## **Electronic Supporting Information**

Exploring the multifunctionality in metal-organic frameworks materials: how do the stilbenedicarboxylate and imidazolyl ligands tune the characteristics of coordination polymers?

Marina O. Barsukova,<sup>a,b</sup> Sergey A. Sapchenko,<sup>\*a,b,c</sup> Konstantin A. Kovalenko,<sup>a,b</sup> Denis G. Samsonenko,<sup>a,b</sup> Andrei S. Potapov,<sup>d</sup> Danil N. Dybtsev,<sup>a,b</sup> Vladimir P. Fedin<sup>a,b</sup>

<sup>a</sup> Nikolaev Institute of Inorganic Chemistry, Siberian Branch of the Russian Academy of Sciences, Acad. Lavrentieva Ave., 3, 630090, Novosibirsk, Russia

<sup>b</sup> Novosibirsk State University, Pirogova st., 2, 630090, Novosibirsk, Russia

<sup>c</sup> School of Chemistry, University of Manchester, Oxford Road, Manchester, M13 9PL, U.K.

<sup>d</sup> National Research Tomsk Polytechnic University, Lenin Ave., 30, 634050, Tomsk, Russia

\* Corresponding author at: Nikolaev Institute of Inorganic Chemistry, Siberian Branch of the Russian Academy of Sciences, Akad. Lavrentiev Ave., 3, 630090 Novosibirsk, Russia. Tel.: +7 383 330 9490; fax: +7 383 330 9489. X-ray crystallography. The fragments of the coordination polymers with disorders for compounds 1, 2, 3, 4, 5 and 6 in thermal ellipsoid mode are presented in Figures S1, S3, S5, S7, S9 and S11, respectively. Additional figures, highlighting the inner surfaces of the frameworks 1, 2, 3, 4, 4a, 5 and 6 are presented in Fig. S2, S4, S6, S8, S8a, S10 and S12, respectively.

**Thermogravimetric analyses.** The TGA plots for the reported compounds are presented in Fig. S13-S14. The TGA plots for the  $I_2$ -adsorbed and activated compounds 2 and 5 are presented in Fig. S15–S16.

**Powder X-ray diffraction analyses.** The PXRD data for the as-synthesythed and activated compounds 1, 2, 3, 4, 5 and 6 are presented on Fig. S17–S22. For the compound 2 and 5 powder X-ray diffraction (PXRD) patterns with adsorbed  $I_2$  are presented in Fig. S23–S24.

 $N_2$  adsorption isotherm measurements. The analyses were performed on Quantochrome's Autosorb iQ at 77K. The gas adsorption/desorption isotherms for the activated compounds 1-6 are provided on Fig.S25.

 $I_2$  sorption from solution. The photos show sample color of activated compound 2 and samples soaked in  $I_2$  solution in methanol, toluene and  $CH_2Cl_2$  for 1 day in Fig. S26.

**Raman spectroscopy.** The Raman spectra of the activated and  $I_2$ -adsorbed compounds 2 and 5 are present in Fig. S27.

**Luminescent properties.** Normalized spectra of emission ( $\lambda_{ex} = 390$  nm) and excitation of obtained as-synthesized and activated compounds, and free H<sub>2</sub>sdc are shown on Fig. S28.

**Tables.** Single crystal X-ray crystallography data and literature references for luminescent properties.



Figure S1. View of node in 1 (hydrogen atoms are omitted; ellipsoids are at the 50% probability level). Symmetry codes for related atoms: i) 1 - x, y, 1.5 - z.



Figure S2. Calculated void surfaces for 1 along a (left), b (center) and c axis (right).



Figure S3. View of node in 2 (hydrogen atoms are omitted; ellipsoids are at the 50% probability level). Symmetry codes for related atoms: i) -x, y, 1.5 - z.



Figure S4. Calculated void surfaces for 2 along a (left), b (center) and c axis (right).



Figure S5. View of node in **3** (hydrogen atoms are omitted; ellipsoids are at the 50% probability level). Symmetry codes for related atoms: i) 1 - x, y, 1.5 - z.



Figure S6. Calculated void surfaces for 3 along a (left), b (center) and c axis (right).



Figure S7. View of node in 4 (hydrogen atoms are omitted; ellipsoids are at the 50% probability level). Symmetry codes for related atoms: i) -2 + x, -1 + y, z; ii) 1.5 - x, -y, -0.5 + z; iii) -0.5 - x, 1 - y, -0.5 + z.



Figure S8. Calculated void surfaces for 4 along a (left), b (center) and c axis (right).



Figure S8a. Calculated void surfaces for 4a along a (left), b (center) and c axis (right).



Figure S9. View of node in 5 (hydrogen atoms are omitted; ellipsoids are at the 50% probability level). Symmetry codes for related atoms: i) 1 - x, y, 1.5 - z.



Figure S10. Calculated void surfaces for 5 along a (left), b (center) and c axis (right).



Figure S11. View of node in 6 (hydrogen atoms are omitted; ellipsoids are at the 50% probability level). Symmetry codes for related atoms: i) 1 - x, y, 0.5 - z.



Figure S12. Calculated void surfaces for 6 along a (left), b (center) and c axis (right).



Figure S13. TGA plots for 1, 2 and 3.



Figure S14. TGA plots for 4, 5 and 6.



Figure S15. TGA plots for activated 2 and  $I_2$ -loaded sample of 2.



Figure S16. TGA plots for activated 5 and  $I_2$ -loaded sample of 5.



Figure S17. PXRD pattern of calculated for 1 (black) in comparison with patterns of assynthesized 1 (blue), and activated 1 (dark blue).



Figure S18. PXRD pattern of calculated for 2 (black) in comparison with patterns of assynthesized 2 (red), and activated 2 (dark red).



Figure S19. PXRD pattern of calculated for 3 (black) in comparison with patterns of assynthesized 3 (green), and activated 3 (dark green).



Figure S20. PXRD pattern of calculated for 4 (black) in comparison with patterns of assynthesized 4 (violet), and activated 4 (dark violet).



Figure S21. PXRD pattern of calculated for **5** (black) in comparison with patterns of assynthesized **5** (light magenta), and activated **5** (magenta).



Figure S22. PXRD pattern of calculated for 6 (black) in comparison with patterns of assynthesized 6 (orange), and activated 6 (brown).



Figure S23. PXRD pattern of calculated for 2 (black) in comparison with patterns of activated 2 (red), 2 with  $I_2$  adsorbed (blue) and 2 after  $I_2$  desorption in methanol (purple).



Figure S24. PXRD pattern of calculated for **5** (black) in comparison with patterns of activated **5** (red), **5** with  $I_2$  adsorbed (blue) and **5** after  $I_2$  desorption in methanol (purple).



Figure S25.  $N_2$  adsorption isotherms for activated 1–6.



Figure S26. Photos of activated **2** (*a*) and samples soaked in  $I_2$  solution in methanol (*b*), toluene (*c*) and  $CH_2Cl_2$  (*d*) for 1 day.



Figure S27. Raman spectra of  $\mathbf{2}$ ,  $\mathbf{5}$  and  $I_2$ -adsorbted  $\mathbf{2}$  and  $\mathbf{5}$ .



Figure S28. Normalized solid-state luminescence spectra of compounds 1, 2, 4, 5 their activated compounds, and stilbenedicarboxylic acid recorded at  $\lambda_{ex} = 390$  nm.

Compound name/Formula	Quantum vield $\phi$ , %	Emission maximum	Excitation wavelength,	Reference
	5 17	wavelength, nm	nm	
$[NH_2Me_2][(Cd_2Cl)_3(TATPT)_4]$	15.1	425	370	1
$[Ir(ppy)_2(bpy)][(Cd_2Cl)_3(TATPT)_4]$	28.7	530	370	
EuN-BDC/m2hmp	31.6	620	370	2
$[Eu(Dpmd)(D_2O)]$	45	5D0	392	3
MOF-253	33	white CCT 5627	395	4
$Zn_2(H_2L)_2(bpy)_2(H_2O)_3 \cdot H_2O$	43	484	376	5
ZrBDC-TCPE0.01%	59	461	363	6

Table S1. Chart of MOFs and MOF-based host-guest systems exhibiting high luminescence quantum yields

LMOF-231(activated)	95.1		365	7
LMOF-241(activated)	92.7		340	
LMOF-251	90.7		400	
Zn-PLA	46%	410	350	8
ZnBDCA	53a	410	360	9
[Zn <sub>2</sub> (bpdc) <sub>2</sub> BTyTPE)]	99	490	365	10
2	70.9			This work
DMF@2	82.0			
2				

<sup>a</sup> quantum effeciency

## Table S2. Crystal data and structure refinement for 1-6.

	1	2	3	4	5	6
Formula	$C_{29}H_{31}CdN_5O_5$	C <sub>29</sub> H <sub>31</sub> N <sub>5</sub> O <sub>5</sub> Zn	C <sub>29</sub> H <sub>31</sub> CoN <sub>5</sub> O <sub>5</sub>	$C_{59}H_{63}Cd_2N_9O_9$	$C_{28}H_{28}N_4O_4Zn$	$C_{28}H_{28}CoN_4O_4$
Formula	641.99	594.96	588.52	1266.98	549.91	543.47
weight						
Crystal	Monoclinic	Monoclinic	Monoclinic	Orthorhombic	Orthorhombic	Orthorhombic
system						
Space	<i>P2/c</i>	<i>P2/c</i>	<i>P2/c</i>	P2 <sub>1</sub> 2 <sub>1</sub> 2 <sub>1</sub>	Pbcn	Pbcn
group						
a, Å	11.2374(3)	11.5569(8)	11.4594(3)	8.8758(2)	8.6880(4)	8.7112(7)
b, Å	6.78252(13)	6.8510(5)	6.5882(2)	24.0175(7)	14.6001(6)	14.5251(18)
<i>c</i> , Å	18.7482(4)	17.4990(9)	18.3365(5)	28.4657(9)	23.4501(11)	23.5166(18)
β, °	91.4013(19)	92.086(7)	92.145(3)	90.0	90.0	90.0
Ζ	2	2	2	4	4	4
<i>V</i> , Å'	1428.52(6)	1384.59(16)	1383.39(7)	6068.2(3)	2974.5(2)	2975.6(5)
$\theta$ range, °	3.51 - 26.36	3.46 - 25.35	3.29 - 25.68	3.43 - 25.35	3.47 - 25.35	3.47 - 25.35
Crystal	0.13×0.09×0.0	0.43×0.13×0	0.44×0.14×0.0	0.39×0.16×0.1	0.28×0.20×0.0	0.37×0.20×0.15
size, mm	6	.11	7	4	8	
$D_c$ , g	1.493	1.427	1.413	1.387	1.228	1.213
cm <sup>-3</sup>						
$\mu$ , mm <sup>-1</sup>	0.811	0.935	0.668	0.761	0.862	0.613
Reflectio	11382/2920	5825/2536	6164/2601	30671/11071	7843/2724	8674/2731
ns						
collected/						
unique						
Reflectio	2530	1817	2157	9566	2037	2313
ns with I						
> 2 <i>o</i> (I)						
R <sub>int</sub>	0.0225	0.0300	0.0214	0.0350	0.0245	0.0296
GOF on	1.101	1.077	1.084	1.085	1.063	1.200
$F^2$						

0.0260, 0.0609	0.0654,	0.0404, 0.1032	0.0768, 0.1875	0.0805, 0.2273	0.1377, 0.3542
	0.1802				
0.0338, 0.0641	0.0884,	0.0526, 0.1094	0.0875, 0.1932	0.1037, 0.2496	0.1512, 0.3638
	0.1969				
0.527/-0.353	0.394/-0.613	0.470/-0.283	3.769/-1.021	0.897/-0.510	0.830/-0.499
	0.0260, 0.0609 0.0338, 0.0641 0.527/-0.353	0.0260, 0.0609 0.0654,   0.1802   0.0338, 0.0641 0.0884,   0.1969   0.527/-0.353 0.394/-0.613	0.0260, 0.0609   0.0654,   0.0404, 0.1032     0.1802   0.0338, 0.0641   0.0884,   0.0526, 0.1094     0.1969   0.527/-0.353   0.394/-0.613   0.470/-0.283	0.0260, 0.0609   0.0654, 0.1802   0.0404, 0.1032   0.0768, 0.1875     0.0338, 0.0641   0.0884, 0.1969   0.0526, 0.1094   0.0875, 0.1932     0.527/-0.353   0.394/-0.613   0.470/-0.283   3.769/-1.021	0.0260, 0.0609   0.0654,   0.0404, 0.1032   0.0768, 0.1875   0.0805, 0.2273     0.1802   0.1802   0.0526, 0.1094   0.0875, 0.1932   0.1037, 0.2496     0.0338, 0.0641   0.0884,   0.0526, 0.1094   0.0875, 0.1932   0.1037, 0.2496     0.1969   0.527/-0.353   0.394/-0.613   0.470/-0.283   3.769/-1.021   0.897/-0.510

## References

1. Sun, C.-Y.; Wang, X.-L.; Zhang, X.; Qin, C.; Li, P.; Su, Z.-M.; Zhu, D.-X.; Shan, G.-G.; Shao, K.-Z.; Wu, H.; Li, J. Efficient and tunable white-light emission of metal-organic frameworks by iridium-complex encapsulation. *Nat. Commun.* **2013**, *4*, 2717.

2. Kyprianidou, E. J.; Lazarides, T.; Kaziannis, S.; Kosmidis, C.; Itskos, G.; Manos, M. J.; Tasiopoulos, A. J. Single crystal coordinating solvent exchange as a general method for the enhancement of the photoluminescence properties of lanthanide MOFs. *J. Mater. Chem. A* **2014**, *2*, 5258.

3. Vilela, S. M. F.; Ananias, D.; Fernandes, J. A.; Silva, P.; Gomes, A. C.; Silva, N. J. O.; Rodrigues, M. O.; Tomé, J. P. C.; Valente, A. A.; Ribeiro-Claro, P.; Carlos, L. D.; Rocha, J.; Almeida Paz, F. A. Multifunctional micro- and nanosized metal–organic frameworks assembled from bisphosphonates and lanthanides. *J. Mater. Chem. C* **2014**, *2*, 3311.

4. Lu, Y.; Yan, B. Lanthanide organic-inorganic hybrids based on functionalized metal-organic frameworks (MOFs) for a near-UV white LED. *Chem. Commun.* **2014**, *50*, 15443.

5. Deng, Y.; Chen, N.; Li, Q.; Wu, X.; Huang, X.; Lin, Z.; Zhao, Y. Highly fluorescent metalorganic frameworks based on a benzene-cored tetraphenylethene derivative with the ability to detect 2,4,6-Trinitrophenol in water. *Cryst. Growth Des.* **2017**, *17*, 3170.

6. Feng, X.; Zeng, L.; Zou, D.; Zhang, Z.; Zhong, G.; Peng, S.; Liu, L.; Chen, L.; Zhang, J. Trace-doped metal-organic gels with remarkably enhanced luminescence. *RSC Adv.* **2017**, *7*, 37194.

7. Lustig, W. P.; Wang, F.; Teat, S. J.; Hu, Z.; Gong, Q.; Li, J. Inorg. Chem. 2016, 55, 7250.

8. Chandrasekhar, P.; Mukhopadhyay, A.; Savitha, G.; Moorthy, J. N. Remarkably selective and enantiodifferentiating sensing of histidine by a fluorescent homochiral Zn-MOF based on pyrene-tetralactic acid. *Chem. Sci.* **2016**, *7*, 3085.

9. Cai, H.; Xu, L.-L.; Lai, H.-Y.; Liu, J.-Y.; Ng, S. W.; Li, D. A highly emissive and stable zinc(II) metal–organic framework as a host–guest chemopalette for approaching white-light-emission. *Chem. Commun.* **2017**, *53*, 7917.

10. Tao, C.-L.; Chen, B.; Liu, X.-G.; Zhou, L.-J.; Zhu, X.-L.; Cao, J.; Gu, Z.-G.; Zhao, Z.; Shena, L.; Tang, B. Z. A highly luminescent entangled metal–organic framework based on pyridine-substituted tetraphenylethene for efficient pesticide detection. *Chem. Commun.* **2017**, *53*, 9975.