

Supporting information to the article:

**Influence of Lanthanides on Spin-Relaxation and Spin-Structure in a Family of  $\text{Fe}_7\text{Ln}_4$  Single Molecule Magnets**

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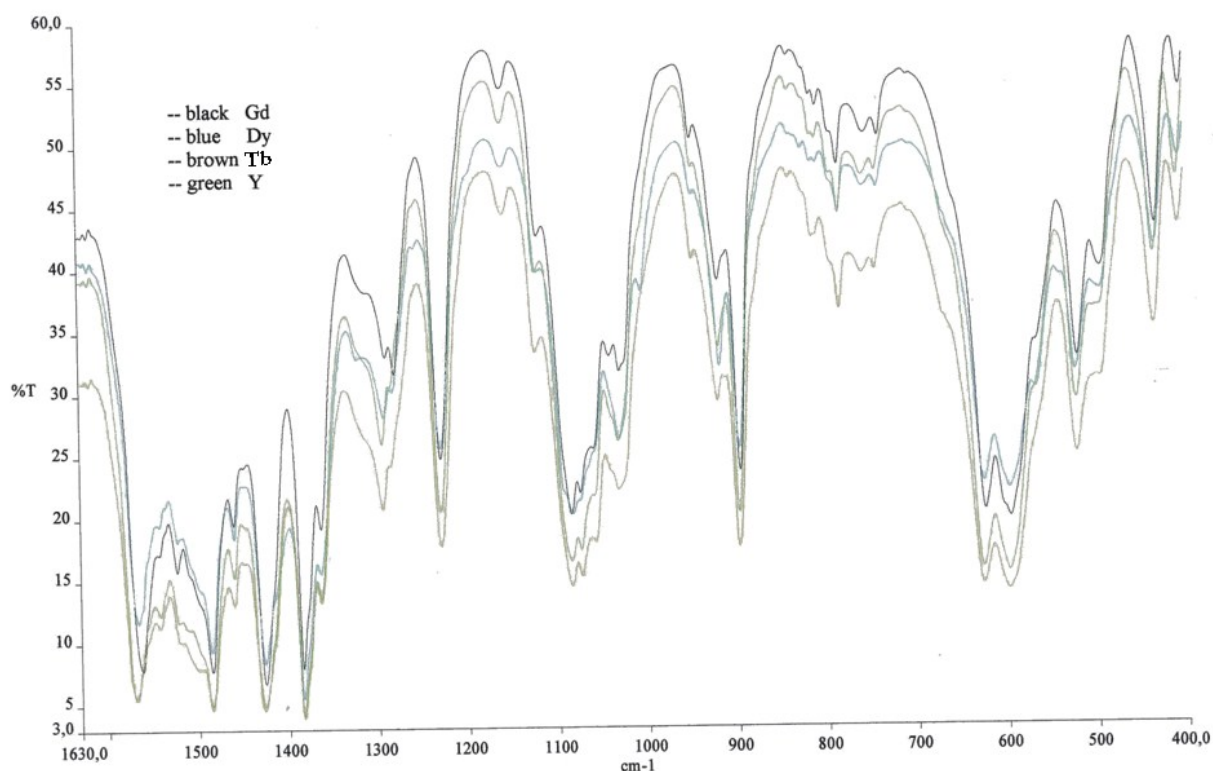


Figure S1: IR spectra for 1-4 clusters.

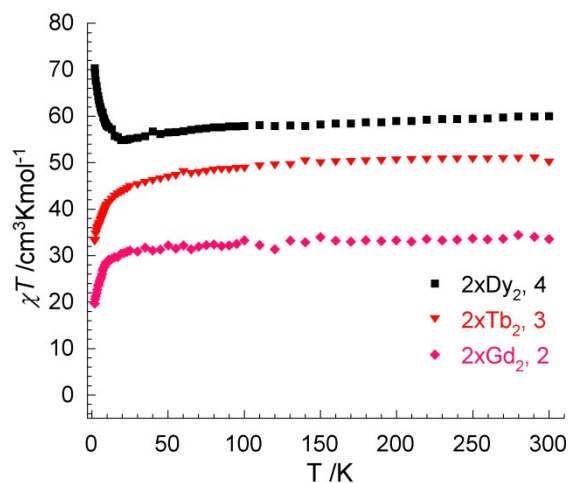
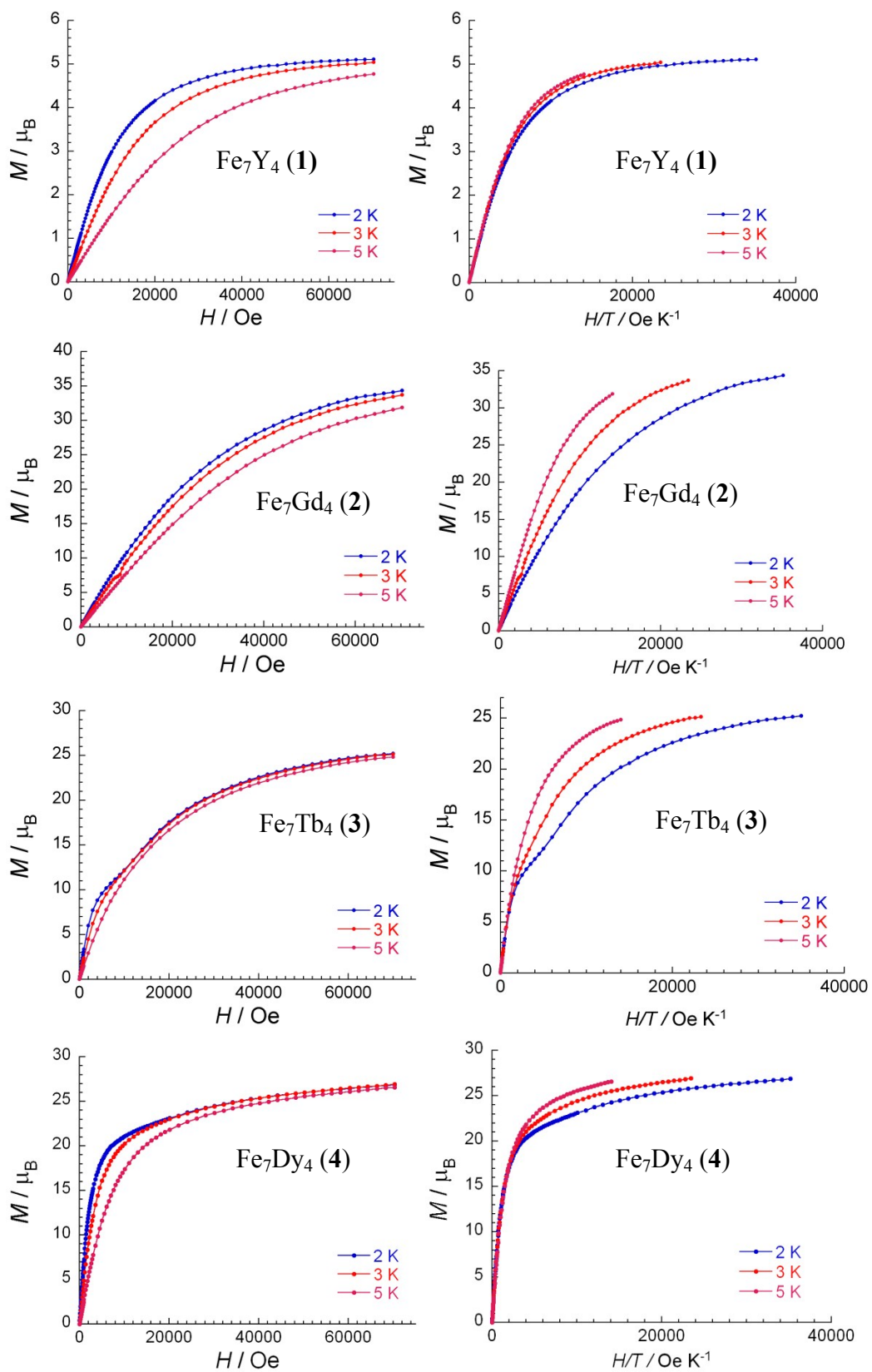
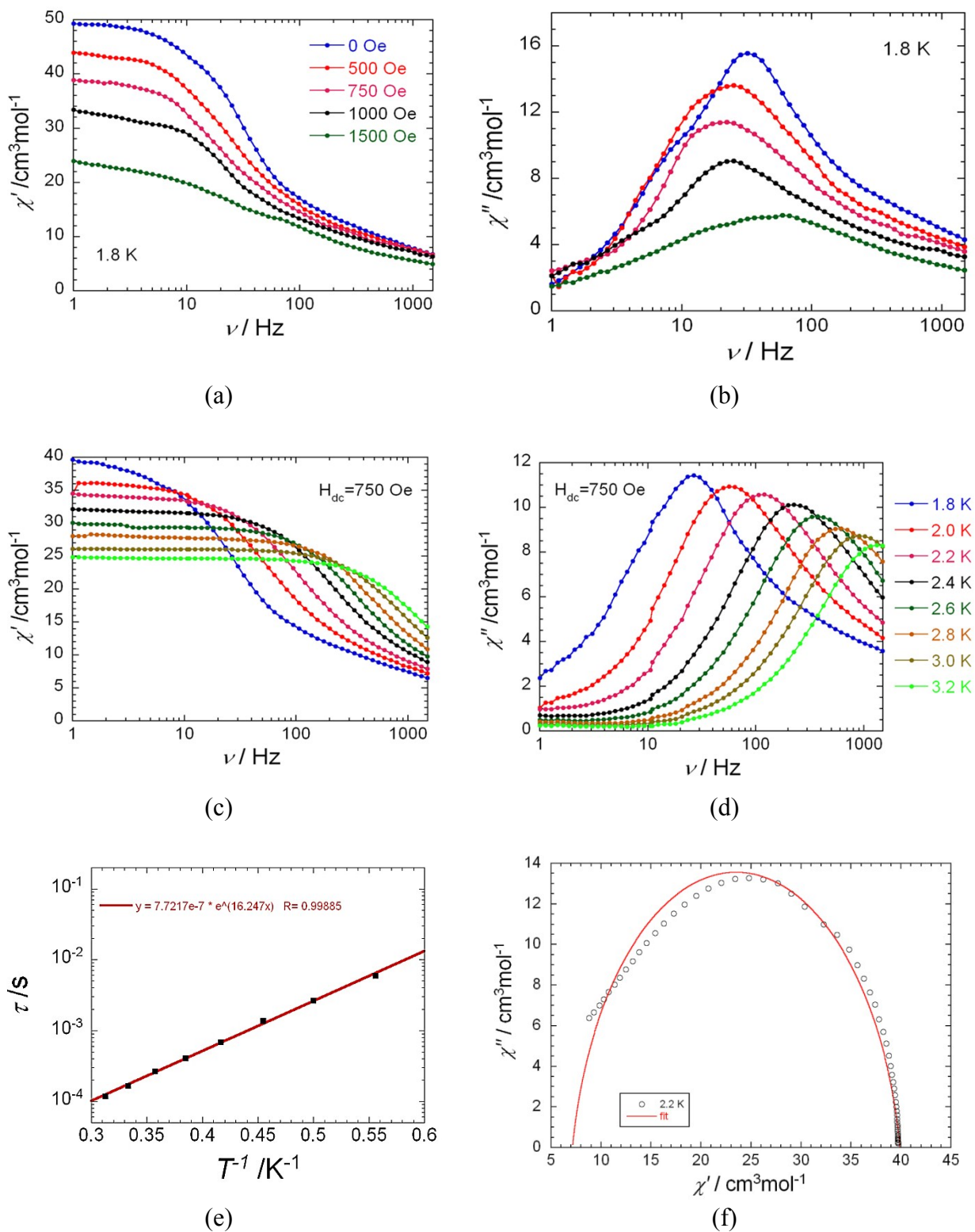


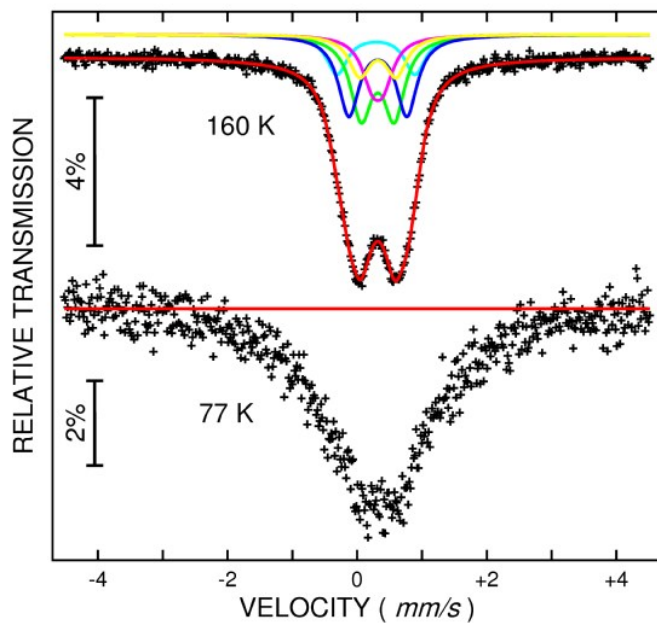
Fig. S2. The  $\chi T$  vs  $T$  plots for 2, 3 and 4 with the data for 1 subtracted, i.e. without the magnetic contribution of the  $\text{Fe}_7$  unit.



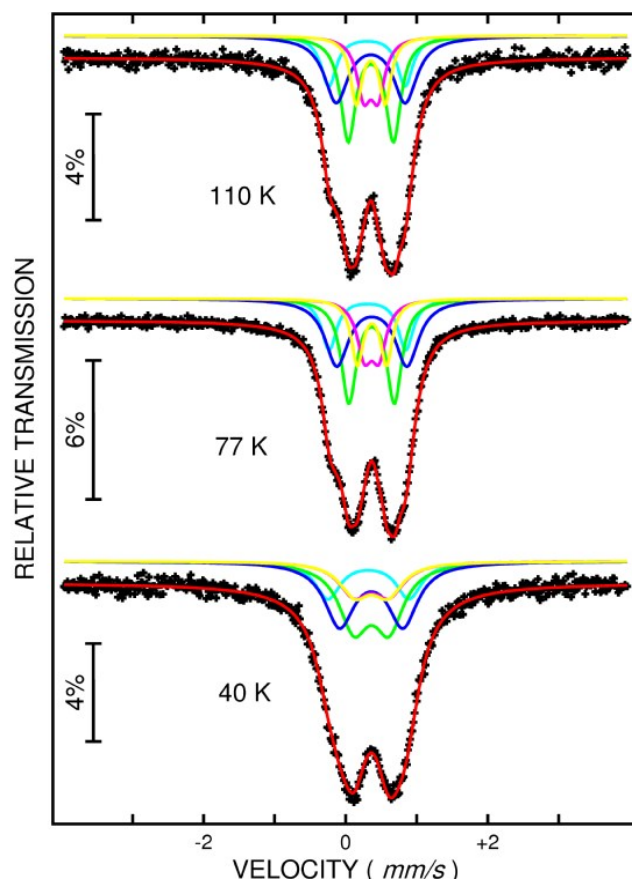
**Figure S3:** Field-dependence of magnetization of compounds 1-4.



**Figure S4:** (a and b) The frequency dependence of the in-phase ( $\chi'$ ) and out-of-phase ( $\chi''$ ) signals under applied dc fields for **4**; (c and d) The frequency dependence of the in-phase ( $\chi'$ ) and out-of-phase ( $\chi''$ ) signals in  $H_{dc} = 750$  Oe at different temperatures for **4**; (e) plot of the magnetization relaxation rate *versus* the inverse of the absolute temperature for **4**. The solid line represents a least-squares fit of the data to the Arrhenius equation; (f) The Cole-Cole plot of  $\chi''$  vs  $\chi'$  at 2.2K for **4**.



**Figure S5:** Mössbauer spectra for **1** ( $\text{Fe}_7\text{Y}_4$ ) at 160 and 77 K in zero-field.



**Figure S6:** Mössbauer spectra for **3** ( $\text{Fe}_7\text{Tb}_4$ ) at 110, 77 and 40 K in zero-field.

**Table S1.** Mössbauer data\* for **1**, **3** and **4** at 3 K in zero-field and at 3 K in applied magnetic fields of 6 or 6.5 T.

Compound, temperature, applied field	Fe sites	$\delta^{[a]}$ , mm/s	$\Delta E_Q$ or $\varepsilon^{[c]}$ , mm/s	$\Gamma^{[d]}$ , mm/s	$\theta^{[d]}$ , °	$B_{\text{eff}}$ , T
<b>(1)</b> Fe <sub>7</sub> Y <sub>4</sub> 3K, 6.5T	Fe(1)	0.49(1)	-0.14(1)	0.34(1)	68.7	46.37
	Fe(2)	0.46(1)	-0.42(1)	0.34(1)	68.7	35.97
	Fe(3)	0.41(3)	-0.42(1)	0.34(1)	68.7	33.27
	Fe(4)	0.49(1)	0.10(1)	0.34(1)	68.7	45.84
	Fe(5)	0.49(1)	0.10(1)	0.34(1)	68.7	45.84
	Fe(6)	0.46(1)	-0.003(3)	0.34(1)	68.7	36.43
	Fe(7)	0.46(1)	-0.003(3)	0.34(1)	68.7	36.43
<b>(3)</b> Fe <sub>7</sub> Tb <sub>4</sub> 3 K	Fe(1)	0.49(1)	-0.23(1)	0.34(1)	57.6	42.45
	Fe(2)	0.46(1)	0.21(1)	0.34(1)	57.6	42.70
	Fe(3)	0.44(1)	0.06(1)	0.34(1)	57.6	41.92
	Fe(4)	0.50(1)	0.11(1)	0.34(1)	57.6	40.77
	Fe(5)	0.50(1)	0.11(1)	0.34(1)	57.6	40.77
	Fe(6)	0.50(2)	0.19(1)	0.34(1)	57.6	44.32
	Fe(7)	0.50(2)	0.19(1)	0.34(1)	57.6	44.32
<b>(3)</b> Fe <sub>7</sub> Tb <sub>4</sub> 3K, 6.5T	Fe(1)	0.49 <sup>[b]</sup>	0.12(1)	0.43(1)	62.7	46.05
	Fe(2)	0.46 <sup>[b]</sup>	0.19(1)	0.43(1)	62.7	39.84
	Fe(3)	0.44 <sup>[b]</sup>	0.29(1)	0.43(1)	62.7	35.65
	Fe(4)	0.50 <sup>[b]</sup>	-0.08(1)	0.43(1)	62.7	45.23
	Fe(5)	0.50 <sup>[b]</sup>	-0.08(1)	0.43(1)	62.7	45.23
	Fe(6)	0.50 <sup>[b]</sup>	0.17(1)	0.43(1)	62.7	37.40
	Fe(7)	0.50 <sup>[b]</sup>	0.17(1)	0.43(1)	62.7	37.40
<b>(4)</b> Fe <sub>7</sub> Dy <sub>4</sub> 3 K	Fe(1)	0.50(1)	0.20(1)	0.43(1)	51.9	39.06
	Fe(2)	0.48(2)	0.15(1)	0.43(1)	51.9	42.52
	Fe(3)	0.45(1)	0.45(1)	0.43(1)	51.9	41.31
	Fe(4)	0.50(1)	0.09(1)	0.43(1)	51.9	39.02
	Fe(5)	0.50(1)	0.09(1)	0.43(1)	51.9	39.02
	Fe(6)	0.48(2)	0.04(1)	0.43(1)	51.9	41.15
	Fe(7)	0.48(2)	0.04(1)	0.43(1)	51.9	41.15
<b>(4)</b> Fe <sub>7</sub> Dy <sub>4</sub> 3K, 6T	Fe(1)	0.50 <sup>[b]</sup>	0.09(1)	0.44(1)	64.1	44.53
	Fe(2)	0.48 <sup>[b]</sup>	0.10(1)	0.44(1)	64.1	37.77
	Fe(3)	0.45 <sup>[b]</sup>	0.11(1)	0.44(1)	64.1	35.53
	Fe(4)	0.50 <sup>[b]</sup>	0.02(1)	0.44(1)	64.1	44.65
	Fe(5)	0.50 <sup>[b]</sup>	0.02(1)	0.44(1)	64.1	44.65
	Fe(6)	0.48 <sup>[b]</sup>	-0.02(1)	0.44(1)	64.1	37.59
	Fe(7)	0.48 <sup>[b]</sup>	-0.02(1)	0.44(1)	64.1	37.59

<sup>[a]</sup> Relative to  $\alpha$ -Fe at room temperature. <sup>[b]</sup> Fixed values. <sup>[c]</sup> For magnetically-split spectra the quadrupole shifts,  $\varepsilon = \frac{1}{2}\Delta E_Q(3\cos^2\varphi - 1)$ .  $\varphi$  - Euler angle between the  $B_{\text{int}}$  and the main electrical field gradient (EFG) principal axis ( $V_{zz}$ ). <sup>[d]</sup> Constrained to the same value for all sites.  $\delta$  is the isomer shift,  $\Delta E_Q$  - quadrupolar splitting,  $\theta$  - angle between  $B_{\text{eff}}$  and direction of  $\gamma$ -rays.  $B_{\text{eff}}$  - hyperfine field. The statistical errors are given in parentheses. The estimated absolute errors on  $B_{\text{eff}}$  are  $\pm 0.05$  mm/s. The relative areas for the doublets and sextets have been constrained to the 1:1:1:2:2 ratio.

\*Given that spectra for all compounds at 40K and higher temperatures gave very similar hyperfine parameters to the published ones (V. Mereacre, et al., *Chem.-Eur. J.*, 2011, **17**, 123) (taking into consideration the compensation of isomeric shifts for second order Doppler shift), their parameters are not given.

**Table S2:** Calculated bond-lengths (Å) and bond-angles (°) of Fe7 core with oxygen in **2** (Fe7Gd4) polynuclear cluster show good agreement with experiments.

<b>2</b> Gd4Fe7	Avg. Bond-length (Å)		<b>2</b> Gd4Fe7	Bond-angle (°)	
	Expt.	Theory		Expt.	Theory
<b>Fe-O</b>	2.06	2.01	<b>Fe1-O-Fe4</b>	97.7 <sup>o</sup> ; 93.9 <sup>o</sup>	95.14 <sup>o</sup> ; 94.52 <sup>o</sup>
			<b>Fe1-O-Fe5</b>	94.5 <sup>o</sup> ; 94.9 <sup>o</sup>	97.88 <sup>o</sup> ; 96.56 <sup>o</sup>
			<b>Fe1-O-Fe(2,3,6,7)/</b> <b>Fe4-O-Fe(2,3,6,7)/</b> <b>Fe5-O-Fe(2,3,6,7)/</b>	119.9 <sup>o</sup> - 134.6 <sup>o</sup> .	121.30 <sup>o</sup> - 137.80 <sup>o</sup>

### Geometry index ( $\tau$ ) for 5-coordinated Fe(3) atom in {Fe<sub>7</sub>Ln<sub>4</sub>} complexes

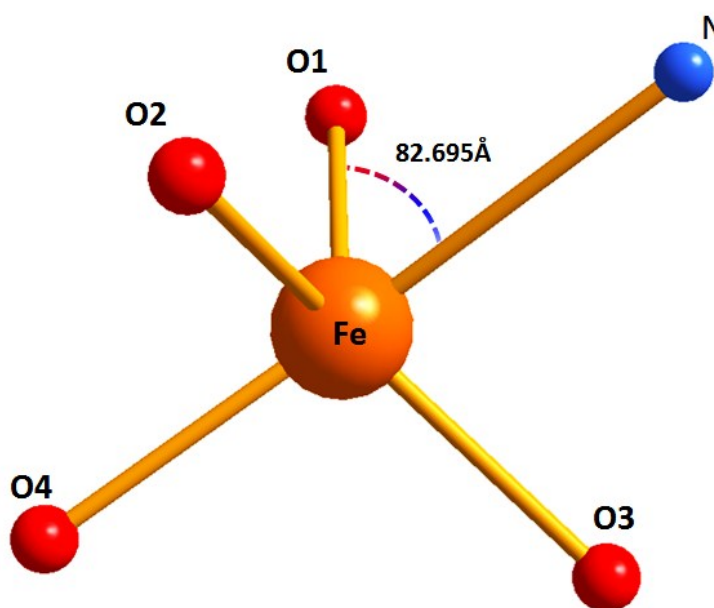
To distinguish whether the geometry of the Fe(3) coordination center is trigonal bipyramidal or square pyramidal, the  $\tau$  parameter was calculated according to Addison et al. [R1]:

$$\tau = \frac{\beta - \alpha}{60^\circ} \approx -0.01667\alpha + 0.01667\beta \quad (1)$$

where:  $\beta > \alpha$  are the two greatest valence angles of the coordination center (Fig. S8). When  $\tau$  is close to 0 the geometry is similar to square pyramidal, while if  $\tau$  is close to 1 the geometry is similar to trigonal bipyramidal.

Angles, Å:

O1-Fe-N = 82.695
O2-Fe-N = 80.264
O3-Fe-N = 80.821
<b>O4-Fe-N = 164.233 (<math>\beta</math>)</b>
O1-Fe-O3 = 115.400
O1-Fe-O2 = 115.887
<b>O2-Fe-O3 = 121.843 (<math>\alpha</math>)</b>
O4-Fe-O1 = 109.827
O4-Fe-O2 = 85.484
O4-Fe-O3 = 101.309



**Figure S7:** Coordination environment of pentacoordinated Fe(3) atom in complex **2**.

$$\tau = -0.01667 * 121.445 + 0.01667 * 164.013 = 0.7096 \text{ (for complex 1, Fe}_7\text{Y}_4\text{)}$$

$$\tau = -0.01667 * 121.843 + 0.01667 * 164.233 = 0.7067 \text{ (for complex 2, Fe}_7\text{Gd}_4\text{)}$$

$$\tau = -0.01667 * 121.816 + 0.01667 * 164.585 = 0.7129 \text{ (for complex 3, Fe}_7\text{Tb}_4\text{)}$$

$$\tau = -0.01667 * 121.544 + 0.01667 * 164.221 = 0.7114 \text{ (for complex 4, Fe}_7\text{Dy}_4\text{)}$$

[R1] A. W. Addison, T. N. Rao, J. Reedijk, J. van Rijn and G. C. Verschoor, *J. Chem. Soc., Dalton Trans.*, 1984, 1349.