

**Supplementary Information to:**

**$\pi$ -Stacked and Unstacked Aggregates Formation of A Near-infrared Dye  
Leading to Characteristic Photophysical Properties**

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## 1. Experimental

### Materials and mixture samples preparation

3,3'-Diethylthiatricarbocyanine iodide (DTCI) and methanol (Wako Chem.) were used as purchased. DTCI solutions in methanol at  $[DTCI] = 4.0 \times 10^{-4}$  M and  $1.0 \times 10^{-4}$  M were prepared by dissolving 5.45 mg of DTCI (10.00  $\mu$ mol) in 25 mL and 100 mL of a solvent, and those at  $[DTCI] = 4.0 \times 10^{-5}$  M,  $1.0 \times 10^{-5}$  M,  $4.0 \times 10^{-6}$  M and  $1.0 \times 10^{-6}$  M were prepared by diluting the  $4.0 \times 10^{-4}$  M and  $1.0 \times 10^{-4}$  M samples.

Aqueous solution samples of DTCI were prepared by adding DTCI to deionized water at the apparent ratios of 0.1 g/L, 0.5 g/L, 1 g/L, and 2 g/L, rigorously stirring the mixtures for 1 h to facilitate dissolution, and standing them for 30 min; the supernatant liquid part was used for measurements. Because DTCI was only partially soluble in water, accurate concentrations of DTCI in water were determined by dissolving 50  $\mu$ L for the supernatant solutions made at 0.5 g/L, and 1 g/L, and 2 g/L or 300  $\mu$ L of the supernatant solution made at 0.1 g/L in 1 mL of MeOH and measuring absorbance spectra of the resulting solution whose methanol content was 95 volume % or 77 volume %. Molar absorptivities of DTCI at 758 nm in pure MeOH follow:  $\epsilon = 367000$  at  $1.0 \times 10^{-6}$  M [absorbance 0.037], 330000 at  $4.0 \times 10^{-6}$  M [absorbance 0.132], 313160 at  $1.0 \times 10^{-5}$  M [absorbance 0.313], 283427 at  $4.0 \times 10^{-5}$  M [absorbance 1.13], and 263220 at  $1.0 \times 10^{-4}$  M [absorbance 2.63]). Non-linear curve fitting led the following equation:

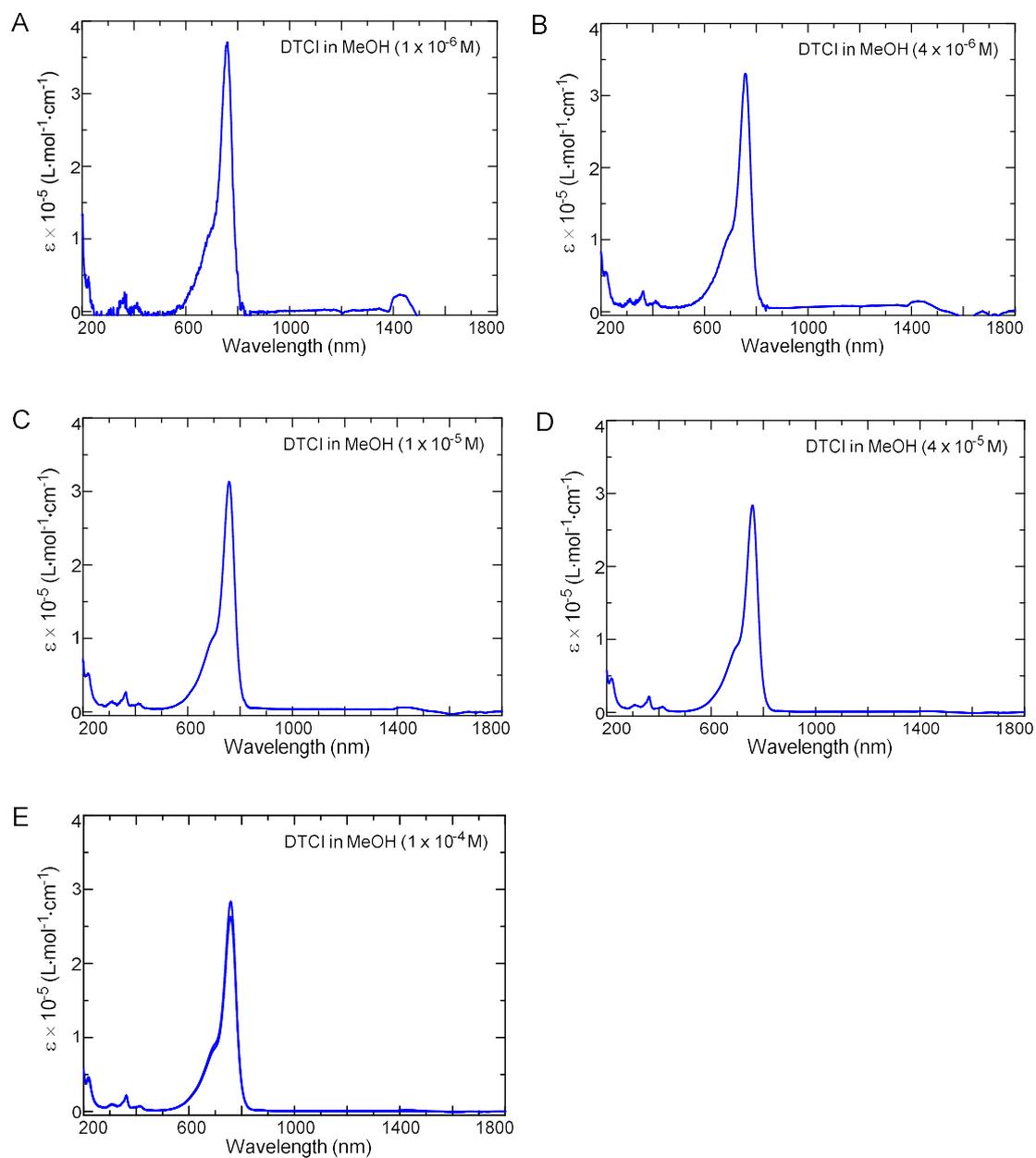
$$[DTCI] = 3.509 \times 10^{-5} \times (\text{Absorbance at 758 nm})^{1.083}$$

On the basis of this equation and absorbance values read from Figure S2 along with the dilution ratios led to the concentrations:  $[DTCI] = 7.2 \times 10^{-5}$  M for 0.1 g/L,  $2.0 \times 10^{-4}$  M for 0.5 g/L,  $4.8 \times 10^{-4}$  M for 1 g/L, and  $8.8 \times 10^{-4}$  M for 2 g/L.

### Instrumentation

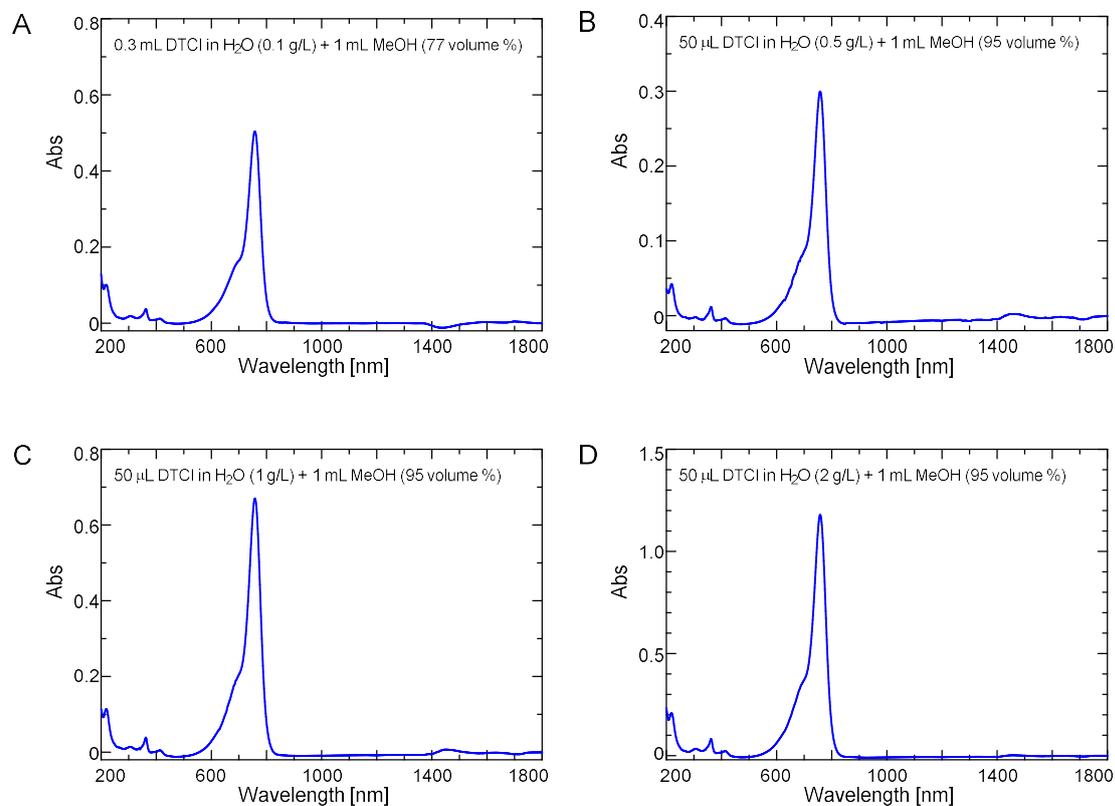
Circular dichroism (CD) spectra were taken with a JASCO-820 spectrometer. UV-vis absorption spectra were measured with a JASCO V-570 spectrophotometer. Emission spectra were taken on a JASCO FP-8500 fluorescence spectrophotometer.

## 2. Absorbance Spectra of DTCl in methanol



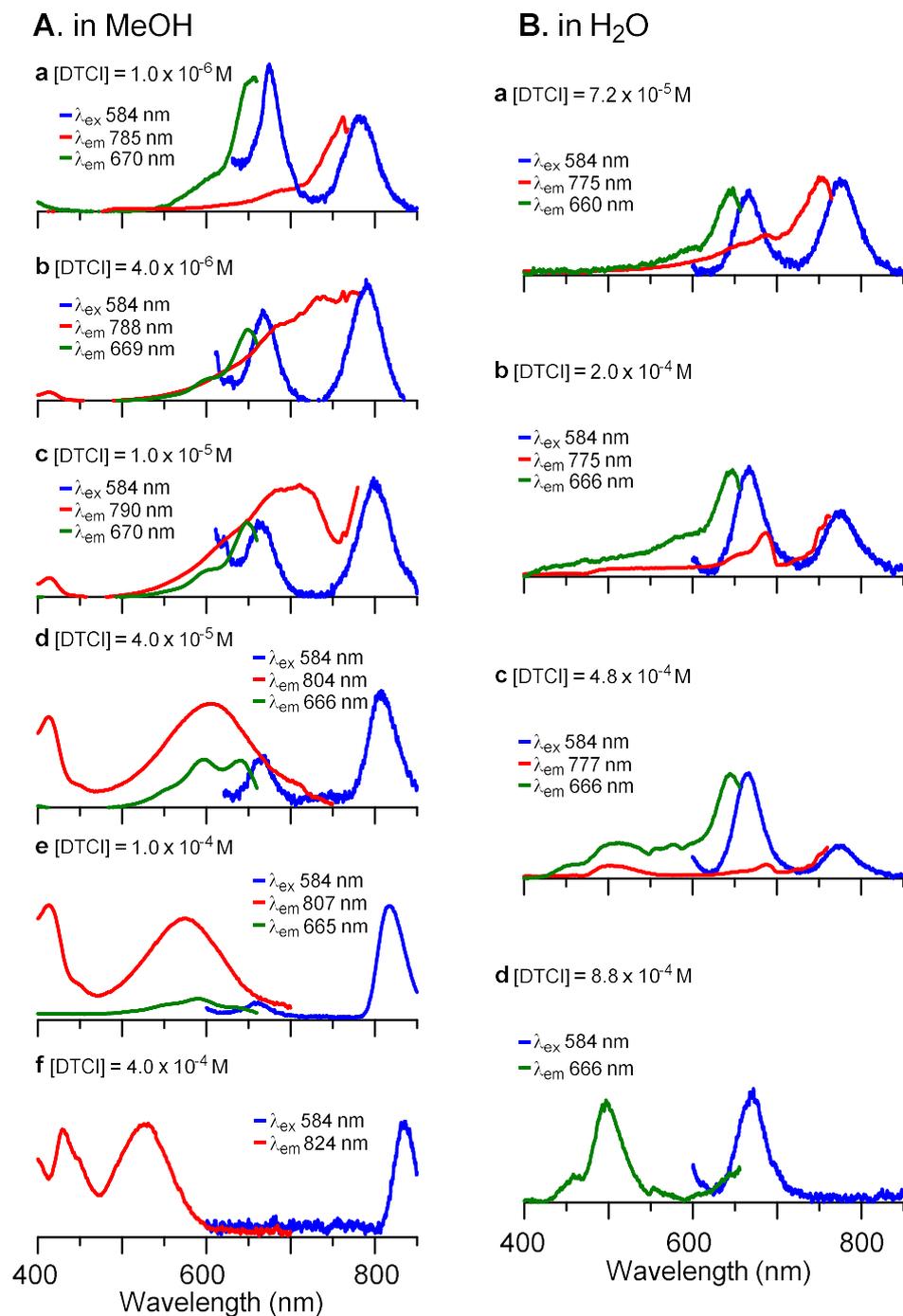
**Figure S1.** Absorbance spectra of DTCl in methanol at  $1.0 \times 10^{-6} \text{ M}$  (A),  $4.0 \times 10^{-6} \text{ M}$  (B),  $1.0 \times 10^{-5} \text{ M}$  (C),  $4.0 \times 10^{-5} \text{ M}$  (D), and  $1.0 \times 10^{-4} \text{ M}$  (D) [room temperature, 1-mm cell].

### 3. Absorbance Spectra of DTCl in methanol containing water



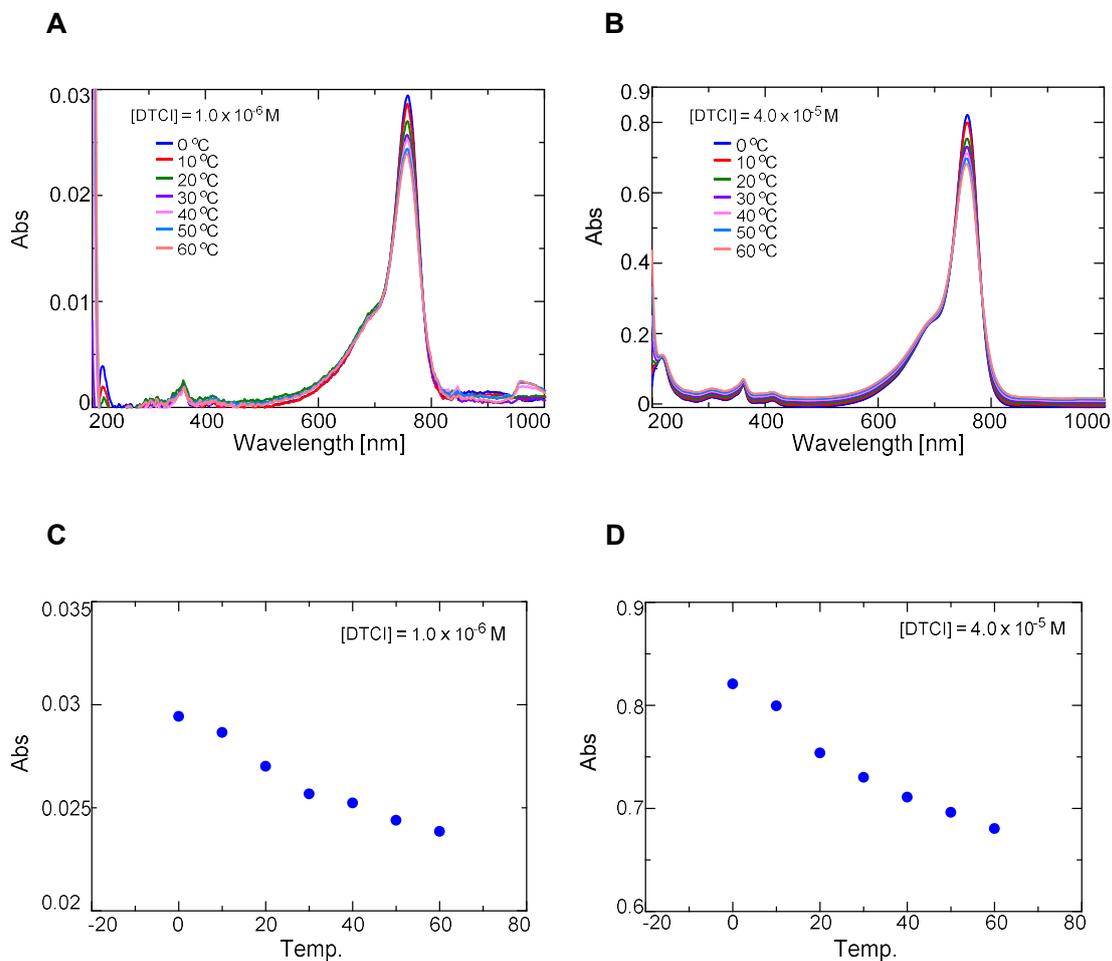
**Figure S2.** Absorbance spectra of DTCl in methanol containing water. The samples were prepared by dissolving supernatant water solution prepared at a ratio of DTCl and water of 0.1 g/L (A), 0.5 g/L (B), 1 g/L (C), and 2 g/L (D) [room temperature, 1-mm cell]. See Experimental text for further details.

**4. Fluorescence and excitation spectra of DTCl in methanol and in water measured through front-face illumination in a triangular cell or a in a square cell set diagonally to the incident light beam**



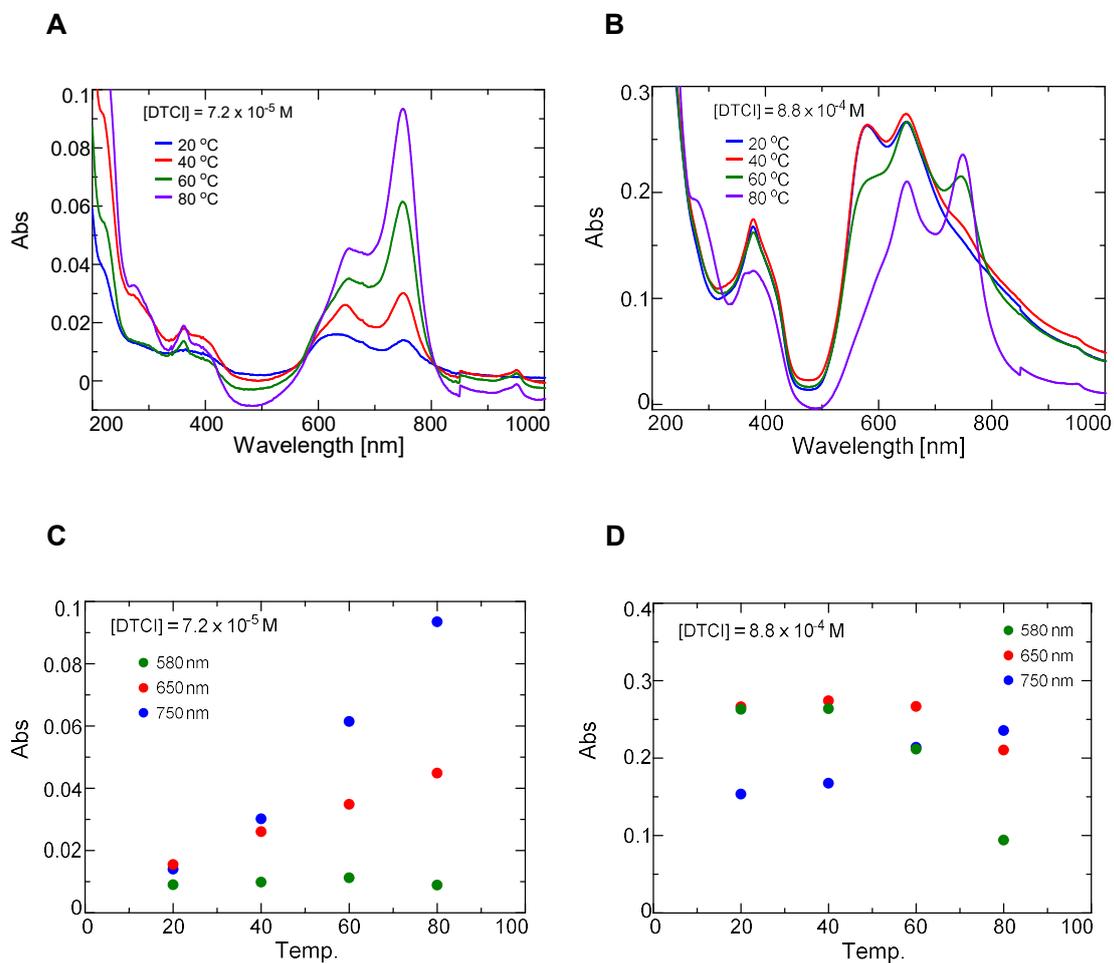
**Figure S3.** Emission (blue) and excitation (red and green) spectra of DTCl in methanol (A) and in water (B) measured through front-face illumination: [DTCl] a=  $1.0 \times 10^{-6}$  M (a),  $4.0 \times 10^{-6}$  M (b),  $1.0 \times 10^{-5}$  M (c),  $4.0 \times 10^{-5}$  M (d),  $1.0 \times 10^{-4}$  M (e), and  $4.0 \times 10^{-4}$  M (f) in A and  $7.2 \times 10^{-5}$  M (a),  $2.0 \times 10^{-4}$  M (b),  $4.8 \times 10^{-4}$  M (c), and  $8.8 \times 10^{-4}$  M (d) in B. Excitation wavelengths for fluorescence emission spectra ( $\lambda_{\text{ex}}$ ) and monitoring wavelengths for excitation spectra ( $\lambda_{\text{em}}$ ) are indicated above relevant spectra [room temperature, 10-mm cell].

### 5. Temperature effects on absorbance spectra in MeOH



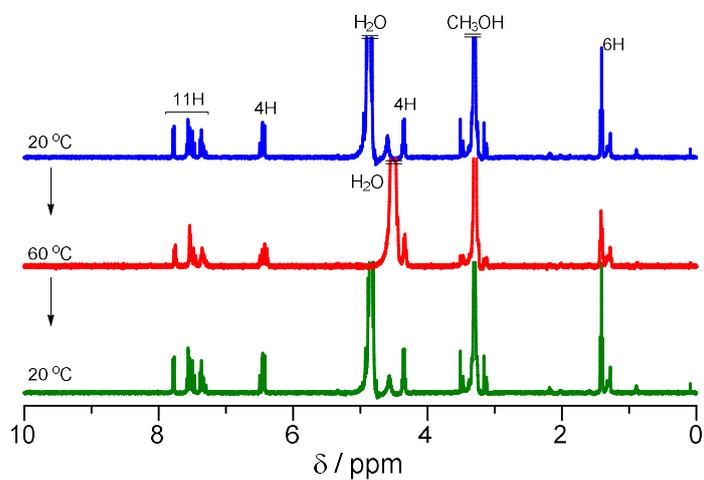
**Figure S4.** Absorbance spectra of DTCl in methanol at  $[DTCl] = 1.0 \times 10^{-6} \text{ M}$  (A) and  $4.0 \times 10^{-5} \text{ M}$  (B) measured at different temperatures and relations between absorbance at 758 nm and temperature at  $1.0 \times 10^{-6} \text{ M}$  (C) and  $4.0 \times 10^{-5} \text{ M}$  (D).

## 6. Temperature effects on absorbance spectra in H<sub>2</sub>O

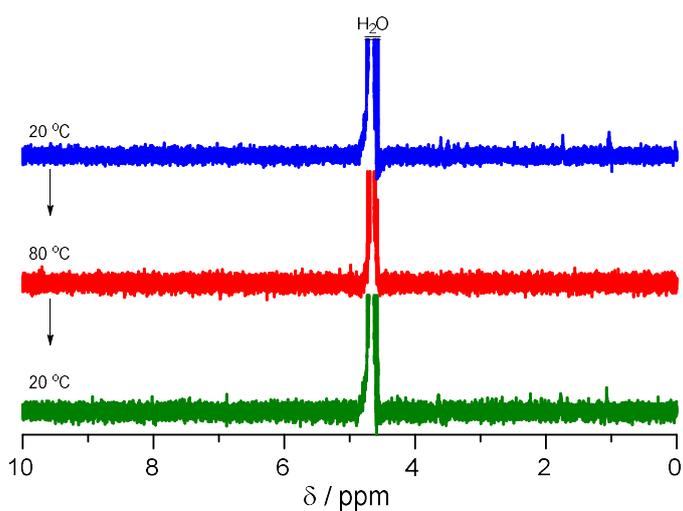


**Figure S5.** Absorbance spectra of DTCl in methanol at  $[DTCl] = 7.2 \times 10^{-5} \text{ M}$  (A) and  $8.8 \times 10^{-4} \text{ M}$  (B) measured at different temperatures and relations of absorbances at 580 nm, 650 nm, and 750 nm and temperature at  $7.2 \times 10^{-5} \text{ M}$  (C) and  $8.8 \times 10^{-4} \text{ M}$  (D).

## 7. Temperature effects on NMR spectra



**Figure S6.**  $^1\text{H}$  NMR spectra of DTICI taken at different temperatures in  $\text{CD}_3\text{OD}$  at  $[\text{DTICI}] = 4.0 \times 10^{-4}$  M.



**Figure S7.**  $^1\text{H}$  NMR spectra of DTICI taken at different temperatures in  $\text{D}_2\text{O}$  at  $[\text{DTICI}] = 8.8 \times 10^{-4}$  M.