

## Supporting Information

### Experimental

**Sample preparation.** Aqueous solution (30 mL) containing  $\text{Na}_3[\text{Co}(\text{CN})_6]$  (2.84 g, 10 mmol) and  $(\text{HIm})\text{Cl}$  (3.14 g, 30 mmol) was evaporated slowly at room temperature over two weeks, affording pale yellow crystals of **Na**. Block yellow crystal of **Rb** was collected by the similar method but using  $\text{Rb}_3[\text{Co}(\text{CN})_6]$  (4.71 g, 10 mmol) instead of  $\text{Na}_3[\text{Co}(\text{CN})_6]$ .

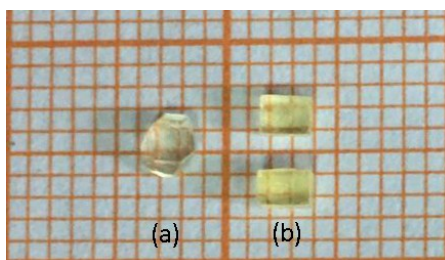
**Na:** yield 90% (based on Co). Elemental analysis calcd (%) for  $\text{C}_{12}\text{H}_{12}\text{CoN}_{10}\text{NaO}$  (394.24): C 35.65, H 5.48, N 34.64; found: C 35.30, H 5.79, N 34.84. FTIR for **Na** (KBr): 3160, 2135, 2116, 1590, 1427, 1198, 1050, 762, 622  $\text{cm}^{-1}$ .

**Rb:** yield 85% (based on Co). Elemental analysis calcd (%) for  $\text{C}_{12}\text{H}_{10}\text{CoN}_{10}\text{Rb}$  (438.70): C 32.12, H 4.49, N 32.21; found: C 32.09, H 4.98, N 32.62. FTIR for **Rb** (KBr): 3165, 2121, 1586, 1090, 1046, 751, 621  $\text{cm}^{-1}$ .

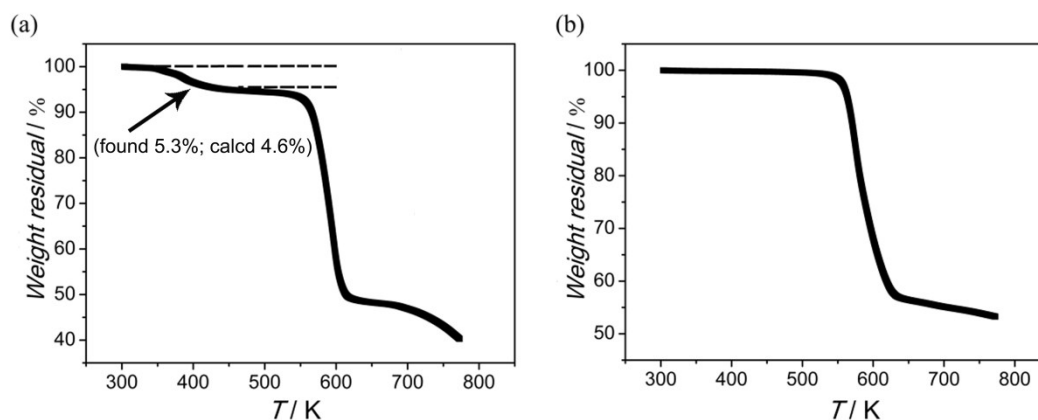
**Measurements.** Elemental analysis of C, H and N was conducted on a vario MICRO analyzer. Thermal gravimetric analysis (TGA) was performed on a TA Instruments Q500 in air flow. A Nicolet 5700 spectrometer was used to receive the infrared (IR) spectra. Powder X-ray diffraction (PXRD) was measured by a Rigaku SmartLab X-ray diffraction instrument to confirm the purity of the samples. Differential scanning calorimetry (DSC) measurements were carried out on a NETZSCH DSC 200F3 instrument with a cooling/heating rate of 5  $\text{K min}^{-1}$  in the temperature ranges of 103–293 K. Dielectric constant curves were collected by a Tonghui TH2828A impedance analyzer within an applied electric field of 1.0 V.

**Single-crystal X-ray crystallography.** A Rigaku Saturn 724<sup>+</sup> diffractometer using  $\text{Mo-K}\alpha$  ( $\lambda = 0.71073 \text{ \AA}$ ) radiation from a graphite monochromator equipped with gas spray cooler device was used to collect single-crystal data of compound **Na** and **Rb**. The bonds of C–H and N–H were placed in calculated positions, with C–H = 0.93  $\text{\AA}$  and  $U_{\text{iso}}(\text{H}) = 1.2 U_{\text{eq}}(\text{C})$  for imidazolium H atoms and N–H = 0.86  $\text{\AA}$  and  $U_{\text{iso}}(\text{H}) =$

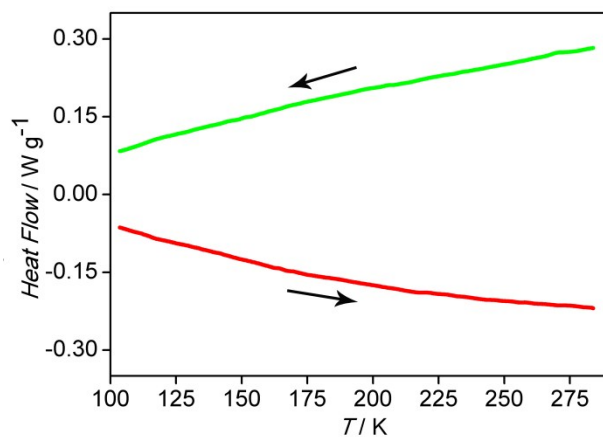
1.2  $U_{\text{eq}}(\text{N})$ . The water H atoms were located in difference Fourier syntheses and refined with an O–H distance restraint of 0.85 Å and with  $U_{\text{iso}}(\text{H}) = 1.5 U_{\text{eq}}(\text{O})$ . Crystallographic data and structural refinement details, selected bond lengths and bond angles, details of hydrogen bonding interactions are given in Table S1–2.



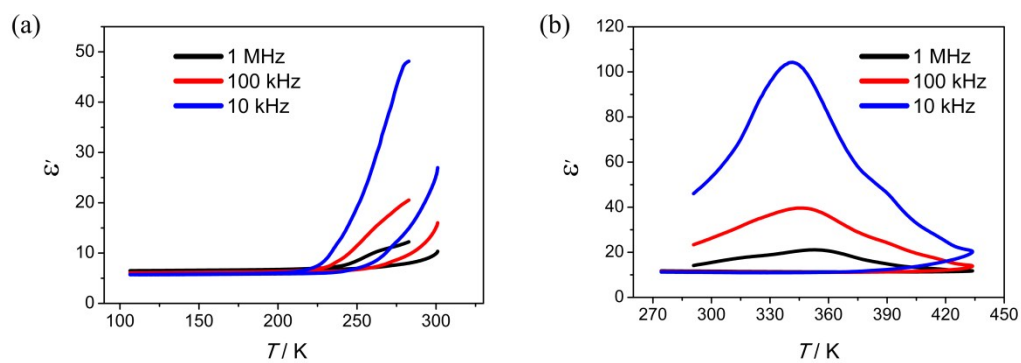
**Figure S1.** Crystal morphology of (a) **Na** and (b) **Rb**.



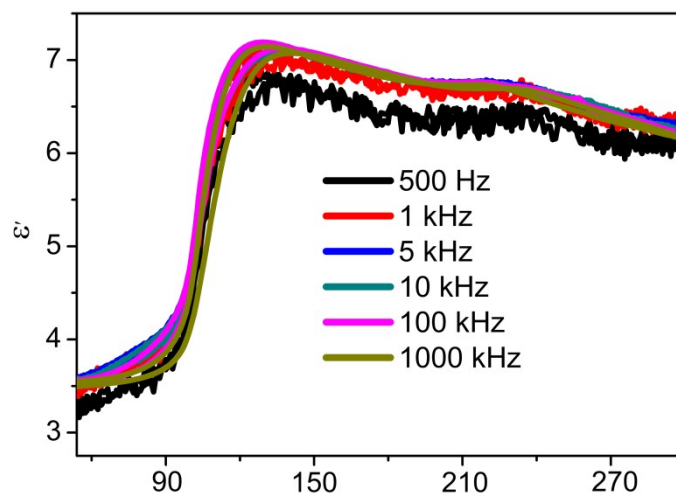
**Figure S2.** TGA curves of (a) **Na** and (b) **Rb**.



**Figure S3.** DSC curves of Na measured in the temperature range 100–290 K.



**Figure S4.** The  $\epsilon''$  vs  $T$  of Na on the power state in the different frequency and temperature ranges of 100–430 K.



**Figure S5.** The  $\epsilon'$  vs  $T$  of **Rb** on the power state in the different frequency and temperature ranges of 63–293 K.

**Table S1.** Bond lengths (Å) and bond angles (°) for **Na and Rb**.

<b>Na (293 K)</b>			
Na(1)–N(1)	2.4140(17)	N(1)#5–Na(1)–N(1)	134.94(9)
Na(1)–N(1)#5	2.4139(17)	N(1)–Na(1)–N(2)#2	102.25(5)
Na(1)–N(2)#2	2.477(2)	N(1)–Na(1)–O(1)	78.46(5)
Na(1)–N(5)#4	2.371(2)	N(2)#2–Na(1)–O(1)	65.08(7)
Na(1)–O(1)	2.955(2)	N(5)#4–Na(1)–N(1)	99.39(5)
		N(5)#4–Na(1)–N(2)#2	121.17(10)
		N(5)#4–Na(1)–O(1)	173.75(8)
Symmetry codes: #2 $-x, -y, -z$ ; #4 $-x+1/2, -y, z-1/2$ ; #5 $x, -y-1/2, z$ .			
<b>Rb (293 K)</b>			
Rb(1)–N(1)	2.994(4)	N(1)–Rb(1)–N(1)#6	83.00(12)
		N(1)–Rb(1)–N(1)#7	97.00(12)
		N(1)–Rb(1)–N(1)#9	180.0
Symmetry codes: #6 $x-y+2/3, x+1/3, -z+4/3$ ; #7 $x-y+2/3, x+1/3, -z+4/3$ ; #9 $x-y+2/3, x+1/3, -z+4/3$ .			
<b>Rb (193 K)</b>			
Rb(1)–N(1)	2.979(6)	N(1)–Rb(1)–N(1)#6	82.36(18)
		N(1)–Rb(1)–N(1)#8	97.64(18)
		N(1)–Rb(1)–N(1)#10	180.0
Symmetry codes: #6 $y-1/3, -x+y+1/3, -z+1/3$ ; #8 $-y+1, x-y+1, z$ ; #10 $-x+2/3, -y+4/3, -z+1/3$ .			
<b>Rb (93 K)</b>			
Rb(1)–N(1)	2.972(6)	N(1)–Rb(1)–N(1)#8	93.2(3)
Rb(1)–N(2)	2.968(5)	N(1)–Rb(1)–N(2)	82.65(18)
Rb(1)–N(3)	2.995(6)	N(1)–Rb(1)–N(2)#8	105.9(2)
		N(1)–Rb(1)–N(3)	95.55(17)
		N(1)–Rb(1)–N(3)#8	169.5(2)
		N(2)–Rb(1)–N(2)#8	167.7(3)
		N(2)–Rb(1)–N(3)	81.1(2)
		N(2)–Rb(1)–N(3)#8	89.28(18)
		N(3)–Rb(1)–N(3)#8	76.4(3)
Symmetry codes: #8 $-x+1, y, -z+1/2$ .			

**Table S2.** Hydrogen bonds (Å, °) for **Na** and **Rb**.

	H···A	D···A	D–H···A
D–H···A			
<b>Na</b> (293 K)			
N7–H7A···N3#3	2.18	2.955(2)	149.0
N8–H8A···O1#4	2.08	2.872(2)	149.9
O1–H1WA···N6#6	1.91(2)	2.765(3)	170(2)
Symmetry codes: #3 $-x+1/2, -y, z+1/2$ ; #4 $-x+1/2, -y, z-1/2$ ; #6 $-x, -y, -z+1$ .			
<b>Rb</b> (93 K)			
N(4)–H(4B)···N(1)#4	2.19	3.015(10)	161
N(5)–H(5B)···N(3)#9	2.60	3.248(11)	133
Symmetry codes: #4 $-x+1/2, -y+1/2, -z$ ; #9 $-x+1/2, y-1/2, -z+1/2$ .			