

Supporting Information for Nanoscale

Sodium Storage in Promising MoS₂-Carbon Anode: Elucidating Structural and Interfacial Transition in Intercalation Process and Conversion Reactions

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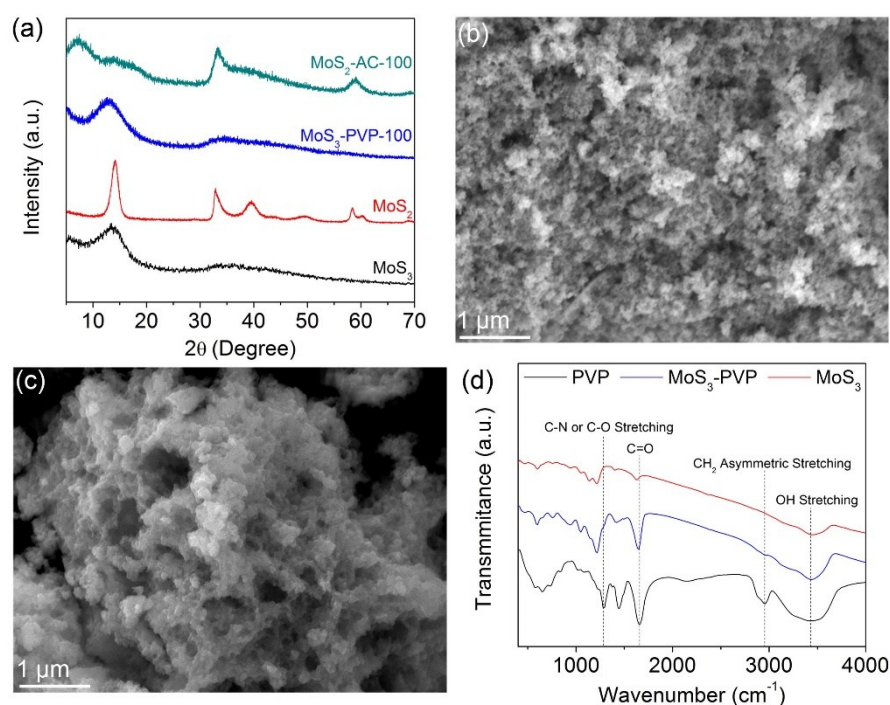


Fig. S1 (a) XRD patterns of MoS₃ and MoS₃-PVP, (b) SEM image of MoS₃, (c) SEM image of MoS₃-PVP, (d) FTIR spectrum of PVP, MoS₃, MoS₃-PVP. MoS₃-PVP represents the precursor for MoS₃-AC-100 sample. The XRD patterns of MoS₃ and MoS₃-PVP show the typical profile of amorphous MoS₃. MoS₃ and MoS₃-PVP precursors display the granular morphology in nanosize.^{S1,S2} The grain size of MoS₃-PVP is slightly larger than that of MoS₃, most likely due to PVP coating. The absorption peaks of FTIR from the chemical groups of PVP including C-N, C-O and CH₂ can be observed in MoS₃-PVP samples prepared with additional PVP.^{S3}

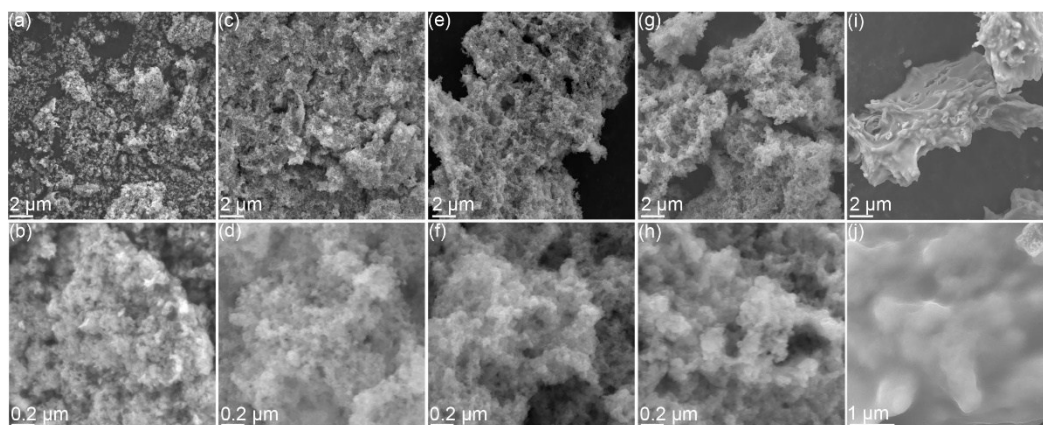


Fig. S2 SEM images for morphology characterization. (a) and (b) for MoS₂, (c) and (d) for MoS₂-AC-20, (e) and (f) for MoS₂-AC-50, (g) and (h) for MoS₂-AC-100, (i) and (j) for MoS₂-AC-200.

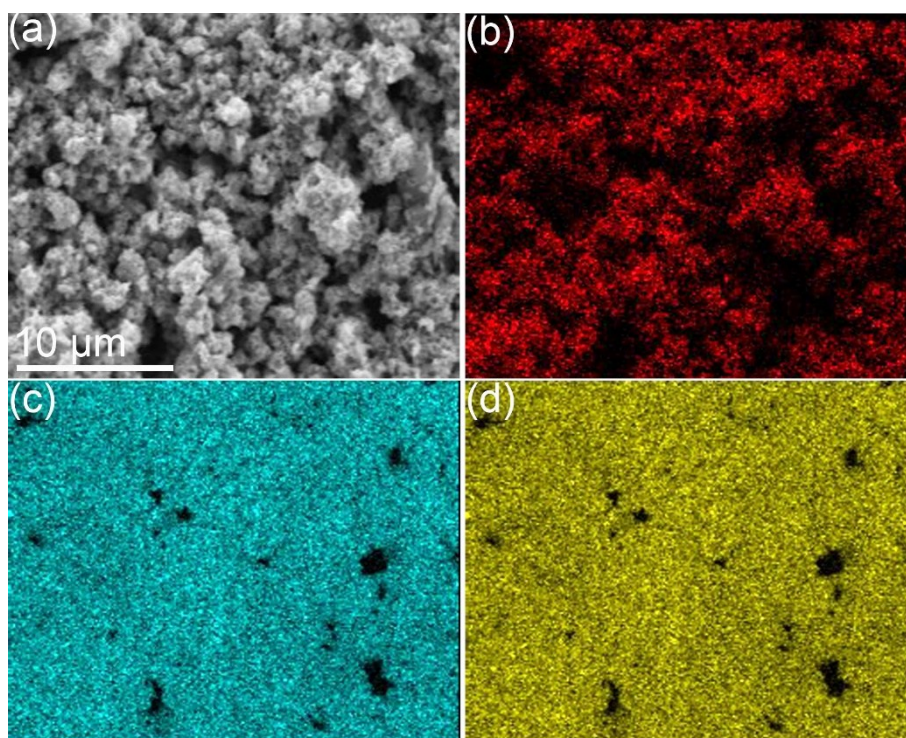


Fig. S3 (a) SEM image of selected area for EDS test for MoS₂-PVP-100 and the corresponding element mappings: (b) C element, (c) Mo element and (d) S element.

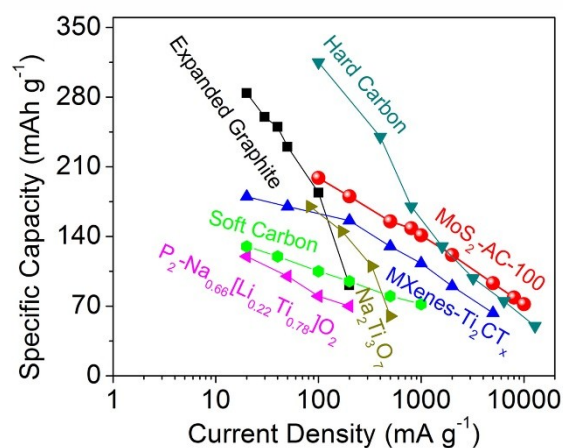


Fig. S4 Rate capability for MoS₂-AC compared to other intercalated anodes in reported literatures: hard carbon,^{S4} MXenes,^{S5} expanded graphite,^{S6} Na₂Ti₃O₇,^{S7} P₂-Na_{0.66}Li_{0.22}Ti_{0.78}O₂,^{S8} and soft carbon.^{S9}

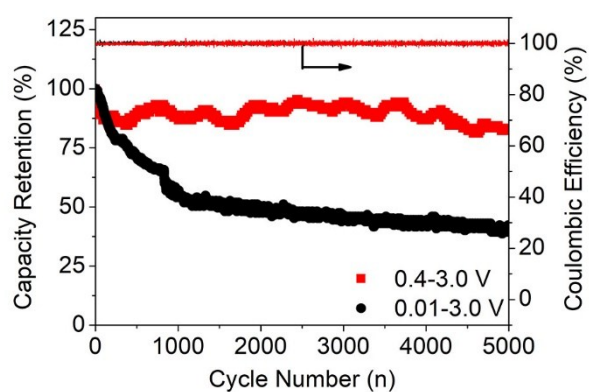


Fig. S5 Long-term cycling performance of MoS₂-AC-100 at a current density of 5 A g⁻¹ within two potential ranges of 0.4-3.0 V and 0.01-3.0 V.

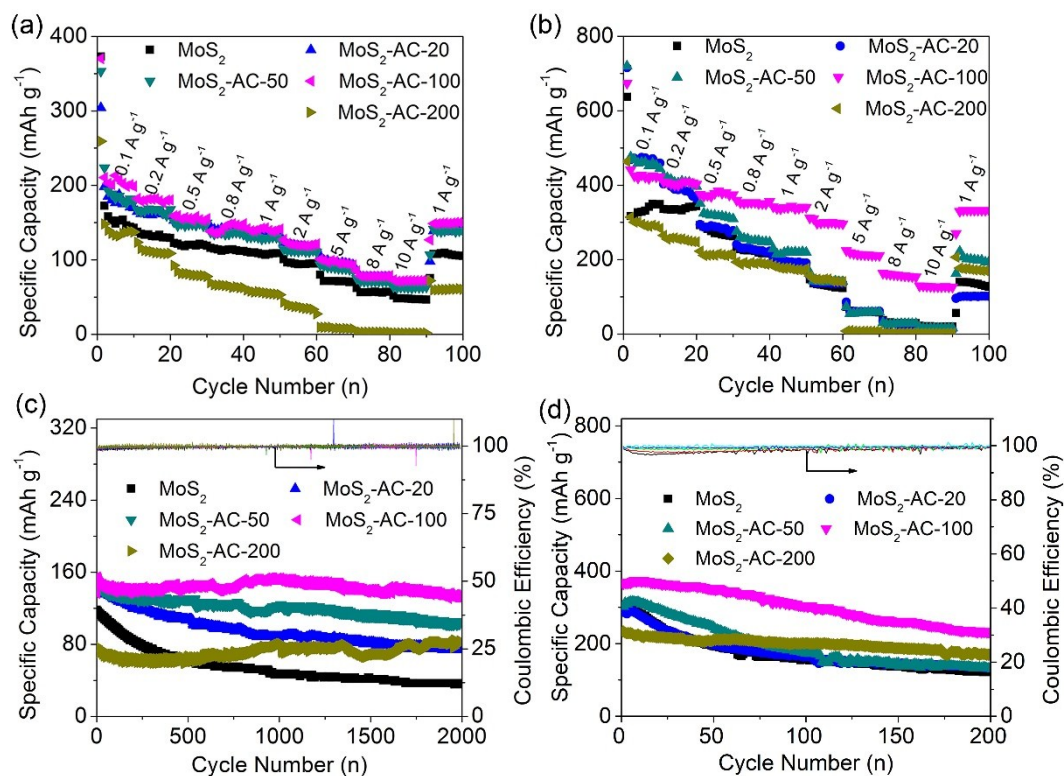


Fig. S6 (a) Rate performance for MoS₂-AC electrodes at various current densities from 0.1 to 10 A g⁻¹ within a working potential range of 0.4-3.0 V. (b) Rate performance for MoS₂-AC electrodes at various current densities from 0.1 to 10 A g⁻¹ within a working potential range of 0.01-3.0 V. (c) and (d) Long-term cycling performance for MoS₂-AC electrodes at 1 A g⁻¹ within two working potential ranges: 0.4-3.0 V and 0.01-3.0 V, respectively.

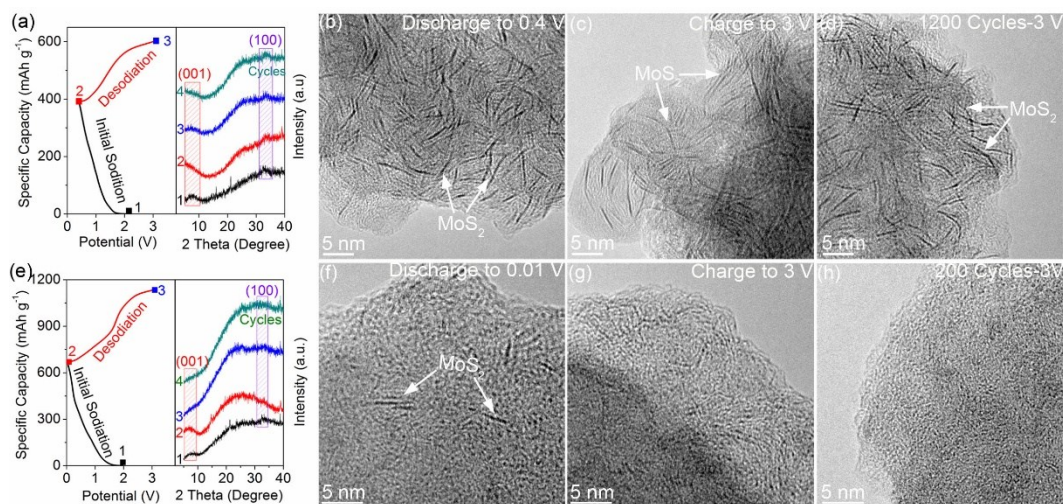


Fig. S7 Initial charge/discharge process for MoS₂-AC-100 at a current density of 0.1 A g⁻¹ and the corresponding ex-situ XRD patterns for MoS₂-AC-100 during the cycles within two trade-off potentials: (a) 0.4 V and (e) 0.01 V. The notes: 1-pristine material, 2-initial sodiation, 3-initial desodiation. The note for 4 is the desodiated state after long cycles of 1200 cycles for (a) and 200 cycles for (b) at a current density of 1 A g⁻¹. HRTEM images of MoS₂-AC-100 sample cycled within 0.4-3.0 V: (b) the initial sodiation to 0.4 V (vs. Na/Na⁺), (c) initial desodiation to 3.0 V (vs. Na/Na⁺), (d) 1200th cycle to 3.0 V (vs. Na/Na⁺). HRTEM images of MoS₂-AC-100 sample cycled within 0.01-3.0 V: (f) the initial sodiation to 0.01 V (vs. Na/Na⁺), (g) initial desodiation to 3.0 V (vs. Na/Na⁺), (h) 200th cycle to 3.0 V (vs. Na/Na⁺). The current density is 0.1 A g⁻¹, except for 1 A g⁻¹ for long cycles for (d) and (h).

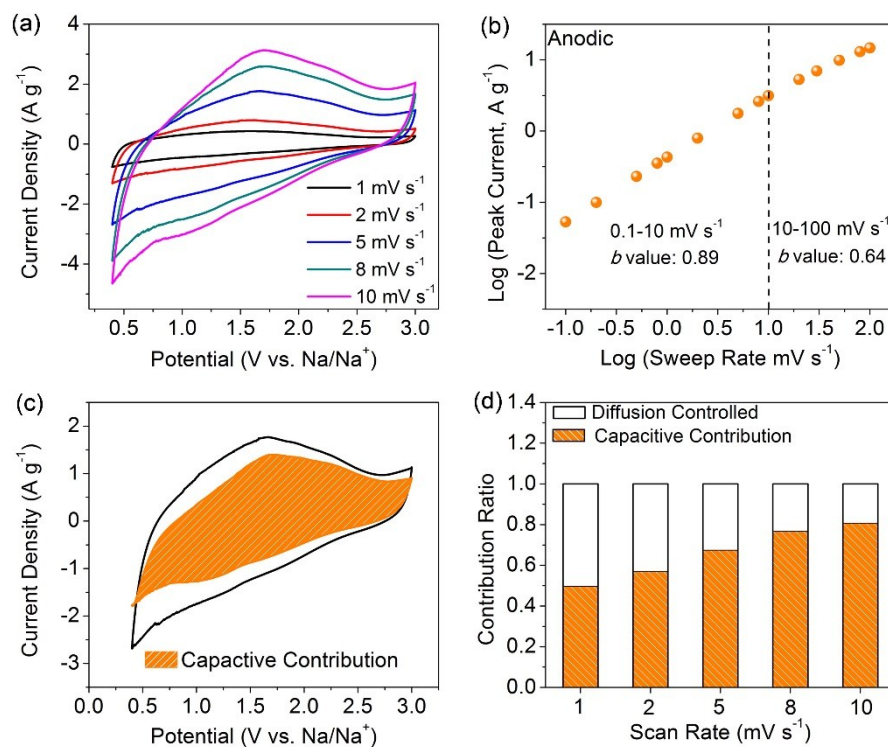


Fig. S8 Electrochemical kinetics analysis of MoS₂-AC-100 electrode. (a) CV curves at various scan rates from 1 to 10 mV s⁻¹. (b) Determination of the *b*-value using the relationship between anodic peak current and scan rate. (c) Separation of the capacitive and diffusion currents at a scan rate of 5 mV s⁻¹, respectively. (d) Contribution ratio of the capacitive charge versus scan rate.

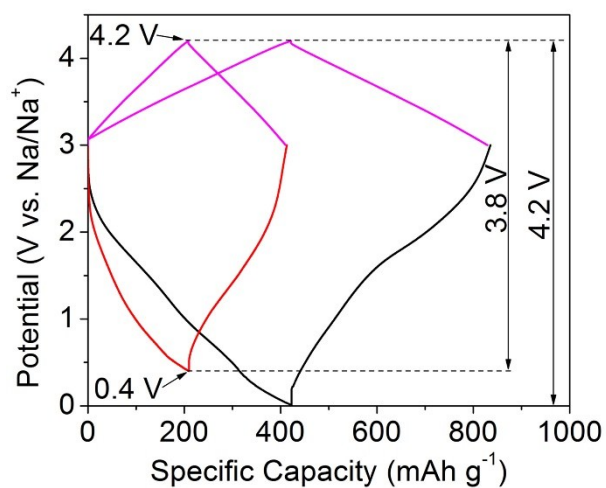


Fig. S9 Charge/discharge curves of MoS₂-AC-100 within two different potential windows (0.4-3.0 V (red line) and 0.01-3.0 V (black line)) and PDPC versus Na/Na⁺, respectively. The specific current density is 1 A g⁻¹.

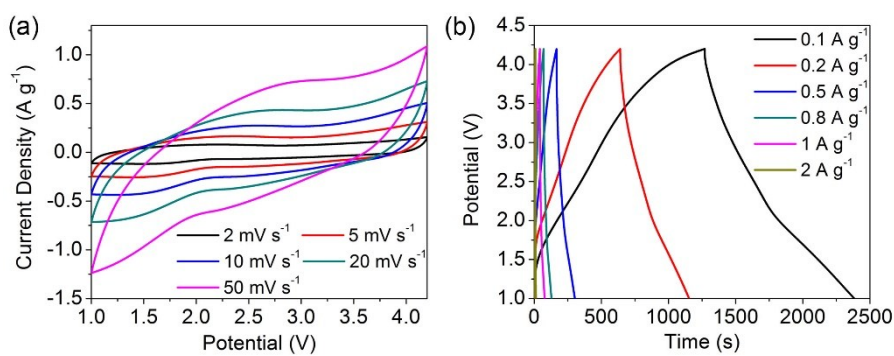


Fig. S10 (a) CVs of as-fabricated MoS₂-AC//PDPC-4.2 V at various scan rates from 2 to 50 mV s⁻¹. (b) Charge/discharge curves of as-fabricated MoS₂-AC//PDPC-4.2 V SIC at the different current densities from 0.1 to 2 A g⁻¹.

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