

Electronic Supporting Information

Describing Oxymethylene Ether Synthesis based on the application of Non-Stoichiometric Gibbs Minimisation

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UNCONSTRAINED GIBBS MINIMIZATION TECHNIQUE

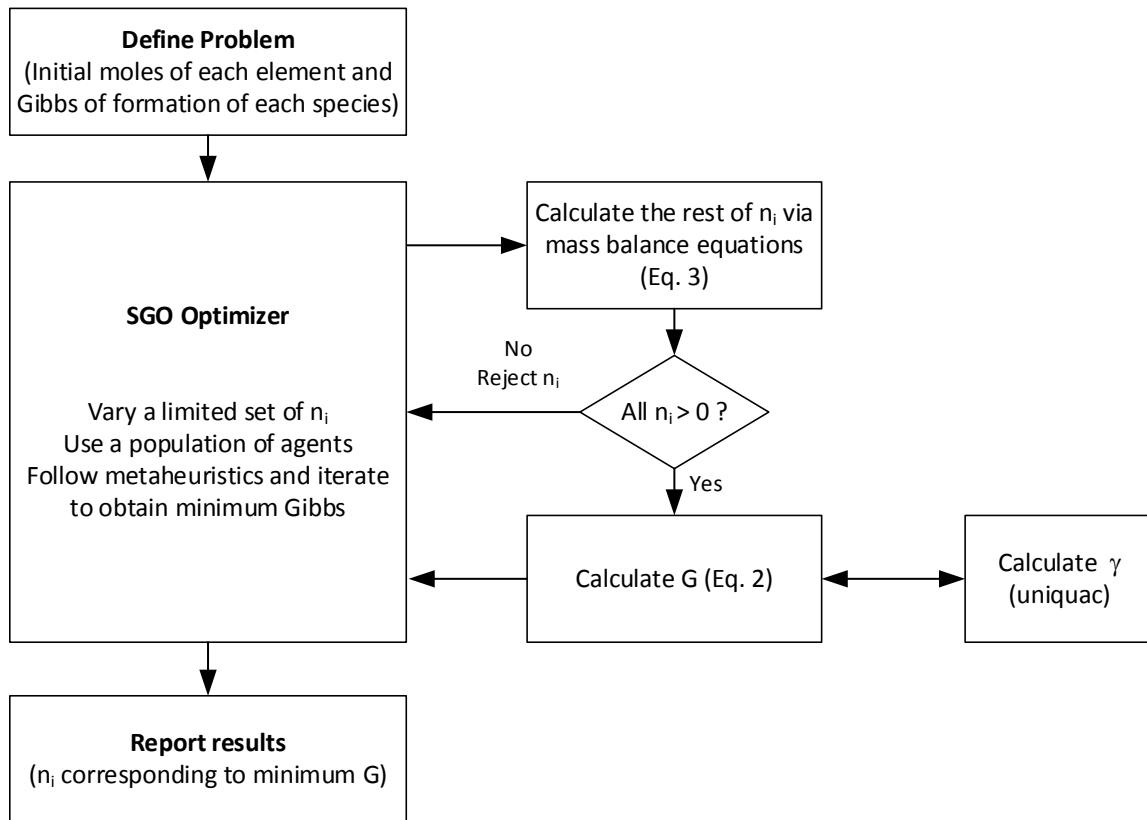


Figure S1 Scheme of the unconstrained Gibbs minimisation approach

IMPLEMENTATION ON OME LIQUID-PHASE REACTION

Table S1 Standard Gibbs of formation values at $T = 333.15$ K for different species in the system calculated using data extraction method

Species	Std. Gibbs Energy of formation (Liquid) at 333.15 K (kJ/mol)
OME ₁	-207.12
OME ₂	-316.79
OME ₃	-426.46
OME ₄	-536.13
OME ₅	-645.80
OME ₆	-755.48
OME ₇	-865.15
OME ₈	-974.82
MG ₁	-336.55
MG ₂	-448.66
MG ₃	-558.37
MG ₄	-668.07
MG ₅	-777.78
MG ₆	-887.49
MG ₇	-997.20
MG ₈	-1106.90
HF ₁	-273.69
HF ₂	-383.12
HF ₃	-491.55
HF ₄	-599.98
HF ₅	-708.42

HF ₆	-816.85
HF ₇	-925.28
HF ₈	-1033.71
FA	-94.81
MeOH	-158.01
H ₂ O	-231.38

Table S2. UNIFAC structural groups with size- and surface parameters (Adapted from Schmitz *et al.* (Ref. 1) and Kuhnert et al. (Ref. 2))

Structural group	Number	R	Q
CH ₂ O	1	0.9183	0.78
H ₂ O	2	0.92	1.4
H ₃ C-O-CH ₂ O-CH ₃	3	2.9644	2.716
HO-CH ₂ O-H	4	2.6744	2.94
OH	5	1	1.2
CH ₂	6	0.6744	0.54
CH ₃ O	7	1.1459	1.088
CH ₂ OH	8	1.2044	1.124
CH ₃ OH	9	1.4311	1.432
(CH ₂ O)OME	10	0.9183	0.78

Table S3. UNIFAC interaction parameters a_{km}/K (Adapted from Schmitz *et al.* (Ref. 1) and Kuhnert et al. (Ref. 2))

<i>k</i>	<i>m</i>									
	1	2	3	4	5	6	7	8	9	10

1	-	867.8	0	189.2	237.7	83.36	0	238.4	238.4	0
2	-254.5	-	a _{2,3} (T)	189.5	-229.1	300	-219.3	a _{2,8} (T)	289.6	a _{2,10} (T)
3	0	a _{3,2} (T)	-	a _{3,2} (T)	237.7	83.36	0	0	410	26
4	59.2	-191.8	a _{2,3} (T)	-	-229.1	300	-142.4	289.6	289.6	59.2
5	28.06	353.5	28.06	353.5	-	156.4	112.8	-137.1	-137.1	28.06
6	251.5	1318	251.5	1318	986.5	-	447.8	697.2	697.2	251.5
7	0	423.8	0	774.8	1164.8	273	-	238.4	238.4	0
8	-128.6	a _{8,2} (T)	0	-181	249.1	16.5	-128.6	-	0	-128.6
9	-128.6	-181	-71.21	-181	249.1	16.5	-128.6	0	-	-128.6
10	0	670.7	141.5	189.2	237.7	83.36	0	238.4	238.4	-

$$a_{2,3}(T) = -225.5 + 0.7205(T/K) ; \quad a_{3,2}(T) = 1031.0 - 1.749(T/K) ; \quad a_{8,2}(T) = -1018.57 + 329900/(T/K) ; \quad a_{2,8}(T) = 451.64 - 114100/(T/K) ; \\ a_{2,10}(T) = 168.9 - 0.8776(T/K)$$

Table S4 UNIFAC group assignment for all components adapted from K (Adapted from Schmitz *et al.* (Ref. 1) and Kuhnert *et al.* (Ref. 2))

Substance/Group	CH ₂ O (FA)	H ₂ O	C ₃ H ₈ O ₂ (OME ₁)	OH(CH ₂ O)H (MG ₁)	OH	CH ₂	CH ₃ O	CH ₂ OH	CH ₃ OH	CH ₂ O (OME)
FA	1									
Water		1								
Methanol									1	
Methylal (OME₁)			1							
MG₁				1						
HF₁							1	1		
MG_{n>2}	n-1				2	1				
HF_{n>2}	n-1						1	1		
OME_{n>2}			1							n-1

EXPERIMENTAL VALIDATION

Table S5 Different educts used for chemical equilibrium of the OME system Adapted from the supporting information of Schmitz et al. (Ref. 2)

Educt	Total Mass of Educt (g)	FA (mass fraction)	Water (mass fraction)	MeOH (mass fraction)
1	820	0.4841	0.0226	0.4933
2	790	0.3669	0.2225	0.4106
3	798	0.4743	0.1866	0.3391
4	776	0.3682	0.3454	0.2865
5	779	0.3447	0.0227	0.6315
6	763	0.2824	0.2181	0.4992
7	763	0.4713	0.0288	0.4996
8	801	0.3970	0.0152	0.5879
9	813	0.5697	0.0235	0.4014
10	791	0.4289	0.113	0.4580
11	799	0.5140	0.1209	0.3651

Table S6 Equilibrium composition of the OME synthesis for different reaction conditions and different feed compositions in comparison to the literature values (Ref. 1 and 2).

EQ (OME 1:TRI)	T [°C]	Exp.	p [bar]	TRI	[wt.%]									
					OME ₁	OME ₂	OME ₃	OME ₄	OME ₅	OME ₆	OME ₇	OME ₈	MEFO	MeOH
3:1	50	B	1.5	1.28	39.12	24.76	15.57	9.17	5.02	2.70	1.39	0.65	0.00	0.13
	50	Own	3	1.22	44.08	24.59	13.99	8.03	3.89	1.79	0.81	0.34	0.01	1.25
	65	B	1.9	1.45	39	24.72	15.53	9.16	5.00	2.71	1.39	0.65	0.00	0.16

	65	Own	4	1.44	44.20	24.33	13.77	8.02	3.91	1.87	0.81	0.34	0.05	1.26
	80	B	2.6	1.62	38.96	24.69	15.43	9.10	4.96	2.67	1.37	0.65	0.09	0.17
	80	Own	6.3	1.61	44.07	24.67	13.84	7.85	3.81	1.82	0.82	0.36	0.40	0.75
2:1	50	B	1.5	1.69	33.83	23.59	16.32	10.64	6.45	3.84	2.21	1.24	0.00	0.03
	50	Own	4	2.07	35.50	22.71	15.32	10.66	6.23	3.63	1.95	1.05	0.02	0.86
	65	B	1.9	1.91	33.85	23.52	16.21	10.57	6.41	3.83	2.19	1.27	0.00	0.03
	65	Own	5	2.26	35.91	22.64	15.19	10.47	6.07	3.44	1.93	1.07	0.15	0.87
	80	B	2.6	2.14	33.67	23.52	16.18	10.55	6.40	3.82	2.17	1.24	0.00	0.05
	80	Own	8	2.62	37.80	22.46	14.50	9.67	5.44	3.08	1.67	0.87	1.21	0.67
2.5:1	50	Own	4	1.56	40.34	23.59	14.67	9.29	4.93	2.59	1.27	0.61	0.00	1.16
	65	Own	5.5	1.84	40.76	23.04	14.40	9.12	4.89	2.60	1.34	0.66	0.13	1.22
	80	Own	9.5	2.15	40.99	23.48	14.22	8.68	4.48	2.29	1.09	0.50	1.31	0.80

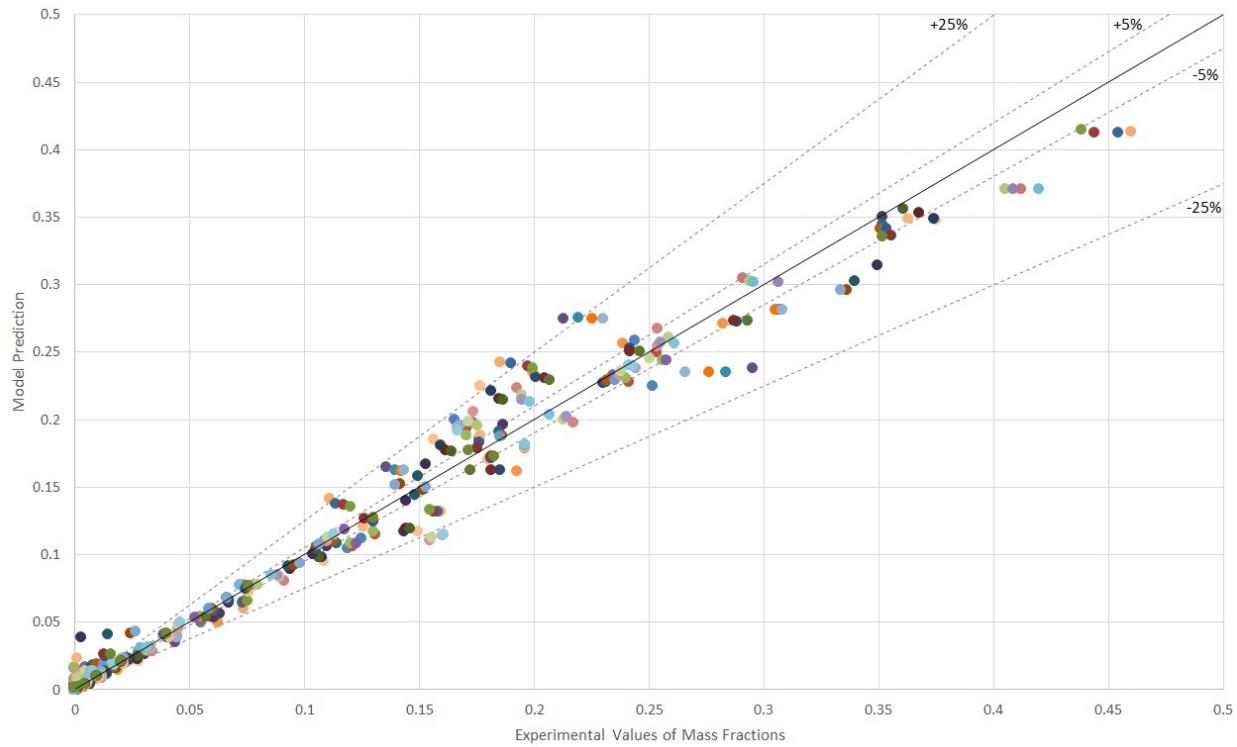


Figure S2 A parity plot of overall mass fractions of FA, MeOH, water and OME of the different educts at equilibrium calculated by the model to the literature experimental values from Schmitz et al. (Ref. 1)

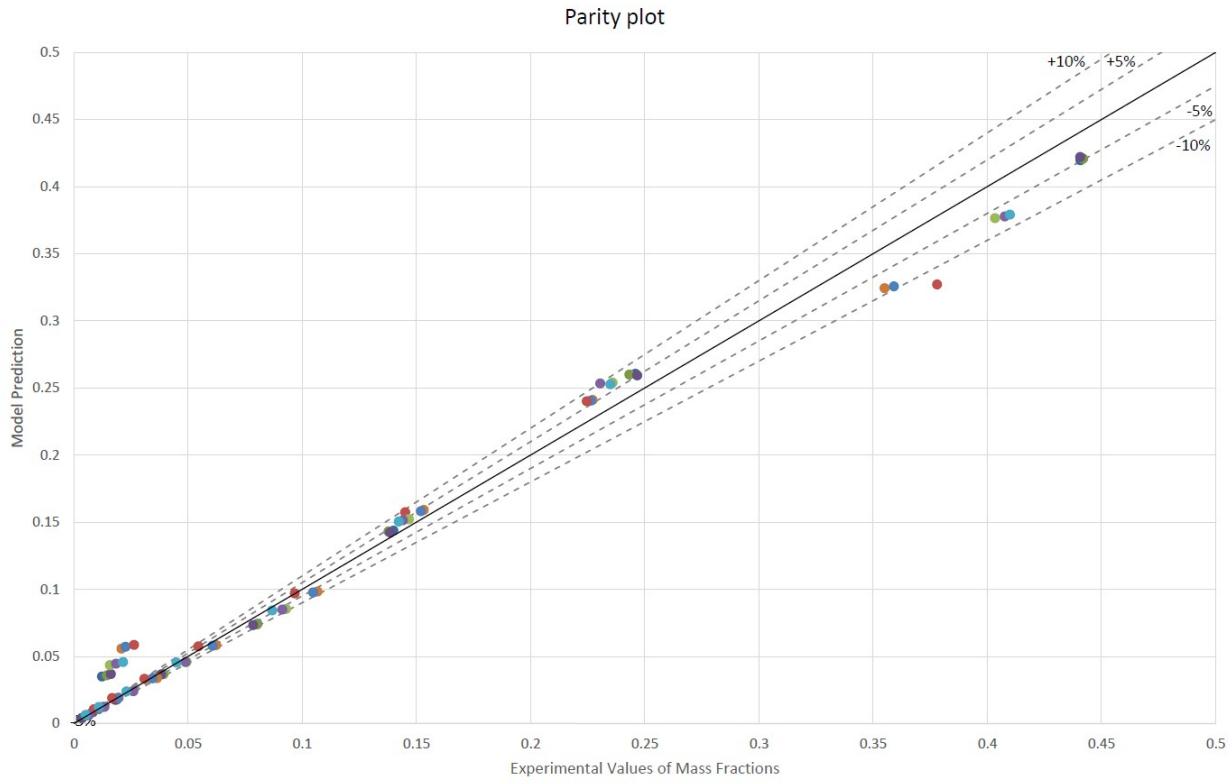


Figure S3 Parity plot between the equilibrium overall mass fractions of FA, MeOH, water and OME of different educts and reaction conditions calculated by the unconstrained Gibbs energy minimisation equilibrium model to the experimental values reported in Table S6

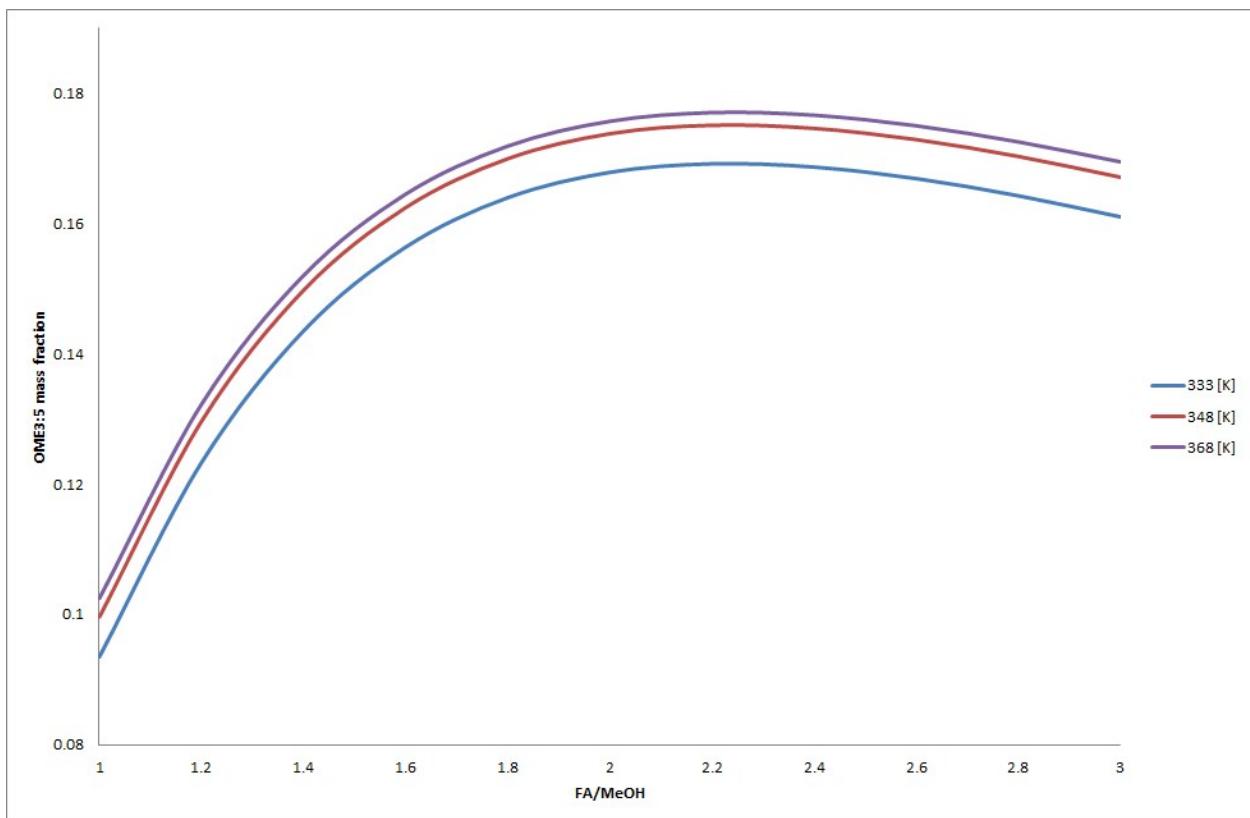


Figure S4 Effect of FA and MeOH ratio for anhydrous feeds on the OME₃₋₅ equilibrium composition

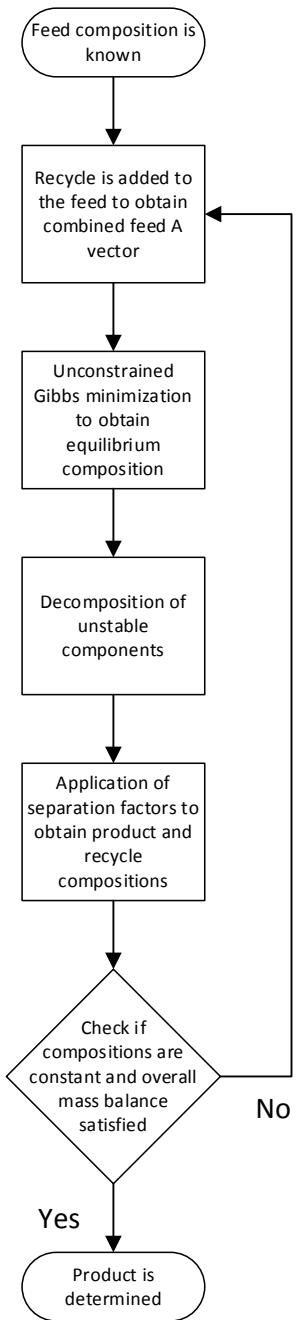


Figure S5 The recycle calculation algorithm

Table S7 Fuel properties of DME, OME and diesel (EN 590) as sourced from Refs. 3, 4, and 5.

Component	M [g/mol]	Oxygen-content [wt. %]	Cetane number	Density at 15°C [kg/m ³]	Melting point [°C]	Boiling point [°C]	v at 40°C [mm ² /s]	LHV [kWh/kg]	Equivalent to diesel [m ³ /m ³]
DME	46.1	34.7	55	668	-141	-25	<0,1	7.5	1.96
OME ₁	76.1	42.1	29	866.8	-105	42.3	0.58 _(20°C)	6.5	1.75
OME ₂	106.1	45.2	63	970	-69.7	105	0.66 _(20°C)	5.4	1.75
OME ₃	136.1	47	67	1,031	-42.5	155.9	1.08	5.4	1.75
OME ₄	166.2	48.1	76	1,075	-9.8	201.8	1.72	5.3	1.73
OME ₅	196.2	48.9	90	1,106	18.3	242.3	2.63	5.1	1.72
OME ₆	226.2	49.5	-	1,135	45.8	273	-	-	-
OME ₇	256.3	49.9	-	1,157	72.4	297	-	-	-
OME ₈	286.3	50.3	-	-	-	-	-	-	-
OME ₃₋₅	166.2	48.8	70-100	1,070	-19	155-242	1.89	5.4	1.8
Diesel	-	-	>51	820-840	-20 _{winter} 0 _{summer}	170-390	2-4.5	11.833	1

References

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