

Supporting Information

From partial to complete neutralization of 2,5-dihydroxyterephthalic acid in the Li-Na system: crystal chemistry and electrochemical behavior of $\text{Na}_2\text{Li}_2\text{C}_8\text{H}_2\text{O}_6$ vs. Li

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Figure S1. Description of the expected dual redox activity in the $p\text{-DHT}^{4-}$ ligand.

Figure S2. Structure of **1** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers.

Figure S3. Structure of **2** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers.

Figure S4. Structure of **3** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers.

Figure S5. Structure of **4** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers.

Figure S6. Photographs of crystals (compounds **1**, **3** and **4**) obtained after a slow evaporation process at room temperature (24°C).

Figure S7. Comparison of FTIR spectra for $\text{Na}_2(\text{Li}_2)\text{-}p\text{-DHT}$ (up) and $\text{Li}_4\text{-}p\text{-DHT}$ (down), respectively.

Table S1. Range of selected bond lengths (Å) and bond angles (°) of compounds **1** – **4**.

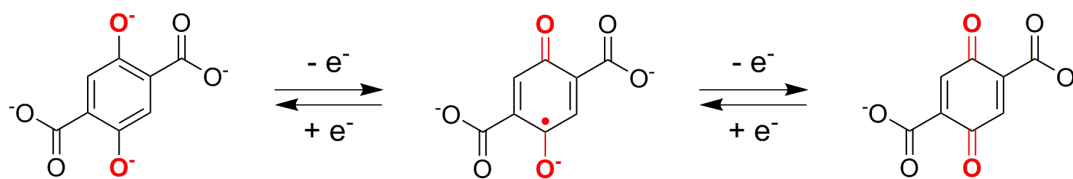
Table S2. Hydrogen-bond geometry (Å, °) for **1**.

Table S3. Hydrogen-bond geometry (Å, °) for **2**.

Table S4. Hydrogen-bond geometry (Å, °) for **3**.

Table S5. Hydrogen-bond geometry (Å, °) for **4**.

a. Diphenolate electrochemical activity in p -DHT $^{4-}$ (high formal redox potential):



b. Dicarboxylate electrochemical activity in p -DHT $^{4-}$ (low formal redox potential):

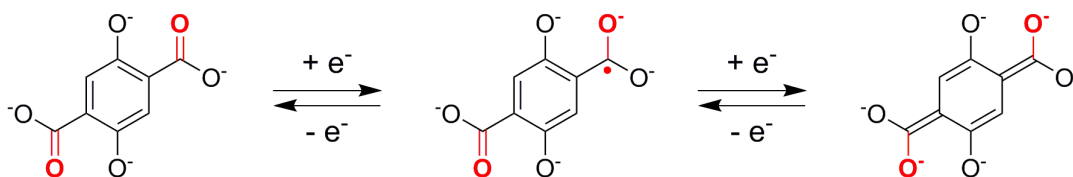


Figure S1. Description of the expected dual redox activity in the p -DHT $^{4-}$ ligand.

Table S1. Range of selected bond lengths (Å) and bond angles (°) of compounds **1** – **4**.

Compound		Distance (Å) and angle (°) ranges
1		
Li1 tetrahedra	Li1–O x4	1.989(4) – 2.003(4)
	O–Li1–O x6	102.23(16) – 113.61(17)
Li2 tetrahedra	Li2–O x4	1.881(3) – 1.965(4)
	O–Li2–O x6	103.48(16) – 113.67(17)
2		
Li1 tetrahedra	Li1–O x4	1.898(3) – 2.038(3)
	O–Li1–O x6	91.48(11) – 118.08(15)
Li2 tetrahedra	Li2–O x4	1.904(3) – 1.990(3)
	O–Li2–O x6	103.87(13) – 117.51(15)
3		
Na1A octahedra	Na1A–O x6	2.3519(19) – 2.5325(18)
	O–Na1A–O <i>cis</i> x12	75.51(7) – 112.26(7)
	O–Na1A–O <i>trans</i> x3	154.00(7) – 175.57(7)
Na2A octahedra	Na2A–O x6	2.3466(19) – 2.5121(18)
	O–Na2A–O <i>cis</i> x12	75.91(7) – 108.63(7)
	O–Na2A–O <i>trans</i> x3	153.07(7) – 177.38(7)
Na1B octahedra	Na1B–O x6	2.3272(19) – 2.4836(18)
	O–Na1B–O <i>cis</i> x12	75.85(7) – 106.40(7)
	O–Na1B–O <i>trans</i> x3	151.34(7) – 177.56(7)
Na2B octahedra	Na2B–O x6	2.3669(19) – 2.5647(18)
	O–Na2B–O <i>cis</i> x12	75.97(7) – 110.19(7)
	O–Na2B–O <i>trans</i> x3	156.14(7) – 173.83(7)
4		
Li1 tetrahedra	Li1–O x4	1.8995(18) – 1.989(2)
	O–Li1–O x6	95.48(8) – 118.75(10)
Na1 octahedra	Na1–O x6	2.3115(10) – 2.6521(10)
	O–Na1–O <i>cis</i> x12	74.67(3) – 114.05(3)
	O–Na1–O <i>trans</i> x3	154.24(3) – 167.14(4)

Table S2. Hydrogen-bond geometry (Å, °) for **1**.

$O-H\cdots O$	$O-H$	$H\cdots O$	$O\cdots O$	$O-H\cdots O$
$O3-H3\cdots O41$	0.91 (3)	1.70 (3)	2.5468 (19)	153 (3)
$O6-H6\cdots O12$	0.95 (3)	1.58 (3)	2.4853 (18)	157 (3)
$O1-H11\cdots O42^v$	0.87 (3)	1.93 (3)	2.797 (2)	172 (3)
$O1-H12\cdots O5^{iii}$	0.80 (3)	2.09 (3)	2.864 (2)	161 (3)
$O2-H21\cdots O4^{vi}$	0.87 (2)	1.98 (2)	2.829 (2)	165 (4)
$O2-H22\cdots O11^{vii}$	0.83 (4)	2.00 (4)	2.808 (2)	164 (3)
$O4-H41\cdots O2^i$	0.86 (2)	1.99 (2)	2.833 (2)	168 (5)
$O4-H42\cdots O3^{vi}$	0.85 (3)	1.99 (3)	2.836 (2)	172 (3)
$O5-H51\cdots O1^{viii}$	0.83 (3)	2.13 (3)	2.935 (2)	162 (2)
$O5-H52\cdots O12^v$	0.95 (3)	1.78 (4)	2.711 (2)	165 (3)

Symmetry codes: (i) $x, y, z-1$; (iii) $x, y, z+1$; (v) $-x+1, -y+1, -z+1$; (vi) $-x+2, -y+2, -z+1$; (vii) $-x+2, -y+1, -z+2$; (viii) $-x+1, -y+2, -z+1$.

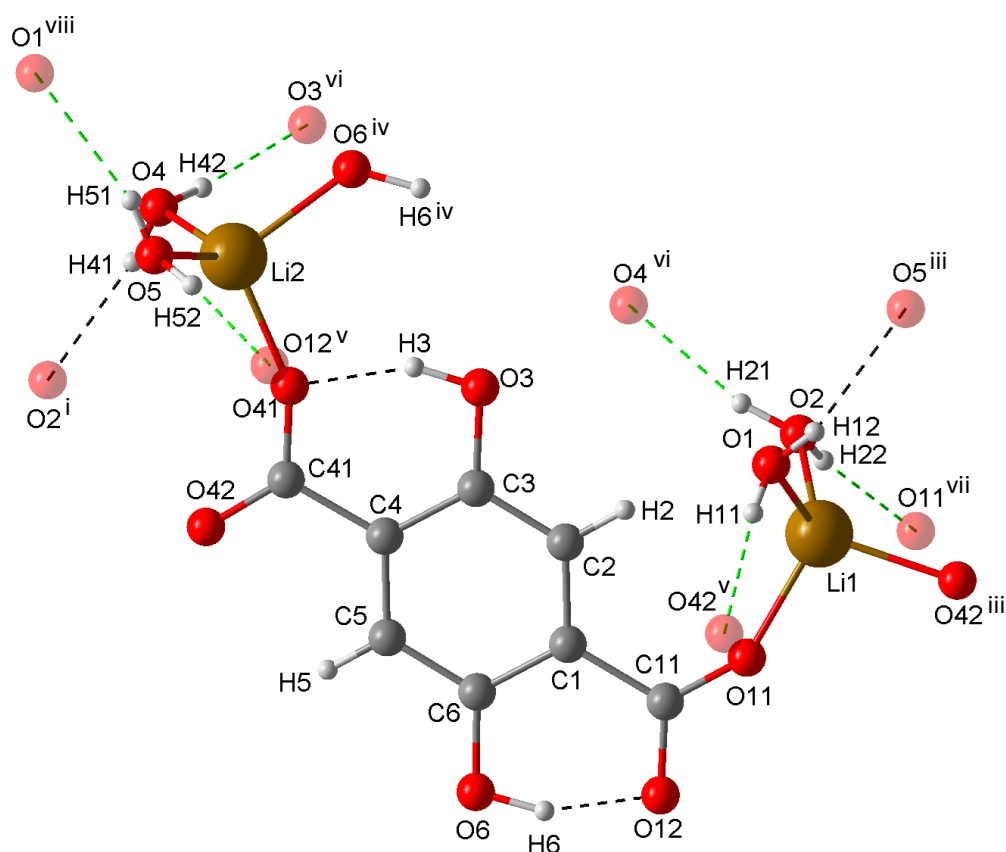


Figure S2. Structure of **1** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers. Symmetry codes: (i) $x, y, z-1$; (iii) $x, y, z+1$; (iv) $x, y+1, z$; (v) $-x+1, -y+1, -z+1$; (vi) $-x+2, -y+2, -z+1$; (vii) $-x+2, -y+1, -z+2$; (viii) $-x+1, -y+2, -z+1$.

Table S3. Hydrogen-bond geometry (Å, °) for **2**.

$O-H\cdots O$	$O-H$	$H\cdots O$	$O\cdots O$	$O-H\cdots O$
$O3-H3\cdots O41$	0.95 (3)	1.66 (3)	2.5466 (15)	155 (3)
$O6-H6\cdots O12$	0.88 (3)	1.77 (3)	2.5609 (16)	148 (3)
$O1-H11\cdots O9^{iii}$	0.79 (3)	2.10 (3)	2.870 (2)	165 (3)
$O1-H12\cdots O8^{iv}$	0.85 (3)	1.97 (3)	2.8031 (18)	168 (2)
$O7-H71\cdots O11^v$	0.85 (2)	2.02 (2)	2.8548 (17)	169 (2)
$O7-H72\cdots O42^{vi}$	0.85 (3)	1.94 (3)	2.7868 (17)	173 (3)
$O8-H81\cdots O12^{vii}$	0.80 (3)	2.01 (3)	2.7895 (17)	165 (3)
$O8-H82\cdots O3^{iv}$	0.86 (3)	1.89 (3)	2.7473 (18)	175 (2)
$O9-H91\cdots O6^{viii}$	0.85 (3)	2.00 (3)	2.8432 (18)	173 (3)
$O9-H92\cdots O7^{ix}$	0.80 (3)	2.08 (3)	2.8531 (17)	162 (2)

Symmetry codes: (iii) $x, y, z+1$; (iv) $-x+1, -y+2, -z+1$; (v) $x, y, z-1$; (vi) $-x, -y+1, -z$; (vii) $-x, -y+1, -z+1$; (viii) $x+1, y+1, z$; (ix) $-x+1, -y+2, -z$.

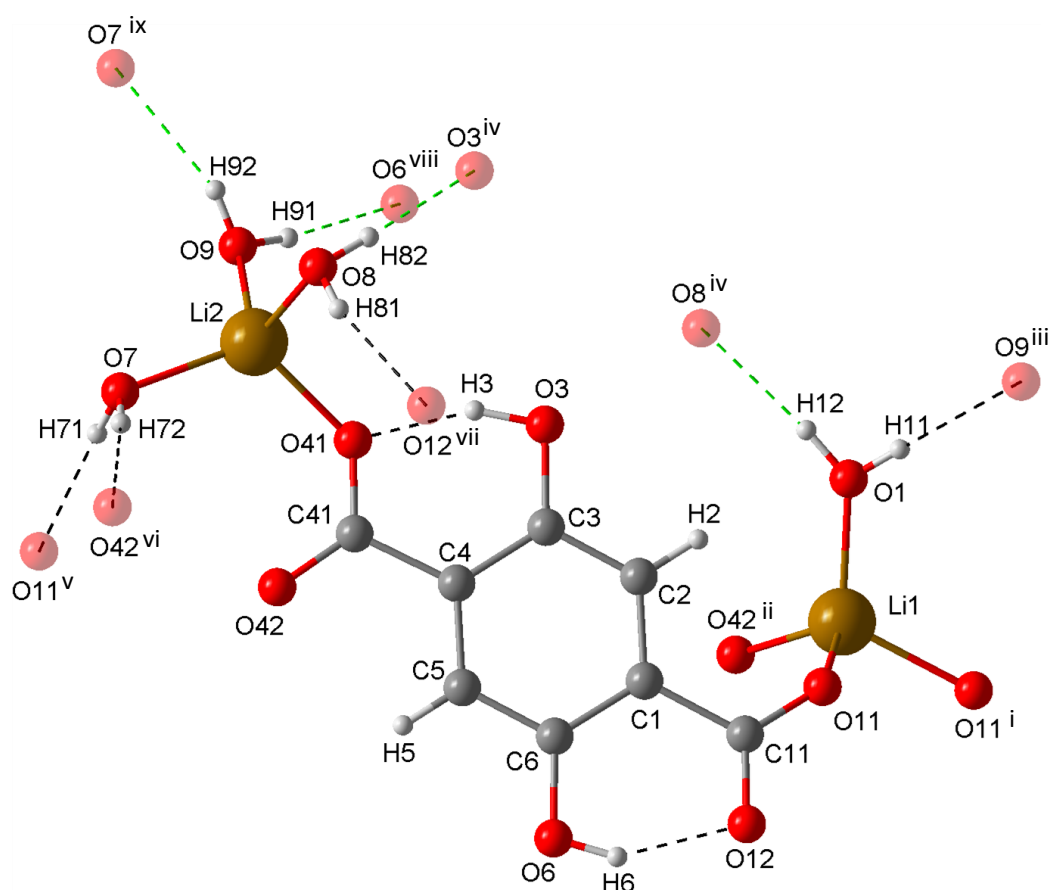


Figure S3. Structure of **2** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers. Symmetry codes: (i) $-x+1, -y+1, -z+2$; (ii) $-x+1, -y+1, -z+1$; (iii) $x, y, z+1$; (iv) $-x+1, -y+2, -z+1$; (v) $x, y, z-1$; (vi) $-x, -y+1, -z$; (vii) $-x, -y+1, -z+1$; (viii) $x+1, y+1, z$; (ix) $-x+1, -y+2, -z$.

Table S4. Hydrogen-bond geometry (Å, °) for **3**.

<i>O</i> — <i>H</i> ... <i>O</i>	<i>O</i> — <i>H</i>	<i>H</i> ... <i>O</i>	<i>O</i> ... <i>O</i>	<i>O</i> — <i>H</i> ... <i>O</i>
O2A—H2A...O11A	0.84 (4)	1.71 (3)	2.487 (2)	155 (3)
O5A—H5A...O42A	0.89 (4)	1.64 (4)	2.476 (2)	156 (3)
O1A—H11A...O12B ^{xi}	0.79 (3)	2.12 (3)	2.852 (2)	156 (3)
O3A—H31A...O12B ^{xi}	0.93 (4)	1.92 (4)	2.826 (3)	166 (3)
O3A—H32A...O4A ^{ix}	0.88 (4)	1.98 (4)	2.856 (3)	177 (4)
O4A—H41A...O4B ^{xii}	0.90 (4)	1.79 (4)	2.692 (2)	177 (4)
O4A—H42A...O41B ⁱⁱⁱ	0.87 (4)	1.93 (4)	2.722 (2)	151 (3)
O6A—H61A...O41B ^v	0.77 (3)	2.21 (3)	2.933 (2)	155 (3)
O6A—H62A...O41A	0.79 (3)	1.95 (4)	2.730 (2)	169 (3)
O2B—H2B...O11B	0.84 (3)	1.70 (3)	2.492 (2)	156 (3)
O5B—H5B...O42B	0.84 (4)	1.68 (3)	2.469 (2)	156 (3)
O1B—H11B...O12A ^{iv}	0.77 (3)	2.09 (3)	2.840 (2)	165 (3)
O3B—H31B...O3A ^{xiii}	0.93 (4)	1.76 (4)	2.687 (3)	174 (4)
O3B—H32B...O41A ^v	0.87 (4)	1.93 (4)	2.710 (2)	149 (3)
O4B—H41B...O3B ⁱ	0.91 (5)	1.95 (5)	2.852 (3)	174 (4)
O4B—H42B...O12A ^{iv}	0.88 (4)	1.98 (4)	2.841 (3)	166 (4)
O6B—H61B...O41A ^v	0.78 (3)	2.25 (3)	2.956 (2)	152 (3)
O6B—H62B...O12B ^{xiv}	0.83 (4)	1.94 (4)	2.756 (2)	168 (3)

Symmetry codes: (i) $x+1, y, z$; (iii) $-x+1, -y, -z+1$; (iv) $x, y-1, z$; (v) $-x, -y, -z+1$; (ix) $x-1, y, z$; (xi) $x, y-1, z-1$; (xii) $-x+1, -y-1, -z+1$; (xiii) $-x, -y-1, -z+1$; (xiv) $-x, -y+1, -z+2$.

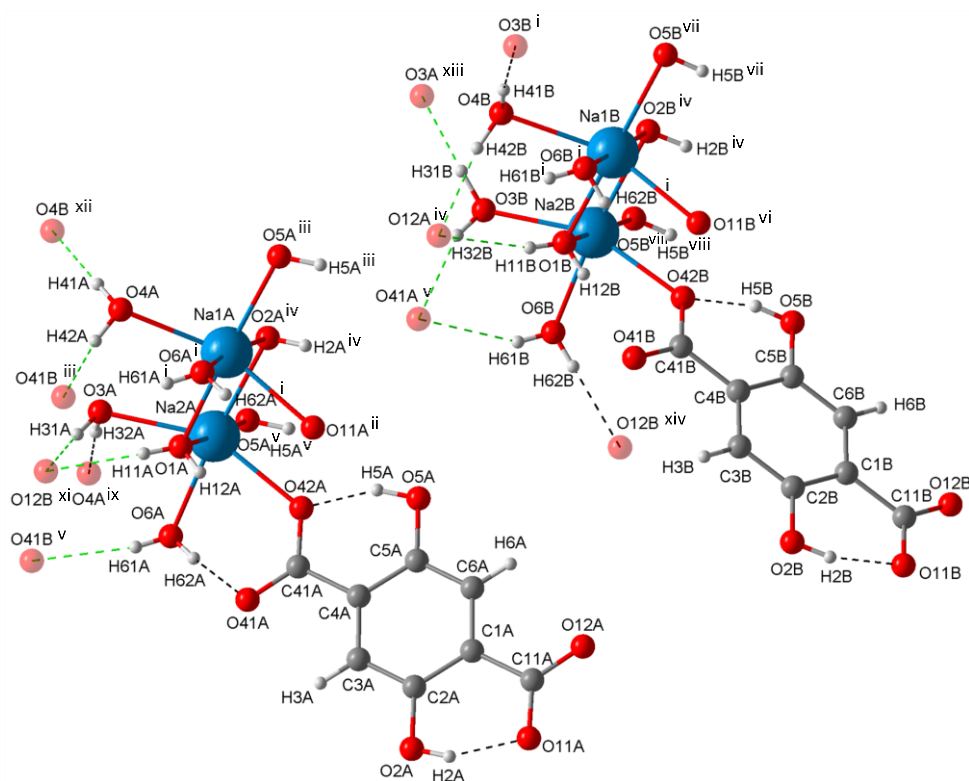


Figure S4. Structure of **3** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers. Symmetry codes: (i) $x+1, y, z$; (ii) $-x+1, -y+1, -z+1$; (iii) $-x+1, -y, -z+1$; (iv) $x, y-1, z$; (v) $-x, -y, -z+1$; (vi) $-x+1, -y+1, -z+2$; (vii) $-x+1, -y, -z+2$; (viii) $-x, -y, -z+2$; (ix) $x-1, y, z$; (xi) $x, y-1, z-1$; (xii) $-x+1, -y-1, -z+1$; (xiii) $-x, -y-1, -z+1$; (xiv) $-x, -y+1, -z+2$.

Table S5. Hydrogen-bond geometry (Å, °) for **4**.

$O-H\cdots O$	$O-H$	$H\cdots O$	$O\cdots O$	$O-H\cdots O$
$O1-H11\cdots O5^v$	0.811 (19)	2.068 (19)	2.8583 (13)	164.7 (17)
$O1-H12\cdots O3^{vi}$	0.869 (18)	1.855 (18)	2.7214 (11)	174.8 (17)
$O4-H41\cdots O3^{vii}$	0.84 (2)	2.03 (2)	2.8625 (13)	171.1 (19)
$O4-H42\cdots O2^{vii}$	0.82 (2)	1.96 (2)	2.7800 (13)	176.5 (18)
$O5-H51\cdots O3^{viii}$	0.84 (2)	1.86 (2)	2.7042 (12)	177.8 (19)
$O5-H52\cdots O11^{ix}$	0.843 (19)	2.11 (2)	2.9432 (13)	169.8 (17)

Symmetry codes: (v) $x, y, z+1$; (vi) $x-1, y, z+1$; (vii) $-x+2, -y+1, -z+1$; (viii) $-x+2, -y, -z$; (ix) $-x+1, -y, -z+1$.

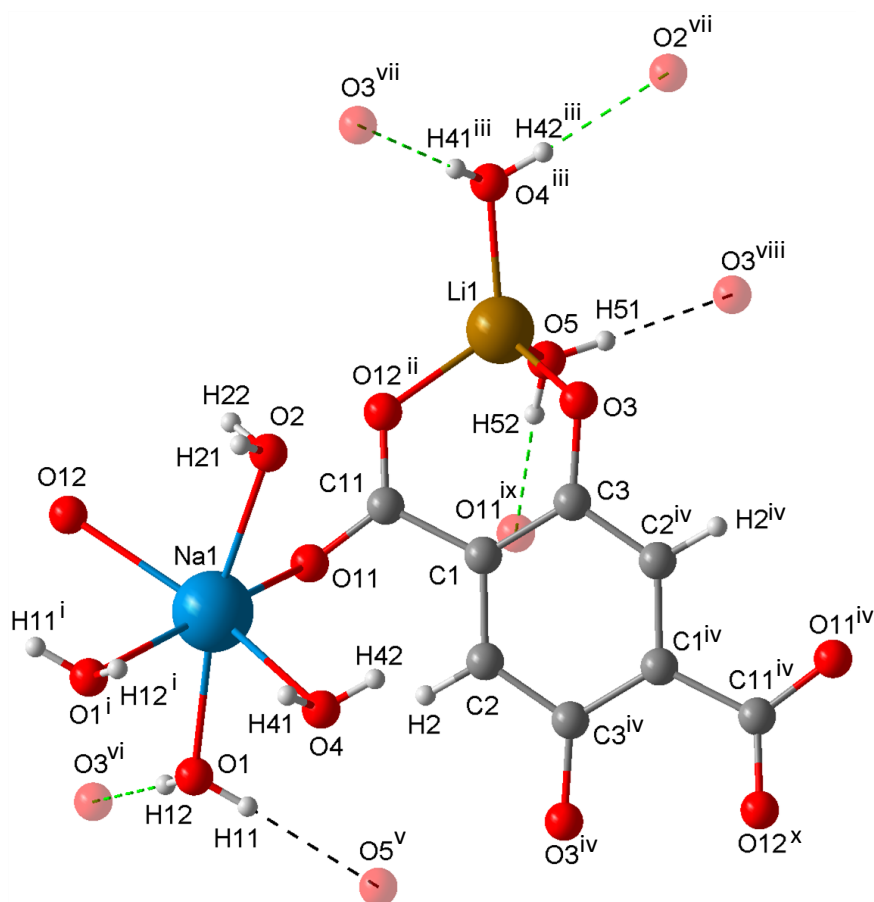


Figure S5. Structure of **4** with hydrogen bonds in dashed lines, in green, H-bonds between 2-D layers. Symmetry codes: (i) $-x+1, -y+1, -z+2$; (ii) $-x+1, -y+1, -z+1$; (iii) $x, y, z-1$; (iv) $-x+2, -y, -z+1$; (v) $x, y, z+1$; (vi) $x-1, y, z+1$; (vii) $-x+2, -y+1, -z+1$; (viii) $-x+2, -y, -z$; (ix) $-x+1, -y, -z+1$; (x) $x+1, y-1, z$.

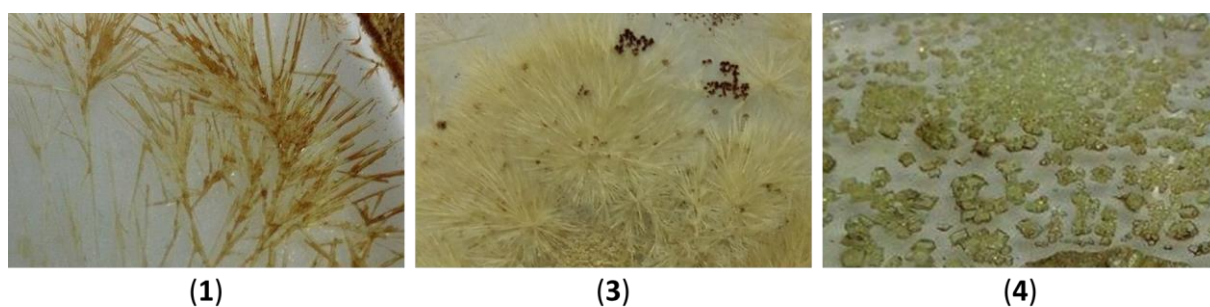


Figure S6. Photographs of crystals (compounds **1**, **3** and **4**) obtained after a slow evaporation process at room temperature (24°C).

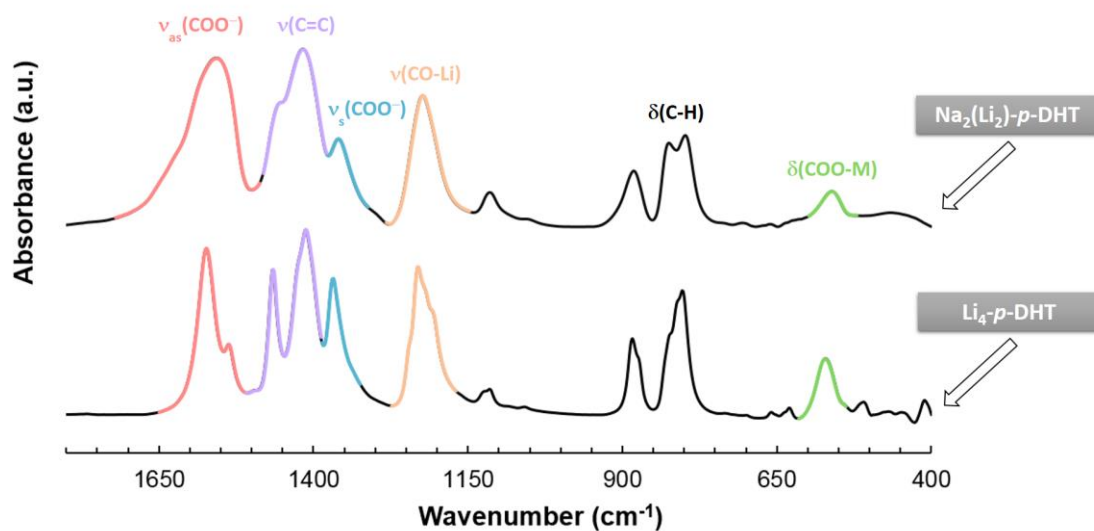


Figure S7. Comparison of FTIR spectra for $\text{Na}_2(\text{Li}_2)\text{-p-DHT}$ (up) and $\text{Li}_4\text{-p-DHT}$ (down), respectively.