

Supporting information

MOF-derived nanoporous multifunctional fillers enhancing performances of polymer electrolytes for solid-state lithium batteries

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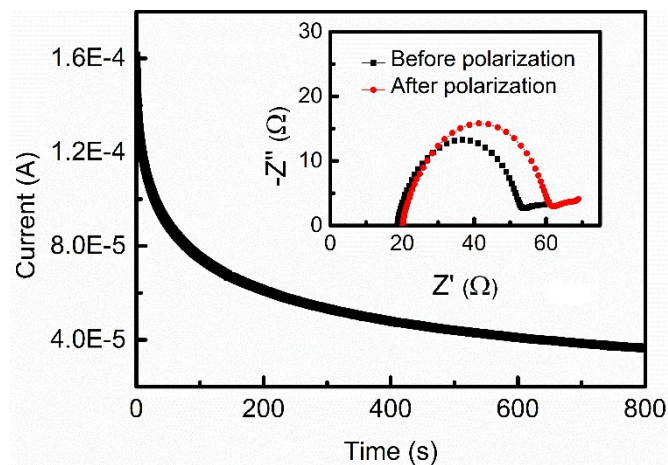


Figure S1 DC polarization curve of UIO/Li-IL. The inset shows AC impedance spectra of symmetric cell before and after DC polarization.

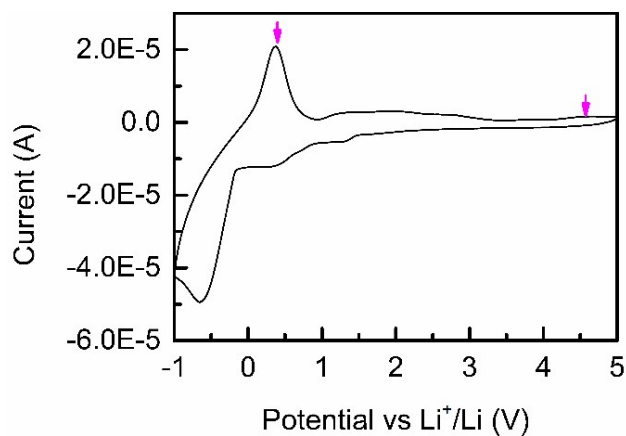


Figure S2 Cyclic voltammetry curve of the PEO solid electrolyte at a scan rate of 0.5 mV s^{-1} .

In the CV curve of the PEO solid electrolyte (Figure S2), there shows a strong peak in the region of 0~0.5 V, which can be ascribed to the Li plating/stripping, and a weak peak at about 4.5 V, which is due to the oxidation of the solid electrolyte.

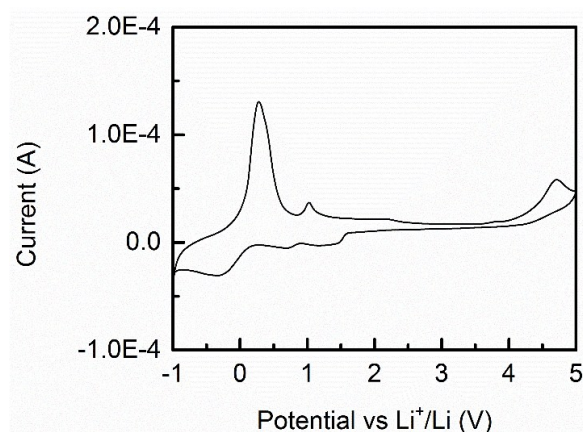


Figure S3 Cyclic voltammetry curve of the Li-IL electrolyte at a scan rate of 0.5 mV s⁻¹.

In the CV curve of the Li-IL electrolyte (Figure S3), there shows a strong peak in the region of 0~0.5 V, which can be ascribed to the Li plating/stripping, and two peaks at 1 V and about 4.5 V, which are due to the reaction between Li-IL and the Li electrode.

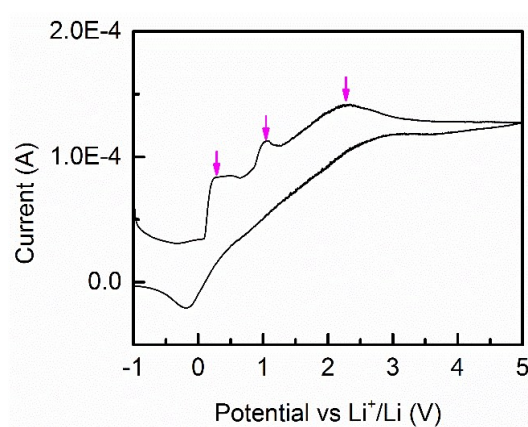


Figure S4 Cyclic voltammetry curve of the UIO/Li-IL (15/16) at a scan rate of 0.5 mV s⁻¹.

To evaluate the stability of UIO-66, the UIO/Li-IL with high UIO-66 content is chosen for the CV test. In the CV curve of UIO/Li-IL (15/16) (Figure S4), there shows a peak in the region of 0~0.5 V, which can be ascribed to the Li plating/stripping, a peak at 1 V, which is due to the reaction of Li-IL, and a peak in the region of 1.5~3 V, originating from the reaction of UIO-66 with metallic Li.

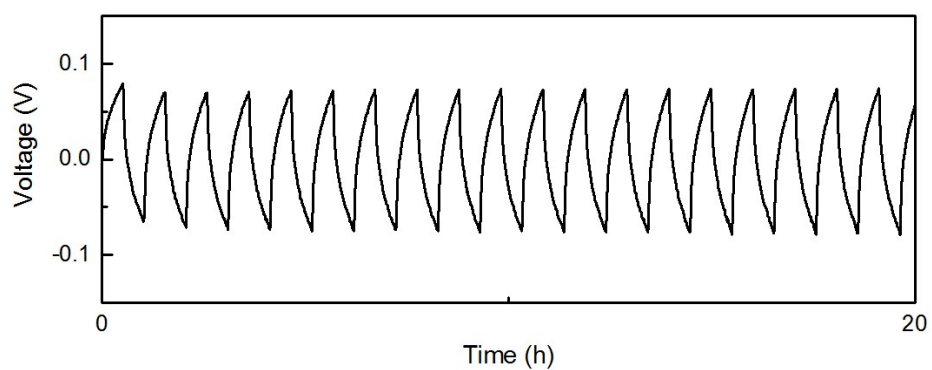


Figure S5 Voltage profile of the Li/UIO/Li-IL(3/4)/Li cell at 60 °C.

As shown in Figure S5, the voltages of the cell in individual cycles are identical, demonstrating stable Li plating/stripping in UIO/Li-IL, and no electronic conduction.