

Supporting Information

Robust non-covalent and covalent anchored *N,N,N',N'*-tetramethyl-*p*-phenylenediamine derivative on electrode surface via spontaneous physical immobilization and in situ generated aryldiazonium ion electro-grafting: implication for on-surface chemistry and electro-catalytic determinations

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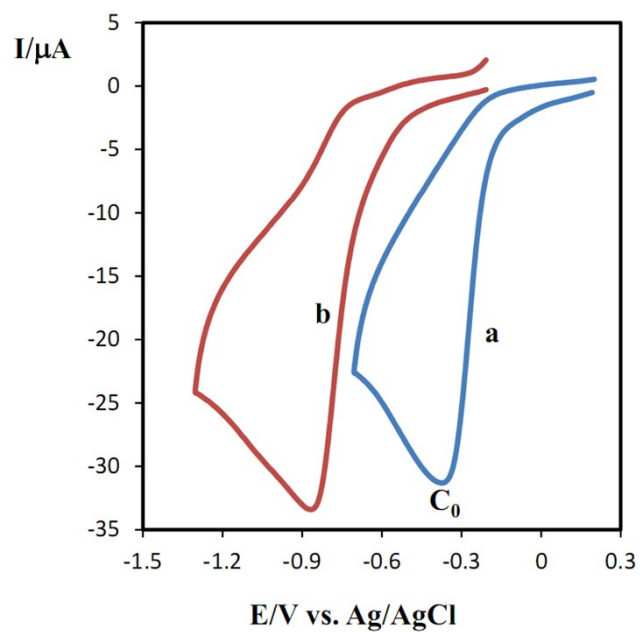


Fig. 1S: Cyclic voltammograms of 0.5 mM MNPD at a glassy carbon electrode in acetonitrile/aqueous solutions (4 ml ACN/6 ml buffer solution, 0.2 M), with pH values of (a) 2 and (b) 7. Scan rate: 100 mV/s.

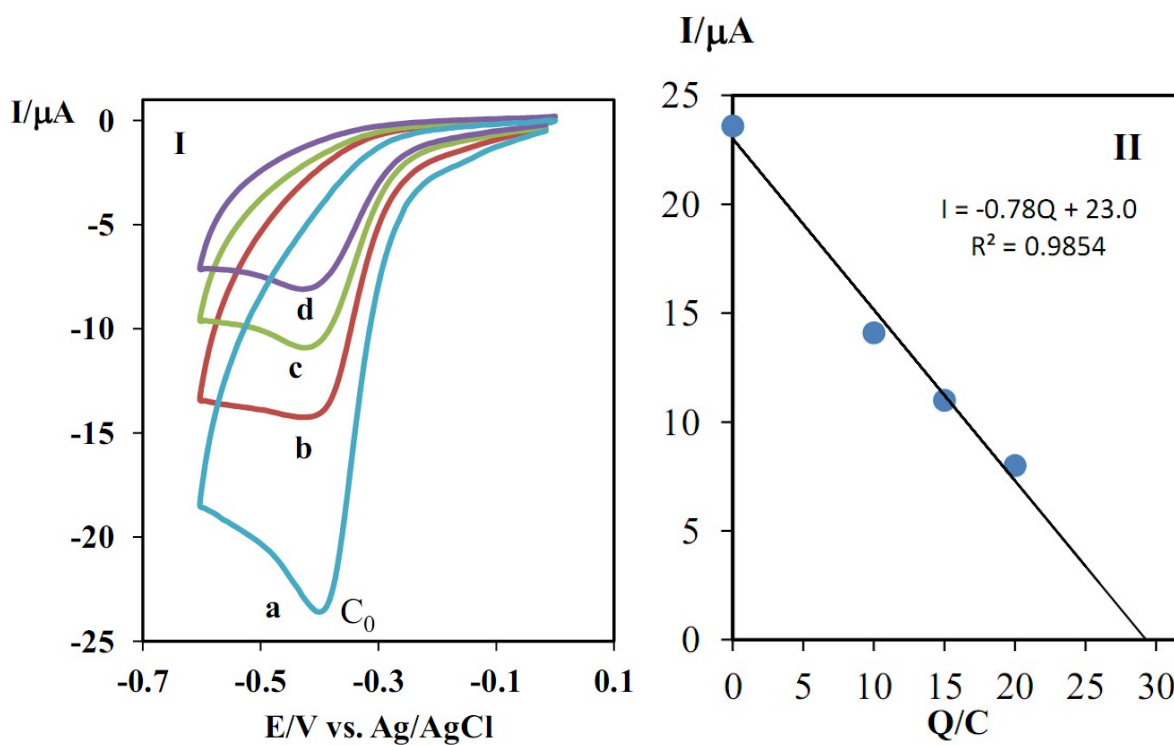


Fig. 2S. Part I: Cyclic voltammograms of 0.04 mmol MNPd, in acetonitrile/aqueous phosphate buffer solution with $p\text{H}=2.0$ (45 ml ACN/45 ml buffer solution) at glassy carbon electrode during controlled-potential coulometry at -0.6 V versus Ag/AgCl. After consumption of: (a) 0, (b) 10, (c) 15 and (d) 20 C. Part II: variation of peak current (I_{pC0}) versus charge consumed. Scan rate 100 mVs^{-1} .

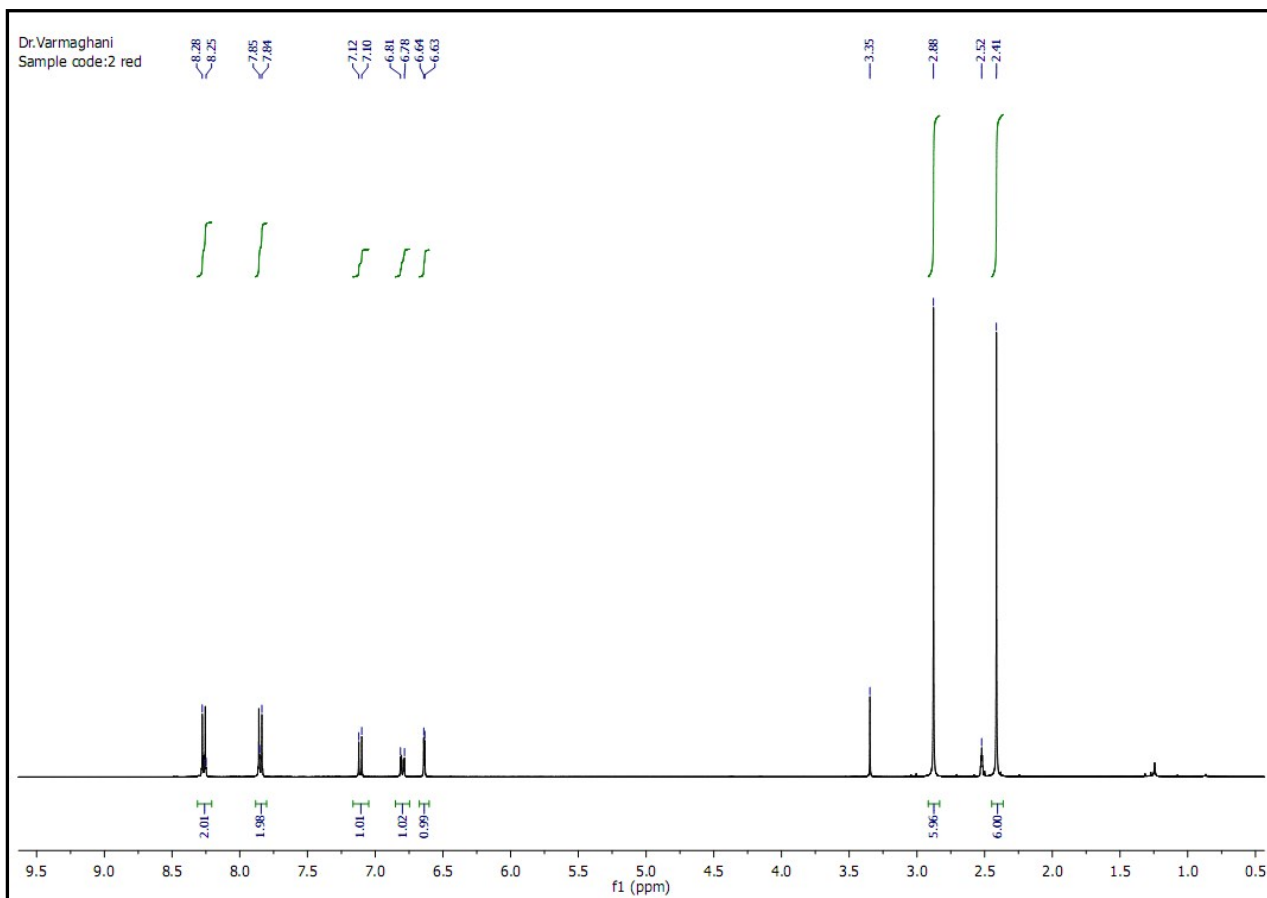


Figure 3S: ^1H NMR of MNPD.

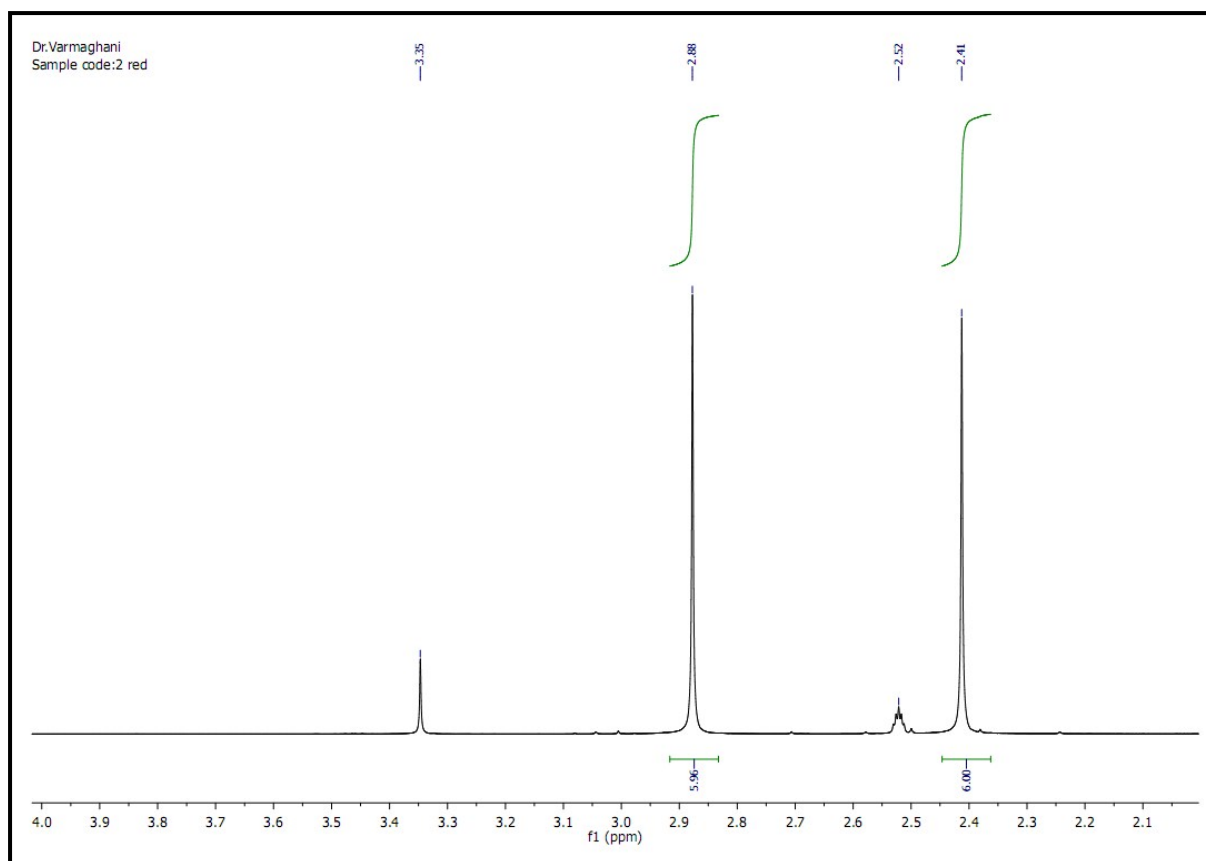


Figure 4S: Expanded ^1H NMR of MNPD in the range of 2-4 ppm

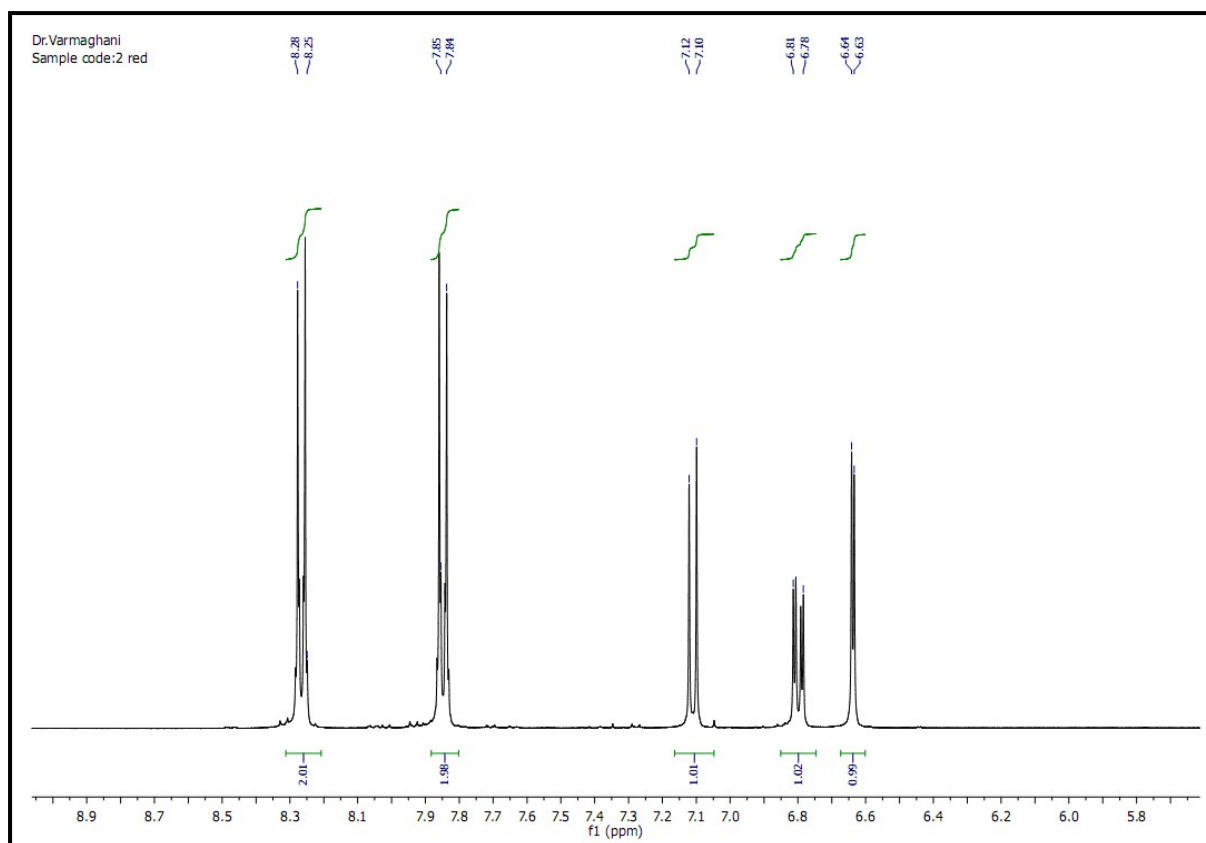


Figure 5S: Expanded ^1H NMR of MNPD in the range of 5-9 ppm

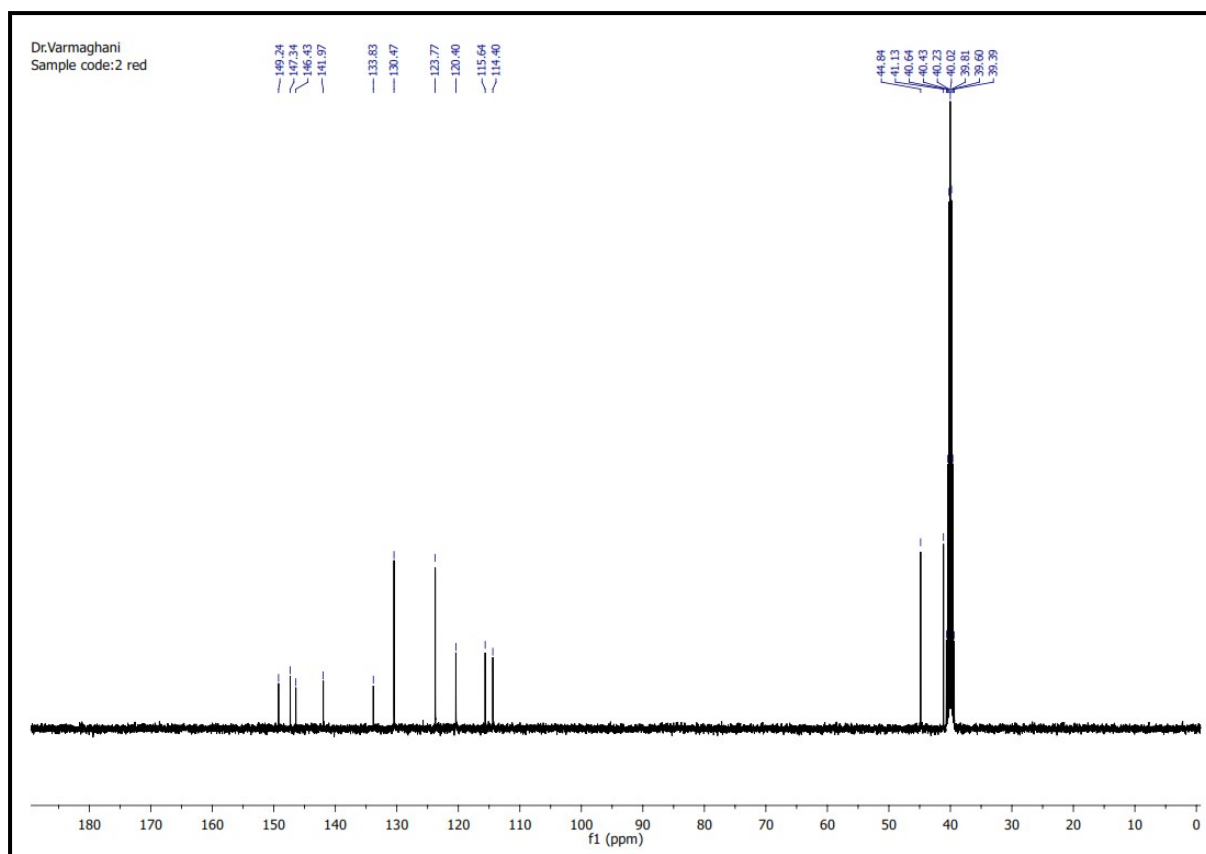


Figure 6S: ^{13}C NMR of MNPD

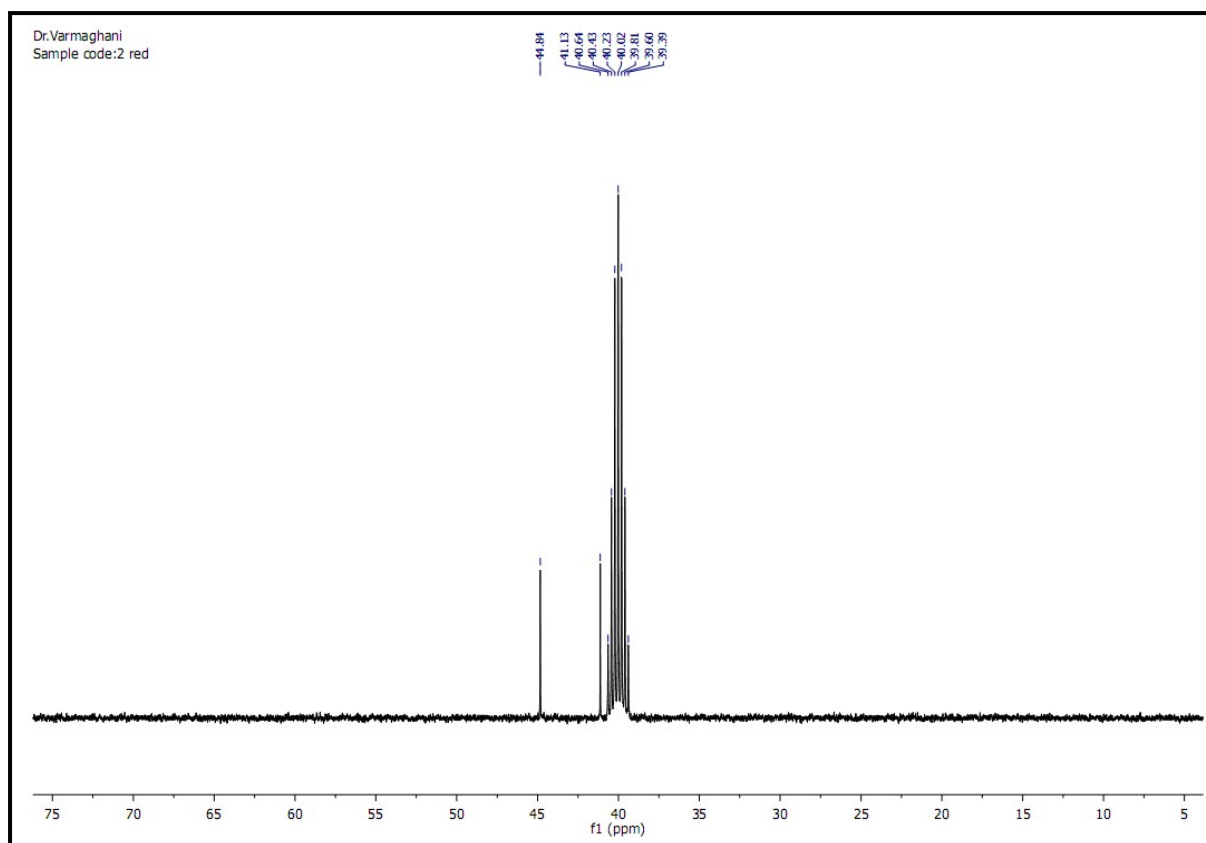


Figure 7S: Expanded ^{13}C NMR of MNPD in the range of 0 to 75 ppm

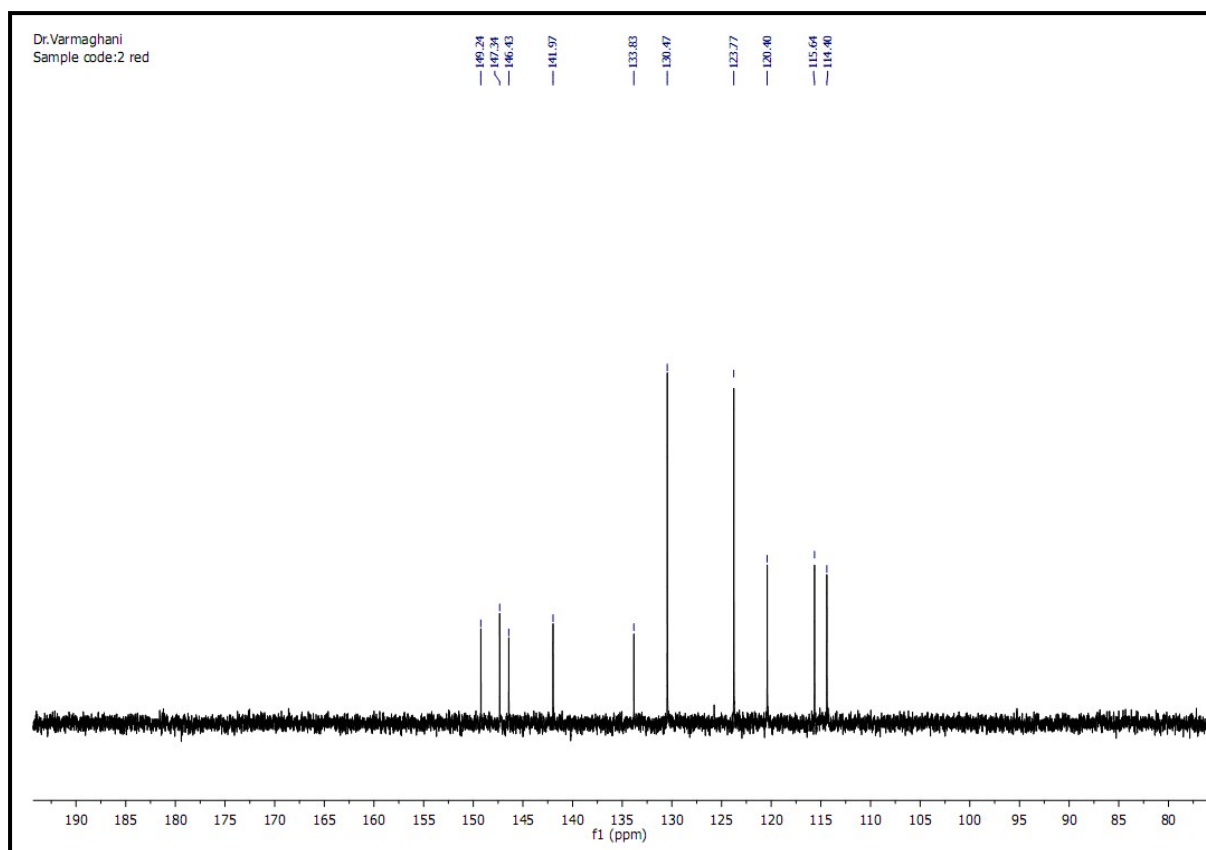


Figure 8S: Expanded ^{13}C NMR of MNPD in the range of 75 to 200 ppm

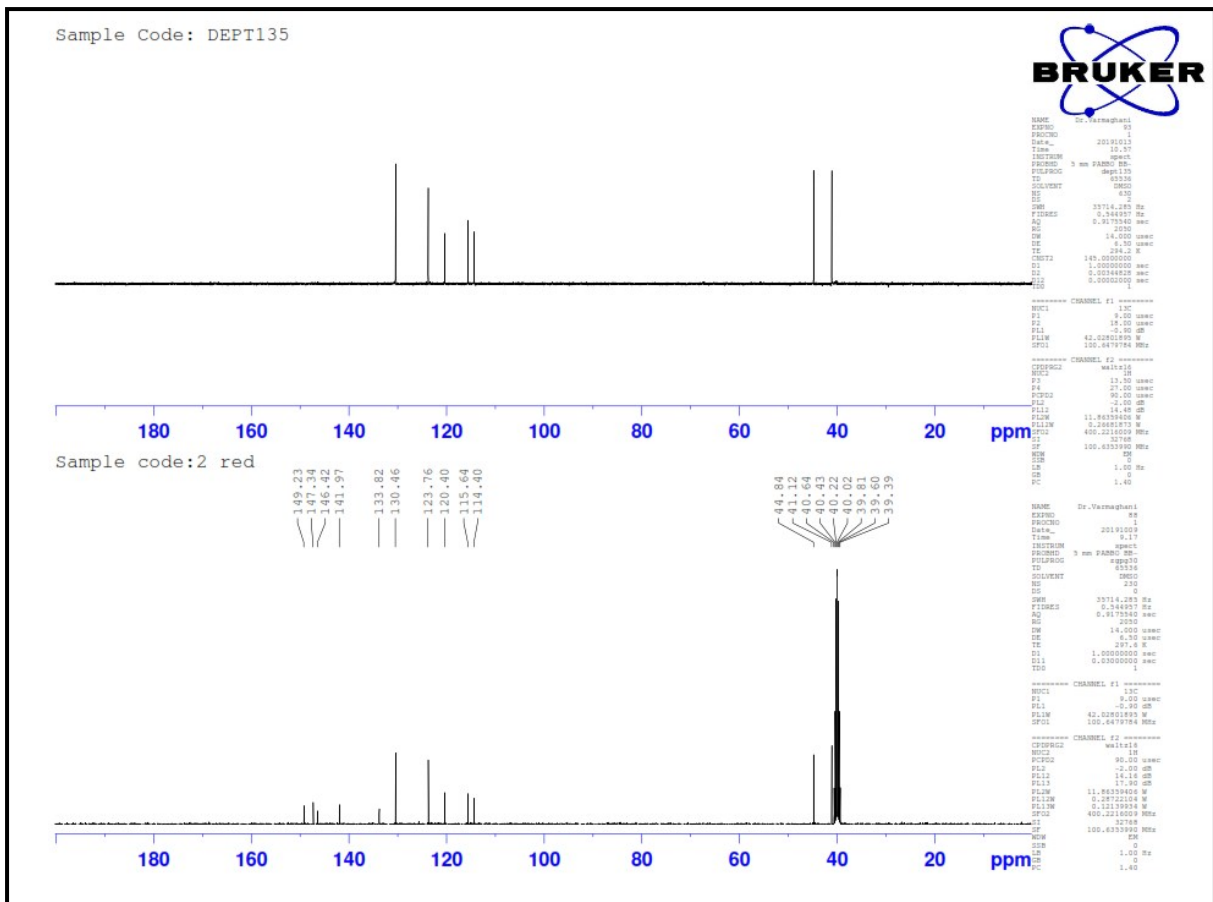


Figure 9S: Comparison of ^{13}C NMR and DEPT 135 $^{\circ}$ of MNPD

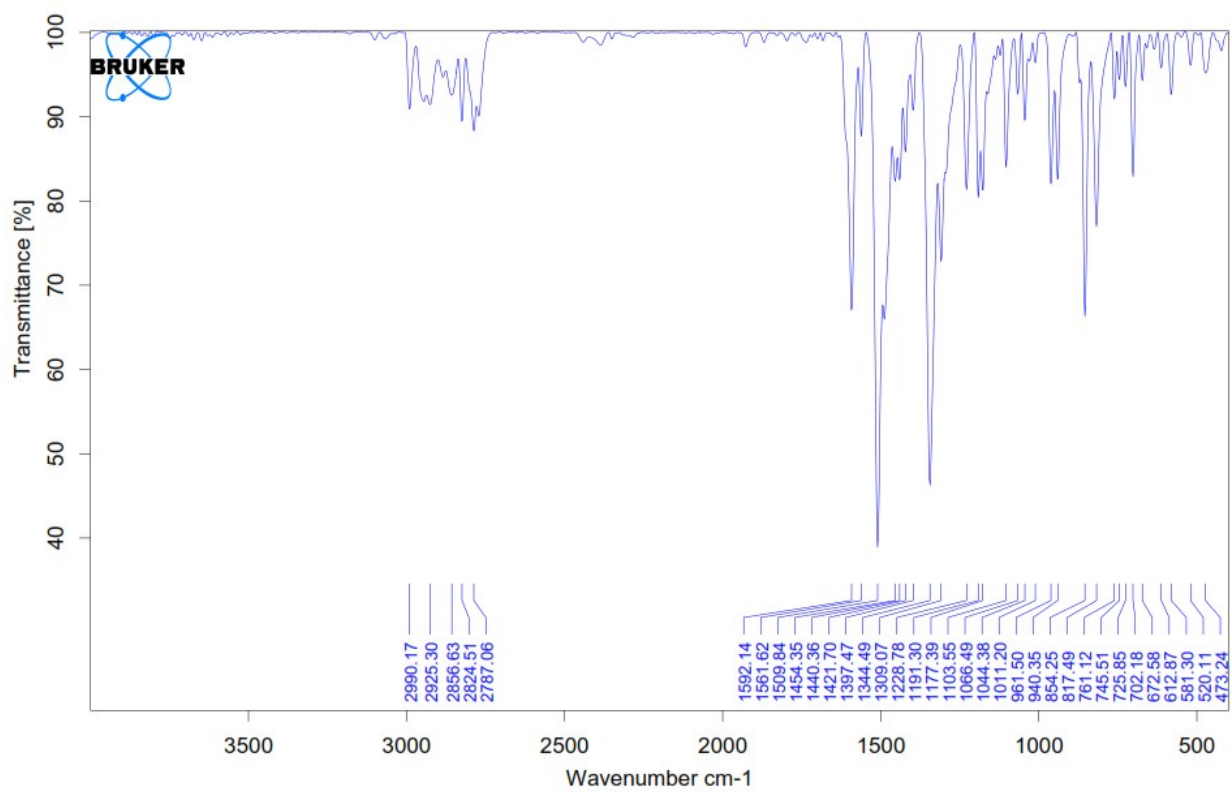


Figure 10S: FTIR of MNPD