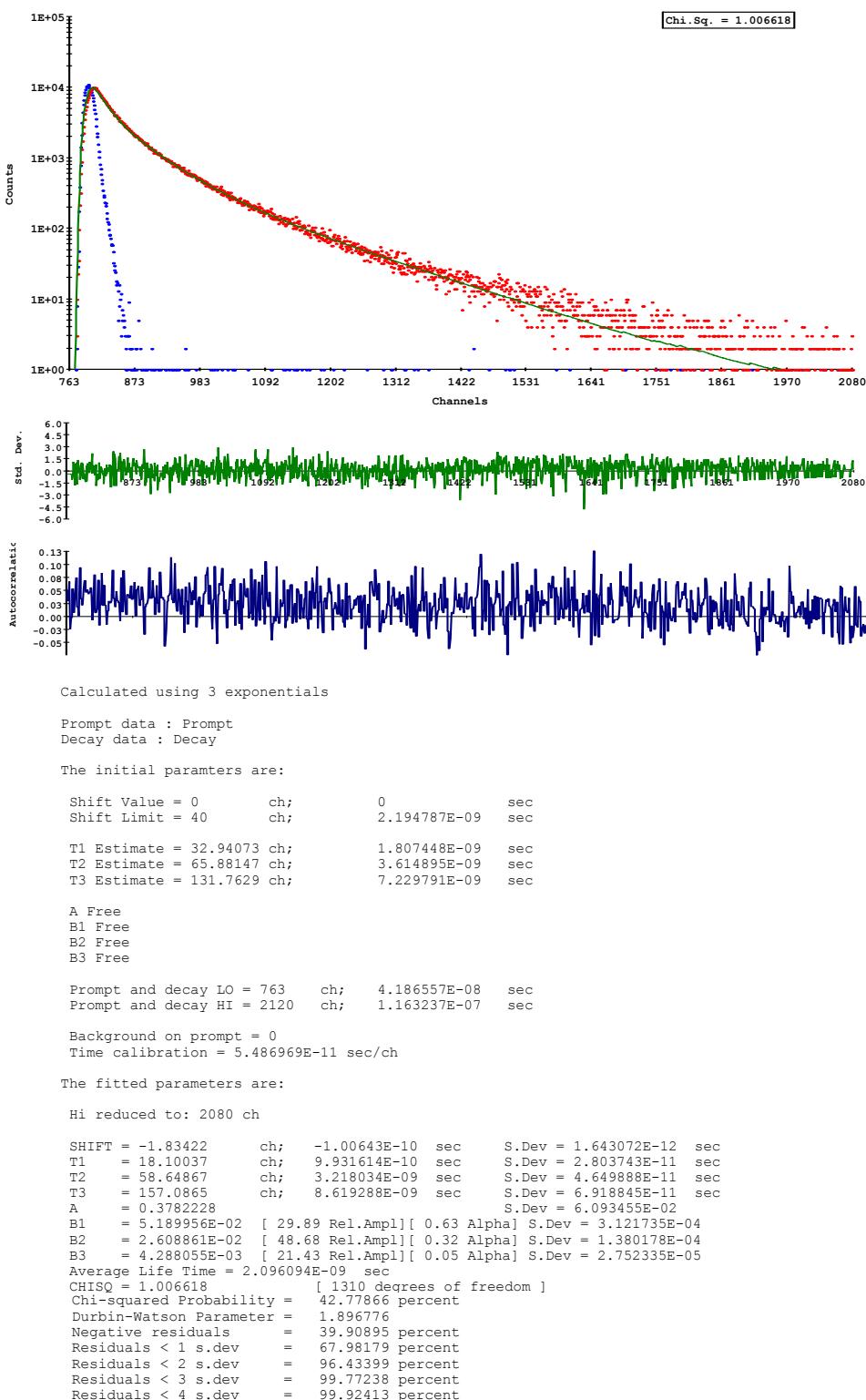
**Fig. S28.** (a) UV-vis absorption and (b) fluorescence spectra of 34  $\mu M$  TriMeOTQTACN in DMF-H<sub>2</sub>O (1:1) at 25 °C in the presence of 2 equiv. of various metal ions ( $\lambda_{ex} = 343$  nm).



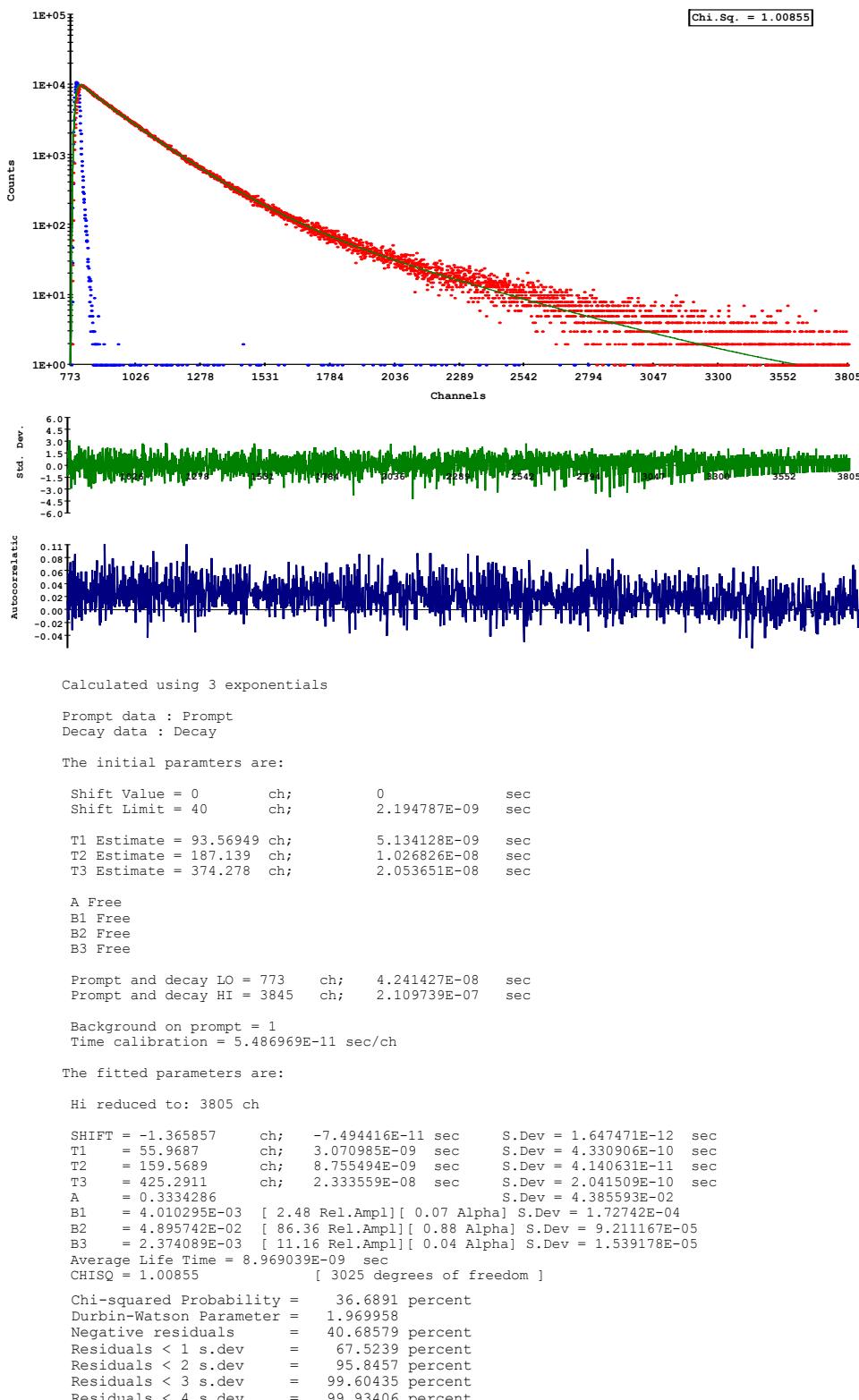
**Fig. S29.** Fluorescent lifetime measurement of 34  $\mu$ M TQTACN in the presence of 2 equiv. of Zn<sup>2+</sup> in DMF-H<sub>2</sub>O (1:1) at 400 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



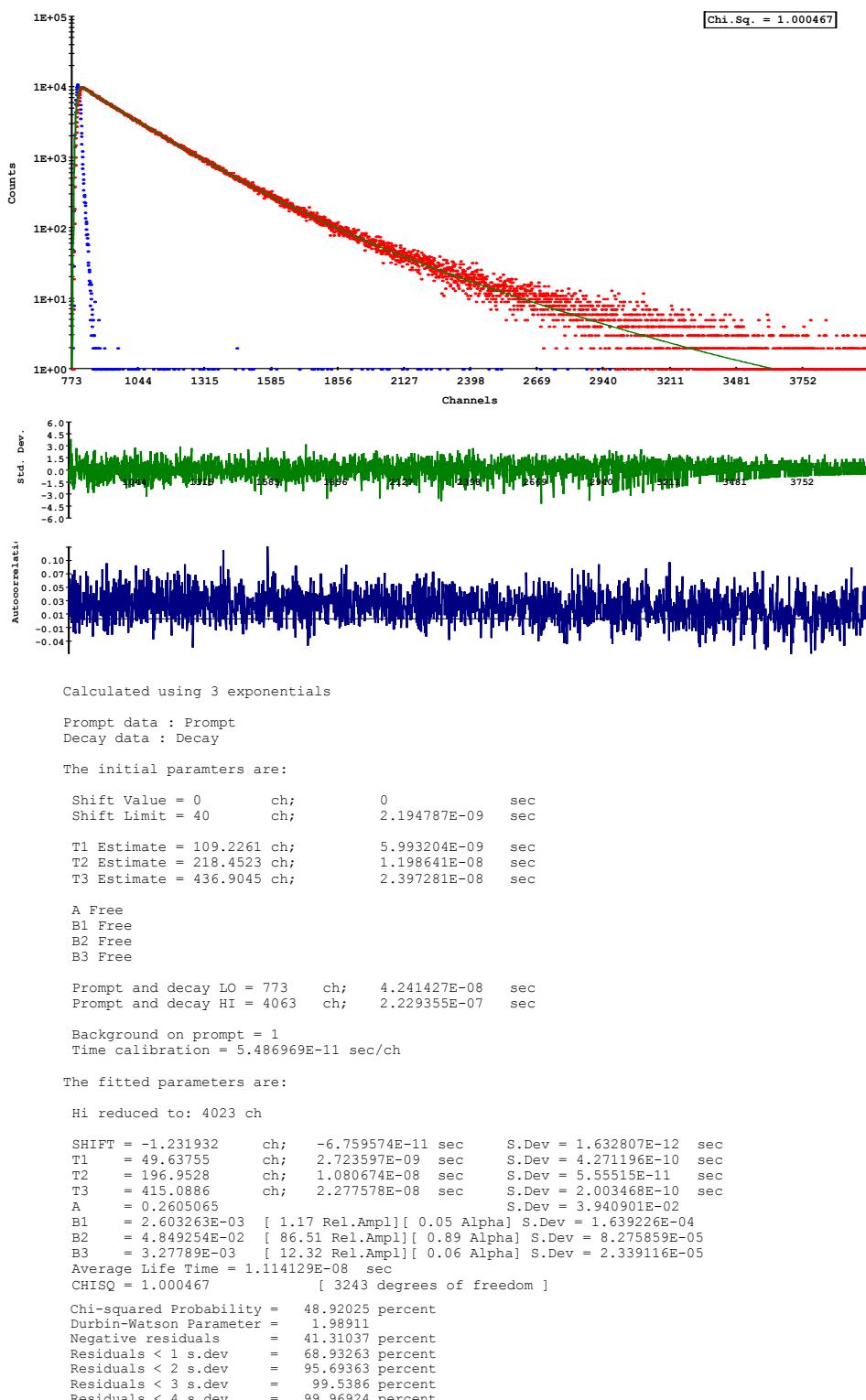
**Fig. S30.** Fluorescent lifetime measurement of 34  $\mu$ M 6-MeOTQTACN in the presence of 2 equiv. of  $Zn^{2+}$  in DMF-H<sub>2</sub>O (1:1) at 430 nm at 25 °C ( $\lambda_{ex} = 331$  nm).



**Fig. S31.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  6-MeOTQTACN in the presence of 2 equiv. of  $\text{Cd}^{2+}$  in DMF- $\text{H}_2\text{O}$  (1:1) at 430 nm at 25 °C ( $\lambda_{\text{ex}} = 331 \text{ nm}$ ).



**Fig. S32.** Fluorescent lifetime measurement of 34  $\mu\text{M}$  TriMeOTQTACN in the presence of 2 equiv. of  $\text{Zn}^{2+}$  in DMF- $\text{H}_2\text{O}$  (1:1) at 490 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).



**Fig. S33.** Fluorescent lifetime measurement of 34  $\mu$ M TriMeOTQTACN in the presence of 2 equiv. of Cd<sup>2+</sup> in DMF-H<sub>2</sub>O (1:1) at 490 nm at 25 °C ( $\lambda_{\text{ex}} = 331$  nm).

EGTQ

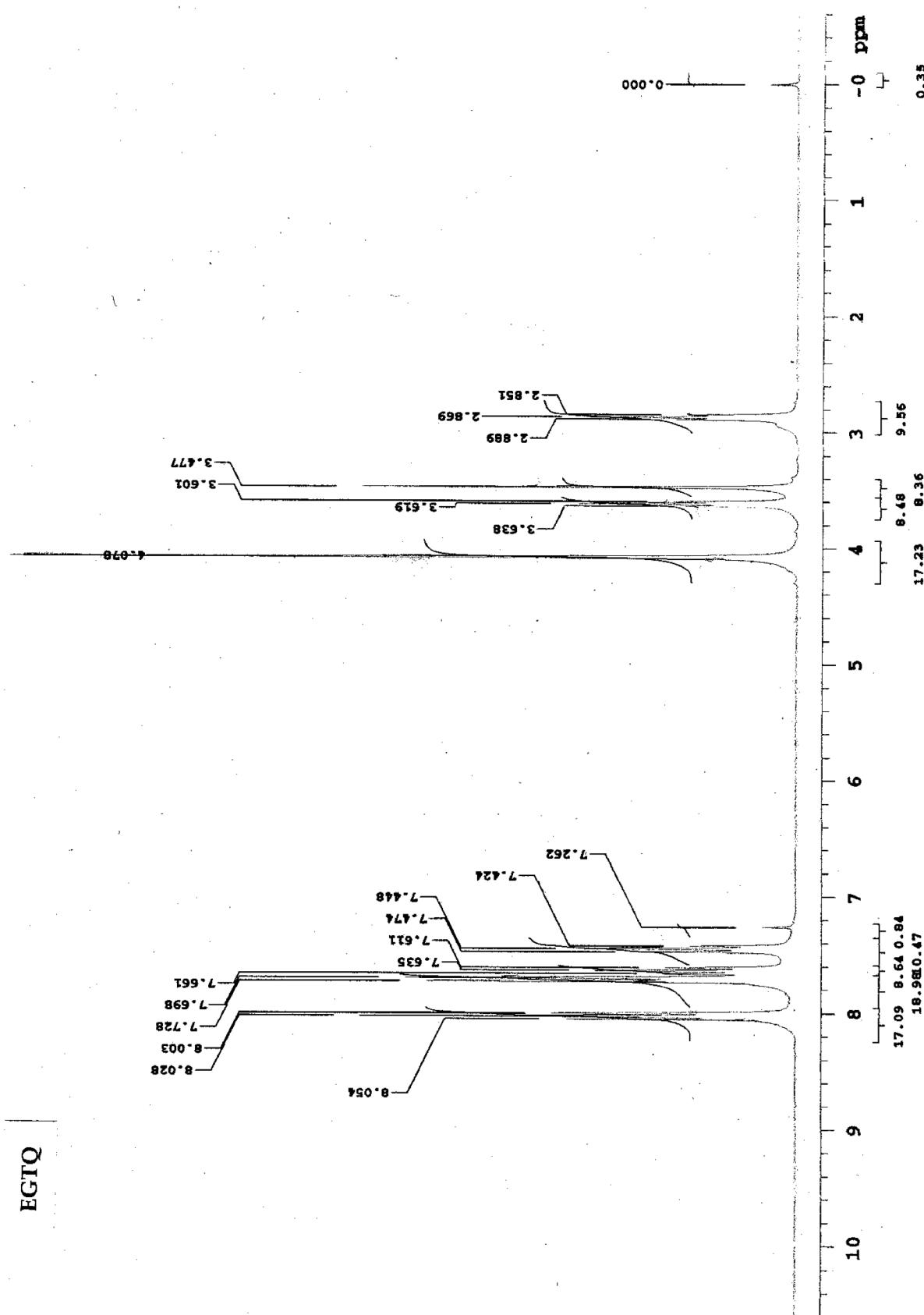


Fig. S34.  $^1\text{H}$  NMR spectrum for EGTQ in  $\text{CDCl}_3$ .

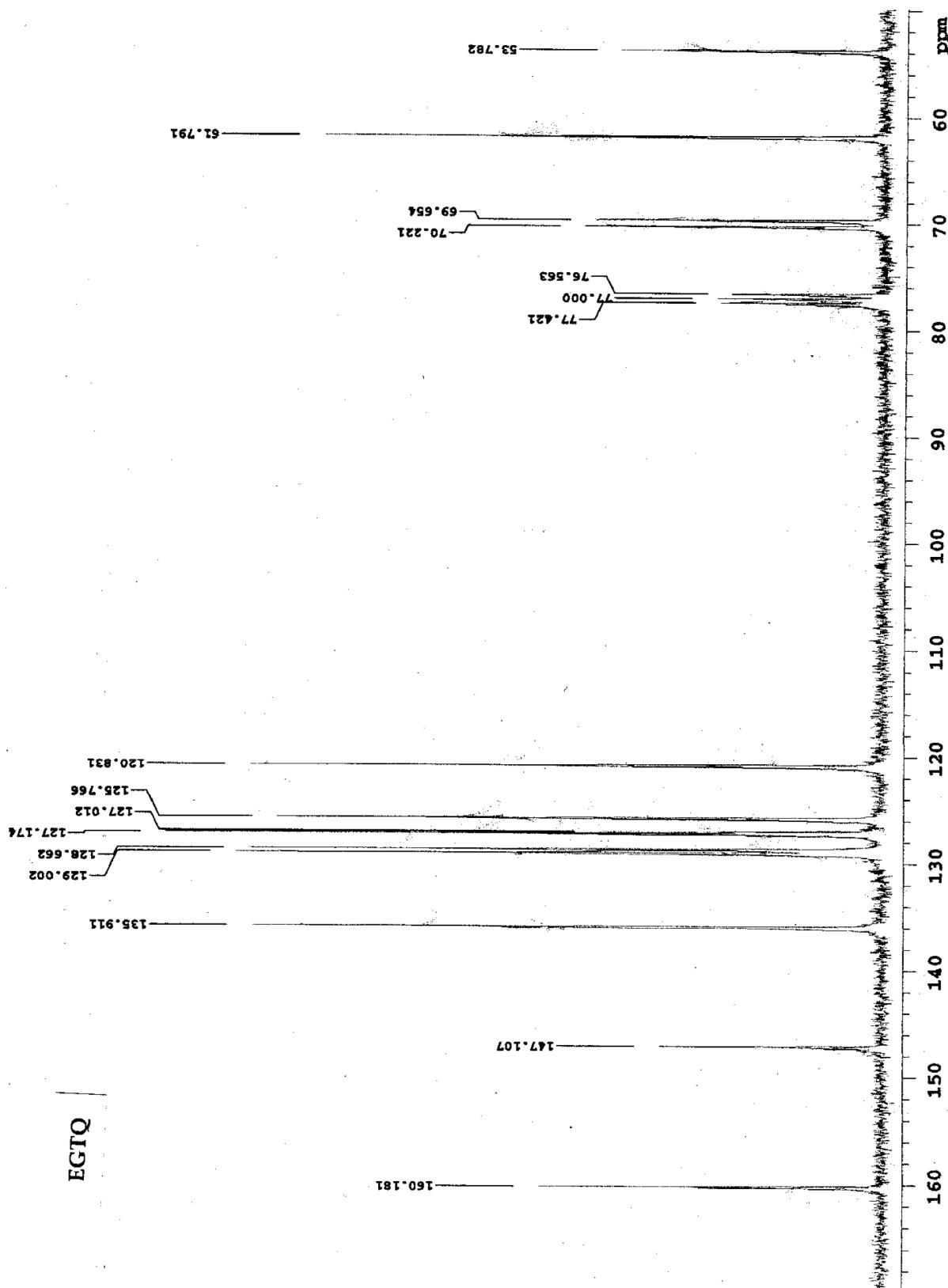
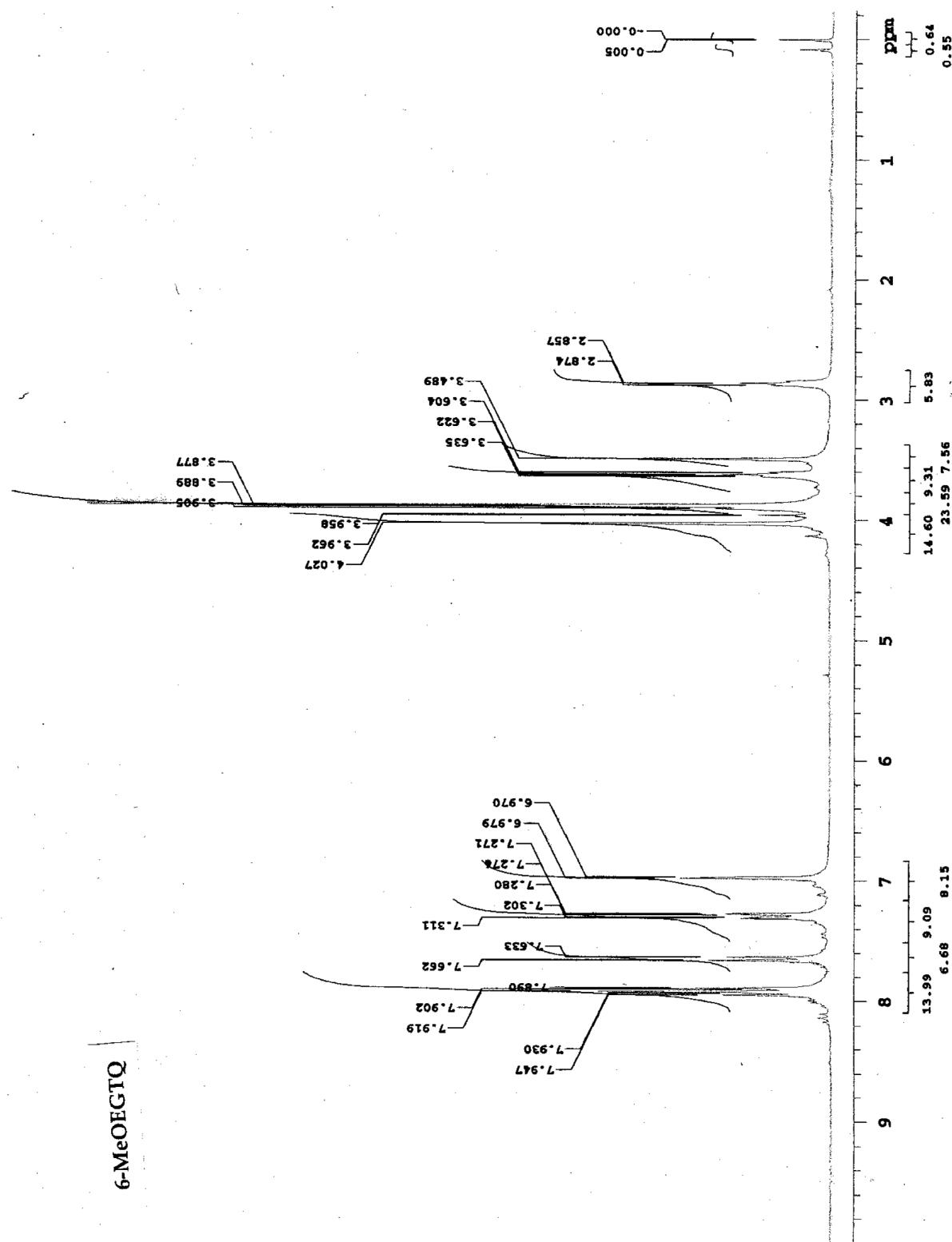
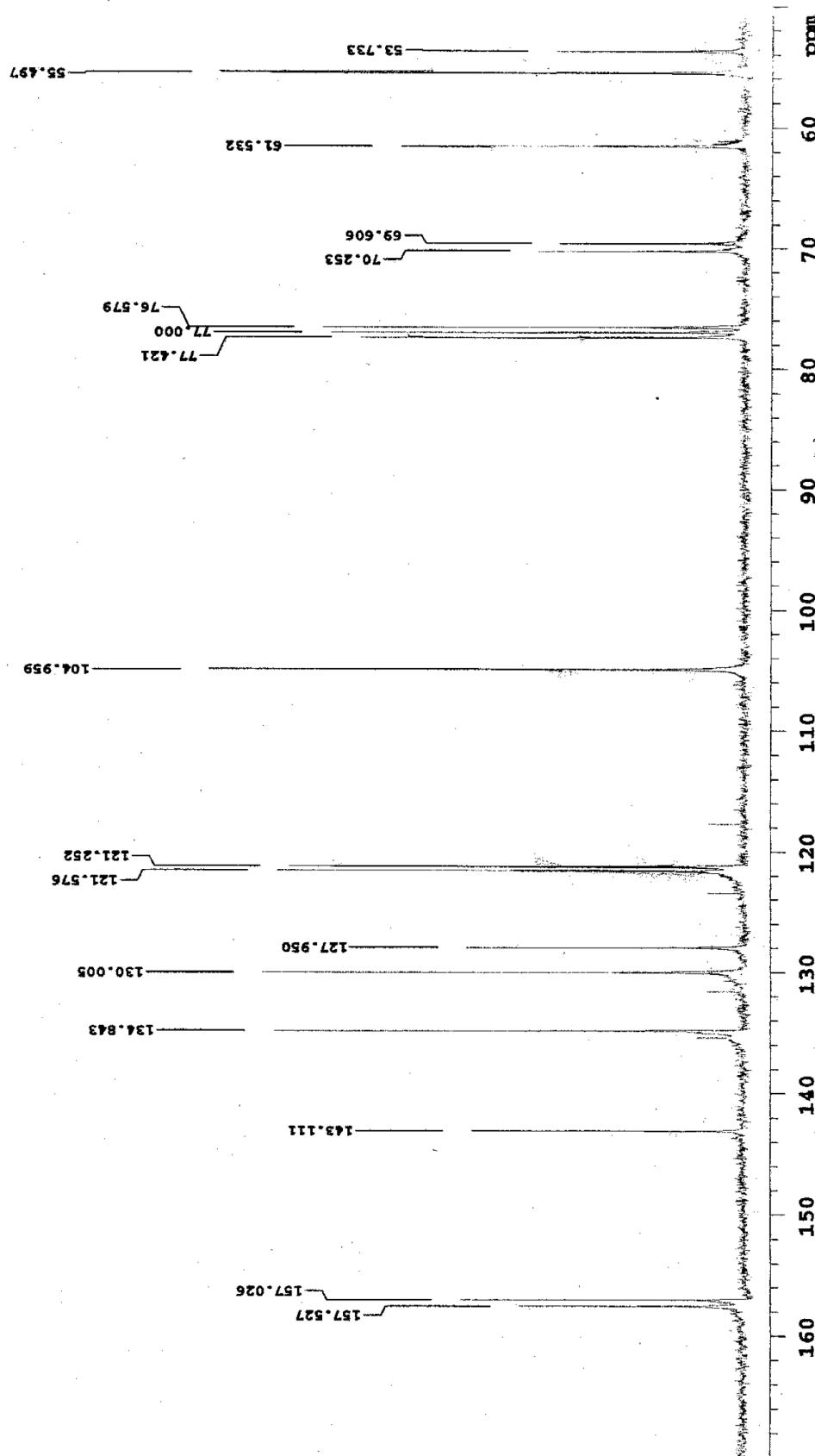


Fig. S35.  $^{13}\text{C}$  NMR spectrum for EGTQ in  $\text{CDCl}_3$ .



**Fig. S36.**  $^1\text{H}$  NMR spectrum for 6-MeOEGTQ in  $\text{CDCl}_3$ .

6-MeOEGTQ



**Fig. S37.**  $^{13}\text{C}$  NMR spectrum for 6-MeOEGTQ in  $\text{CDCl}_3$ .

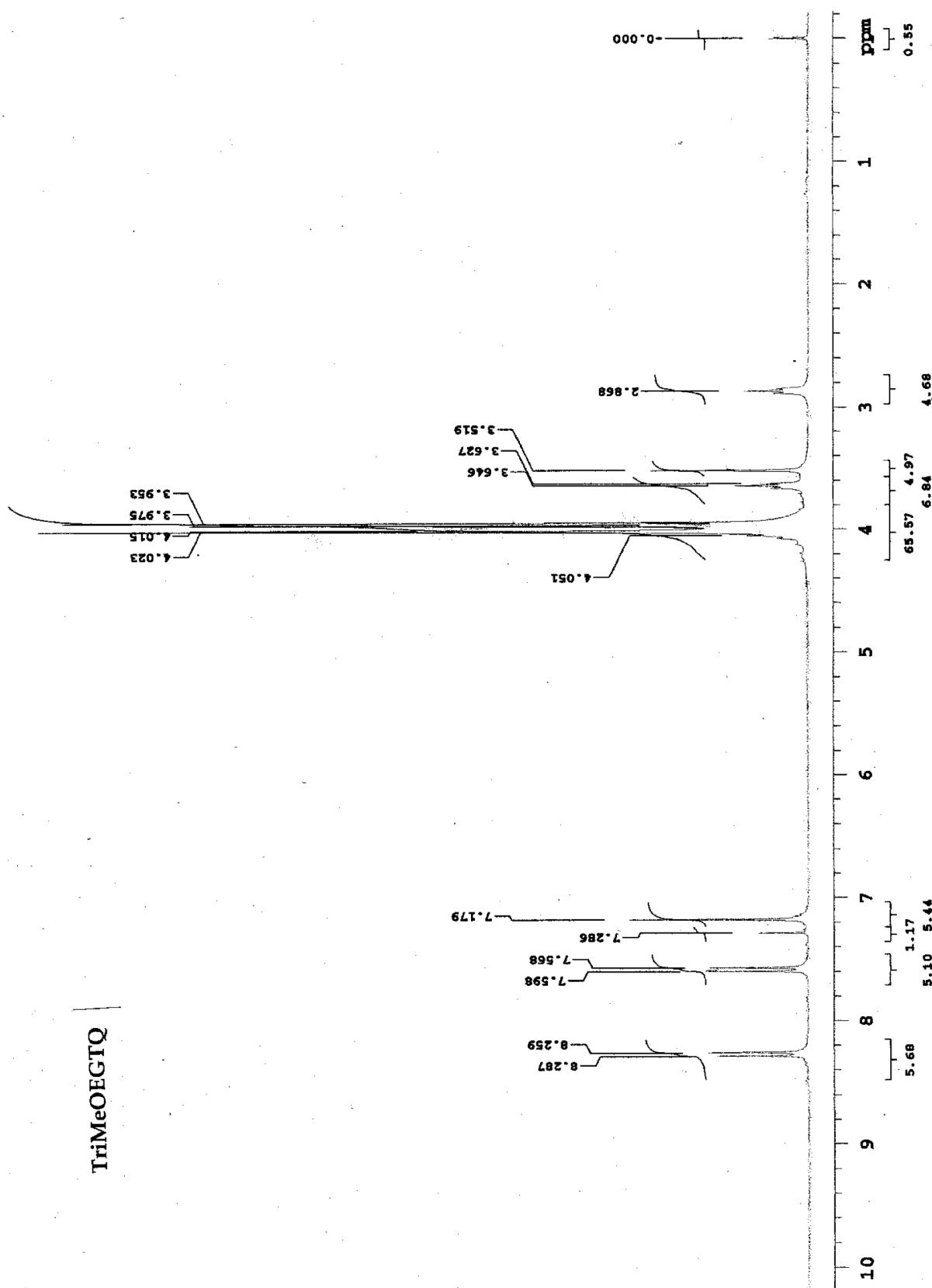


Fig. S38.  $^1\text{H}$  NMR spectrum for TriMeOEGTQ in  $\text{CDCl}_3$ .

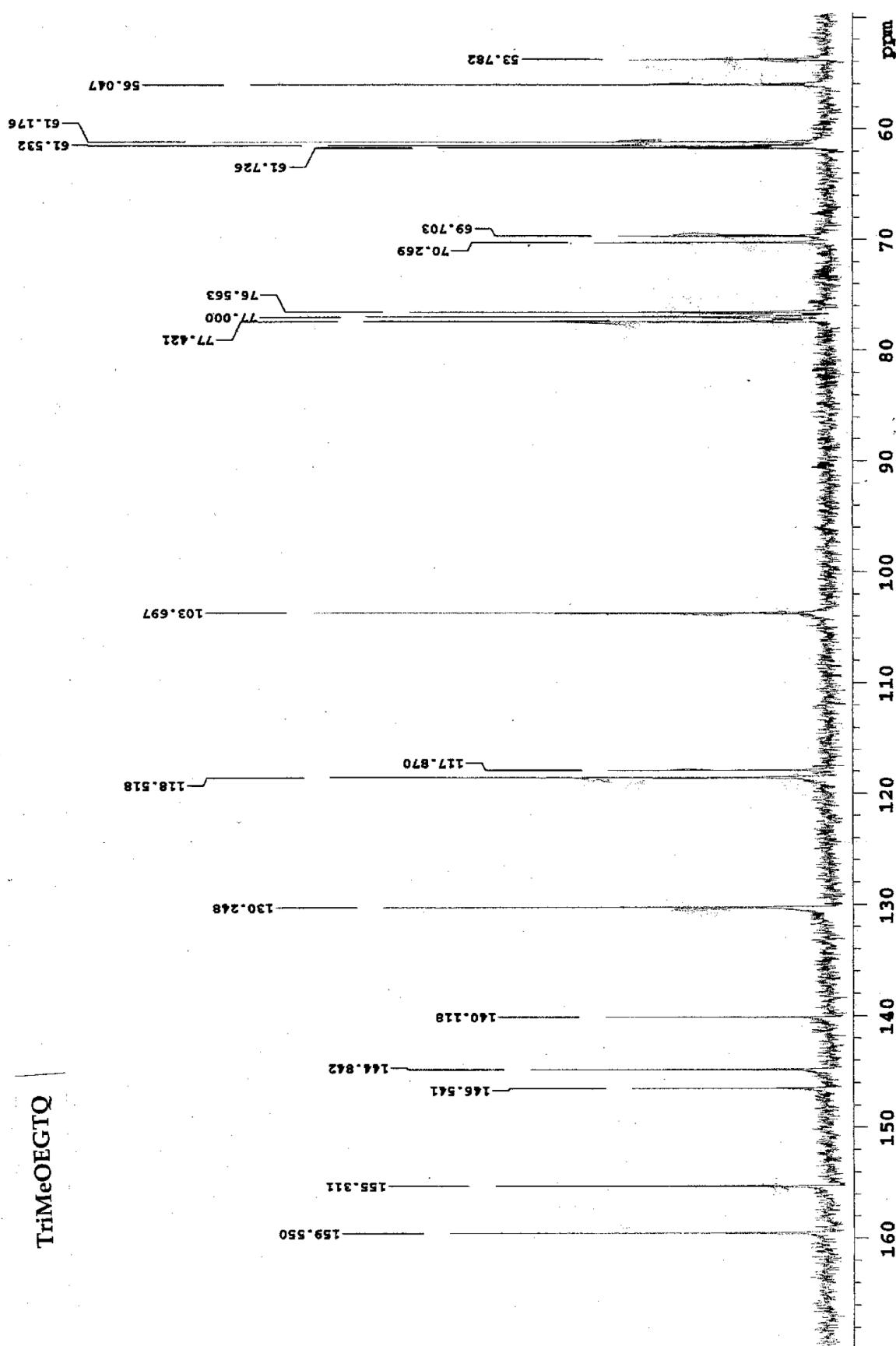
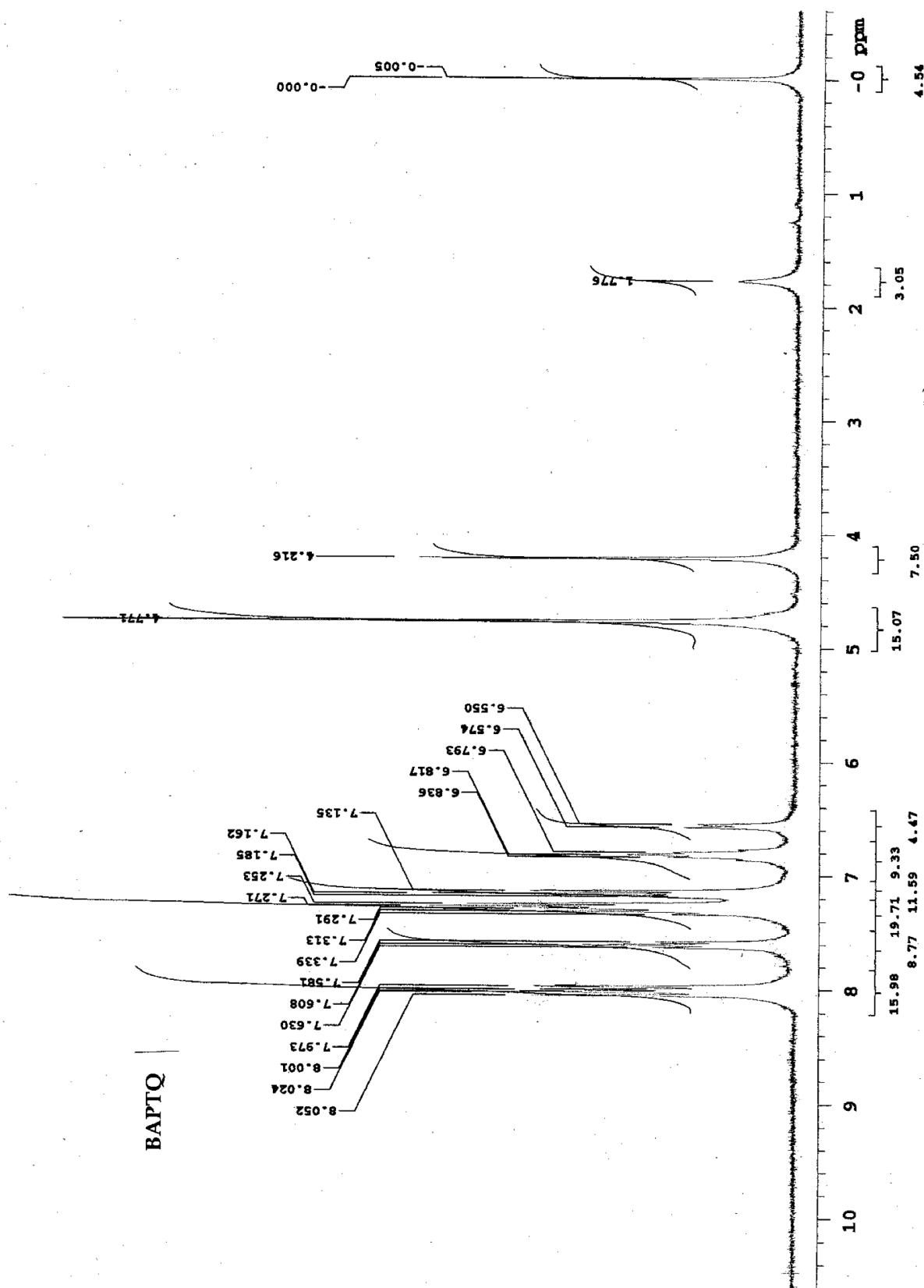


Fig. S39.  $^{13}\text{C}$  NMR spectrum for TriMeOEGTQ in  $\text{CDCl}_3$ .



**Fig. S40.**  $^1\text{H}$  NMR spectrum for BAPTQ in  $\text{CDCl}_3$ .

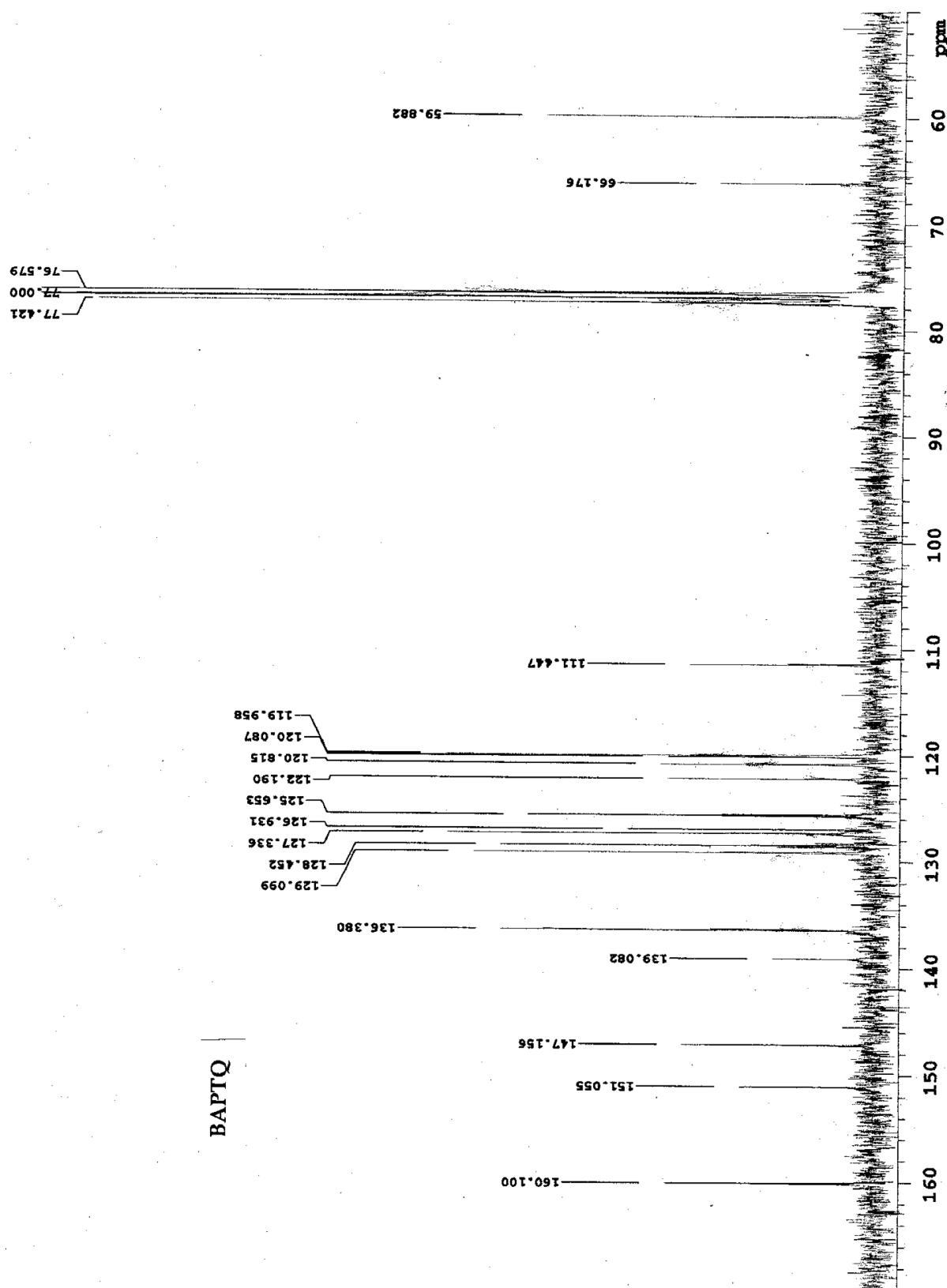
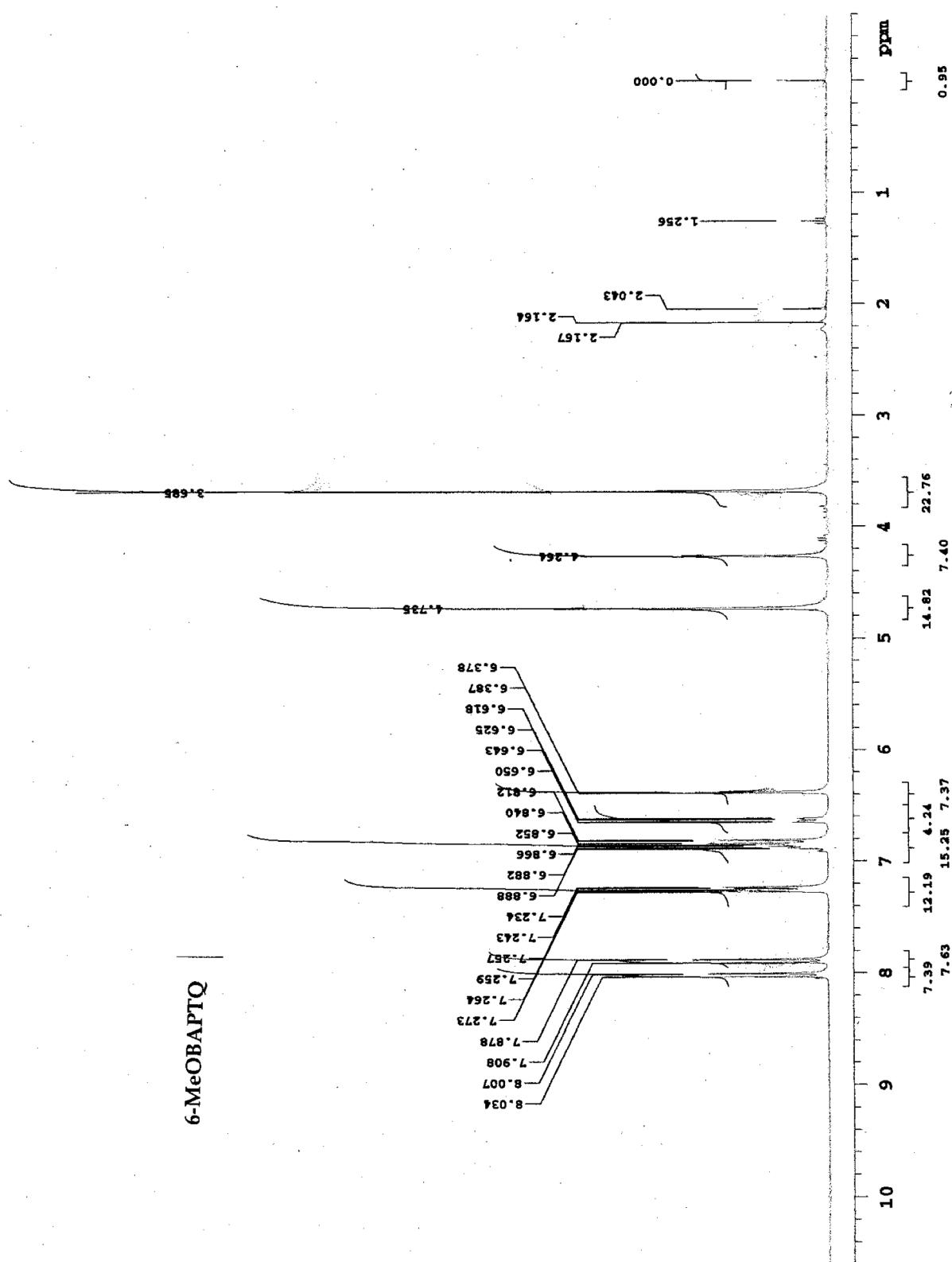


Fig. S41.  $^{13}\text{C}$  NMR spectrum for BAPTQ in  $\text{CDCl}_3$ .



**Fig. S42.**  $^1\text{H}$  NMR spectrum for 6-MeOBAPTQ in  $\text{CDCl}_3$ .

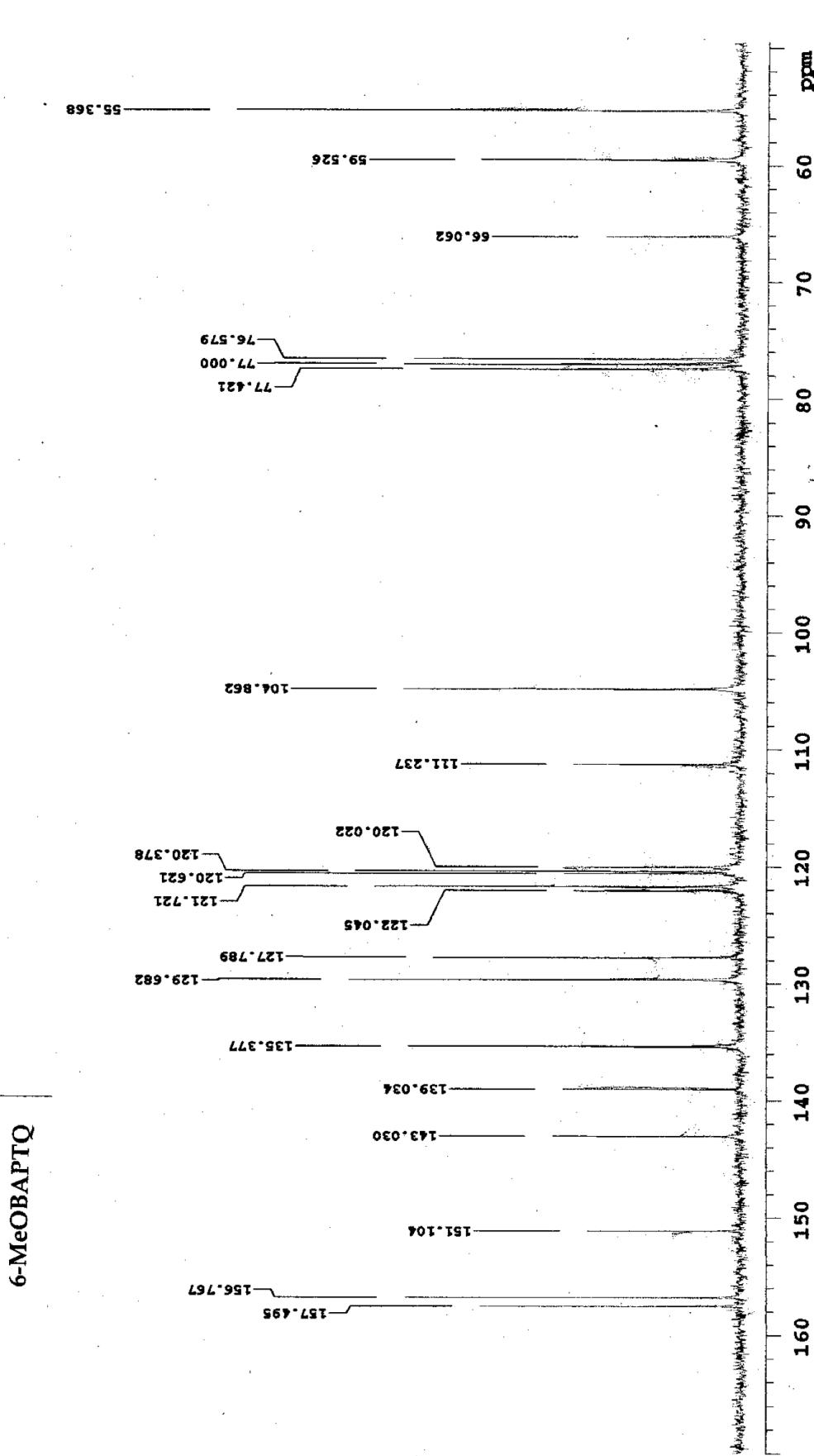
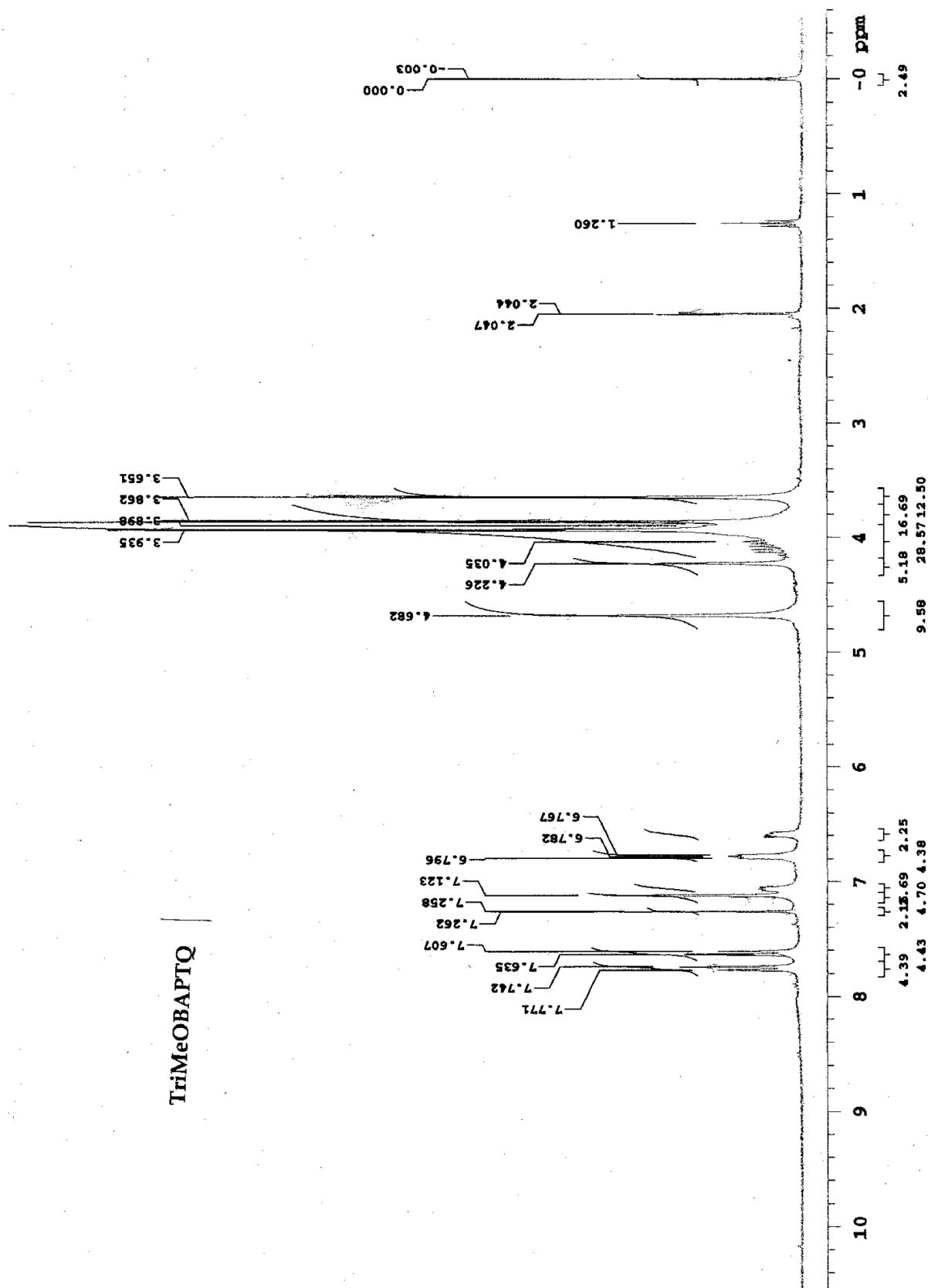


Fig. S43.  $^{13}\text{C}$  NMR spectrum for 6-MeOBAPTQ in  $\text{CDCl}_3$ .



**Fig. S44.** <sup>1</sup>H NMR spectrum for TriMeOBAPTQ in CDCl<sub>3</sub>.

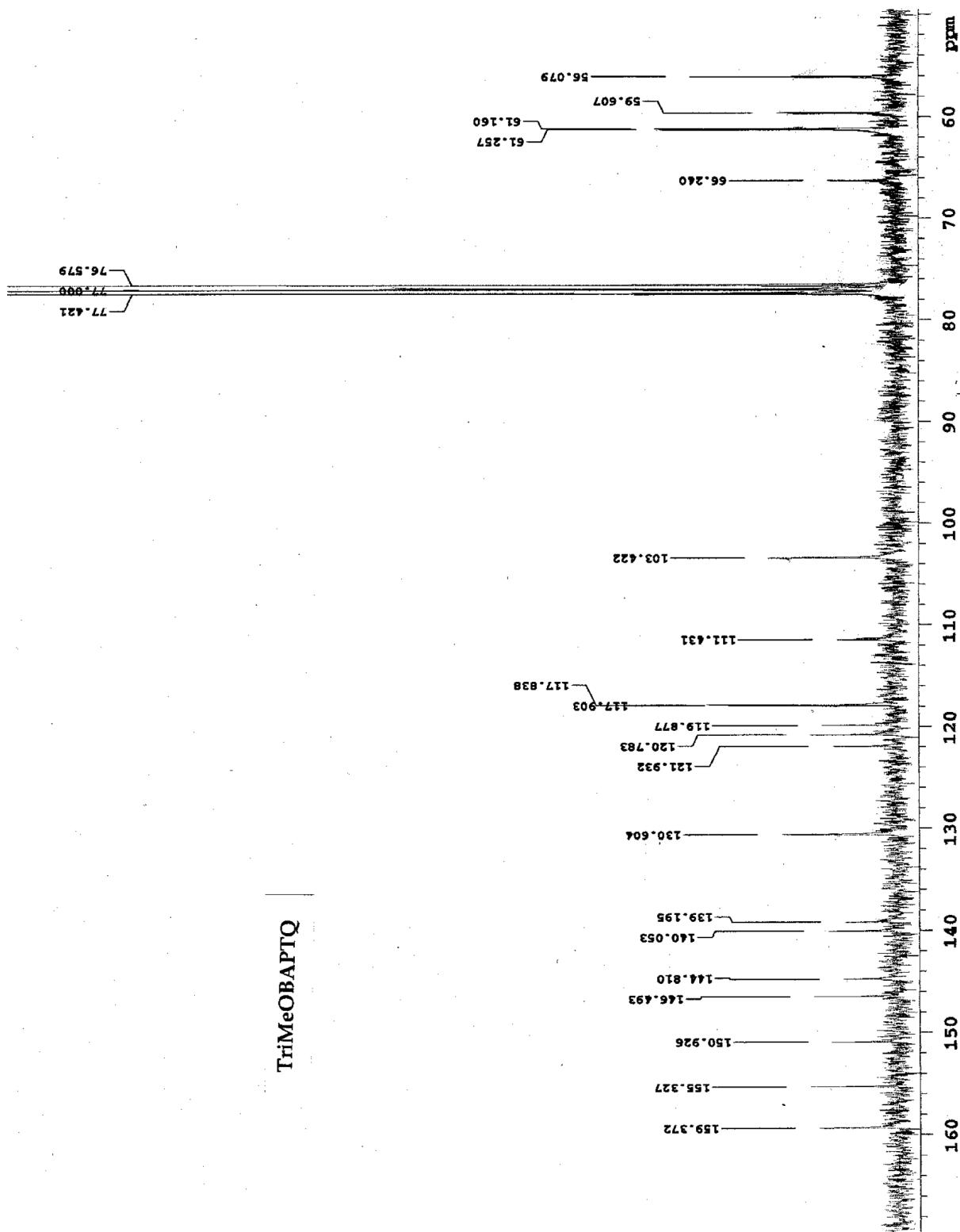
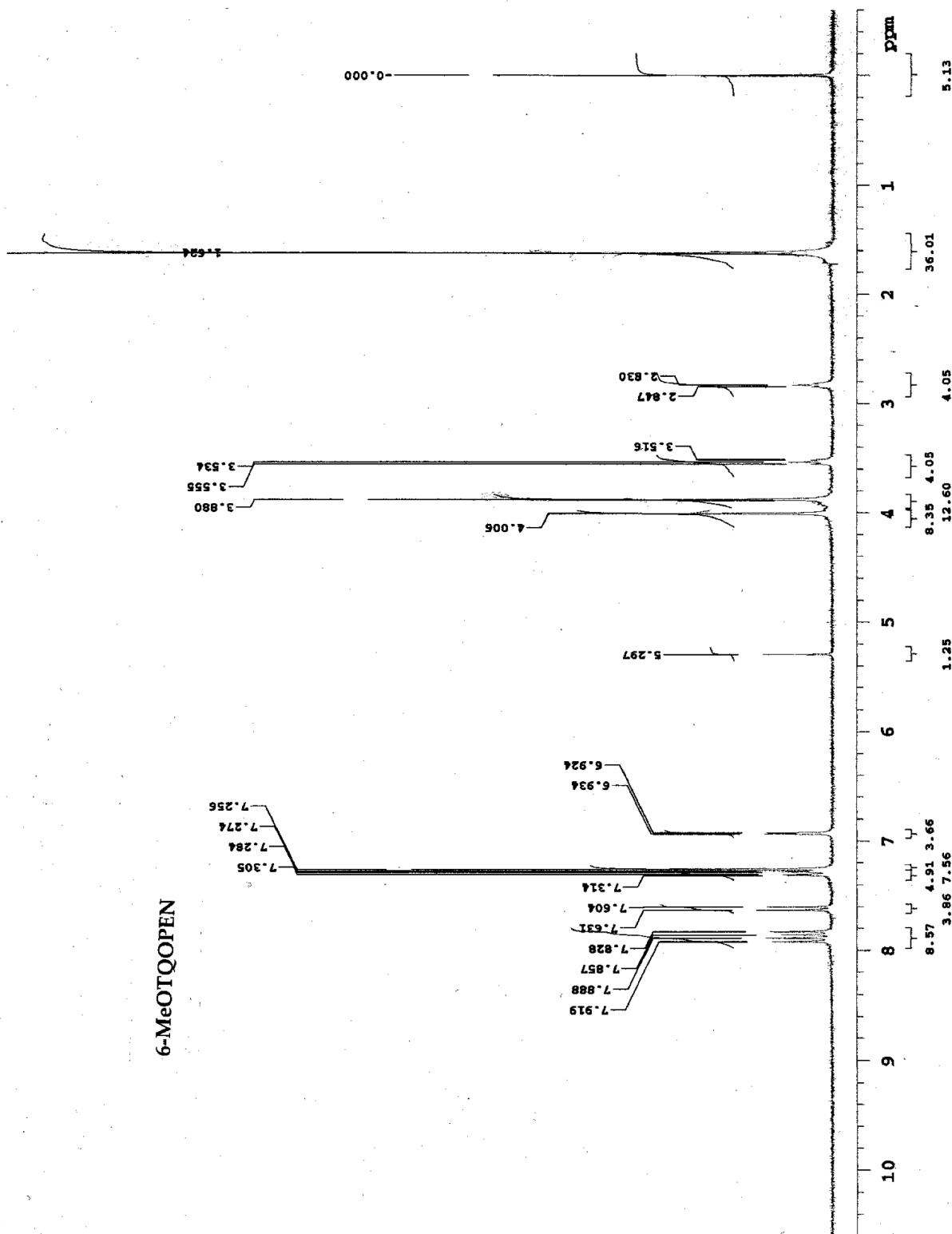


Fig. S45.  $^{13}\text{C}$  NMR spectrum for TriMeOBAPTO in  $\text{CDCl}_3$ .



**Fig. S46.** <sup>1</sup>H NMR spectrum for 6-MeOTQOPEN in CDCl<sub>3</sub>.

## 6-MeOTQOPEN

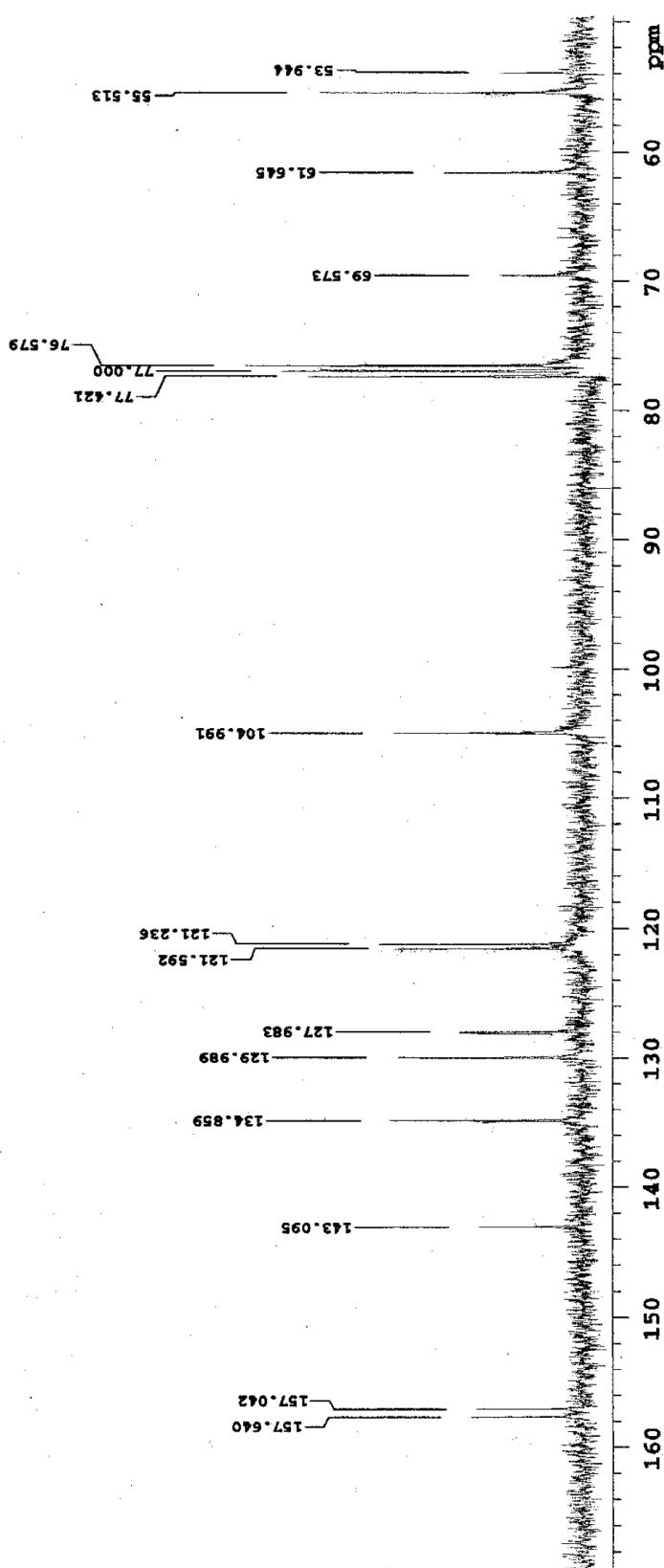
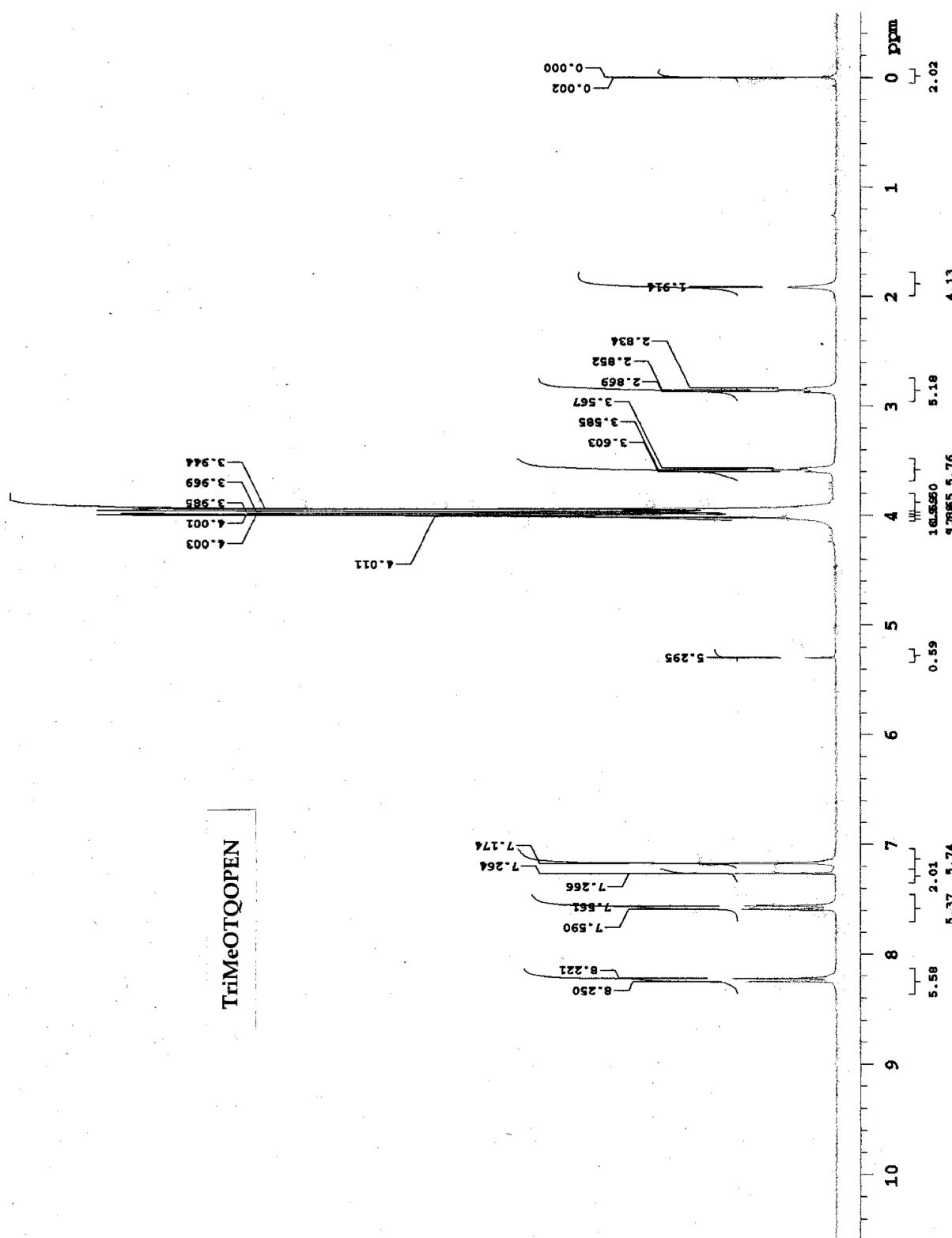


Fig. S47.  $^{13}\text{C}$  NMR spectrum for 6-MeOTQOPEN in  $\text{CDCl}_3$ .



**Fig. S48.** <sup>1</sup>H NMR spectrum for TriMeOTQOPEN in CDCl<sub>3</sub>.

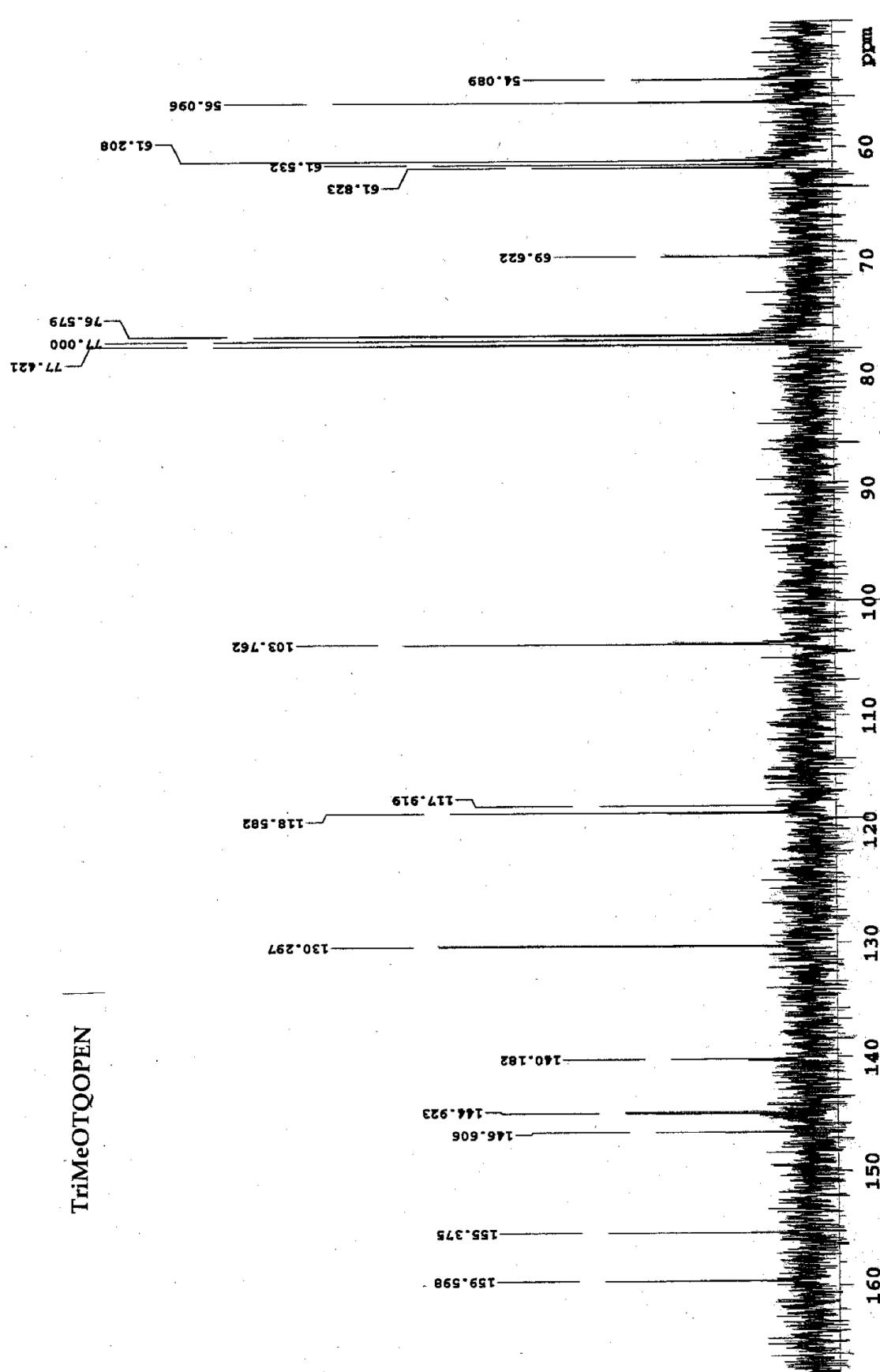


Fig. S49.  $^{13}\text{C}$  NMR spectrum for TriMeOTQOPEN in  $\text{CDCl}_3$ .

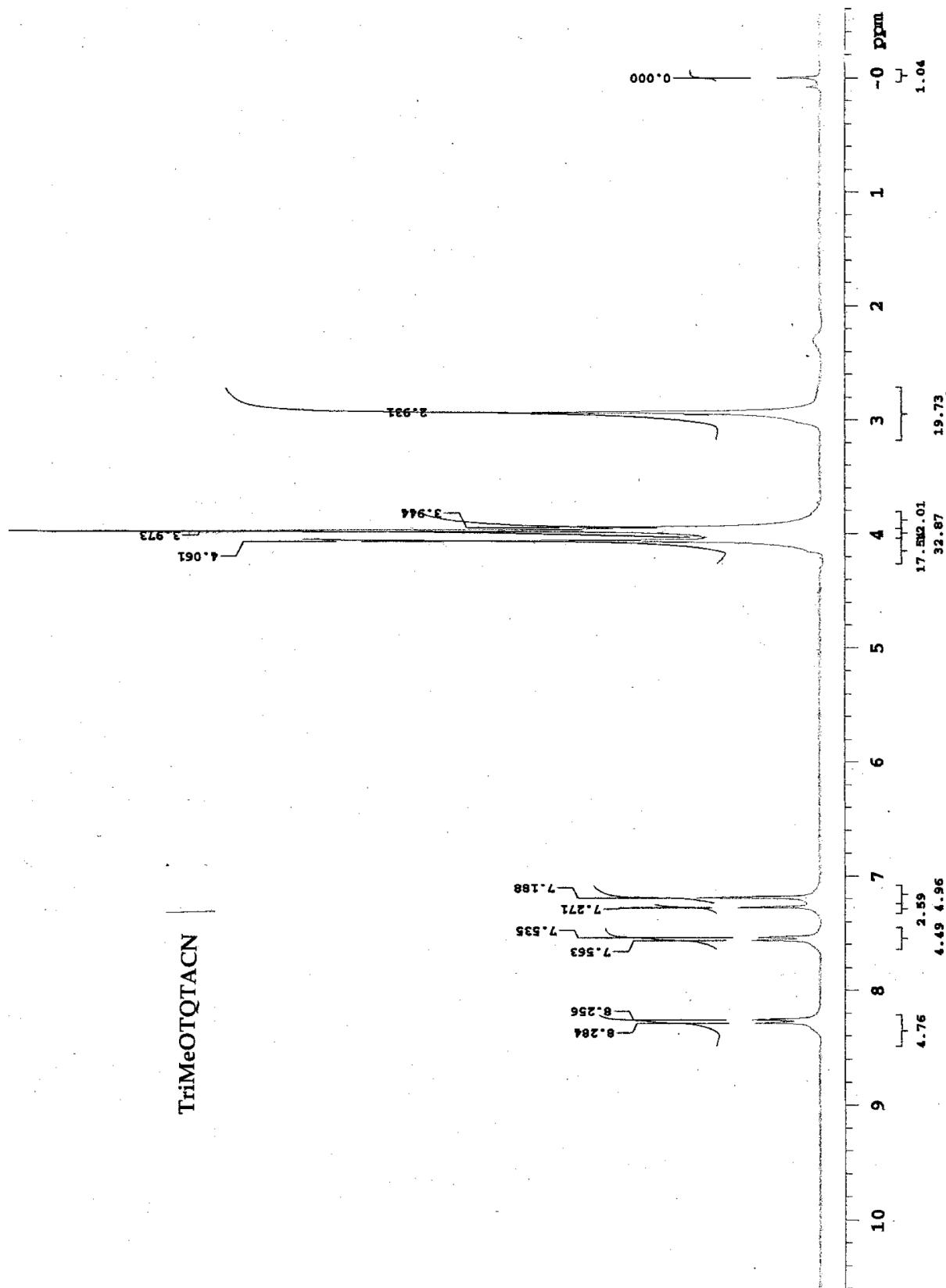


Fig. S50.  $^1\text{H}$  NMR spectrum for TriMeOTQTACN in  $\text{CDCl}_3$ .

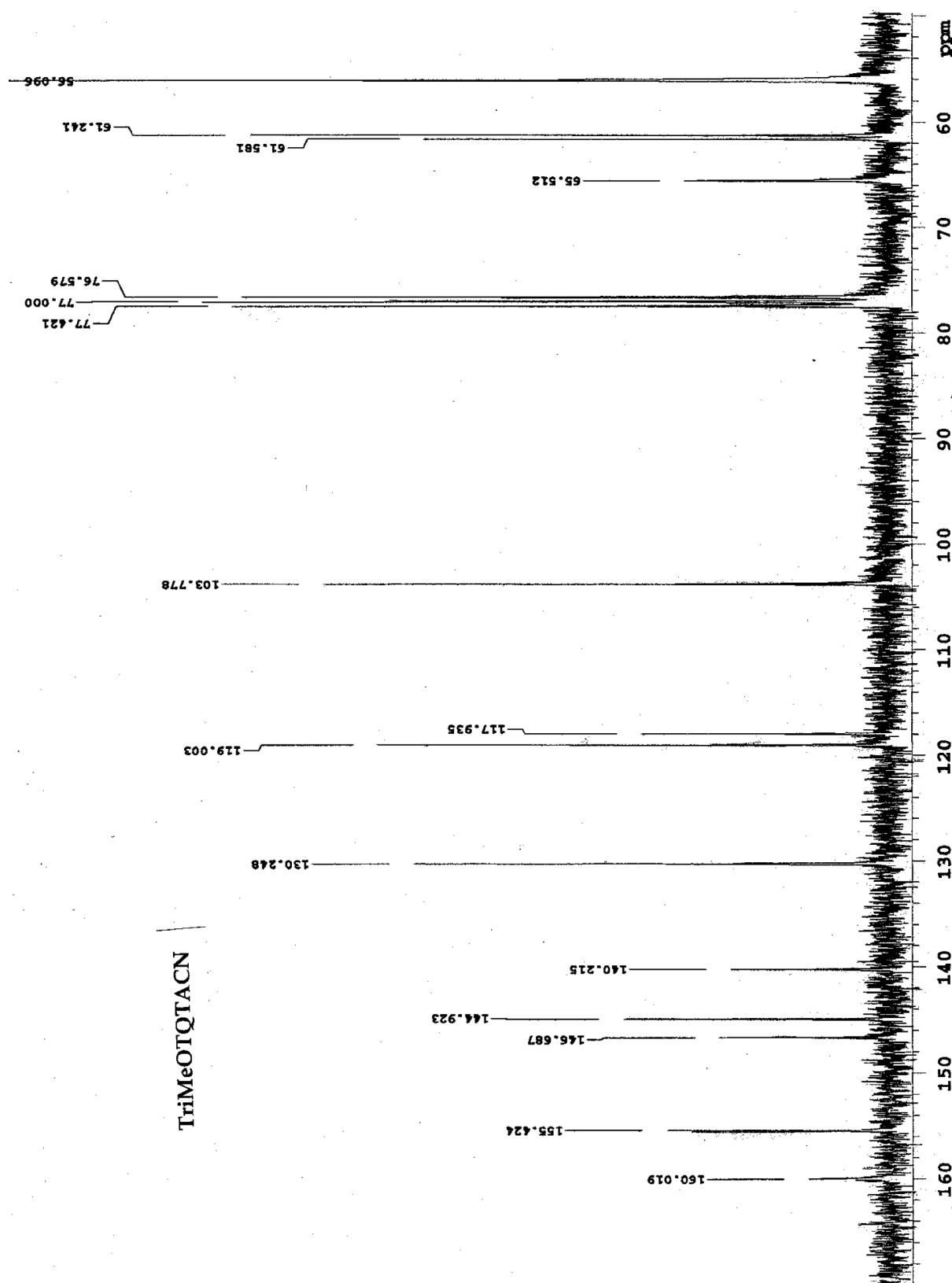


Fig. S51.  $^{13}\text{C}$  NMR spectrum for TriMeOTQTACN in  $\text{CDCl}_3$ .

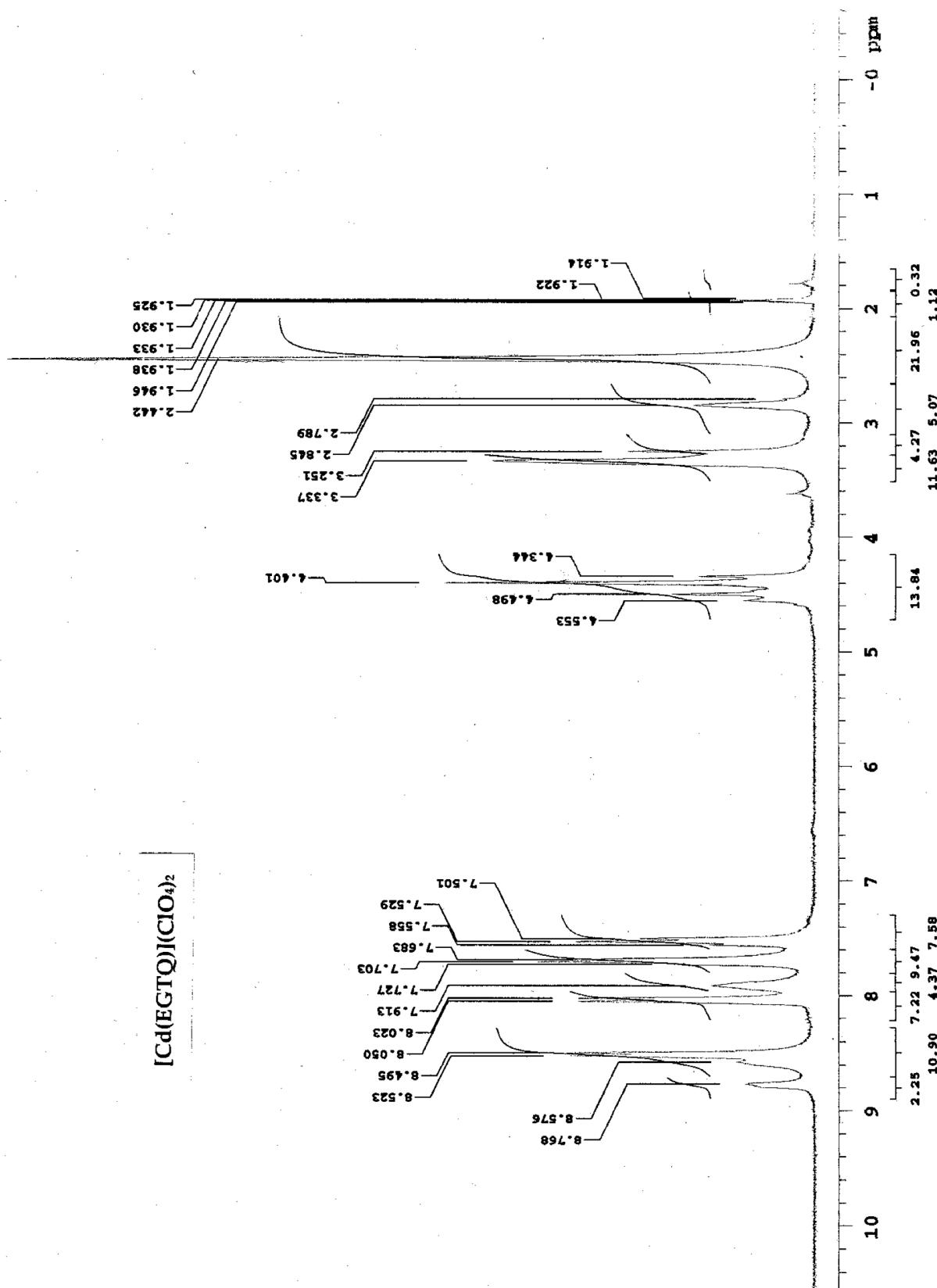
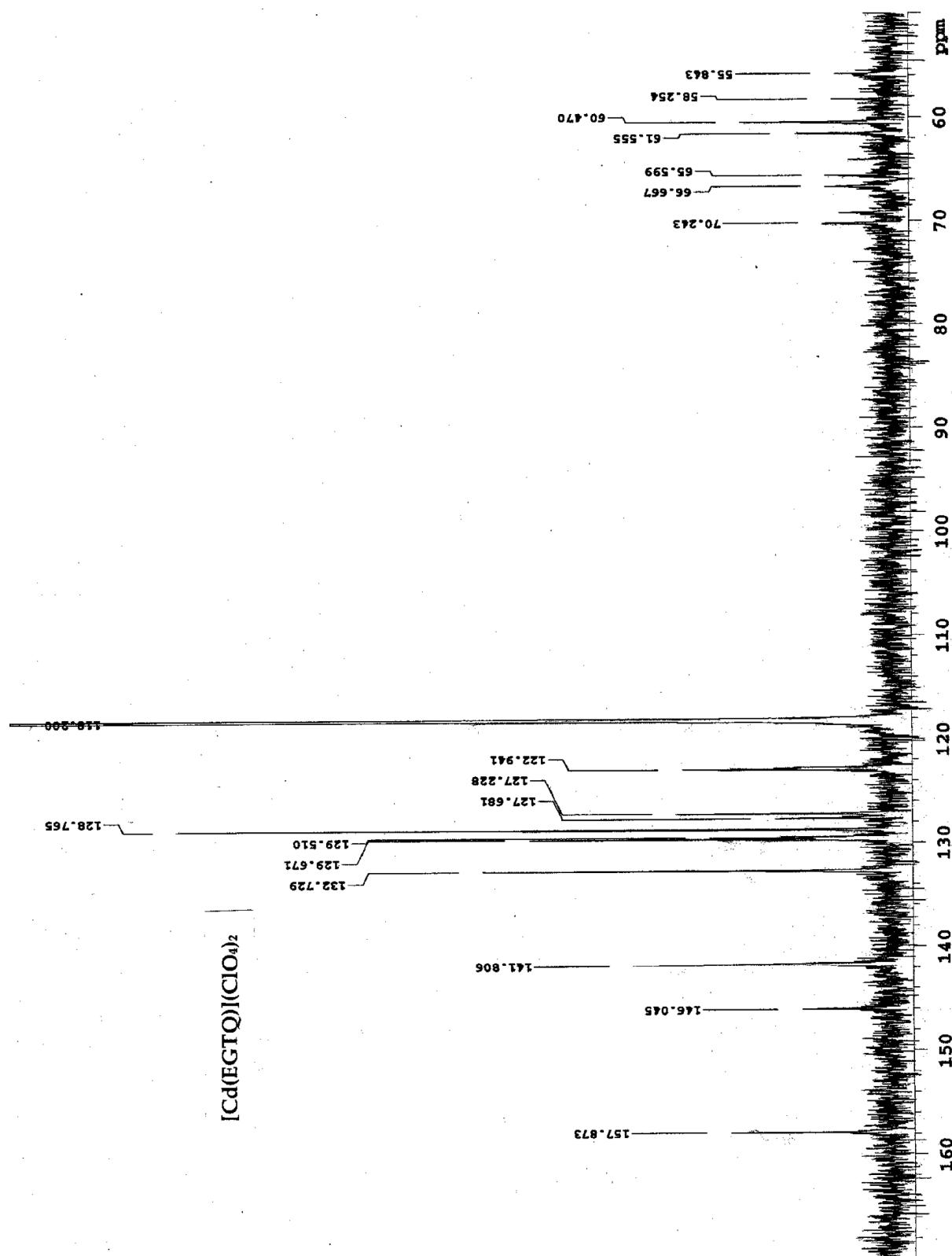


Fig. S52. <sup>1</sup>H NMR spectrum for [Cd(EGTQ)](ClO<sub>4</sub>)<sub>2</sub> in CD<sub>3</sub>CN.



**Fig. S53.**  $^{13}\text{C}$  NMR spectrum for  $[\text{Cd}(\text{EGTQ})](\text{ClO}_4)_2$  in  $\text{CD}_3\text{CN}$ .

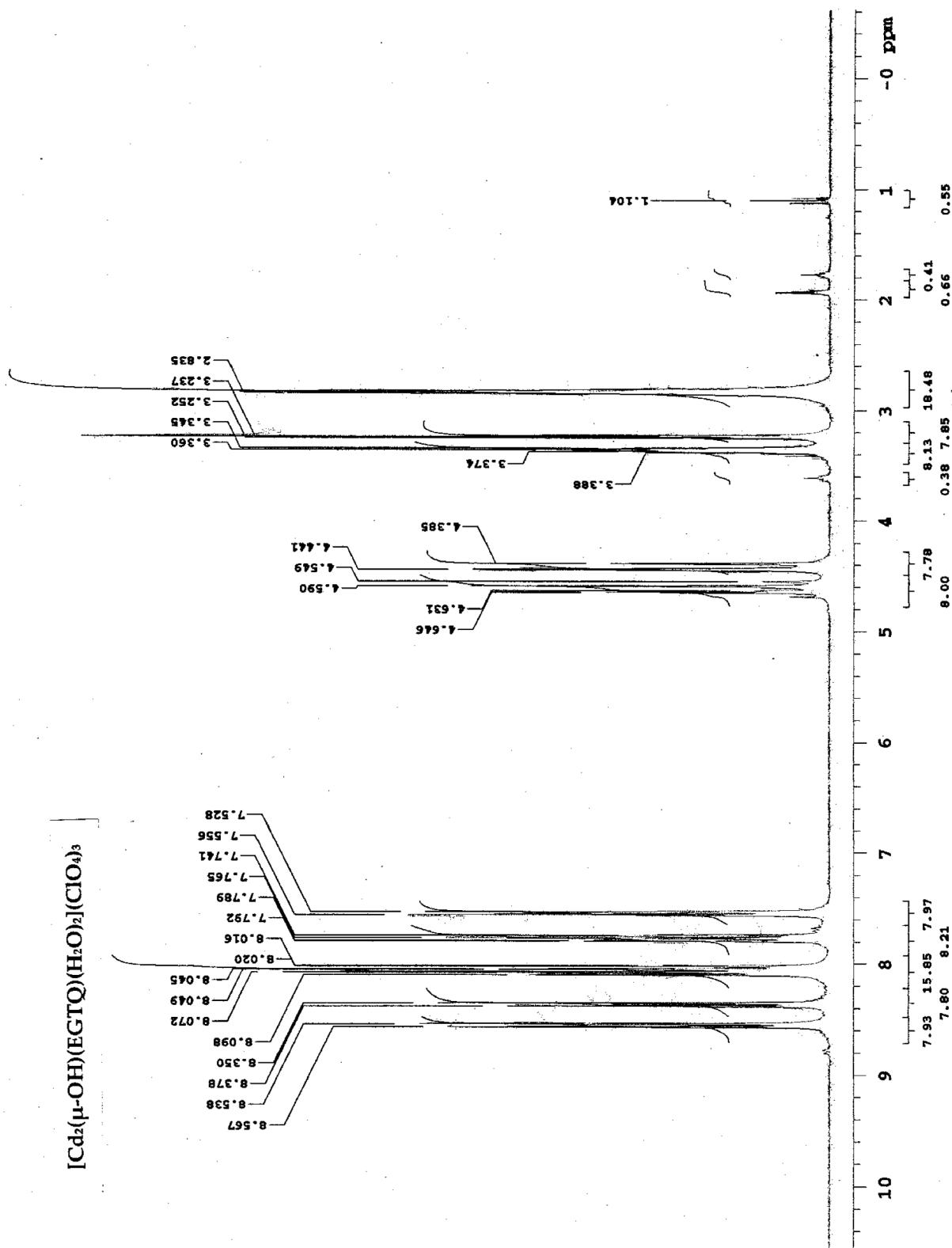
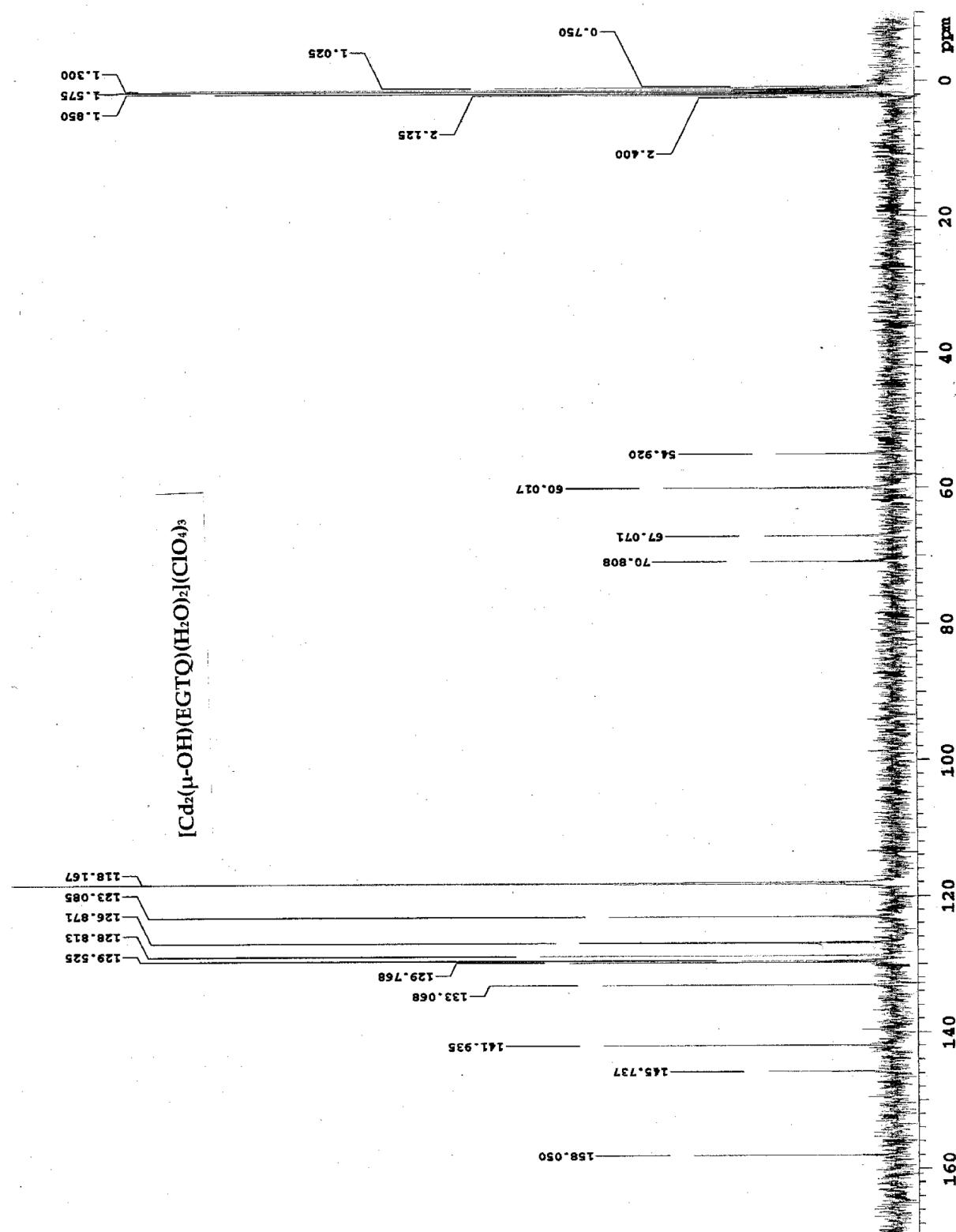
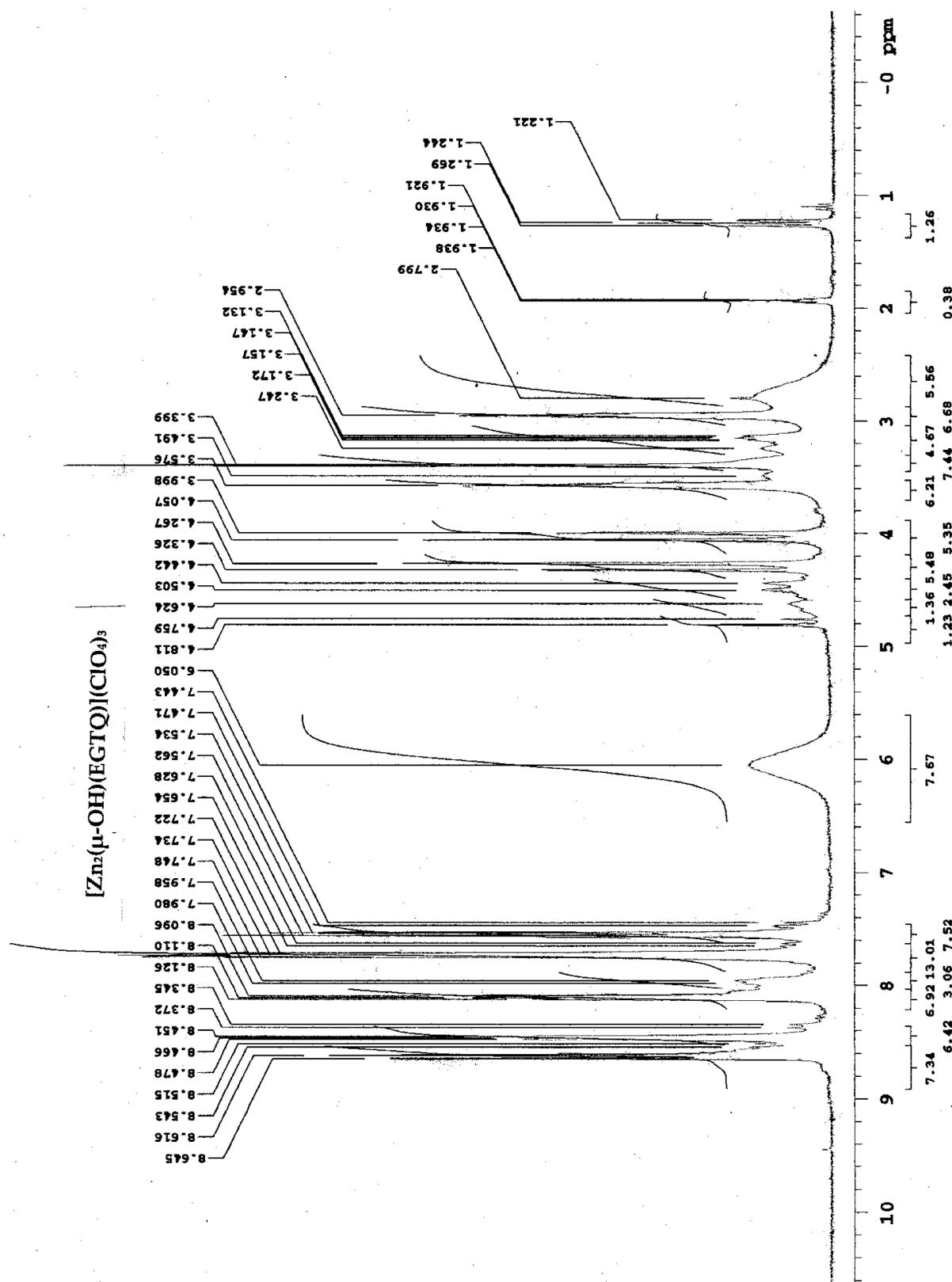


Fig. S54.  $^1\text{H}$  NMR spectrum for  $[\text{Cd}_2(\mu\text{-OH})(\text{EGTQ})(\text{H}_2\text{O})_2](\text{ClO}_4)_3$  in  $\text{CD}_3\text{CN}$ .



**Fig. S55.**  $^{13}\text{C}$  NMR spectrum for  $[\text{Cd}_2(\mu\text{-OH})(\text{EGTQ})(\text{H}_2\text{O})_2](\text{ClO}_4)_3$  in  $\text{CD}_3\text{CN}$ .



**Fig. S56.**  $^1\text{H}$  NMR spectrum for  $[\text{Zn}_2(\mu\text{-OH})(\text{EGTQ})](\text{ClO}_4)_3$  in  $\text{CD}_3\text{CN}$ .

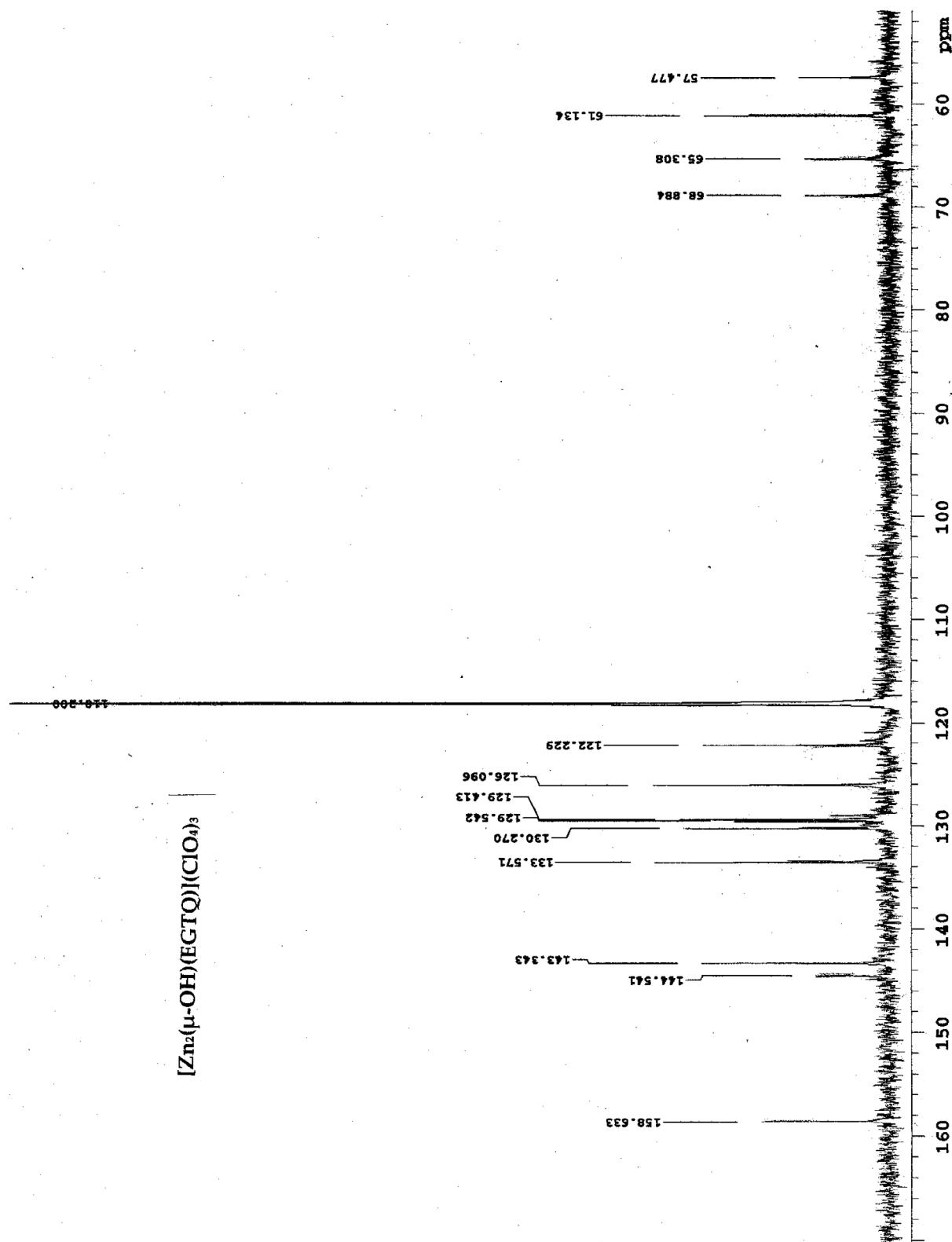
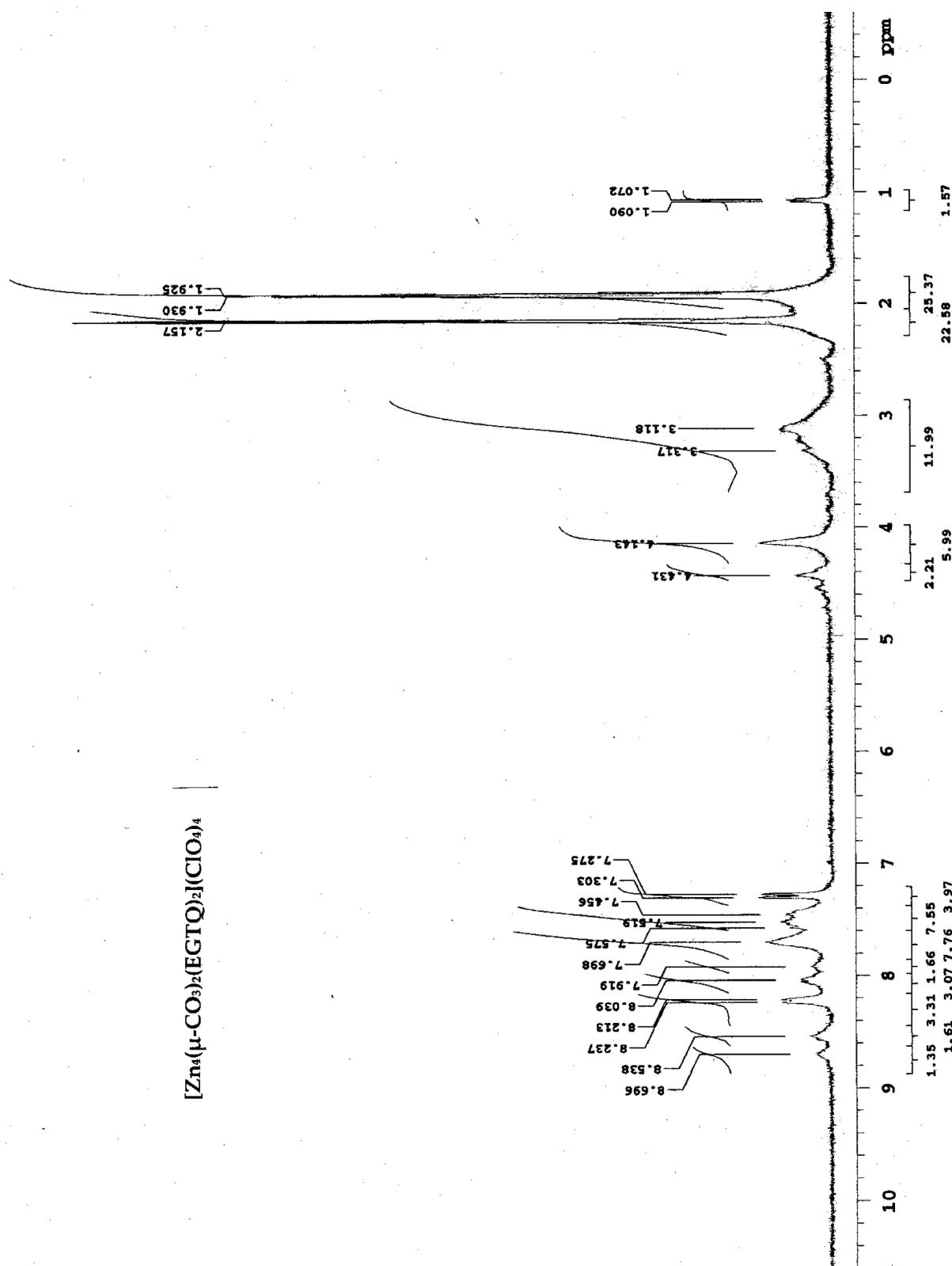
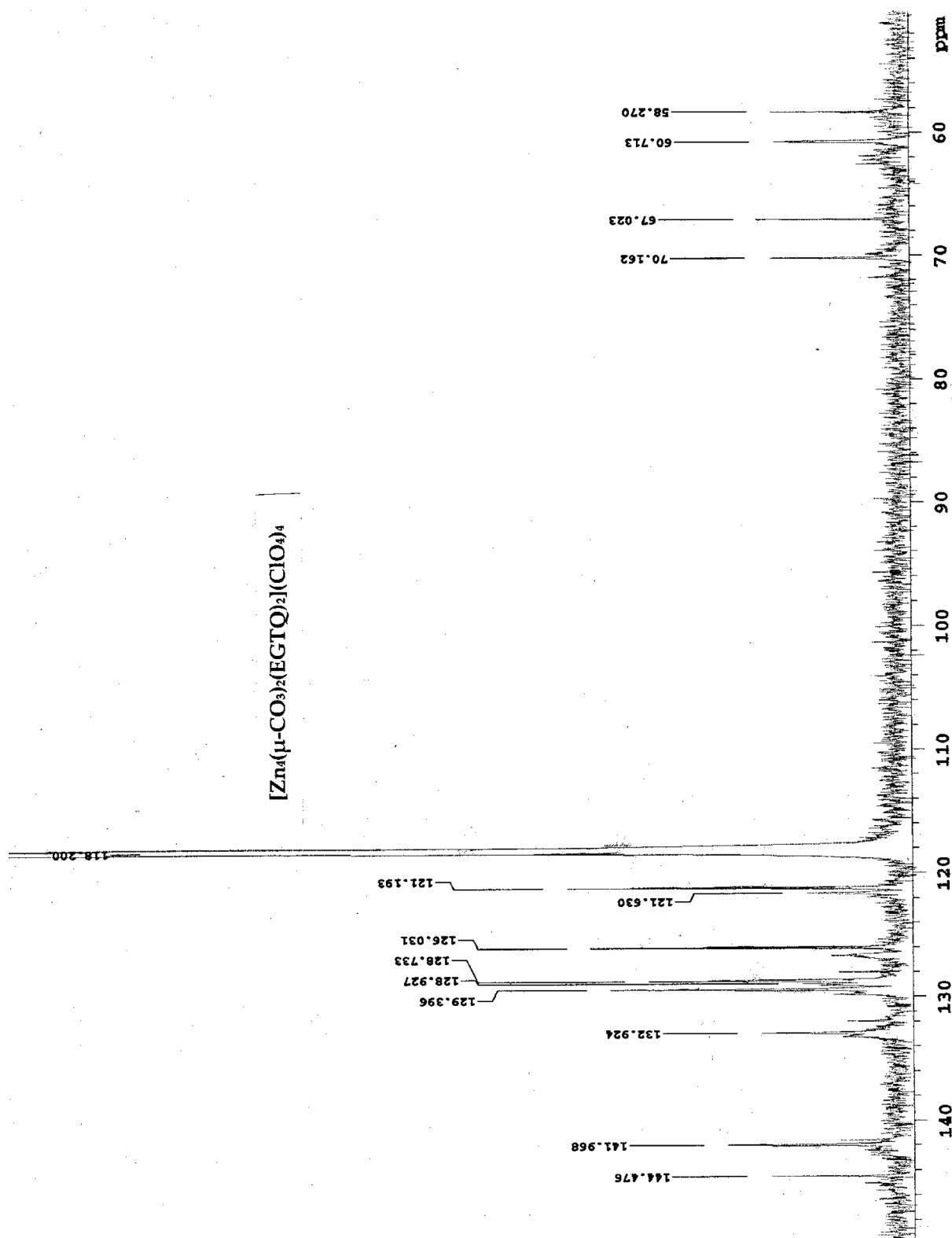


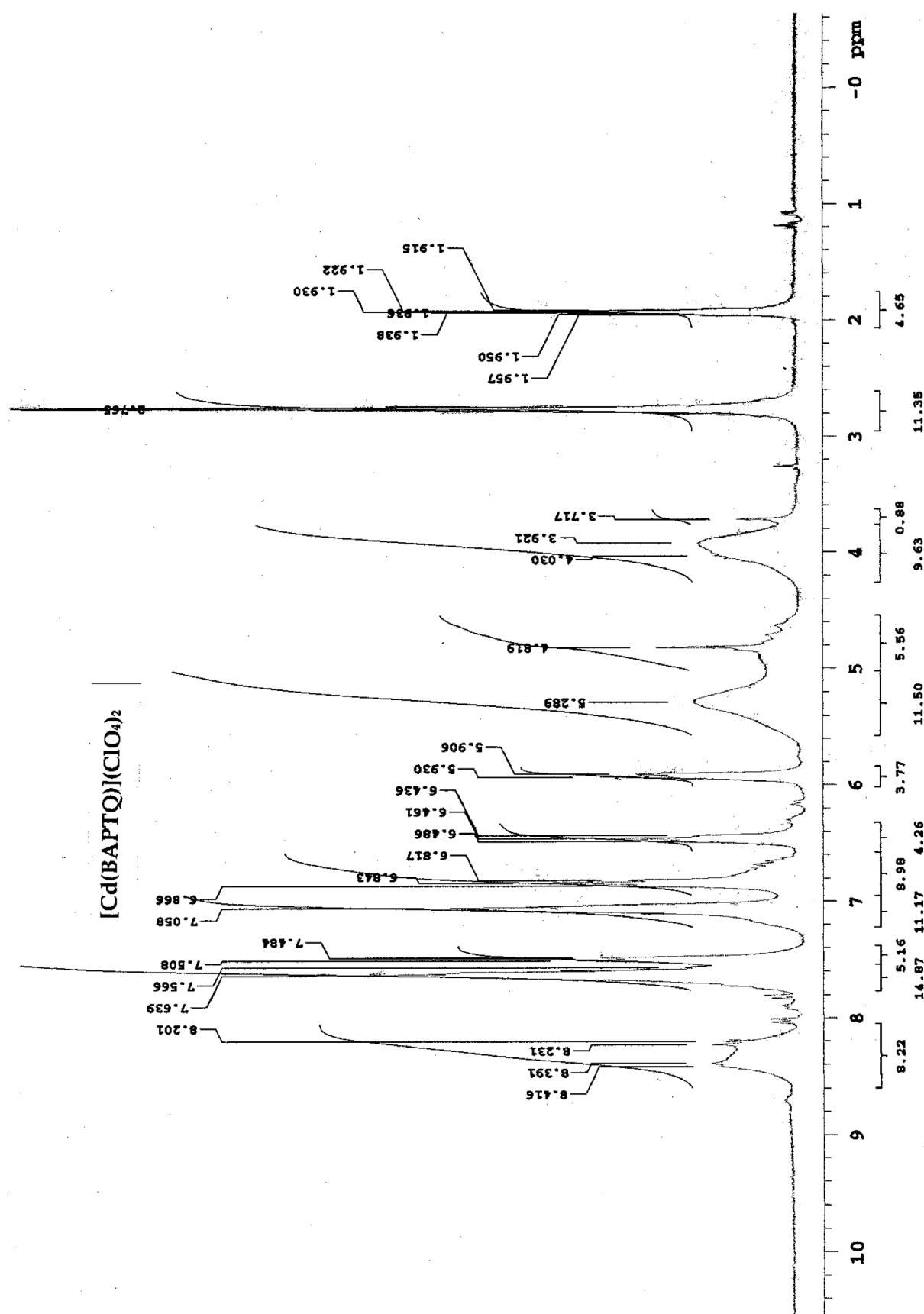
Fig. S57.  $^{13}\text{C}$  NMR spectrum for  $[\text{Zn}_2(\mu\text{-OH})(\text{EGTQ})](\text{ClO}_4)_3$  in  $\text{CD}_3\text{CN}$ .



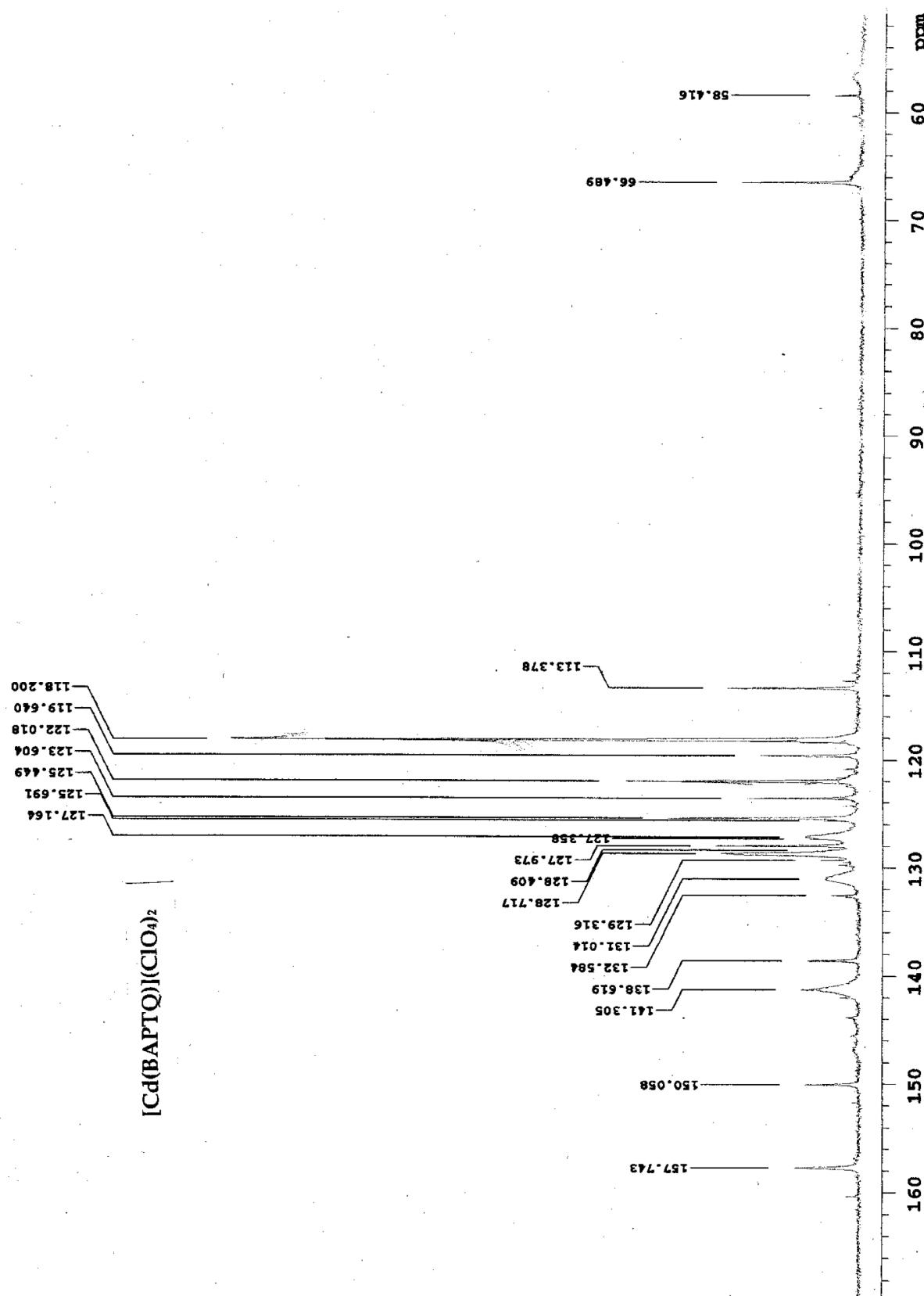
**Fig. S58.**  $^1\text{H}$  NMR spectrum for  $[\text{Zn}_4(\mu\text{-CO}_3)_2(\text{EGTQ})_2](\text{ClO}_4)_4$  in  $\text{CD}_3\text{CN}$ .



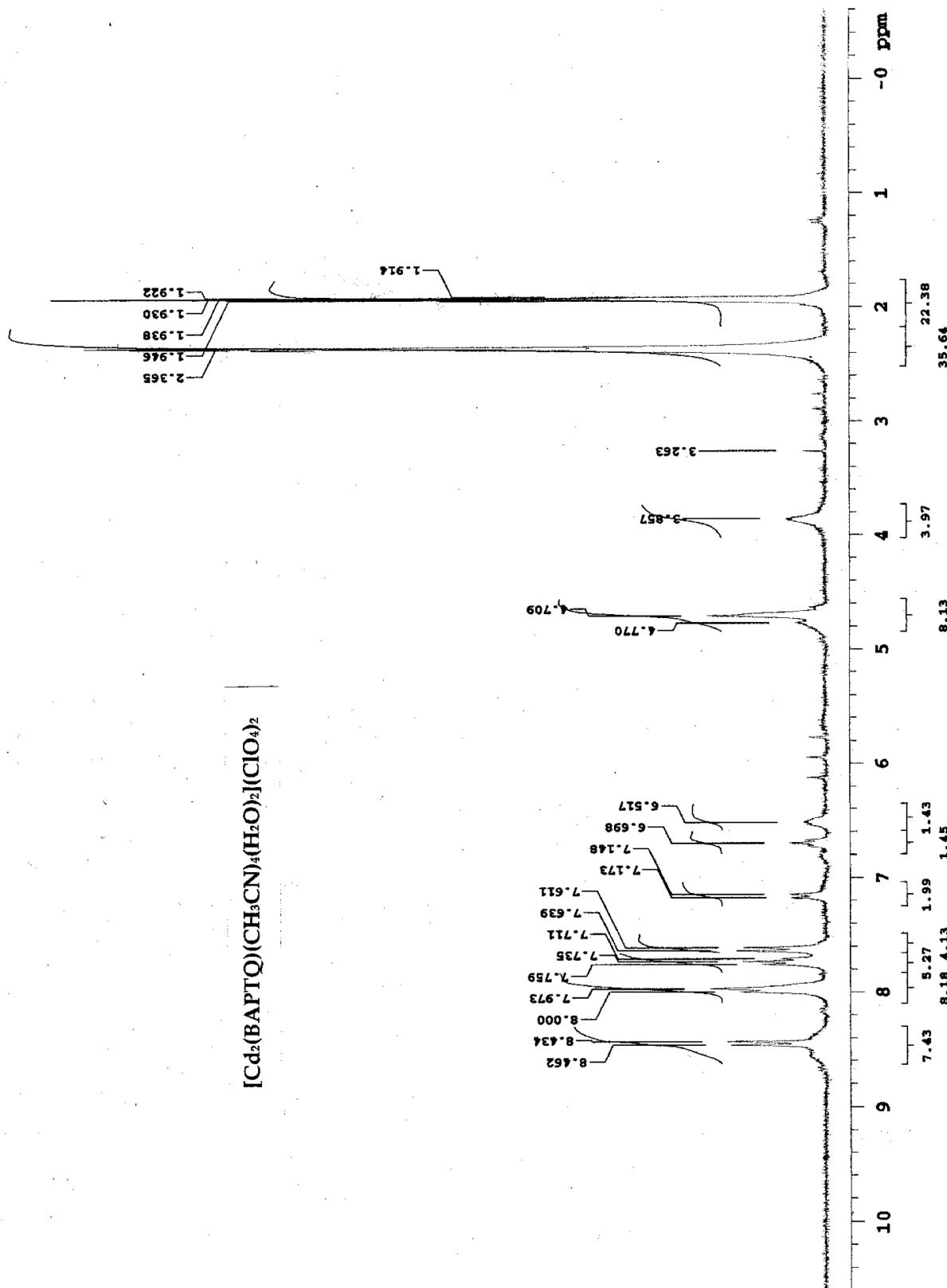
**Fig. S59.**  $^{13}\text{C}$  NMR spectrum for  $[\text{Zn}_4(\mu\text{-CO}_3)_2(\text{EGTQ})_2](\text{ClO}_4)_4$  in  $\text{CD}_3\text{CN}$ .



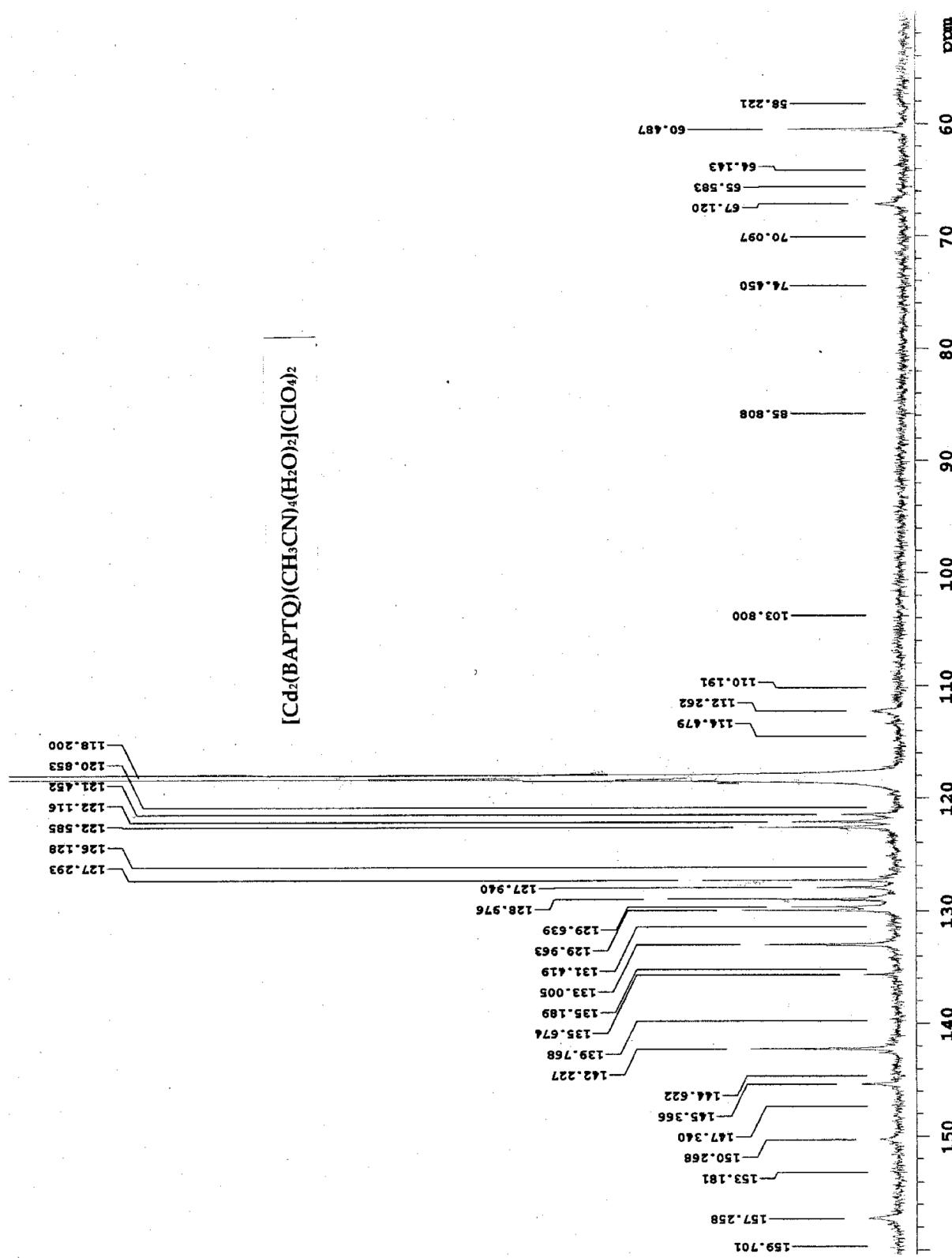
**Fig. S60.**  $^1\text{H}$  NMR spectrum for  $[\text{Cd}(\text{BAPTQ})(\text{ClO}_4)_2$  in  $\text{CD}_3\text{CN}$ .



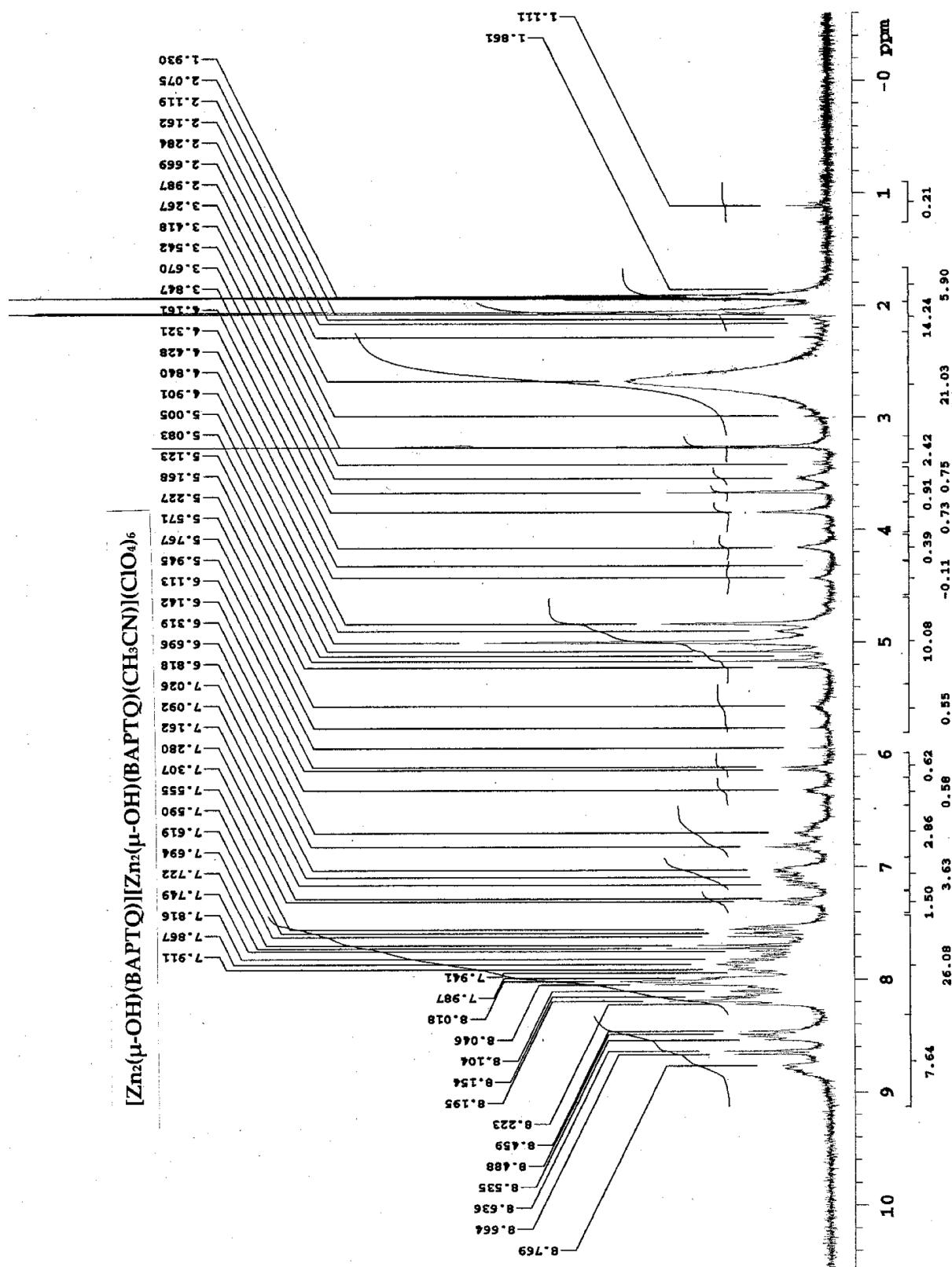
**Fig. S61.**  $^{13}\text{C}$  NMR spectrum for  $[\text{Cd}(\text{BAPTQ})](\text{ClO}_4)_2$  in  $\text{CD}_3\text{CN}$ .



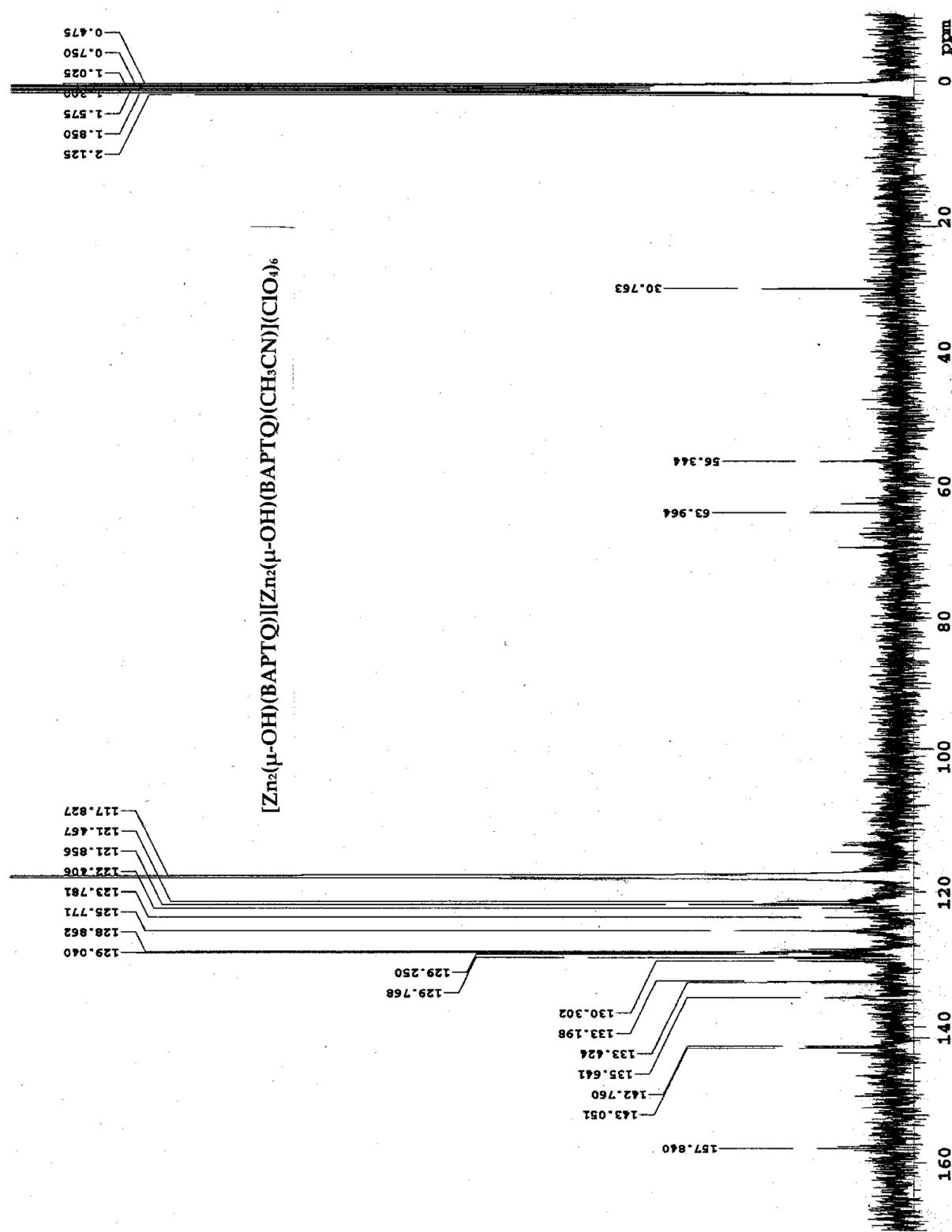
**Fig. S62.**  $^1\text{H}$  NMR spectrum for  $[\text{Cd}_2(\text{BAPTQ})(\text{CH}_3\text{CN})_4(\text{H}_2\text{O})_2](\text{ClO}_4)_2$  in  $\text{CD}_3\text{CN}$ .



**Fig. S63.**  $^{13}\text{C}$  NMR spectrum for  $[\text{Cd}_2(\text{BAPTQ})(\text{CH}_3\text{CN})_4(\text{H}_2\text{O})_2](\text{ClO}_4)_2$  in  $\text{CD}_3\text{CN}$ .



**Fig. S64.**  $^1\text{H}$  NMR spectrum for  $[Zn_2(\mu\text{-OH})(BAPTQ)][Zn_2(\mu\text{-OH})(BAPTQ)(CH_3CN)](ClO_4)_6$  in  $\text{CD}_3\text{CN}$ .



**Fig. S65.**  $^{13}\text{C}$  NMR spectrum for  $[Zn_2(\mu\text{-OH})(BAPTQ)][Zn_2(\mu\text{-OH})(BAPTQ)(CH_3CN)](ClO_4)_6$  in  $\text{CD}_3\text{CN}$ .