

**Core-satellite structured Z-scheme catalyst  $\text{Cd}_{0.5}\text{Zn}_{0.5}\text{S}/\text{BiVO}_4$  for highly efficiency and stable photocatalytic water splitting**

Chao Zeng,<sup>†</sup> Yingmo Hu,<sup>\*,†</sup> Tierui Zhang,<sup>‡</sup> Fan Dong,<sup>§</sup> Yihe Zhang,<sup>†</sup> Hongwei Huang

<sup>\*,†</sup>

<sup>†</sup>Beijing Key Laboratory of Materials Utilization of Nonmetallic Minerals and Solid Wastes, National Laboratory of Mineral Materials, School of Materials Science and Technology, China University of Geosciences, Beijing 100083, China

<sup>‡</sup>Key Laboratory of Photochemical Conversion and Optoelectronic Materials, Technical Institute of Physics and Chemistry, Chinese Academy of Sciences, Beijing 100190, China

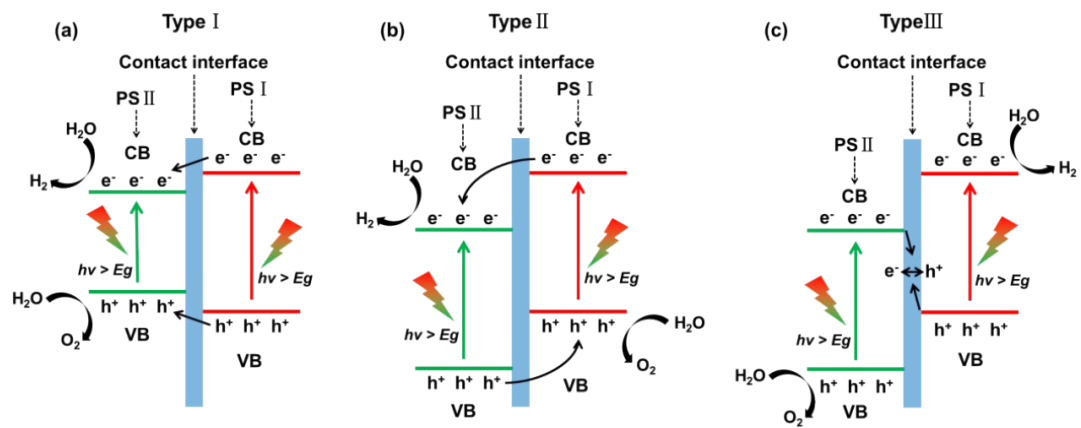
<sup>§</sup>Chongqing Key Laboratory of Catalysis and Functional Organic Molecules, College of Environmental and Biological Engineering, Chongqing Technology and Business University, Chongqing, 400067, China

\*E-mail: [huyingmo@cugb.edu.cn](mailto:huyingmo@cugb.edu.cn), [hhw@cugb.edu.cn](mailto:hhw@cugb.edu.cn)

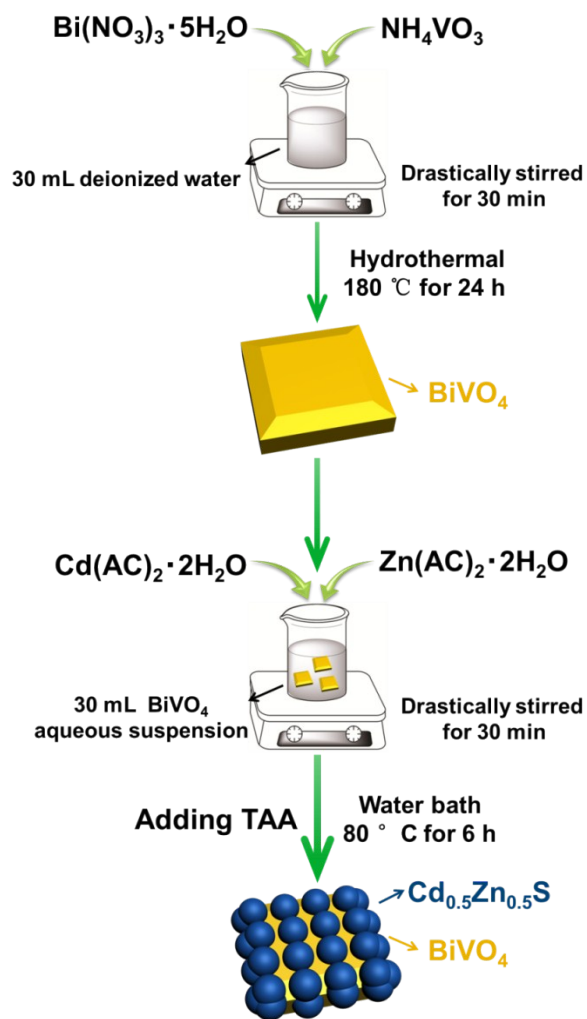
### *Fabrication of $Cd_{0.5}Zn_{0.5}S-CoO_x(0.5\%)/BiVO_4$*

$CoO_x(0.5\text{ wt}\%)/BiVO_4$  was prepared as follows.  $BiVO_4$  was dispersed in a corundum crucible containing 530  $\mu\text{l}$  of  $Co(NO_3)_2$  aqueous (80  $\text{mmol/L}$ ) solution. The crucible was dried at 60  $^\circ\text{C}$  for 3 h in air. Then it was calcined at 400  $^\circ\text{C}$  for 2 h in air. After cooling to room temperature, the sample was ground into powder.

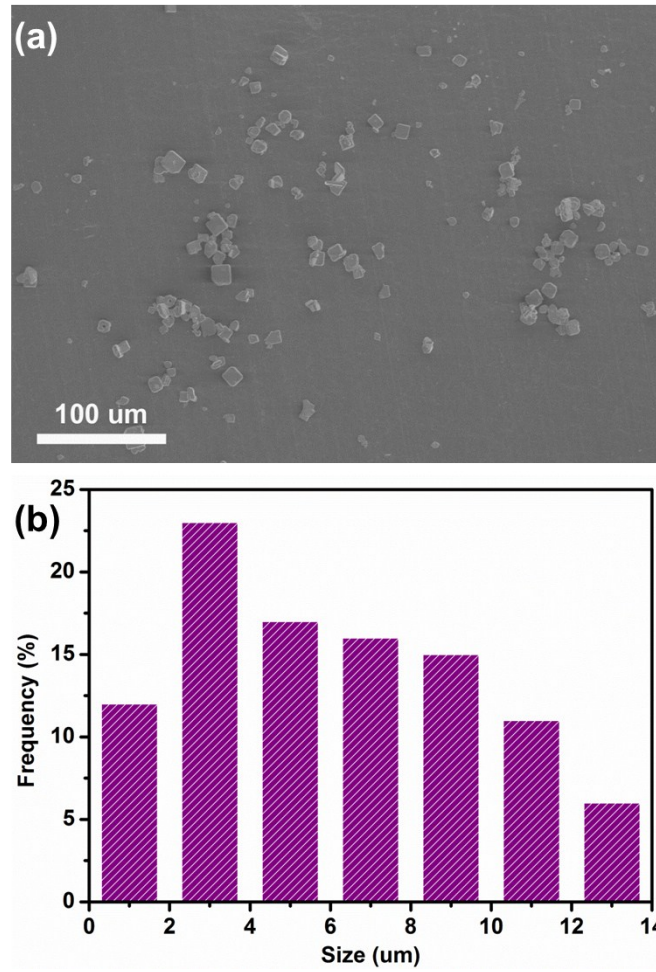
$Cd_{0.5}Zn_{0.5}S-CoO_x(0.5\%)/BiVO_4$  was obtained via the following route.  $CoO_x(0.5\%)/BiVO_4$  was dispersed in a beaker containing 30 ml of deionized water with ultrasonic and then certain amount of  $Cd(AC)_2 \cdot 2H_2O$  and  $Zn(AC)_2 \cdot 2H_2O$  were added into the above suspension liquid. Whereafter, quantitative thioacetamide was added into the above suspension. Then the beaker (sealed with plastic wrap) was put into an 80  $^\circ\text{C}$  water bath for 6 h with magnetic stirring. After natural cooling, the product was collected by centrifugation and rinsed repeatedly with distilled water and ethanol, and then dried at 60  $^\circ\text{C}$  for 10 h in air.



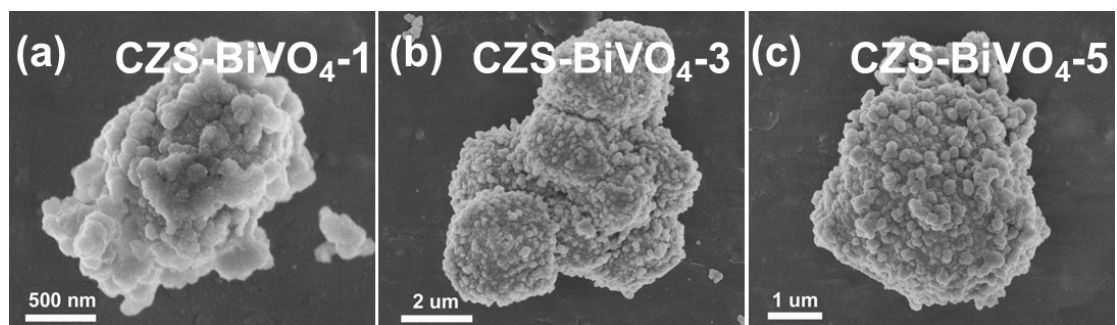
**Scheme. S1** Three types of charge transfer mechanism for heterojunction photocatalysts.



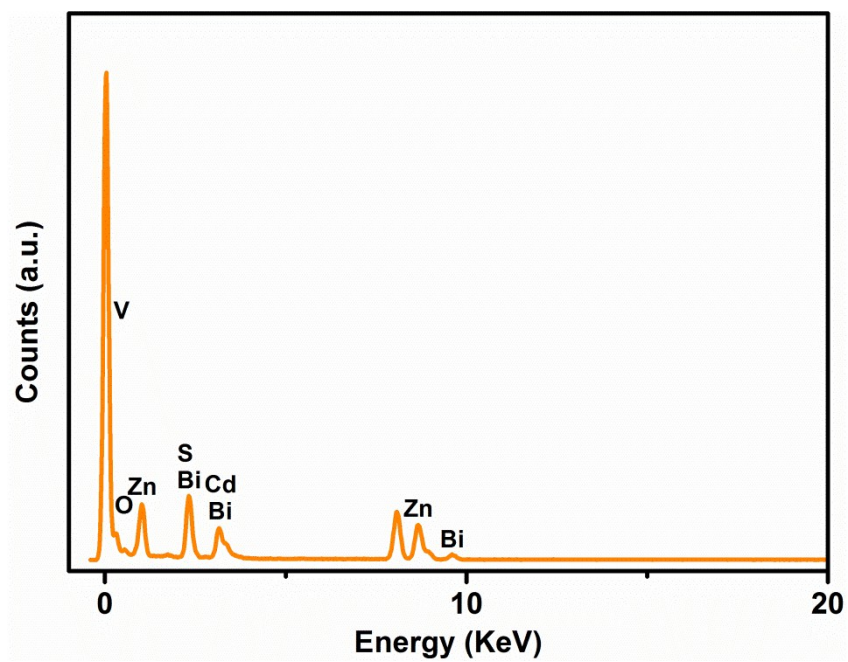
**Scheme. S2** The schematic illustration of the preparation process for CZS-BiVO<sub>4</sub> composite photocatalysts.



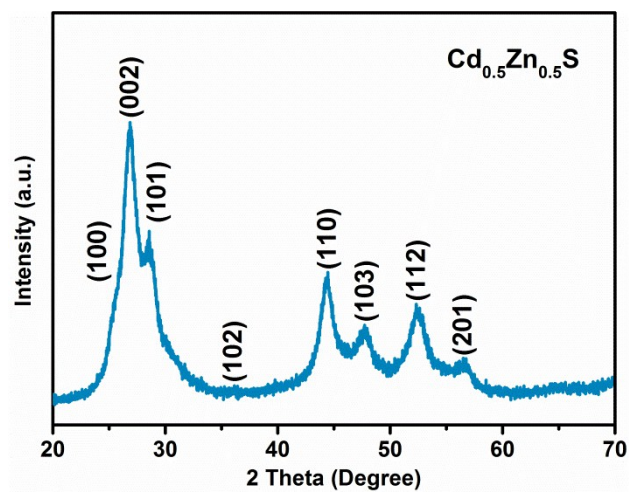
**Fig. S1** SEM image and size distribution histogram (calculated by the software of Image J) of BiVO<sub>4</sub> sample.



**Fig. S2** SEM images of CZS-BiVO<sub>4</sub>-1, CZS-BiVO<sub>4</sub>-3, and CZS-BiVO<sub>4</sub>-5 samples.

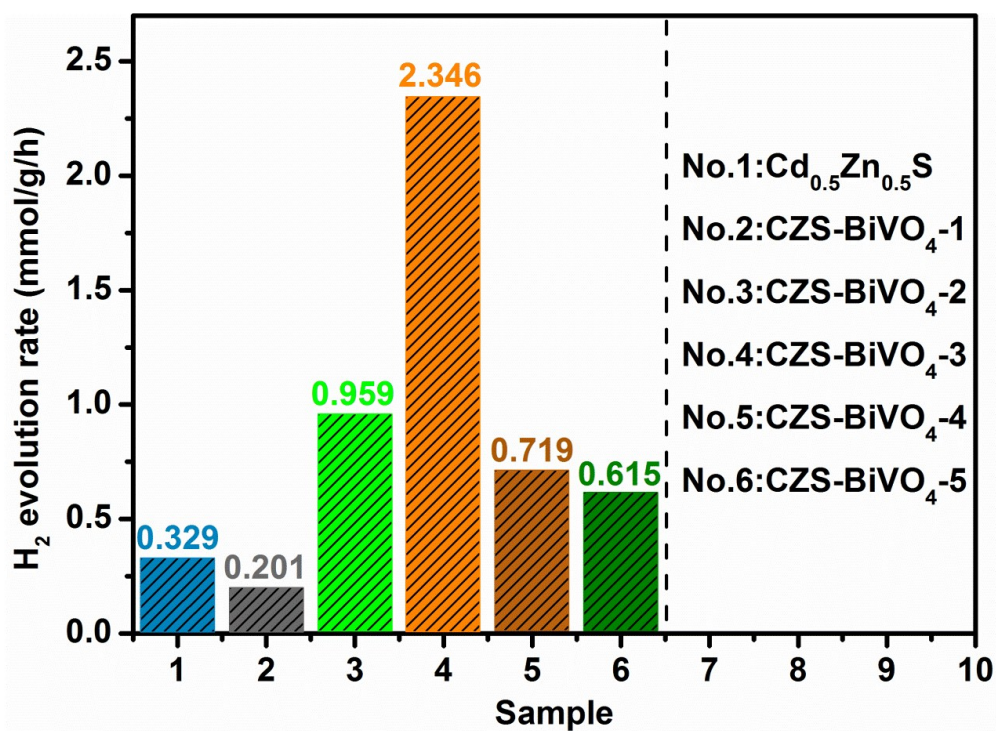


**Fig. S3** EDX spectrum for CZS-BiVO<sub>4</sub>-3 sample.



**Fig. S4** XRD pattern of Cd<sub>0.5</sub>Zn<sub>0.5</sub>S.



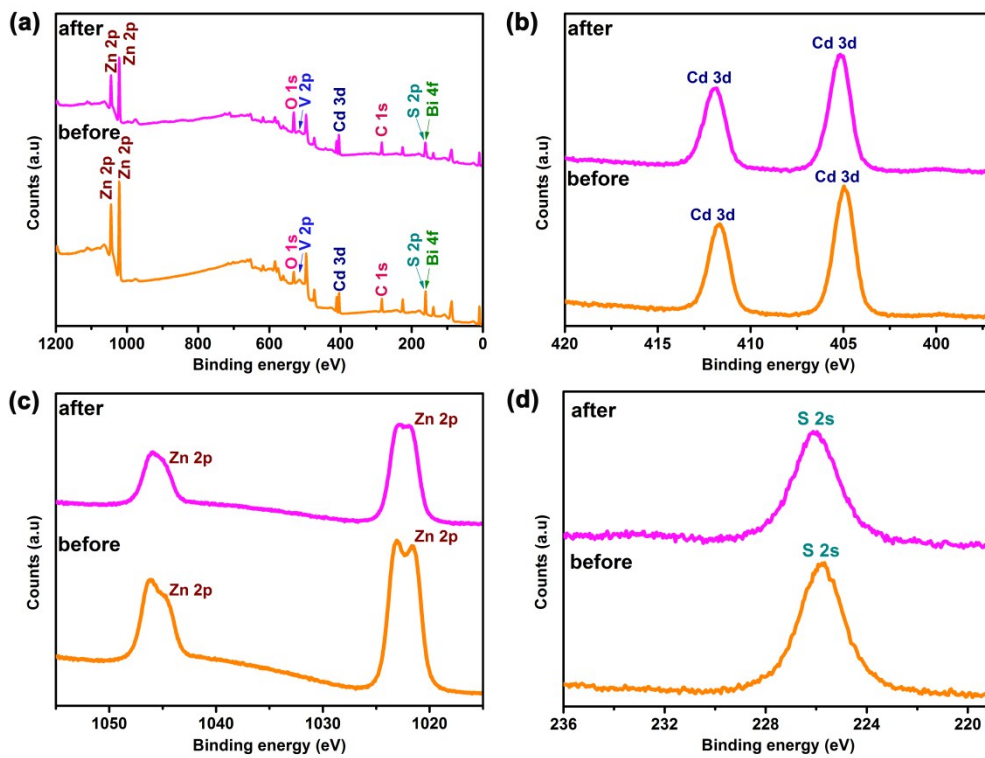


**Fig. S5** Apparent rate constants of H<sub>2</sub> evolution for Cd<sub>0.5</sub>Zn<sub>0.5</sub>S, and CZS-BiVO<sub>4</sub> composite samples under visible light illumination.

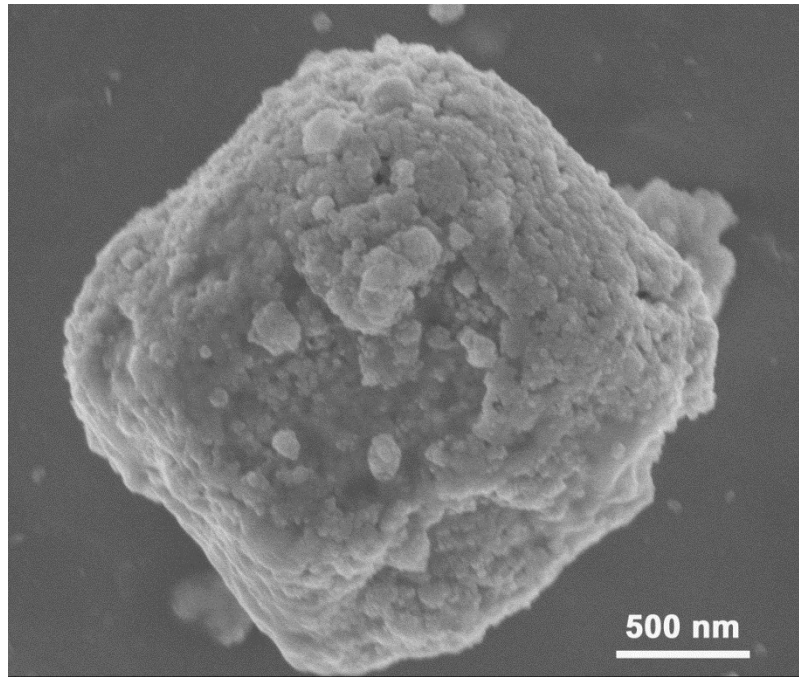
The TOF of Pt atoms for CZS-BiVO<sub>4</sub>-3 sample under visible light irradiation for 4 h is calculated as following:

$$[\text{TON}] = [9.01165 * 10^{-3} * 0.05 * 4] / [0.05 * 3\% / 195.084] = 234.4$$

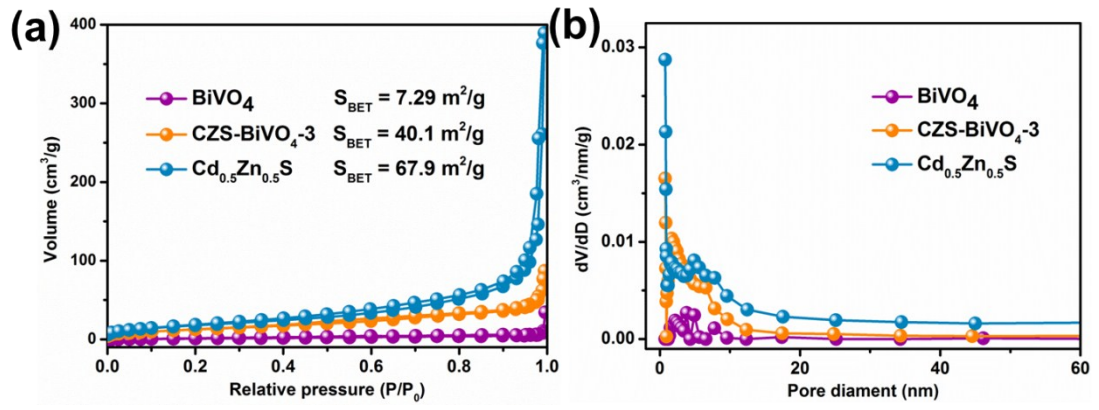
$$[\text{TOF}] = [\text{TON}] / t = 234.4 / 4 = 58.6 \text{ h}^{-1}$$



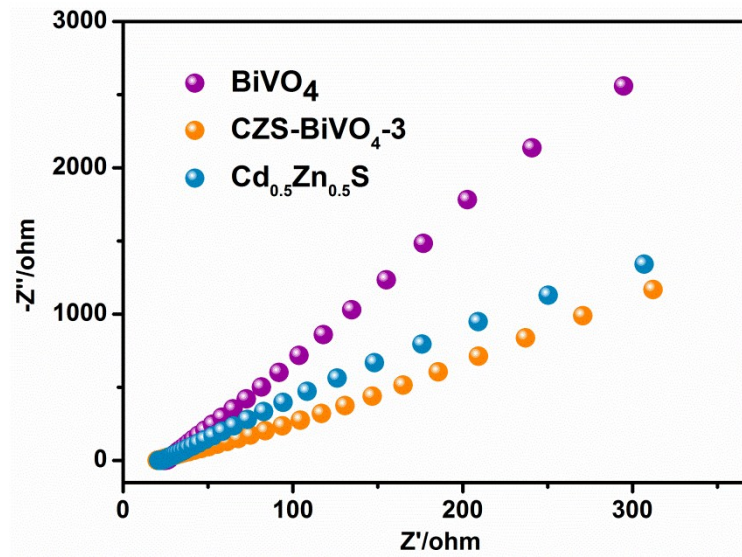
**Fig. S6** Typical XPS survey spectra (a), high-resolution XPS spectra of Cd (b), Zn (c), and S (d), for CZS-BiVO<sub>4</sub>-3 sample before and after photocatalytic reaction.



**Fig. S7** SEM image of CZS-BiVO<sub>4</sub>-3 sample after photocatalytic reaction.



**Fig. S8** Nitrogen adsorption/desorption isotherms **(a)** and pore diameter distribution **(b)** of BiVO<sub>4</sub>, CZS-BiVO<sub>4</sub>-3, and Cd<sub>0.5</sub>Zn<sub>0.5</sub>S samples.



**Fig. S9** EIS Nyquist plots for BiVO<sub>4</sub>, CZS-BiVO<sub>4</sub>-3, and Cd<sub>0.5</sub>Zn<sub>0.5</sub>S samples.