

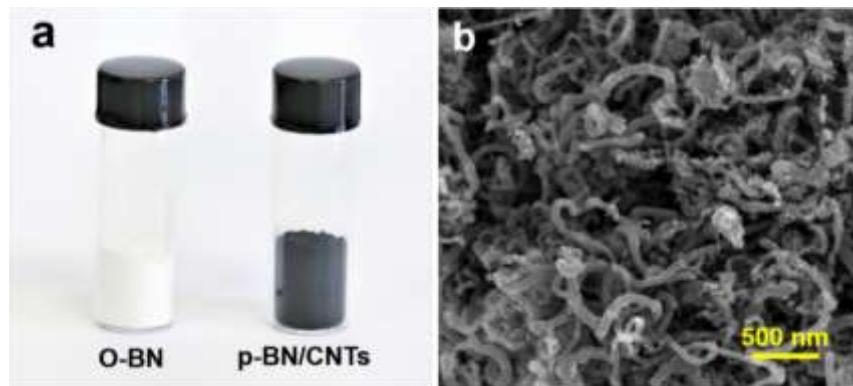
## Supporting Information

### Paragenesis BN/CNTs Hybrid as a Monoclinic Sulfur Host for High Rate and Ultralong Life Lithium-Sulfur Battery

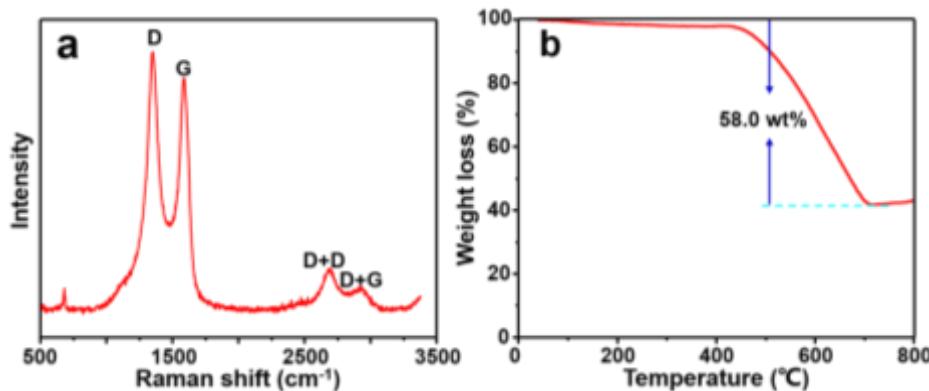
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**Table S1.** The specific surface area and pore volume of polar materials used as sulfur host for lithium-sulfur batteries in the literature.

Host materials	S <sub>BET</sub> (m <sup>2</sup> g <sup>-1</sup> )	V <sub>total</sub> ( cm <sup>3</sup> g <sup>-1</sup> )	Refs.
Co <sub>4</sub> N	48.4	0.237	1
TiN	69.6	0.32	2
NbC	22.37	0.078	3
Co <sub>3</sub> O <sub>4</sub> nanosheets	80.4		4
TiO <sub>2</sub> nanosheet	92	0.27	5
Co <sub>3</sub> O <sub>4</sub> nanoneedle	75.6	0.26	6
SnO <sub>2</sub> shells	66.7		7
hollow TiO <sub>2</sub> sphere	76	0.15	8
TiO <sub>2</sub> nanotube	134.9		9
MoO <sub>2</sub>	70	0.3	10
Rutile TiO <sub>2</sub>	73.6	0.213	11
Co <sub>3</sub> S <sub>4</sub>	31		12
VN/C	38.9	0.084	13
C/TiO <sub>2</sub>	148	0.29	14
La <sub>0.6</sub> Sr <sub>0.4</sub> CoO <sub>3-δ</sub>	70.3		15
CC/TiO <sub>2</sub>	44.1		16
SnO <sub>2</sub>	29	0.125	17
C/SnO <sub>2</sub> nanosheets	89.6		18
CNTs/MnO	51.3		19
BaTiO <sub>3</sub>	12		20
p-BN/CNTs	168	0.33	This work



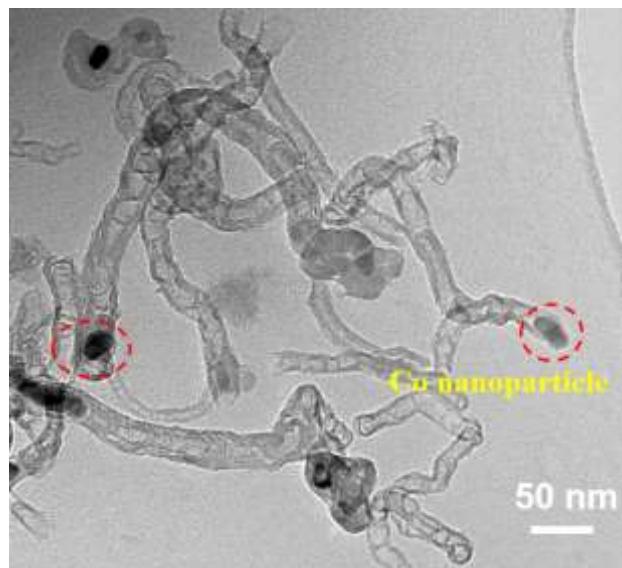
**Figure S1.** (a) The photograph of O-BN and p-BN/CNTs. (b) SEM image of p-BN/CNTs.



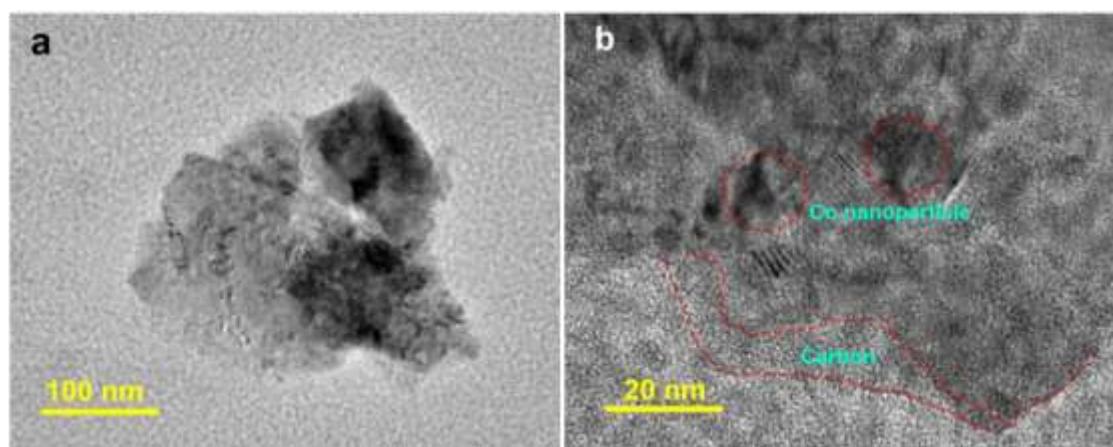
**Figure S2.** (a) Raman spectrum, (b) Thermogravimetric analysis of p-BN/CNTs.

Raman spectrum of p-BN/CNTs (**Figure S2a**) exhibits two intense peaks centered at approximately 1344 and 1582 cm<sup>-1</sup> and two small peaks centered at approximately 2620 and 2900 cm<sup>-1</sup>. The peak at 1582 cm<sup>-1</sup> is generally observed in single crystalline graphite and attributed to the in-plane bond stretching of sp<sup>2</sup> C pairs. As the D peak of carbon and the peak assigned to E<sub>2g</sub> vibration mode of BN are so close, thus the peak at 1344 cm<sup>-1</sup> is maybe the superimposition of the above two peaks. The other peaks, located at ~2620 and 2900 cm<sup>-1</sup>, are called 2D (D+D) and D+G bands and correspond to the second-order of Raman spectrum in overtone and combination modes, respectively.

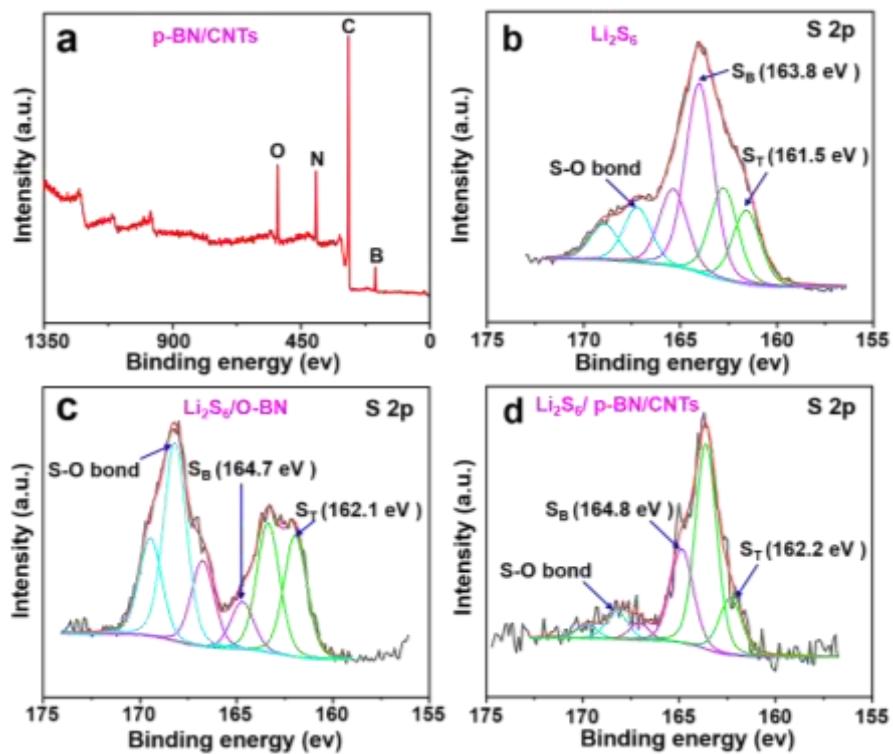
The 58% weight loss of p-BN/CNTs at 400-700 °C is due to combustion of carbon nanotubes. The Co was oxidized into Co<sub>3</sub>O<sub>4</sub> at the high temperature of 400-700 °C under air and 7.9 wt% Co can be generated 10.8 wt% Co<sub>3</sub>O<sub>4</sub>, thus the BN content is approximately 31.2 wt% (58%+10.8%+31.2%=100%).



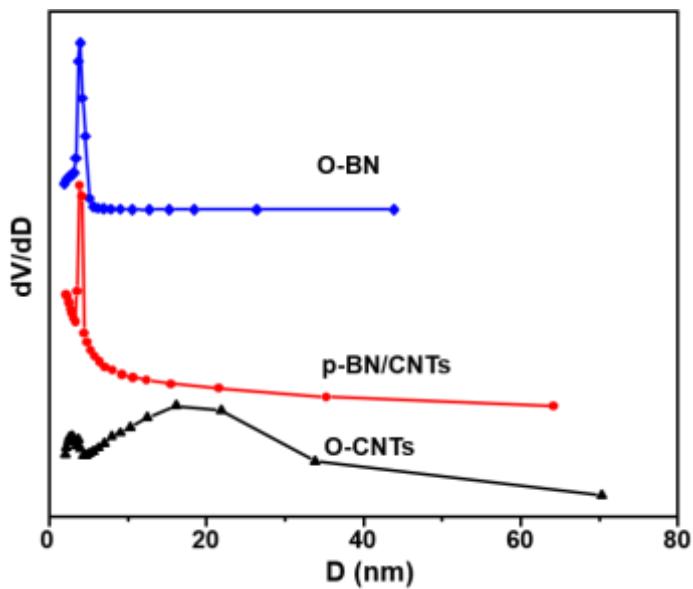
**Figure S3.** TEM images of p-BN/CNTs.



**Figure S4.** TEM images of control sample prepared without  $\text{H}_3\text{BO}_3$



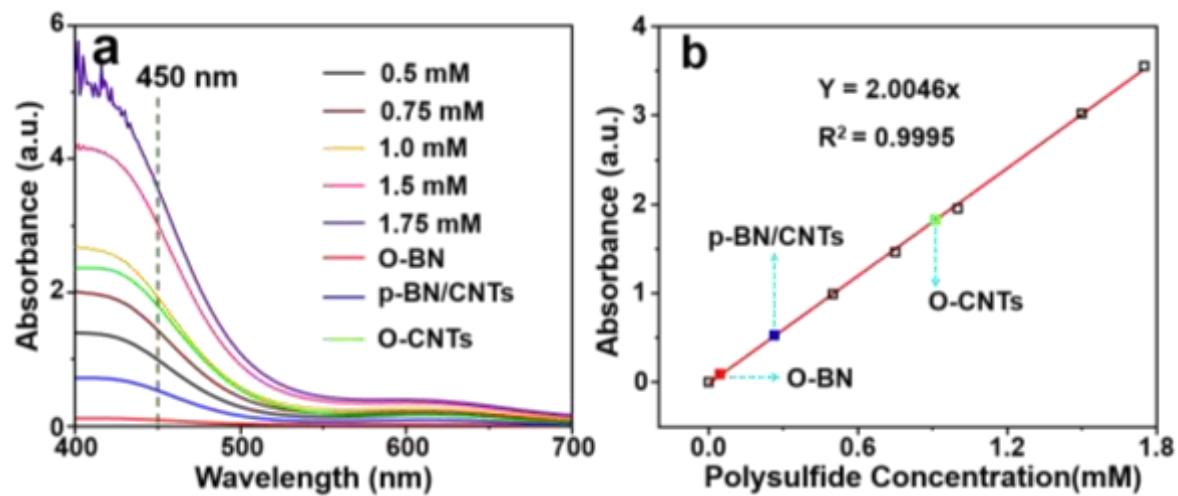
**Figure S5.** (a) XPS spectrum of p-BN/CNTs. (b-d) High-resolution XPS S 2p spectra of (b) Li<sub>2</sub>S<sub>6</sub>, (c) Li<sub>2</sub>S<sub>6</sub>/O-BN and (d) Li<sub>2</sub>S<sub>6</sub>/p-BN/CNTs.



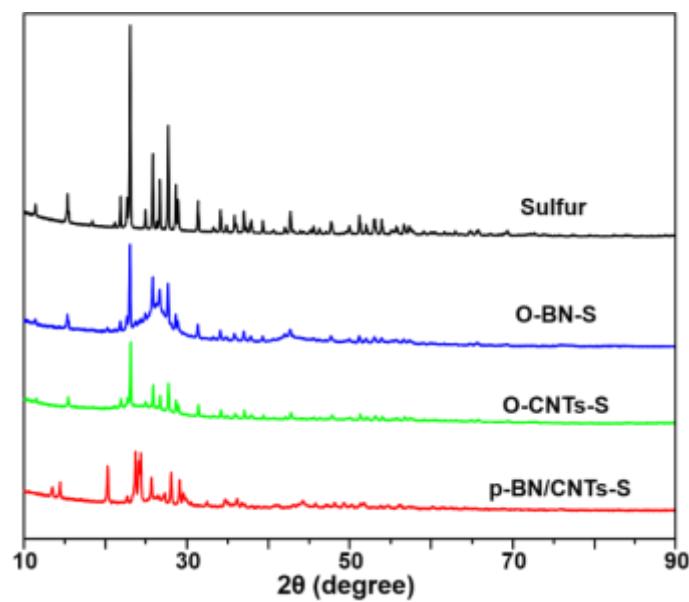
**Figure S6.** Pore size distributions of O-BN, p-BN/CNTs and O-CNTs.



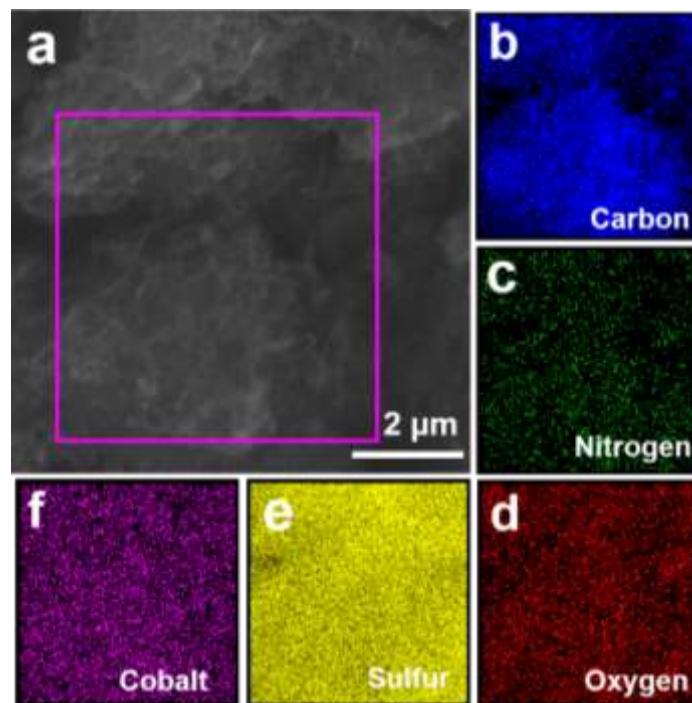
**Figure S7.** Static adsorption test of O-CNTs, O-BN and p-BN/CNTs with Li<sub>2</sub>S<sub>6</sub> solution.



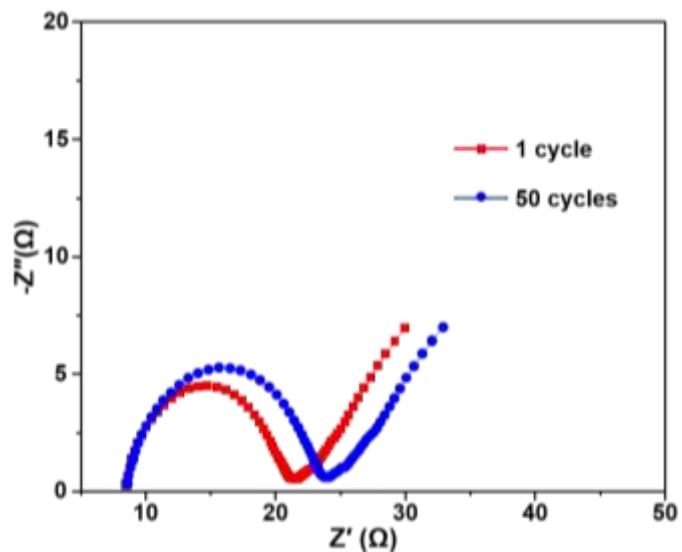
**Figure S8.** (a) UV/Vis-spectra of Li<sub>2</sub>S<sub>6</sub> solution between 0.5 and 1.75 mM. (b) Linear calibration of the absorbance at 450 nm of Li<sub>2</sub>S<sub>6</sub> solutions among 0.5 and 1.75 mM.



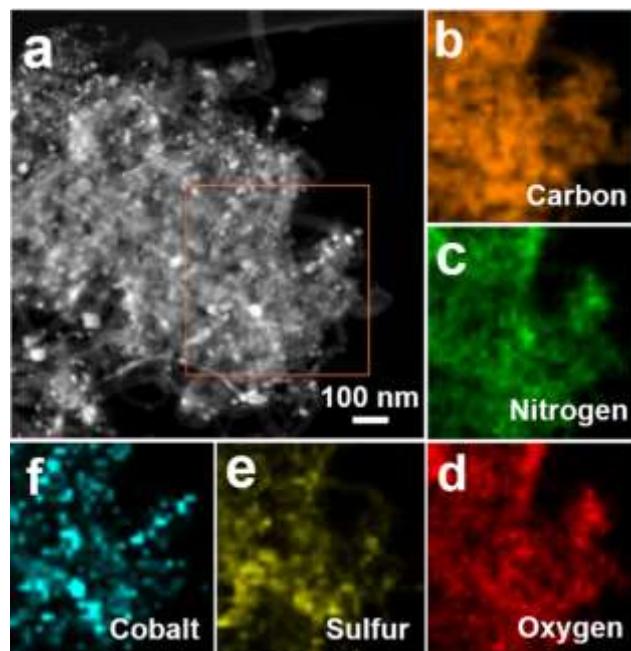
**Figure S9.** XRD patterns of p-BN/CNTs-S after storing it for two months, and sulfur, O-BN-S, O-CNTs.



**Figure S10.** (a) SEM image and EDS mapping of p-BN/CNTs-S. (b-f) EDS elemental maps of (b) carbon, (c) nitrogen, (d) oxygen, (e) sulfur and (f) cobalt.



**Figure S11.** Electrochemical impedance spectra of p-BN/CNTs-S after the first cycle and 50 cycles.



**Figure S12.** (a) STEM image and EDS elemental mapping of p-BN/CNTs-S after 50 cycles. (b-f) EDS elemental maps of (b) carbon, (c) nitrogen, (d) oxygen, (e) sulfur and (f) cobalt.

**Table S2.** The cycle and rate performances of the Li-S batteries in the literature

Materials	Sulfur content (%)	Sulfur loading (mg cm <sup>-2</sup> )	1 C			Rate performance Specific capacity (mAh g <sup>-1</sup> )	Refs.
			Cycle number	Last cycle (mAh g <sup>-1</sup> )	Decay rates (%)		
p-BN/CNTs	66	1.5-2.0	500	816	0.045	840 (4C)	This work
Hollow Co <sub>3</sub> S <sub>4</sub>	53	2.5	450	610	0.08	752 (2C)	21
VN/graphene	56	3	200	917	0.0935	701 (3C)	22
CMK-3/polymer	48	1.0-1.5	100	838	0.3829	850 (3 C)	23
PCNTs@Gra/DTT	63	0.49	400	880	0.05	750 (3C)	24
Carbon/Celgard	70	0.7	200	721	0.20	450 (4 C)	25
Carbon Rods	78.9	0.93	300	700	0.0927	770 (3 C)	26
Polypyrrole-MnO <sub>2</sub>	70	1.0~2.0	500	550	0.071	350 (4 C)	27
Carbon	70	0.9–1.2	250	588	0.1386	480 (3 C)	28
Graphene Oxide	70	1–1.2	400	750	0.08	800 (1 C)	29
Carbon	50.5		150	558	0.13	696.5 (1 C)	30
Co-N-GC	70	1.5-2.0	500	625	0.09	685 (2 C)	31
GN-CNT	76.4	1.3–1.6	500	476	0.09	535 (2 C)	32
Ti <sub>4</sub> O <sub>7</sub>	70	0.4-0.6				861 (2 C)	33
TiO-G	65	1.0				831 (2 C)	34
MoS <sub>2</sub>	75	1.5				850 (2 C)	35
SnO/CNT	70	1.0–1.3				773 (3C)	36
MIL-100(V)/rGO	50	0.9–1.0				600 (0.5 C)	37
MoS <sub>2</sub> /Celgard	65					770 (1 C)	38
Cobalt Hydroxide	75	3				500 (1 C)	39
Nb <sub>2</sub> O <sub>5</sub>	60	1.5				741 (3 C)	40
Carbon nanotubes	50					439 (2 C)	41
RGO–CNTs	73	1.1				712 (2 C)	42
Si/SiO <sub>2</sub> carbon	70					614 (2 C)	43
Carbon	68.3	1-1.5				738 (2 C)	44
Carbon nanotube	70	1.2				300 (4 C)	45
TiO <sub>2-x</sub>	70	1.5				655 (2 C)	46
Carbon	70	1.3				900 (2 C)	47
CNT-Graphene	73	1.3–1.6				696 (2 C)	48
Ti <sub>3</sub> C <sub>2</sub> Nanoribbon	68	0.7-1.0				403 (4 C)	49
Ti <sub>3</sub> C <sub>2</sub> T <sub>x</sub> @Meso-C	72.8	2.0				544.3 (4 C)	50

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