

A multi-metal PtAgAu@CeO₂ core-shell nanocatalyst with improved catalytic performance

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Characterization

Transmission electron microscopy (TEM) experiments were conducted on a JEM-1230 microscope operated at 100 kV. The TEM samples were prepared by transferring one drop of sample dispersion in ethanol onto a carbon-coated copper grid and then dried in air. Energy dispersion X-ray analysis (EDX) were conducted on a JEM-1230 microscope operated at 100 kV.

Catalytic evaluation

The reduction of 4-NP to 4-AP was chosen as a model reaction to study the catalytic properties. In a typical run, 4-NP (0.03 mL, 0.01 M, aqueous solution), H₂O (2 mL) and nanocatalyst (0.5 mL, 1 mg/mL) were mixed with freshly prepared NaBH₄ (0.5 mL, 0.25 M, aqueous solution) in a quartz cell (3.0 mL). The absorption spectra were recorded on a UV-Vis spectrophotometer at a regular interval in the range of 250-700 nm.

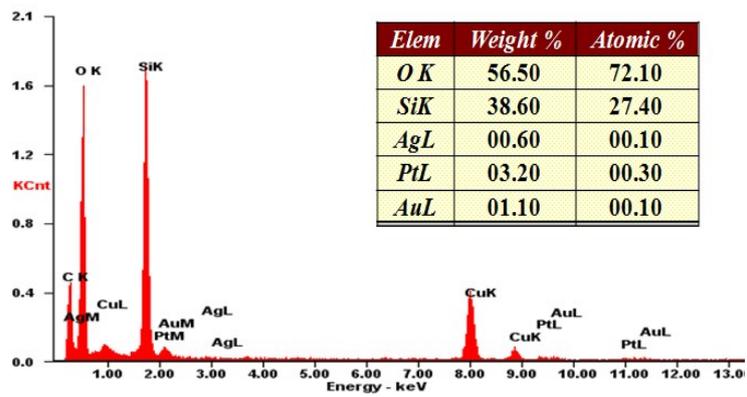


Fig. S1 The EDX analysis of Pt-Ag-Au@SiO₂.

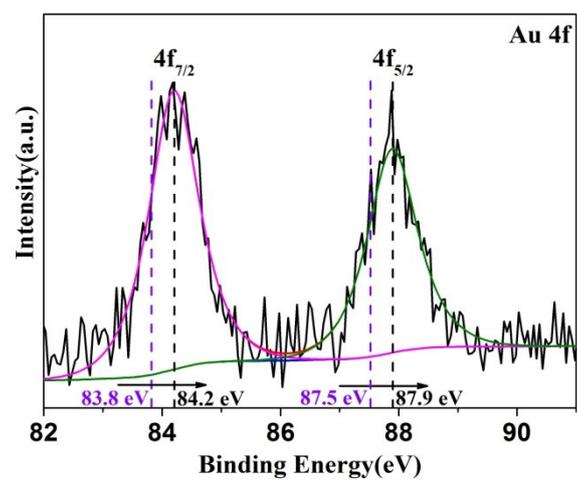


Fig. S2 XPS peaks of Au 4f in PtAgAu@SiO₂ nanospheres.

Table S1 ICP analysis of PtAgAu@CeO₂.

Element	Content (wt %)
Au	1.3142
Ag	0.7574
Pt	3.2816

Table S2 Comparison of rate constant for the catalytic reduction of 4-NP by NaBH₄ using Pt-based nanocatalysts.

Catalyst	Initial concentration of 4-NP (mM)	Amount of noble NPs (nmol)	k_{app} per noble NPs content (min⁻¹μmol⁻¹)	The multiple of k_{napp}
PtAgAu@CeO ₂ (This work)	0.1	152	1.65	1
PtCo/NaY ¹	7.2	579.5	1.0332	1.6
Au@SiO ₂ ²	0.1	135.9	0.84	1.96
Fe ₃ O ₄ @SiO ₂ -Au@mSiO ₂ ³	0.24	335	1.044	1.58
porous AuPt particles ⁴	0.24	2564	1.287	1.28
ultra-small Pt NPs ⁵	0.14	510	2.472	0.67

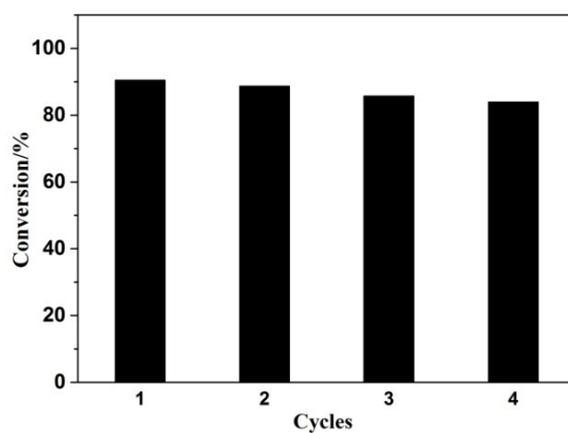


Fig. S3 Conversion of 4-NP in four successive cycles

with PtAgAu@CeO₂ nanocatalyst.

References

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