Electronic Supplementary Information

Hierarchical Fe-Doped Ni₃Se₄ Ultrathin Nanosheets as an Efficient Electrocatalyst for Oxygen Evolution Reaction

Jing Du^{a,b*}, Zehua Zou^a, Chen Liu,^c and Cailing Xu^{a*}

^aState Key Laboratory of Applied Organic Chemistry, Laboratory of Special Function
Materials and Structure Design of the Ministry of Education, College of Chemistry and
Chemical Engineering, Lanzhou University, Lanzhou 730000, China
^bKey Laboratory of Advanced Energy Materials Chemistry (Ministry of Education), Nankai
University, Tianjin 300071, China

^cSchool of Physical Science and Technology, Lanzhou University, Lanzhou 730000, China



Fig. S1. XRD pattern of LDH precursor.



Fig. S2. (a) TEM image of Ni₃Se₄ nanosheets. (b) TEM image of (Ni,Fe)₃Se₄ nanosheets.



Fig. S3. The energy dispersive X-ray spectrum obtained from the selected square in the TEM image.



Fig. S4. (a) N_2 adsorption-desorption isotherms for $(Ni,Fe)_3Se_4$ nanosheets. (b) The characteristic BJH pore size distributions.



Fig. S5. Electrochemical double-layer capacitance measurements. The cyclic voltammograms (CVs) measurements with various scan rates for (a) $(Ni,Fe)_3Se_4$ and (b) Ni_3Se_4 nanosheets in 1.0 M KOH.



Fig. S6. (a) Specific and (b) mass activities of $(Ni,Fe)_3Se_4$ and Ni_3Se_4 electrodes for the OER in 1.0 M KOH.



Fig. S7. XPS spectra of the (a) Ni 2p, (b) O 1s, and (c) Se 3d peaks of (Ni,Fe)₃Se₄ nanosheets

after OER measurements.



Fig. S8. EDX maps of (Ni,Fe)₃Se₄ nanosheets after OER measurements.



Fig. S9. HRTEM image of (Ni,Fe)₃Se₄ nanosheets after OER measurements.



Fig. S10. Polarization curves of GC, Ni_3Se_4 and $(Ni_4Fe)_3Se_4$ nanosheets in 1.0 M KOH at a potential sweep rate of 5 mV/s.



Voltage (V) Fig. S11. Polarization curve of overall water splitting by using Pt/C-IrO₂ couple with a scan rate of 5 mV s⁻¹ in 1 M KOH.