

## Electronic Supplementary Information

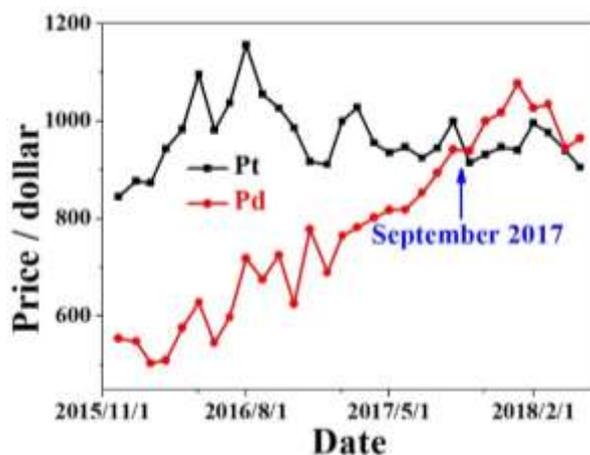
### **From monometallic Au nanowires to trimetallic AuPtRh nanowires: interface control for the formic acid electrooxidation†**

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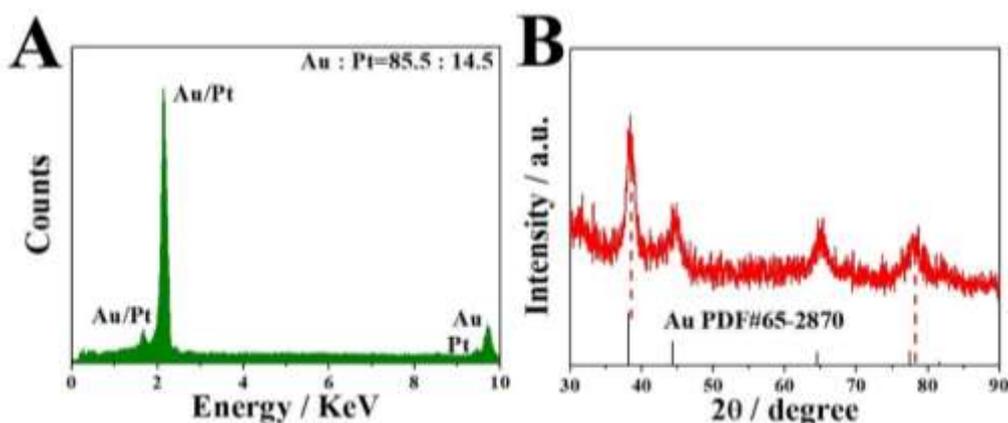
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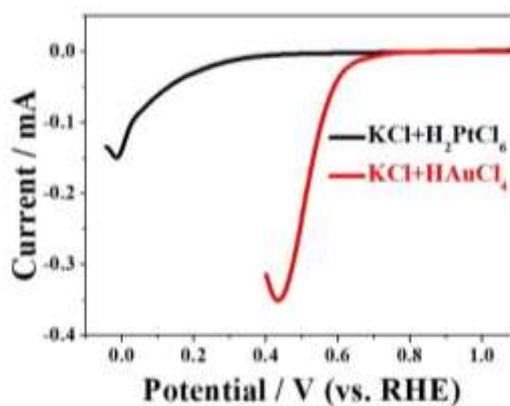
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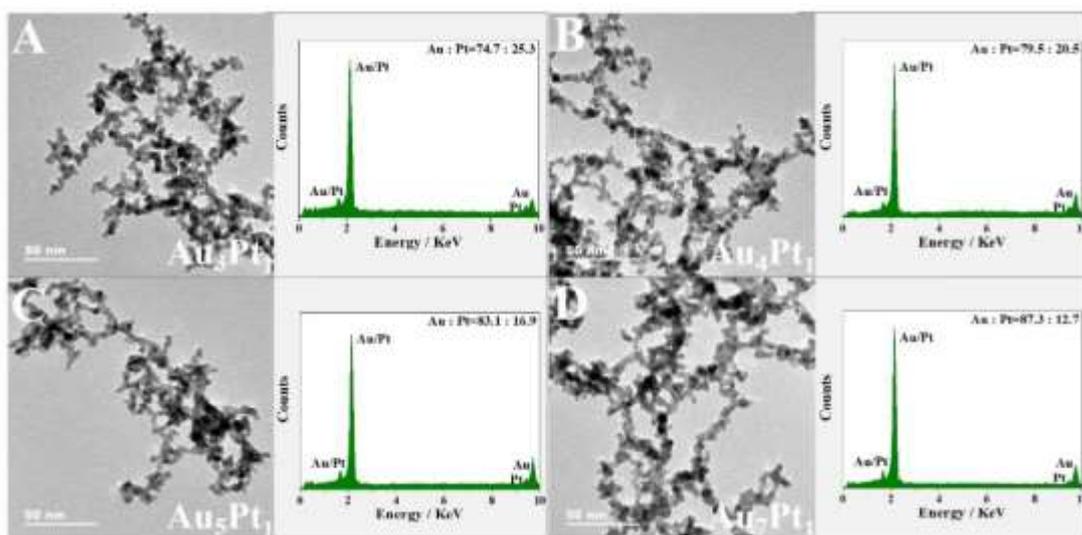
**Fig. S1** Johnson Matthey base price chart of Pt and Pd between December 2015 and May 2018.



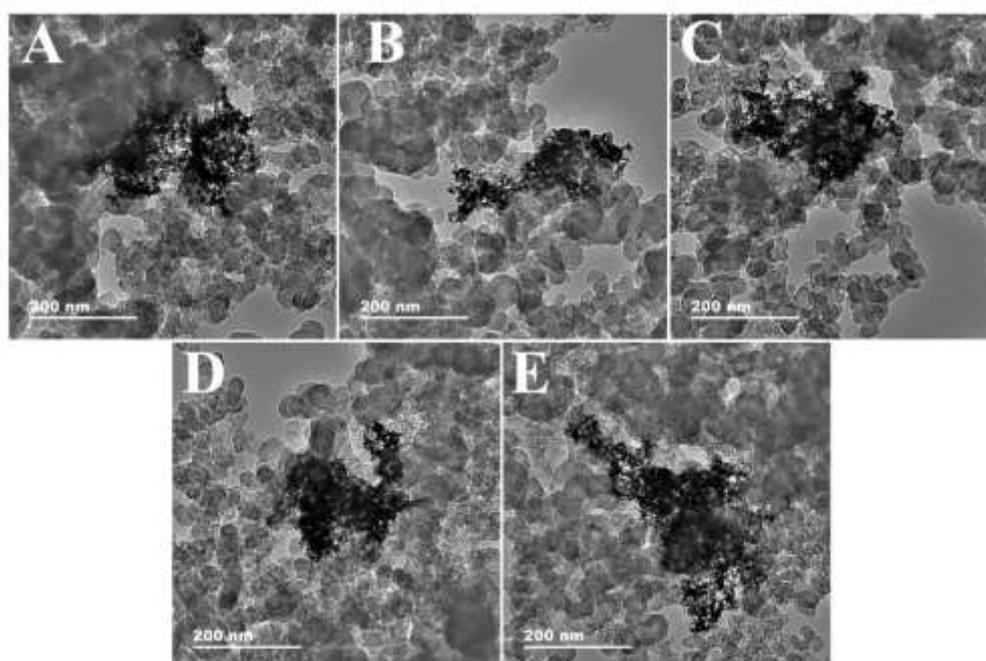
**Fig. S2** (A) EDX spectrum and (B) XRD pattern of Au<sub>6</sub>Pt<sub>1</sub> NWs.



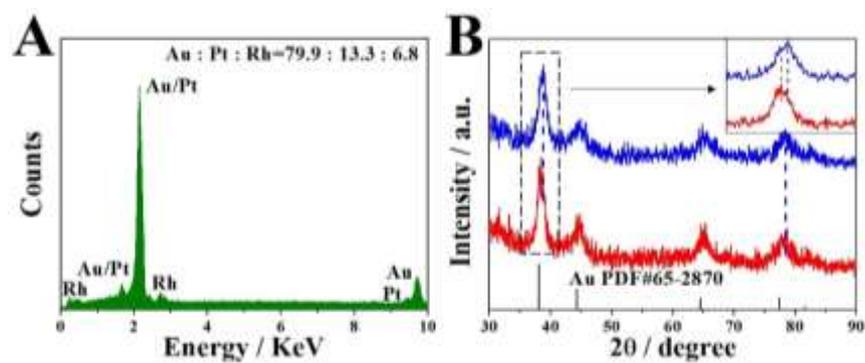
**Fig. S3** Linear sweep voltammetry curves of the mixed solution of 0.1 M KCl + 0.01 M H<sub>2</sub>PtCl<sub>6</sub> and 0.1 M KCl + 0.01M HAuCl<sub>4</sub> solution on the glassy carbon electrode at 50 mV s<sup>-1</sup>.



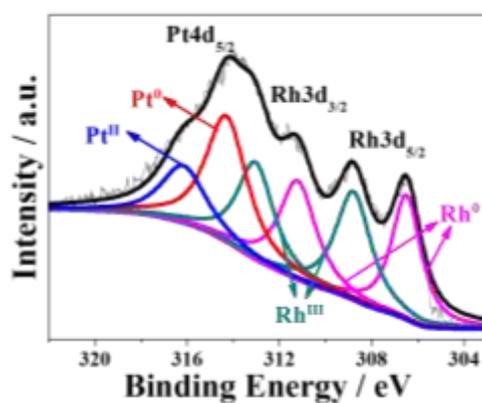
**Fig. S4** TEM images and EDX spectra of (A)  $\text{Au}_3\text{Pt}_1$  NWs, (B)  $\text{Au}_4\text{Pt}_1$  NWs, (C)  $\text{Au}_5\text{Pt}_1$  NWs, and (D)  $\text{Au}_7\text{Pt}_1$  NWs.



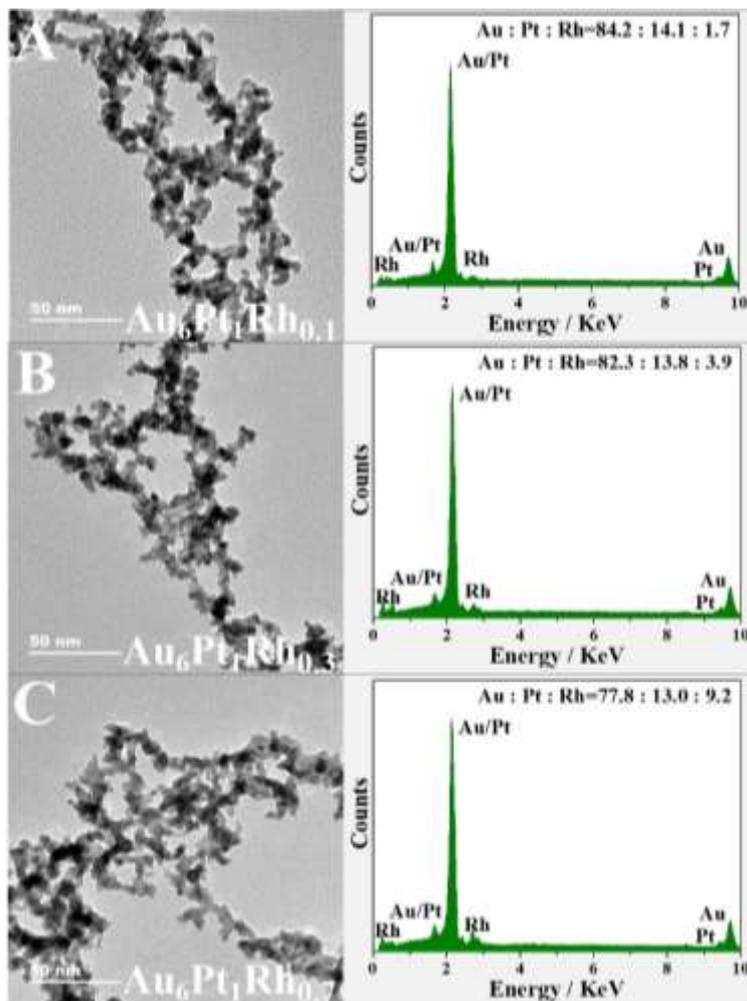
**Fig. S5** TEM images of (A)  $\text{Au}_3\text{Pt}_1/\text{C}$ , (B)  $\text{Au}_4\text{Pt}_1/\text{C}$ , (C)  $\text{Au}_5\text{Pt}_1/\text{C}$ , (D)  $\text{Au}_6\text{Pt}_1/\text{C}$ , and (E)  $\text{Au}_7\text{Pt}_1/\text{C}$ .



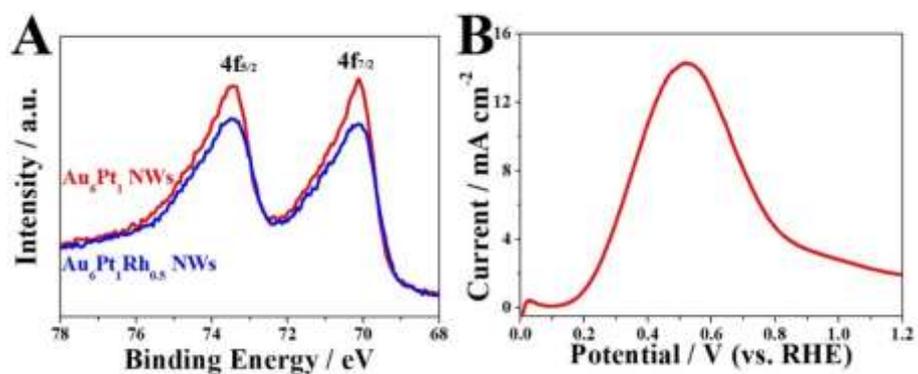
**Fig. S6** (A) EDX spectrum of  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.5}$  NWs. (B) XRD patterns of  $\text{Au}_6\text{Pt}_1$  NWs and  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.5}$  NWs.



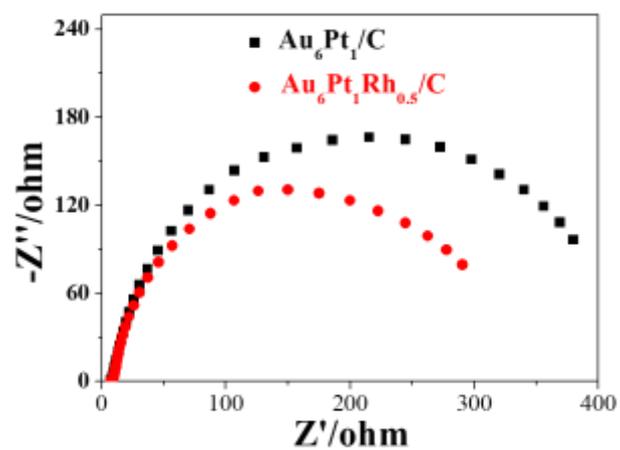
**Fig. S7** Rh 3d XPS spectra of  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.5}$ .



**Fig. S8** TEM images and EDX spectra of (A)  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.1}$  NWs, (B)  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.3}$  NWs, and (C)  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.7}$  NWs.



**Fig. S9** (A) Pt 4f XPS spectra of  $\text{Au}_6\text{Pt}_1$  NWs and  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.5}$  NWs. (B) ECSA-normalized positive direction CV scan curves of  $\text{Au}_6\text{Pt}_1\text{Rh}_{0.5}/\text{C}$  electrocatalysts in 0.5 M  $\text{HCOOH}$  + 0.5 M  $\text{H}_2\text{SO}_4$  solution at scan rate of  $50 \text{ mV s}^{-1}$ .



**Fig. S10** EIS curves of Au<sub>6</sub>Pt<sub>1</sub>Rh<sub>0.5</sub>/C and commercial Au<sub>6</sub>Pt<sub>1</sub>/C in 0.5 M HCOOH + 0.5 M H<sub>2</sub>SO<sub>4</sub> solution (applied potential: 0.4 V vs. RHE).

**Table S1.** The catalytic activity of recently reported Pt-based or Pd-based electrocatalysts for the FAOR.

Electrocatalysts	Electrolyte	Mass activity (A mg <sub>Pt or Pd</sub> <sup>-1</sup> )	Specific activity (mA cm <sub>ECSA</sub> <sup>-2</sup> )	Refs.	Year
Au <sub>6</sub> Pt <sub>1</sub> Rh <sub>0.5</sub> /C	0.5 M HCOOH + 0.5 M H <sub>2</sub> SO <sub>4</sub>	8.05	14.3	This work	2018
	0.25 M HCOOH + 0.5 M H <sub>2</sub> SO <sub>4</sub>	4.47	7.92	This work	2018
	0.1 M HCOOH + 0.5 M H <sub>2</sub> SO <sub>4</sub>	2.24	3.97	This work	2018
CuPd/WO <sub>2.72</sub>	0.1 M HCOOH + 0.1 M HClO <sub>4</sub>	2.09	n.a.	1	2018
Au@Pt-graphene	0.1 M HCOOH + 0.1 M H <sub>2</sub> SO <sub>4</sub>	0.464	0.443	2	2018
PdH <sub>x</sub> nanocatalysts	0.25 M HCOOH + 0.5 M H <sub>2</sub> SO <sub>4</sub>	1.06	5.12	3	2018
AgPt nanowires	1.0 M HCOOH + 0.5 M H <sub>2</sub> SO <sub>4</sub>	0.152	1.03	4	2018
Pt <sub>3</sub> Ni tetrahedral nanoframes/C	0.5 M HCOOH + 0.1 M HClO <sub>4</sub>	n.a.	~3.0	5	2017
PdCu nanosheets	0.25 M HCOOH + 0.5 M H <sub>2</sub> SO <sub>4</sub>	1.66	1.18	6	2017
Pt <sub>4</sub> PdCu <sub>0.4</sub> nanoframes	0.1 M HCOOH + 0.1 M HClO <sub>4</sub>	~1.29	~3.0	7	2017
Porous Pd nanosheets	0.5 M HCOOH + 0.5 M H <sub>2</sub> SO <sub>4</sub>	0.409	3.17	8	2017

## References

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