

Electronic Supplementary Information

Chemically Soft Solid Electrolyte Interphase Forming Additives for Lithium-ion Batteries

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mpj-12a

Parameter	Value
1 Origin	Varian
2 Spectrometer	nmrns
3 Solvent	cd2cl2
4 Temperature	25.0
5 Pulse Sequence	s2pul
6 Acquisition Time	2.0447
7 Spectrometer Frequency	499.93
8 Spectral Width	8012.8
9 Lowest Frequency	-1013.3
10 Nucleus	¹ H

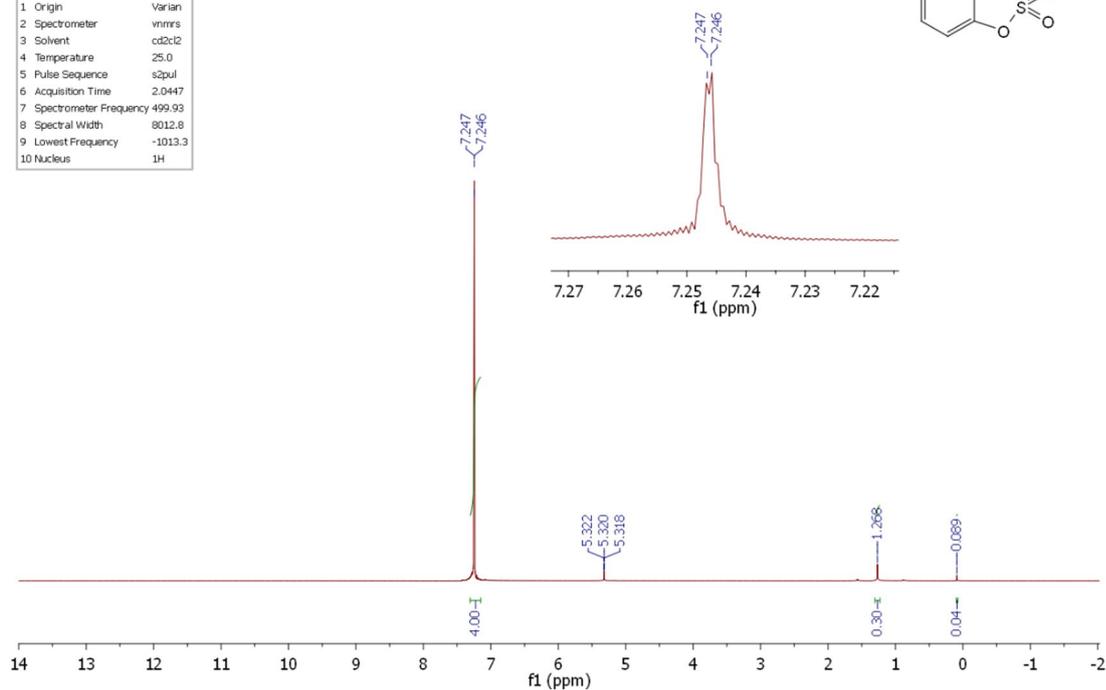
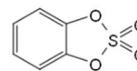


Figure S1. ¹H NMR spectrum of DTDPh.

mpj-12a-c

Parameter	Value
1 Origin	Varian
2 Spectrometer	nmrns
3 Solvent	cd2cl2
4 Temperature	25.0
5 Pulse Sequence	s2pul
6 Acquisition Time	1.0224
7 Spectrometer Frequency	125.72
8 Spectral Width	32051.3
9 Lowest Frequency	-2197.7
10 Nucleus	¹³ C

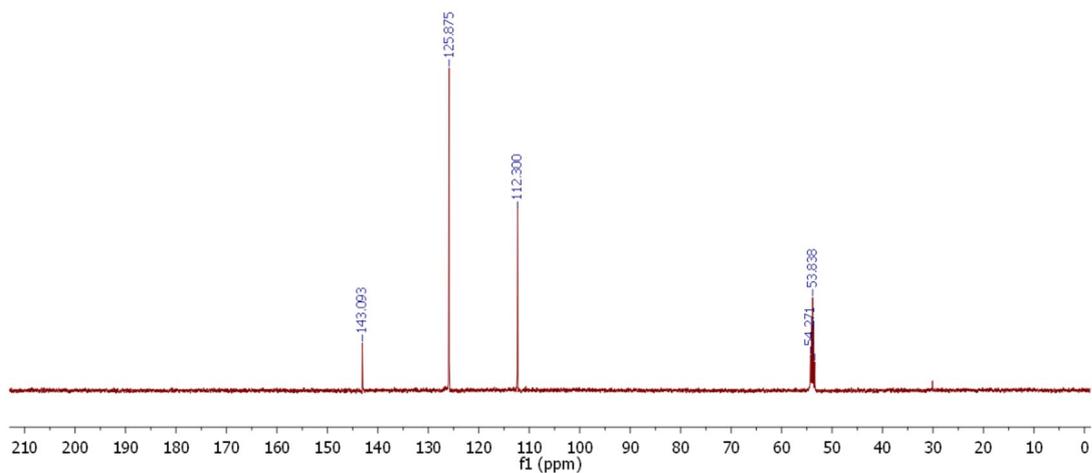
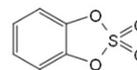


Figure S2. ¹³C NMR spectrum of DTDPh.

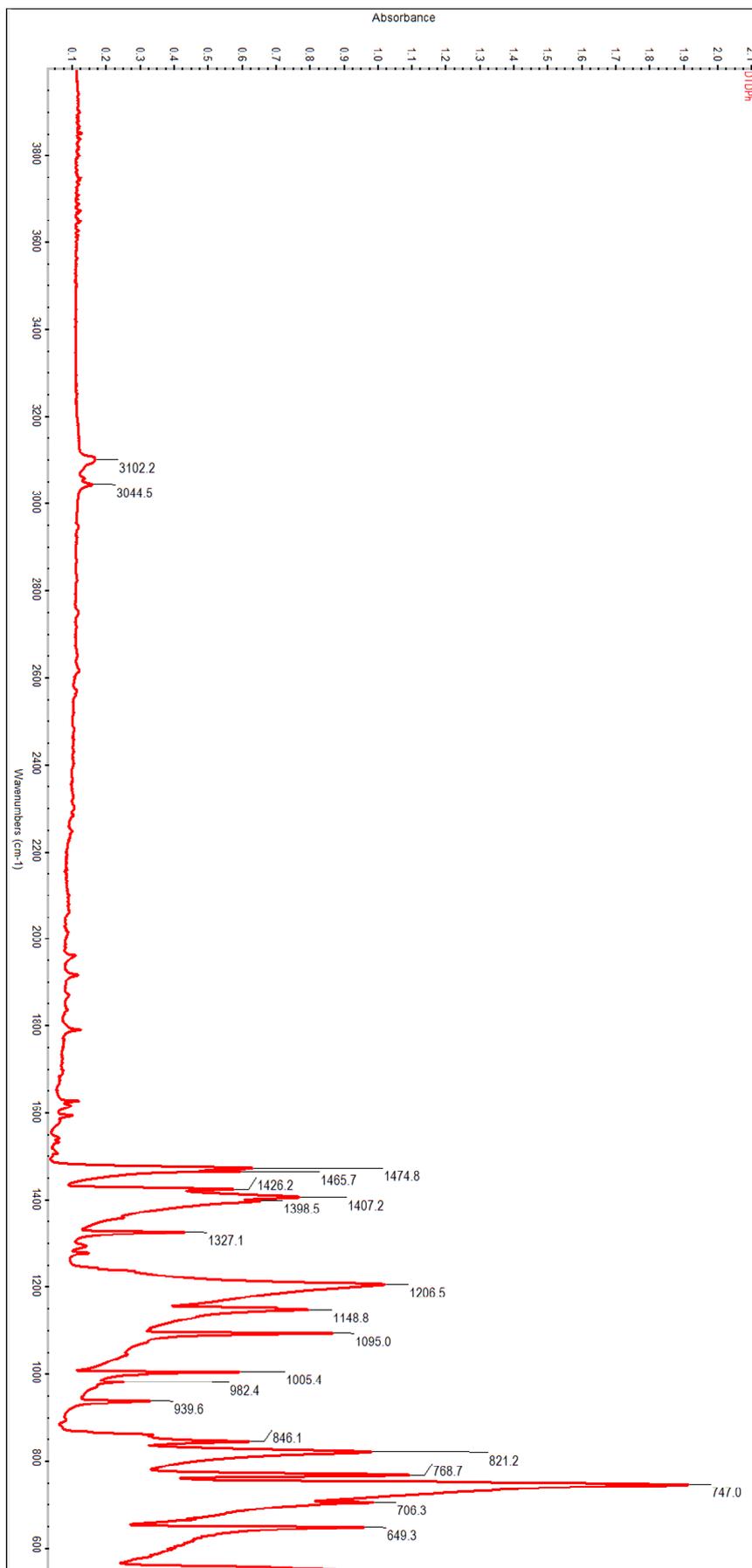


Figure S3. FTIR spectrum of DTDPH.

mpj11a

Parameter	Value
1 Origin	Varian
2 Spectrometer	nmrns
3 Solvent	dmsd
4 Temperature	25.0
5 Pulse Sequence	squl
6 Acquisition Time	2.0447
7 Spectrometer Frequency	499.94
8 Spectral Width	8012.8
9 Lowest Frequency	-1006.8
10 Nucleus	¹ H

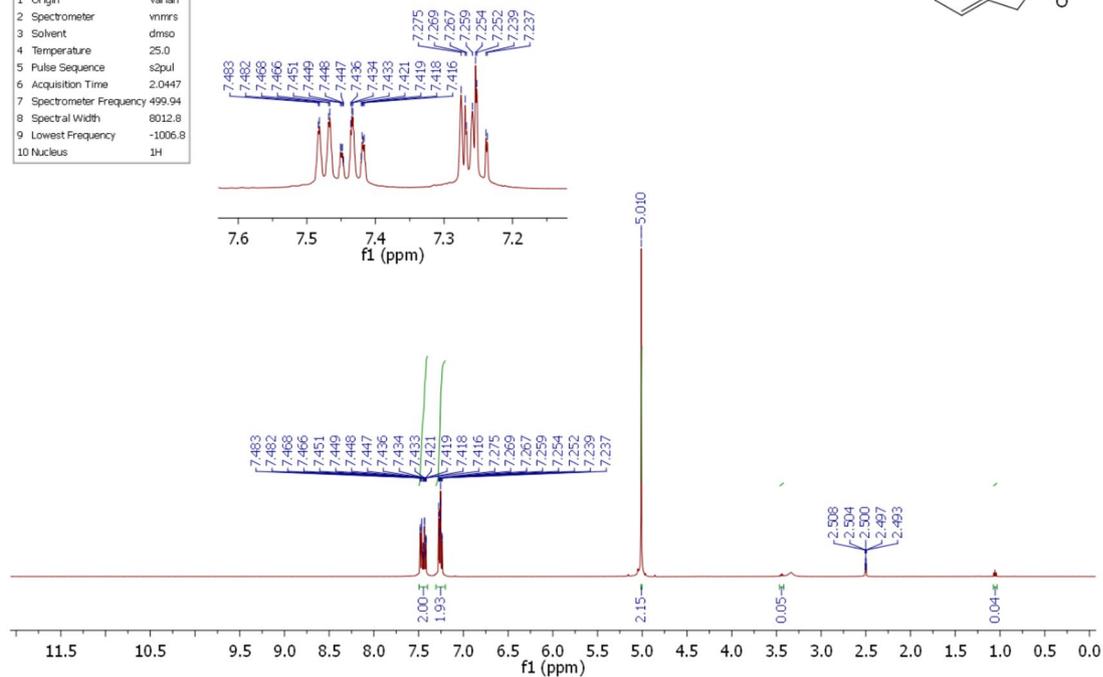
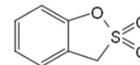


Figure S4. ¹H NMR spectrum of PSPH.

mpj-11-c

Parameter	Value
1 Origin	Varian
2 Spectrometer	nmrns
3 Solvent	dmsd
4 Temperature	25.0
5 Pulse Sequence	squl
6 Acquisition Time	1.0224
7 Spectrometer Frequency	125.72
8 Spectral Width	32051.3
9 Lowest Frequency	-2255.1
10 Nucleus	¹³ C

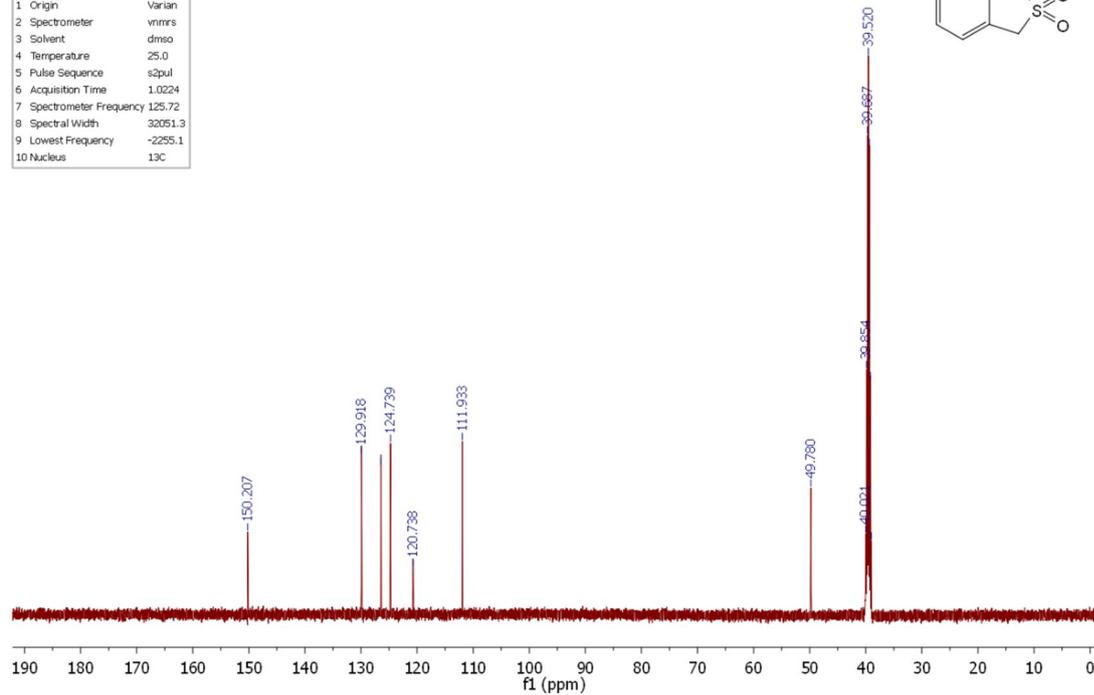
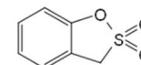


Figure S5. ¹³C NMR spectrum of PSPH.

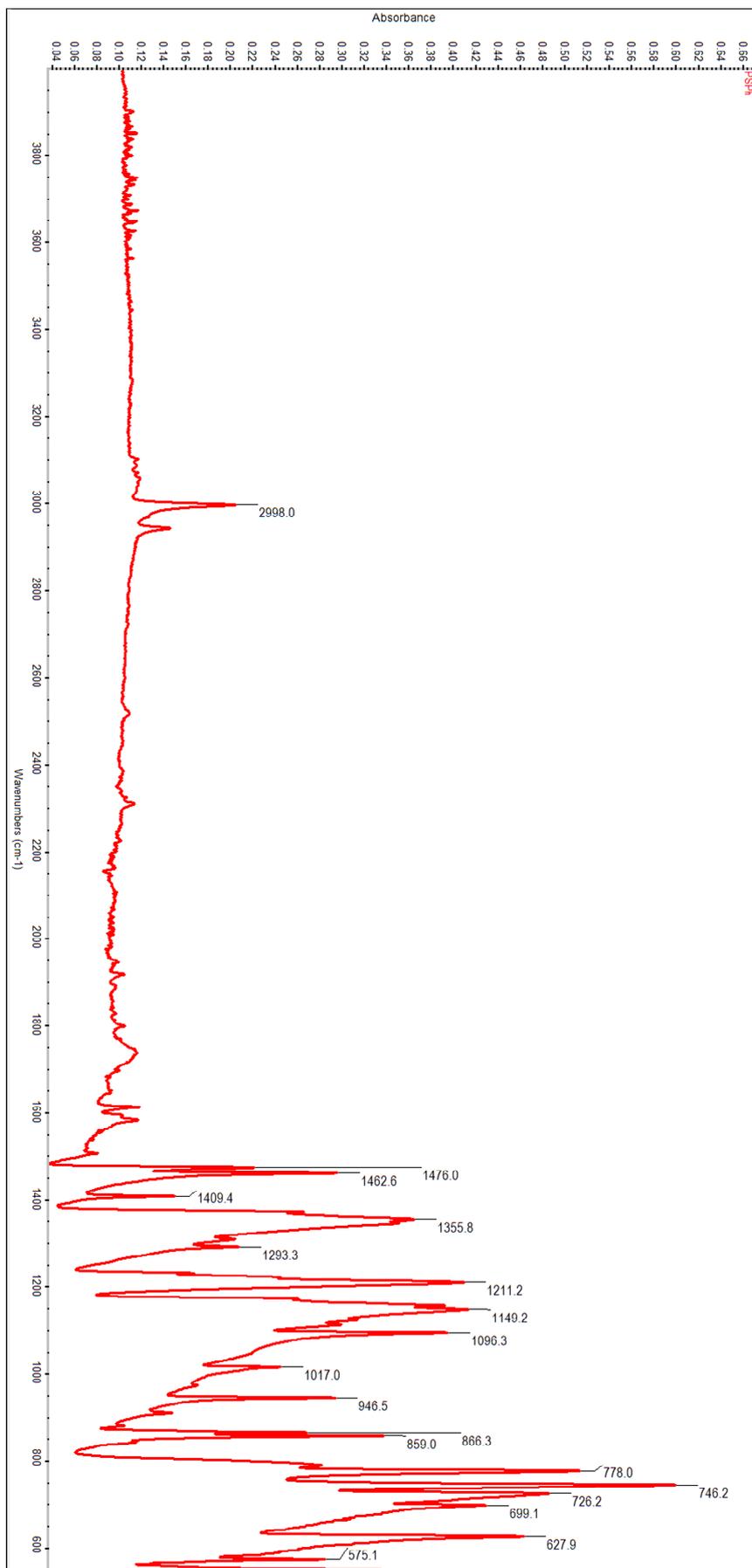


Figure S6. FTIR spectrum of PSPH.

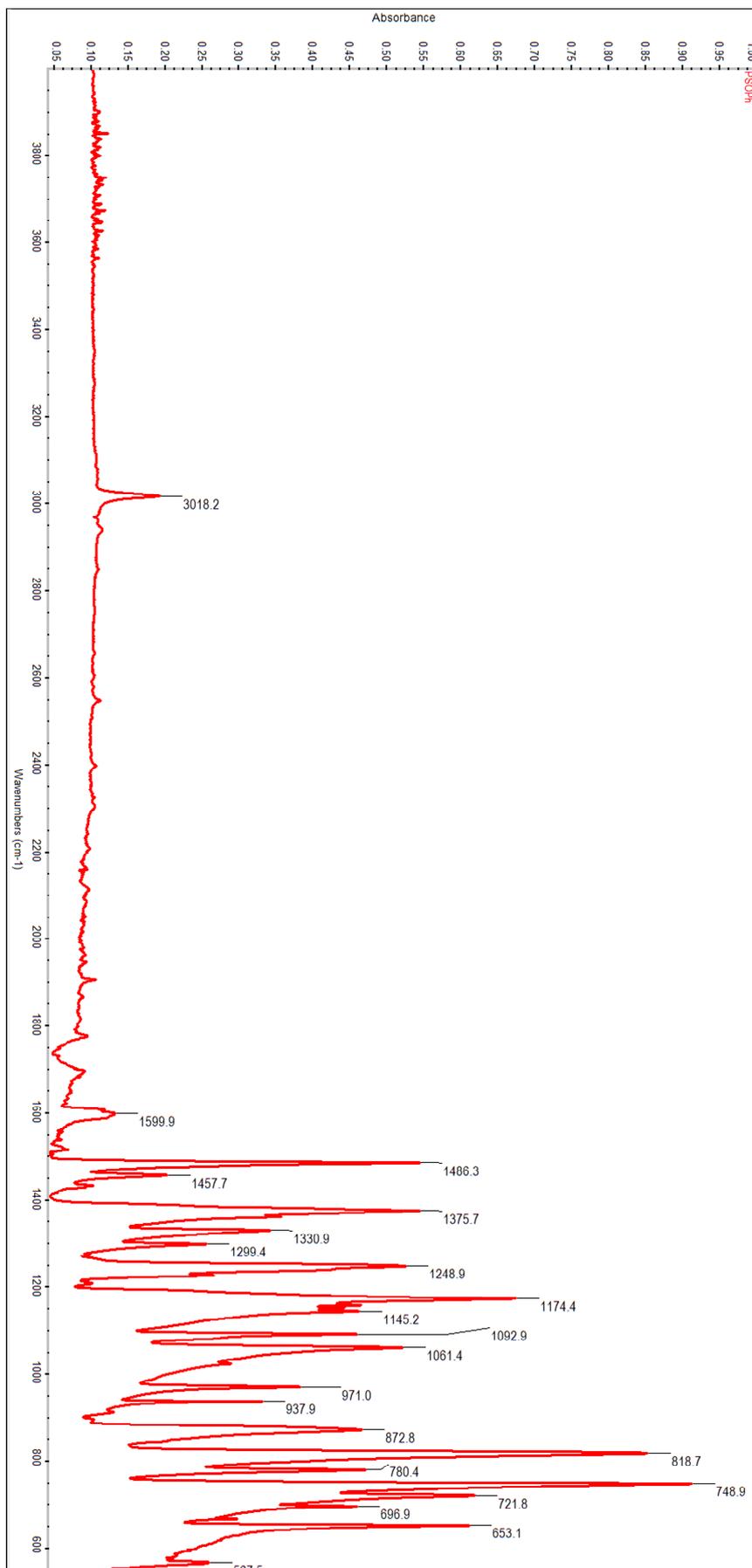


Figure S9. FTIR spectrum of PSOPh.

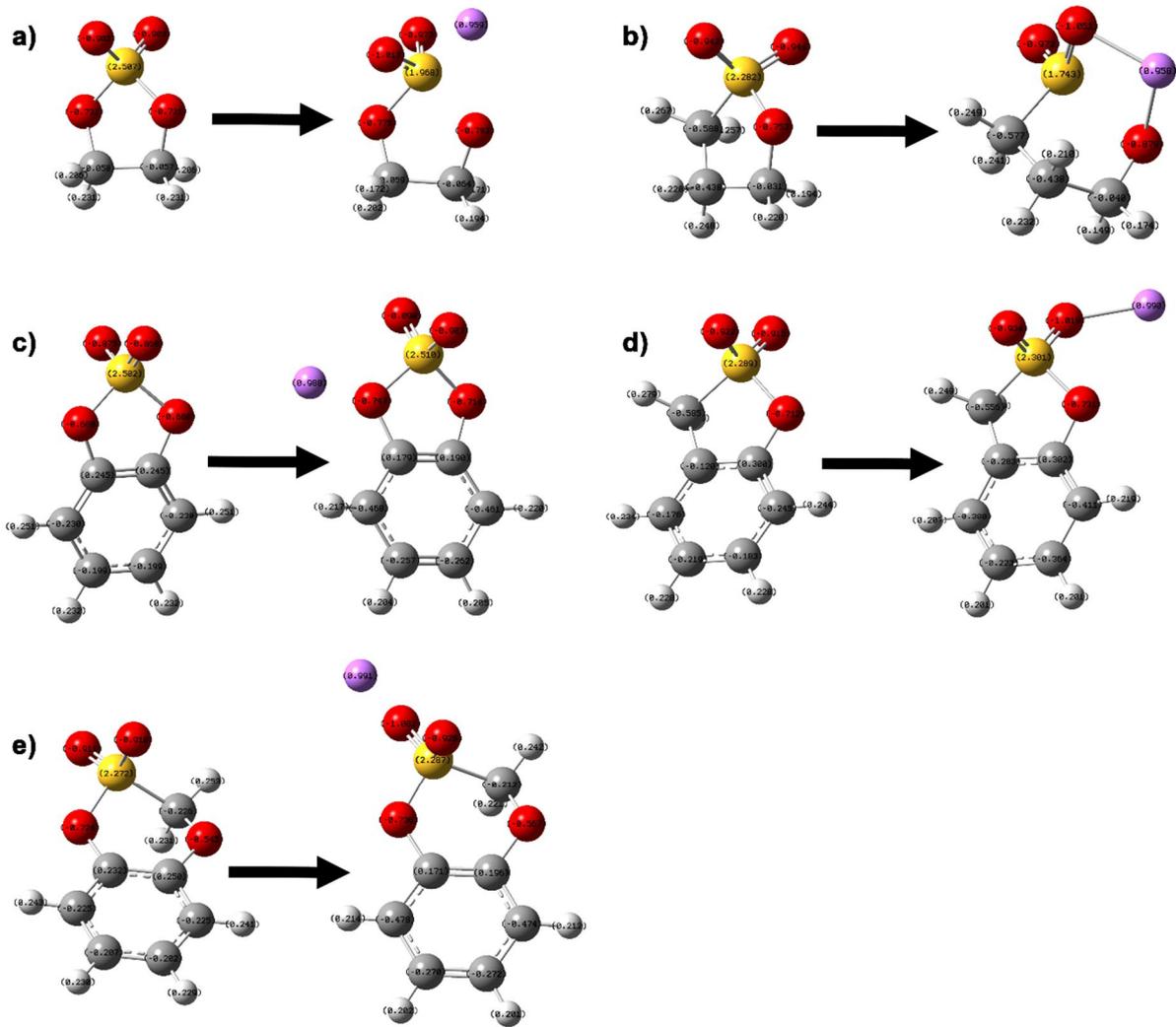


Figure S10. Changes in charge distribution by reduction of additives: DTD (a), PS (b), DTDPh (c), PSpH (d) and PSOpH (e).

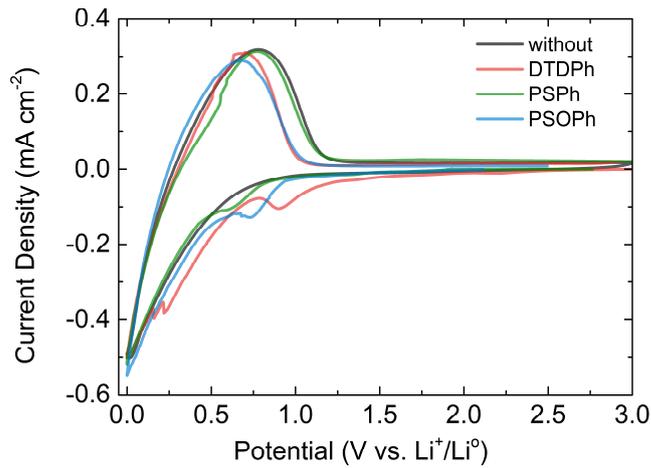


Figure S11. Cyclic voltammogram of graphite electrode with commercial style electrolyte without/with 5 wt% additive; sweep rate 1 mV s⁻¹.

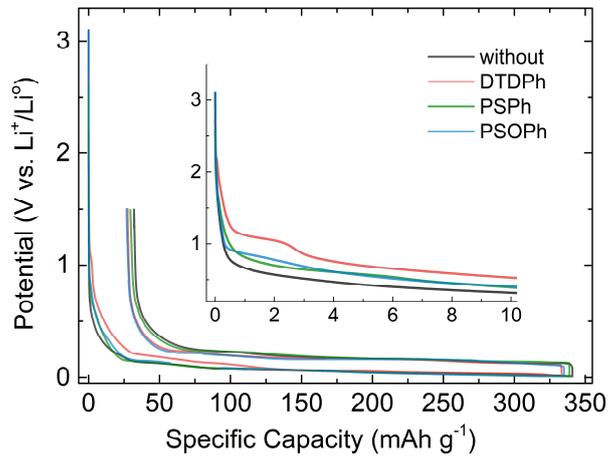


Figure S12. Potential profile of the first charge-discharge cycle of half-cell with commercial style electrolyte without/with 5 wt% of additive at C/20 rate. Inset shows in detail the initial stage of the charge cycle.

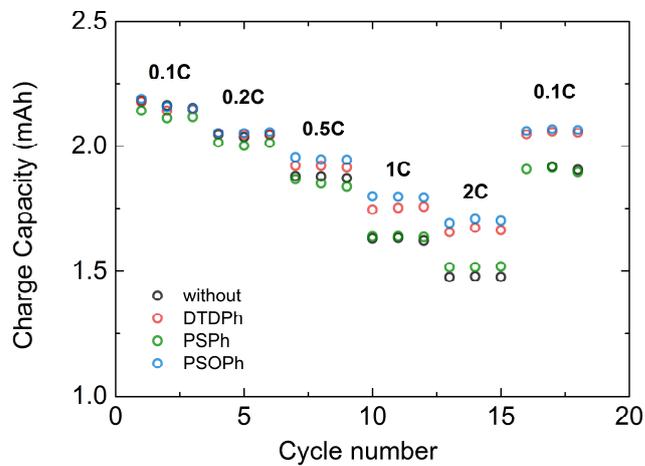


Figure S13. Cycling performance of full-cells without/with 1% of additive at different current rates.

Table S1. Comparison of electrochemical performance of cells with LiTDI-based electrolyte.

Study	Electrodes	Electrolyte	Results
L. Niedzicki et al. <i>J. Power Sources</i> , 2011, 196, 8696	Li/LMO	1M LiTDI in EC:DMC (1:1)	96% capacity retention after 22 cycles at 1C rate
S. Paillet et al. <i>J. Power Sources</i> , 2015, 299, 309	Graphite/Li	1 M LiTDI in EC:DEC (3:7)	Reduction of salt at 0.9-0.7 V vs. Li ⁺ /Li ⁰ (twin peak)
S. Paillet et al. <i>J. Power Sources</i> , 2015, 299, 309	LTO/LFP	1 M LiTDI in EC:DEC (3:7)	85.4% capacity retention after 900 cycles at 1C rate
S. Paillet et al. <i>J. Power Sources</i> , 2015, 299, 309	LTO/NMC	1 M LiTDI in EC:DEC (3:7)	85.8% capacity retention after 450 cycles at 0.25C rate
S.A. Delp et al. <i>Electrochimica Acta</i> , 2016, 209, 498	GC/Li	1 M LiTDI in EC:EMC (3:7)	Reduction of salt at 1.6 V vs. Li ⁺ /Li ⁰
G.G. Eshetu et al. <i>ACS Appl. Mater. Interfaces</i> 2016, 8, 16087	Graphite/Li	1 M LiTDI in EC:DMC (1:2)	Reduction of salt at 0.9V vs. Li ⁺ /Li ⁰
I.A. Shkrob et al. <i>J. Phys. Chem. C</i> , 2016, 120, 28463	Graphite/Li	0.5 M LiTDI in EC:EMC (3:7)	~35% capacity retention after 7 cycles at 0.05C rate; reduction of salt at 0.68-0.74 V vs. Li (twin peak)
This work	Graphite/Li	1 M LiTDI in EC:DMC (1:1)	Reduction of salt at 1.3 and 0.8 V vs. Li ⁺ /Li ⁰
This work	Graphite/LFP	1 M LiTDI in EC:DMC (1:1)	25% capacity retention after 10 cycles (~0% after 30 cycles) at 0.05C rate