Supporting Information

Facile synthesis of graphene nanoribbons from zeolite-templated ultra-small carbon nanotubes for lithium ion storage

Xintong Xu,^{a,b} Shuangchen Ruan,^{a,*} Jianpang Zhai,^a Haiou Zhu,^a Li Yu,^a

Jihong Pei,^b Zhe Cui,^c Rujia Zou,^c Junqing Hu,^c Aijiang Lu,^d Mingyang Yu^e and Zikang Tang^f

^aShenzhen Key Laboratory of Laser Engineering, Shenzhen University,

Shenzhen, 518060, China

^bCollege of Information Engineering, ATR National Defence S&T Key Laboratory, Shenzhen University, Shenzhen, 518060, China

^cState Key Laboratory for Modification of Chemical Fibers and Polymer Materials, College of Materials Science and Engineering, Donghua University, Shanghai, 201620, China

^dCollege of Science, Donghua University, Shanghai, 201620, China
^eInstitute for Fusion Theory and Simulation, Zhejiang University, Haizhou, 310027, China

Institute of Applied Physics & Materials Engineering, Faculty of Science and Technology, University of Macau, Macau, China



Figure S1. Evidence for no structural damage of as-synthesized AELafter carbonization of DPA. (a) SEM image of the AEL crystals beforecarbonization. (b) SEM image of the AEL after carbonization for 3 h at550 °C under Ar atmosphere. (c) Powder XRD patterns of the AELbeforeandaftercarbonization.



Figure S2. The XPS spectrums for DPA@AEL before (a lower curve)and after (an upper curve) carbonization at 550 °C for 3 h. The inset plotisthecurvesnearN1speak.



Figure S3. Raman spectrum of several samples processed under different pyrolysis temperature. As-synthesized DPA@AEL crystal (red one), and the samples after processing at 400 °C (blue one), 550 °C (black one) and 700 °C (olive one), respectively.



Figure S4. Radial distribution functions of initial model and optimized structure. The sharp and discrete peaks turned into broadened and successive ones after the Coarse-Graining (CG) optimization. It indicated the ordered arrangement of the atoms in the initial model was not kept. The ideal free standing (2,2) SWCNT was unstable during the relaxation, and it was easy to break and may transform into the other carbon structured materials.



Figure S5. XPS spectrum of as-synthesized GNRs. (a) Al 2p region, suggesting no signals for Al in the sample. (b) Zn 2p region, suggesting no signals Zn in the sample.



Figure S6. Electron energy loss spectroscopy (EELS) spectrum of as-synthesized GNRs. (a) C K-edge in EELS spectrum. (b) EELS features at410eVcharacteristicofNbonding.



Figure S7. Nitrogen adsorption and desorption isotherms.



Figure S8. A photo picture showing bulk quantity of N-doped GNRs products, in which ~ 1 g N-doped GNRs can be fabricated from ~ 11 g DPA@AEL crystals.



Figure S9. Cycling performance at current density of 0.1C



Figure S10. Cycling performance of Ni foam electrode at the same condition.

Supplementary Table 1. Summary of elemental (CHNO) and ICP analyses of as-synthesized GNRs.

CHNO Analyses (%)				ICP analyses (%)		
С	Н	Ν	0	Zn	Al	Na
91.82	2.01	1.32	4.68	0.01	0.01	0.03