Electronic Supporting Information

Geometric Deconstruction of Core and Electron Activation of π -System in Deformed Porphyrin Series: Mimics of Heme

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1. UV/Vis spectral data

Compound	Soret region	Q region				
1 - <i>s</i>	446 (5.30)	537 (4.13)	573 (3.64)	630 (2.78)	711 (3.58)	
2 - <i>s</i>	438 (5.46)	532 (4.20)	568 (3.82)	611 (2.94)	689 (3.64)	
3 - <i>s</i>	428 (5.63)	522 (4.24)	558 (3.86)	593 (3.11)	663 (3.74)	
4 - <i>s</i>	425 (5.68)	519 (4.24)	552 (3.91)	585 (3.16)	646 (3.53)	
1 - <i>w</i>	457 (5.19) ^{<i>a</i>}	551 (3.95)	586 (3.71)	645 (3.14)	722 (3.45)	
2 - <i>w</i>	436 (5.56)	529 (4.24)	565 (3.86)	610 (3.04)	678 (3.64)	
3- <i>w</i>	428 (5.66)	522 (4.28)	557 (3.86)	592 (3.14)	663 (3.70)	
4 - <i>w</i>	426 (5.69)	520 (4.27)	551 (3.92)	585 (3.17)	656 (3.54)	
1 - <i>r</i>	447 (5.33)	546 (4.02)	588 (3.81)	622 (3.62)	682 (3.19)	
2 - <i>r</i>	440 (5.44)	538 (4.08)	578 (3.92)	613 (3.74)	672 (3.41)	
3 - <i>r</i>	431 (5.57)	527 (4.18)	563 (3.87)	612 (3.73)	660 (3.31)	
4 - <i>r</i>	423 (5.71)	516 (4.31)	550 (4.09)	593 (3.84)	649 (3.72)	

Table S1. UV/Vis spectral data, λ_{max} [nm] (log ε) of all mimics in CHCl₃ at 293 K

^{*a*}: Sample 1-*w* show a maximal spectral red-shift, the shift is up to 39 nm compared with that of regular porphyrin TPP ($\lambda_{max} = 418$ nm).







Figure S1. Q-band comparison

2. The fluorescent comparison



Figure S2. Fluorescent comparison

3. The computation results

Energy Level	E(HOMO)		E(LUMO)		Δ <i>E</i> (LUMO-HOMO)	
material	(a.u.)	(eV)	(a.u.)	(eV)	(a.u.)	(eV)
1 - <i>s</i>	-0.1781	-4.85	-0.0874	-2.38	0.0908	2.47
1 - <i>w</i>	-0.1737	-4.73	-0.0825	-2.24	0.0912	2.48
1 - <i>r</i>	-0.1835	-5.03	-0.0895	-2.45	0.0899	2.58
TPP	-0.1897	-5.16	-0.0913	-2.48	0.0985	2.68

Table S2 HOMO and LUMO energy level of deformed porphyrins 1-s, 1-w and 1-r



Figure S3. LUMO orbitals of materials 1-s, 1-w, 1-r and TPP.

4. Additional Electrochemical and spectroelectrochemical results

4.1 Electrochemical results

Figure S4. All V_{ap} and V_{cp} values

Note: V_{ap} and V_{cp} denote peak potentials of reduction and oxidation, respectively.

4.2 spectroelectrochemical results

Figure S5. Spectroelectrochemical changes upon the macrocycle-centered reductions.

Materials: 1-*w* to 4-*w*, under experimental voltages of -2.2, -1.9, -1.1 V ($a \rightarrow c$) in benzonitrile, 0.1 M TBAPF₆. The inserted equations denote the electrochemical reactions.

Figure S6. Spectroelectrochemical changes upon the macrocycle-centered oxidations

Materials: 1-*w* to 4-*w*, under experimental voltages of +0.9, +0.6 and +0.3 V ($d \rightarrow f$) in benzonitrile, 0.1 M TBAPF₆. The inserted equations denote the electrochemical reactions.

4.3 Setting of experimental voltages

Take the waved porphyrins as examples.

Figure S7. Setting of experimental voltages in spectroelectrochemical measurements.

The large macrocyclic distortion facilitates not only the electronic transition of the center metal, as manifested in our previous reports, [Inorg. Chem. 2018, 57, 277] but also the electron transfer of the macrocycle itself.

4.4 Absorbance changes

Figure S8. Absorbance changes (ΔA) over time at voltage of -2.2, -1.9, -1.1, 0.3, 0.6 and 0.9 V.

Spectroelectrochemistry can well distinguish the reaction rate, the reaction progress, and the difference between the oxidation and reduction of each compound.

5. HR-MS of new compounds

5.1 HR-MS of 1-s

All MS are obtained from ESI source.

5.2 HR-MS of **2**-*s*

5.3 HR-MS of **3**-*s*

5.4 HR-MS of **4**-*s*

5.5 HR-MS of **1**-*w*

5.6 HR-MS of **2**-*w*

5.7 HR-MS of **3**-*w*

5.8 HR-MS of **4**-*w*

6. NMR of new compounds

¹H NMR, HH COSY and ¹³C NMR of **1**-*s* to **4**-*s*;

¹H NMR, HH COSY and ¹³C NMR of **1**-*w* to **4**-*w*;

Note: NMR characterization of 1-*r* to 4-*r* can be reproduced from previous report.

