

SUPPORTING INFORMATION

Electrochemistry of layered metal diborides

V. Mazánek^a, Hindia Nahdi^b, J. Luxa^a, Z. Sofer^a, M. Pumera^{a,b}

^a University of Chemistry and Technology, Department of Inorganic Chemistry, 166 28 Prague 6, Czech Republic. E-mail: zdenek.sofer@vscht.cz, Fax: +420 22431-0422

^b Division of Chemistry & Biological Chemistry, School of Physical and Mathematical Sciences, Nanyang Technological University, Nanyang Link 21, Singapore, 637371, Singapore. Fax: +65 6791-1961, E-mail: pumera@ntu.edu.sg

Table SI1. List of powder diffraction files used for identification of phase composition from diffractograms.

Phase	PDF	Phase	PDF	Phase	PDF
AlB₂	04-010-4305	HfB	04-004-7156	TiB₂	03-065-1073
Al₂O₃	01-076-8057	MgB₂	04-013-1803	TiN	03-065-0414
Al	04-003-4850	MgO	04-012-3469	VB₂	04-006-2024
CrB₂	04-003-6119	MgB₄	03-065-1310	VO₂	04-005-7410
Cr₃B₄	04-003-0080	NbB₂	03-065-0512	ZrB₂	04-004-2991
HfB₂	04-001-1202	TaB₂	01-075-0966	ZrB	03-065-2907

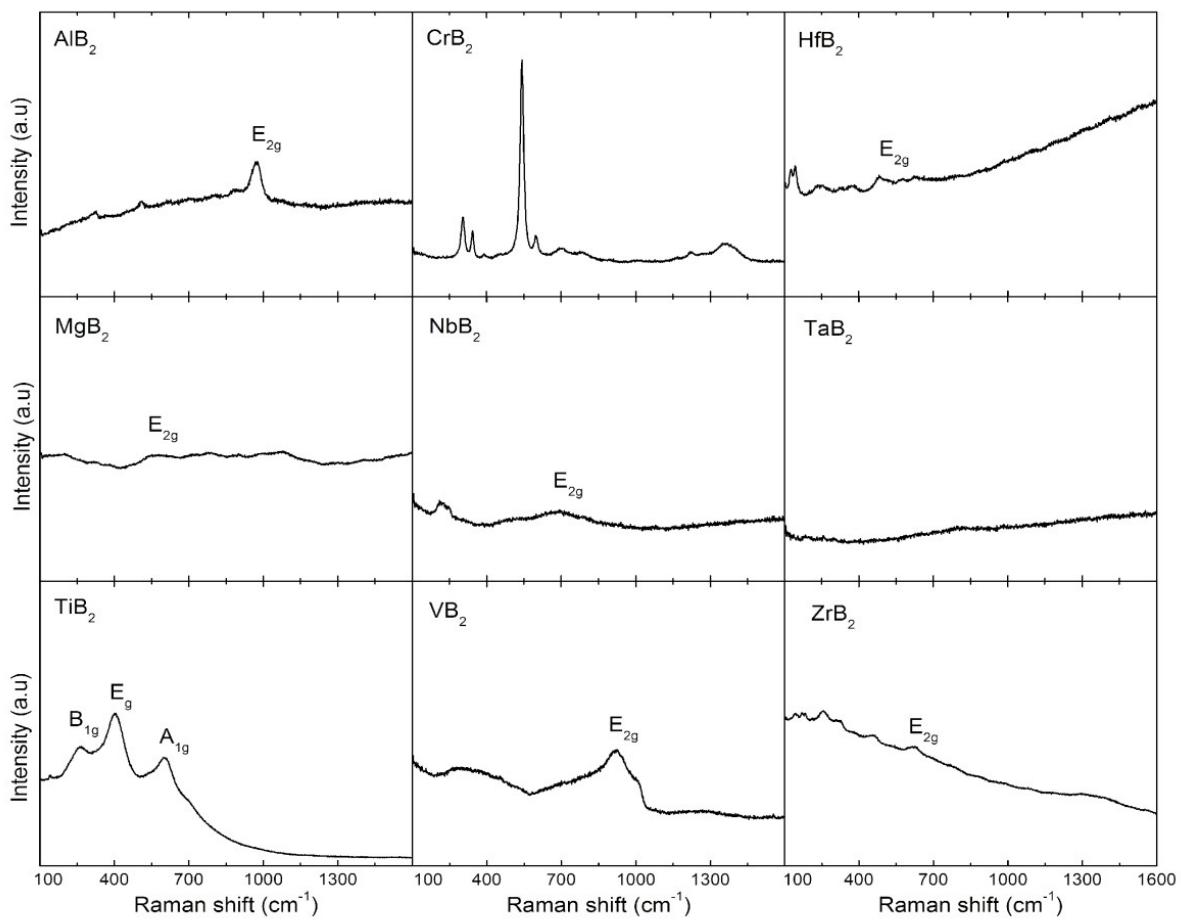


Figure S1. Raman spectra of layered diborides.

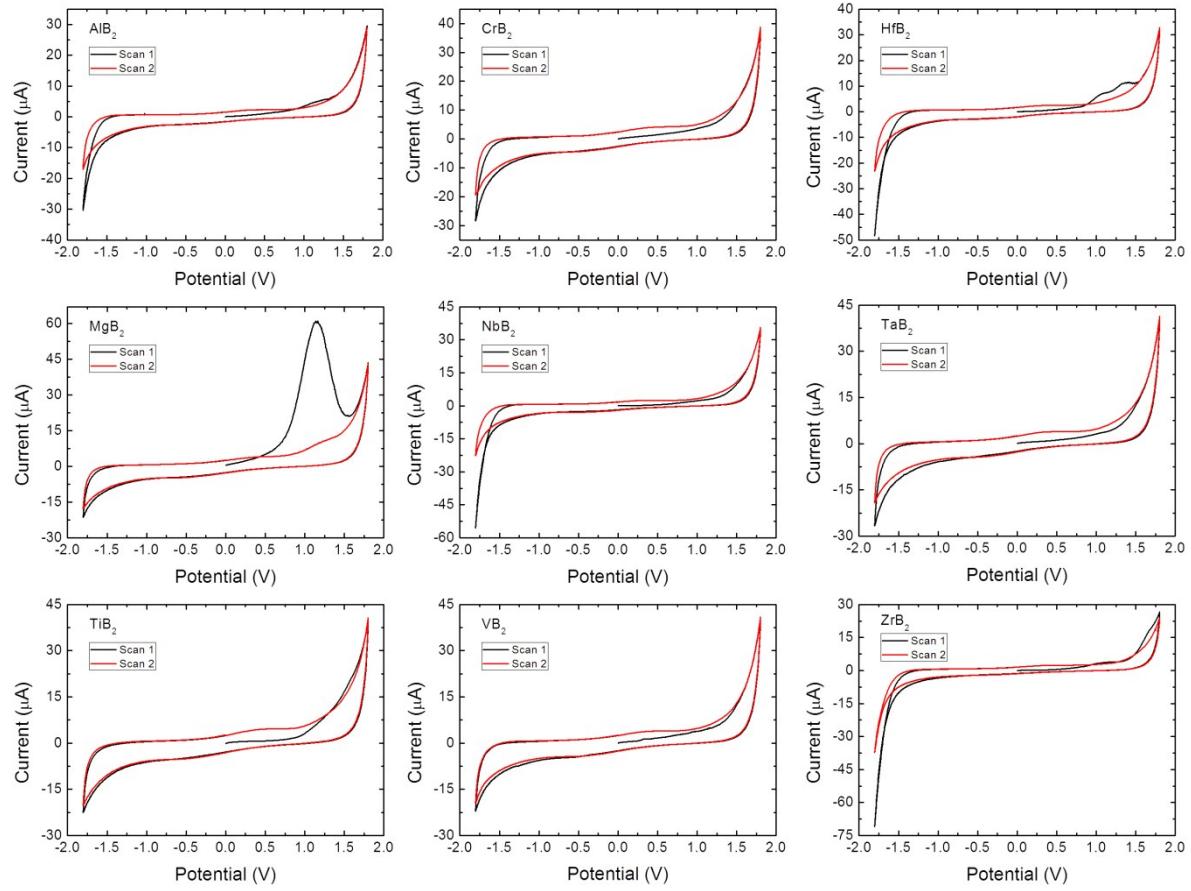


Figure S2. Inherent electrochemistry of layered diborides. The scan start at 0.0 V vs. sat. Ag/AgCl in anodic direction. Conditions: supporting electrolyte 50 mM PBS pH 7.2 scan rate 100 mVs⁻¹.

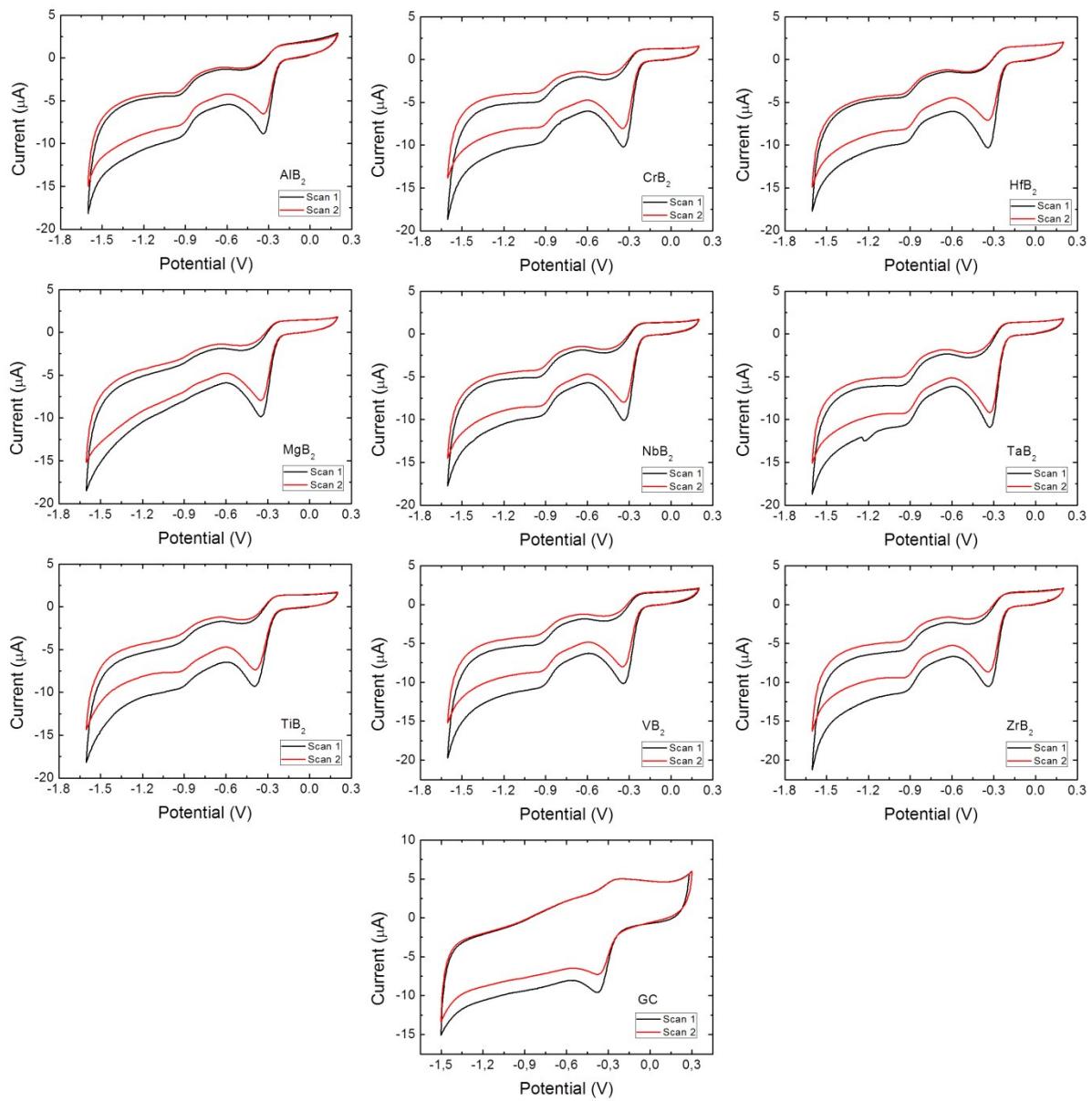


Figure S3. ORR measured by cyclic voltammetry on layered diborides and bare GC electrode.