

## **Electronic Supplementary Information (ESI)**

### **In situ generated electron-deficient metallic copper as the catalytically active site for enhanced hydrogen production from alkaline formaldehyde solution**

**Shipan Liang,<sup>a</sup> Shuang Chen,<sup>a</sup> Ziwei Guo,<sup>a</sup> Zhuohuang Lan,<sup>a</sup>  
Hisayoshi Kobayashi,<sup>b</sup> Xiaoqing Yan,<sup>c</sup> Renhong Li,<sup>\*a</sup>**

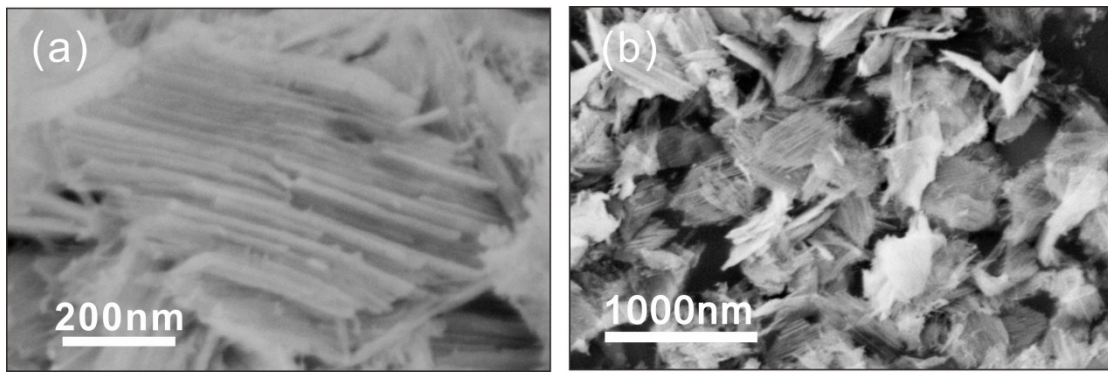
<sup>a</sup> Department of Materials Engineering, College of Material and Textiles, Zhejiang Sci-Tech University, Hangzhou 310018, China

<sup>b</sup> Emeritus Professor of Department of Chemistry and Materials Technology, Kyoto Institute of Technology, Matsugasaki, Sakyo-ku, Kyoto 606-8585, Japan

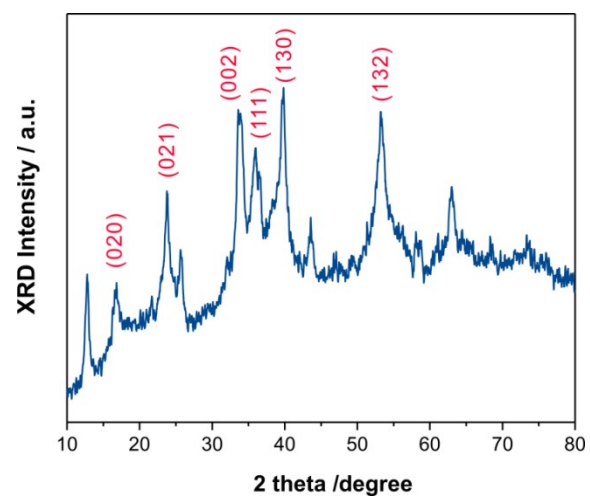
<sup>c</sup> Department of Chemistry, College of Science, Zhejiang Sci-Tech University, Hangzhou 310018, China

#### **Corresponding Author**

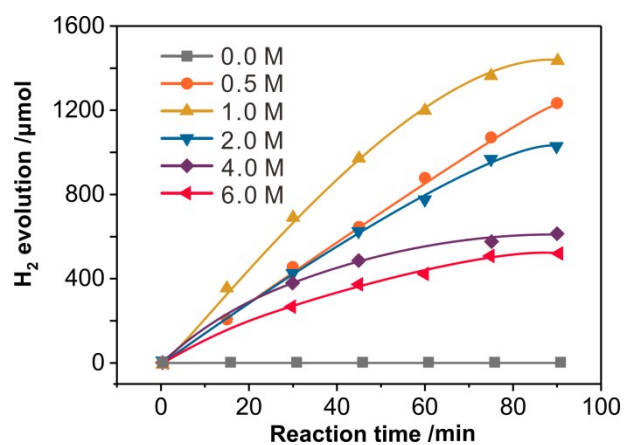
\*E-mail: [lirenhong@zstu.edu.cn](mailto:lirenhong@zstu.edu.cn)



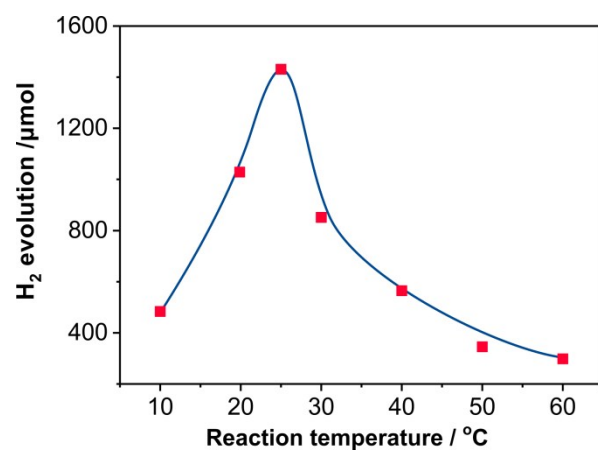
**Fig. S1** SEM images of  $\text{Cu}(\text{OH})_2$  NRs precursor.



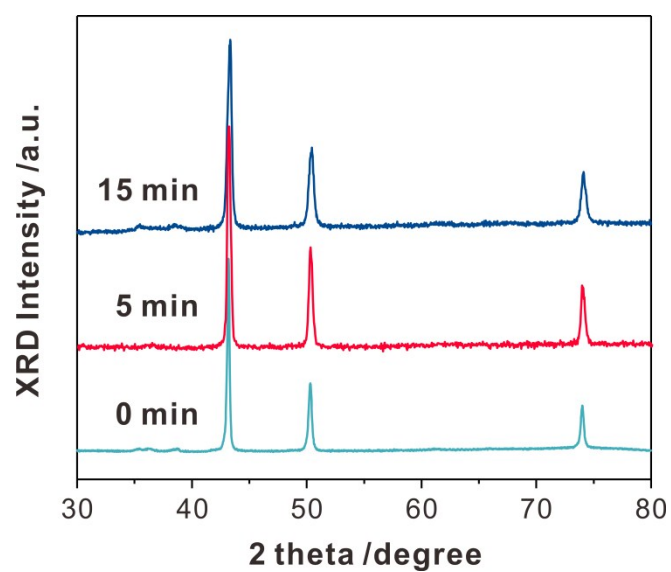
**Fig. S2** XRD pattern of  $\text{Cu}(\text{OH})_2$  NRs precursor.



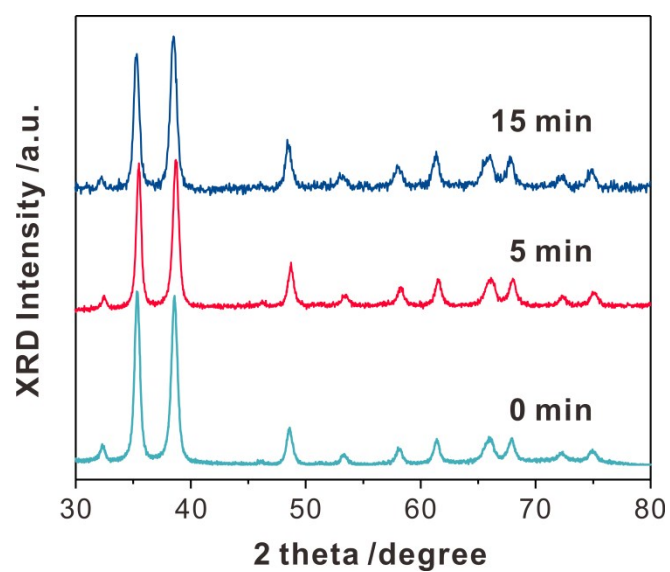
**Fig. S3** The effect of formaldehyde concentration on the rate of H<sub>2</sub> evolution over CuO NPs.



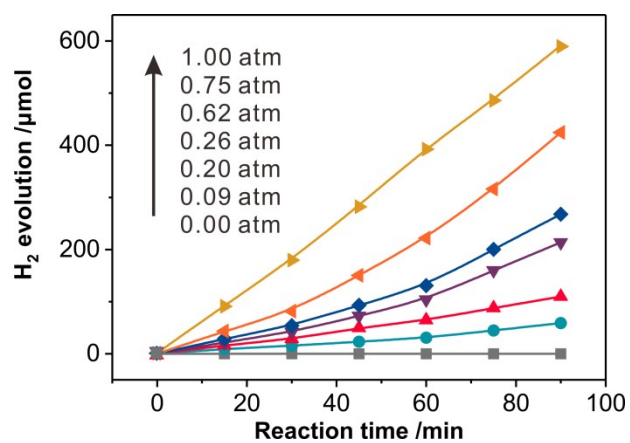
**Fig. S4** The effect of reaction temperature on the rate of H<sub>2</sub> evolution over CuO NPs.



**Fig. S5** *In-situ* XRD analysis of Cu NPs measured in 1M Na<sub>2</sub>CO<sub>3</sub> solution before, after 5 min and 15 min of hydrogen production reaction.

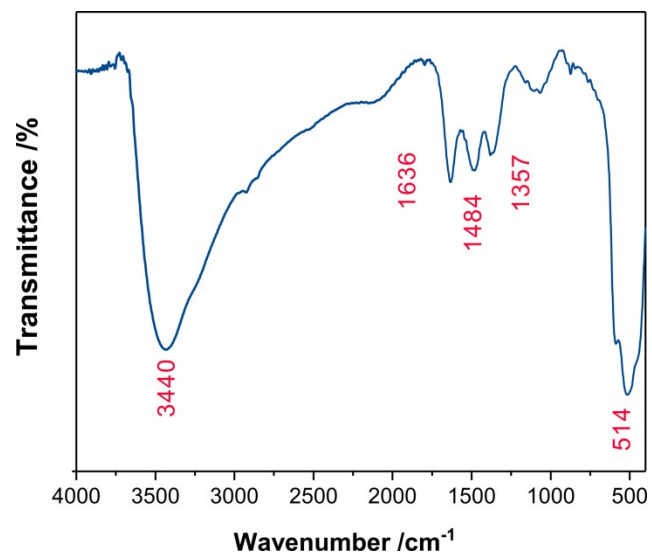


**Fig. S6** *In-situ* XRD analysis of CuO NPs measured in 1M Na<sub>2</sub>CO<sub>3</sub> solution before, after 5 min and 15 min of hydrogen production reaction.

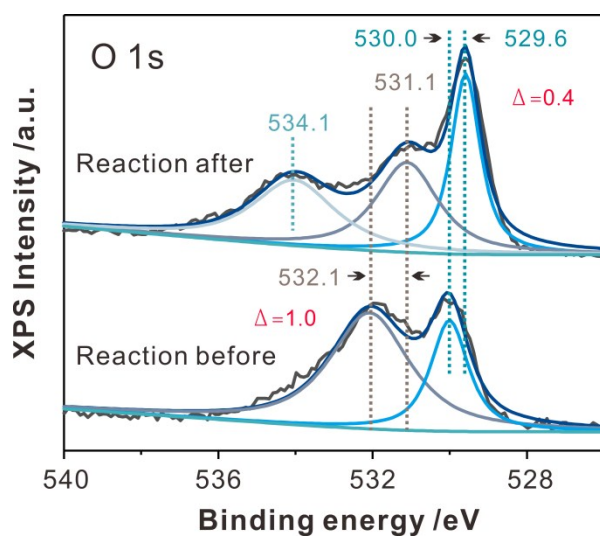


**Fig. S7** The effect of oxygen partial pressure on the rate of H<sub>2</sub> evolution over Cu NPs.

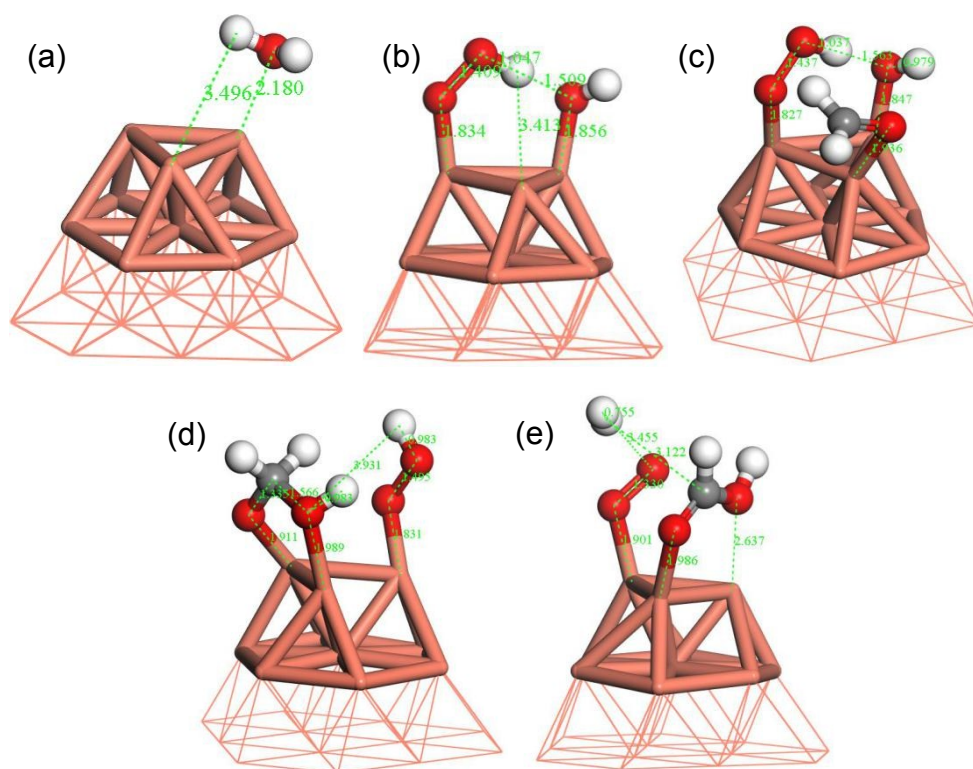




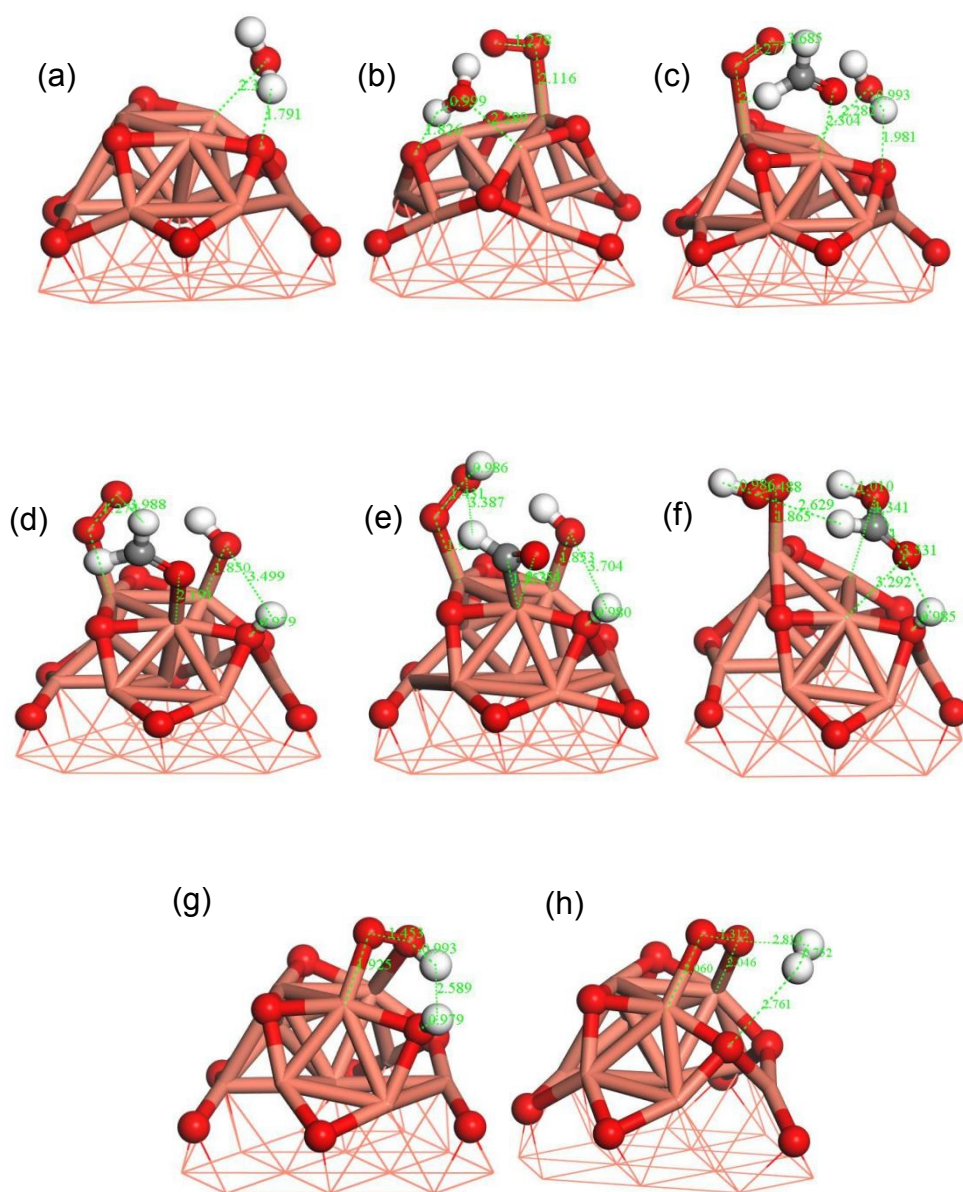
**Fig. S8** FT-IR spectra of the reduced Cu after 15 min of hydrogen evolution reaction.



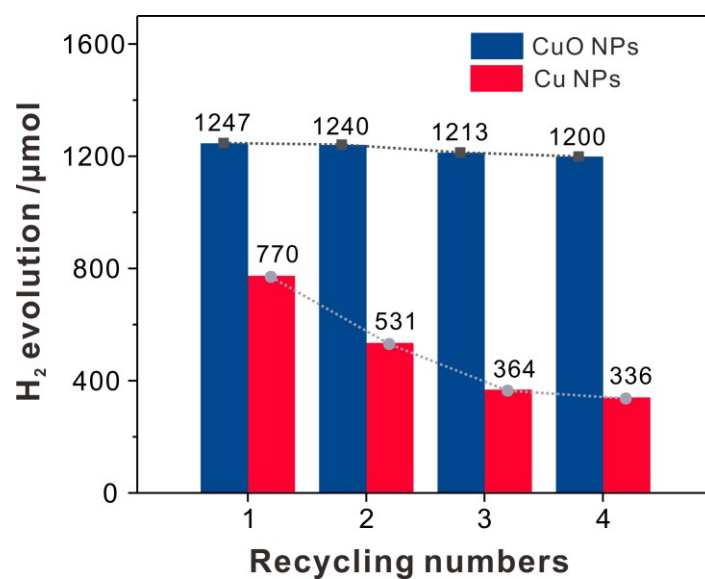
**Fig. S9** O 1s XPS spectra of the reduced Cu before and after reaction.



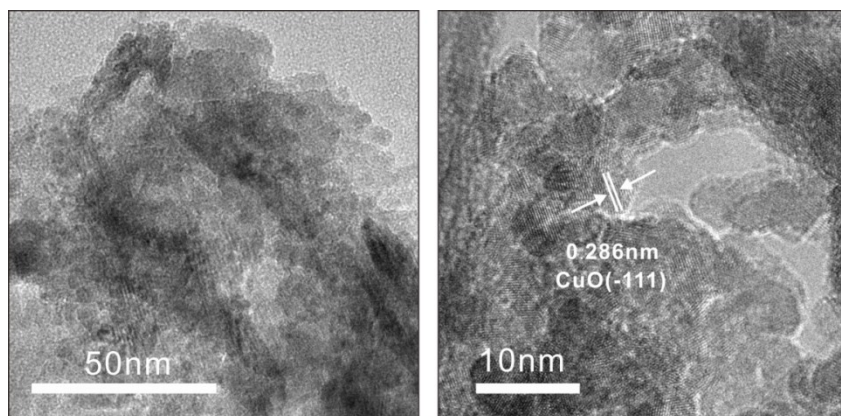
**Fig. S10** (a) Optimized structures for  $\text{Cu}_{22}\text{-H}_2\text{O}$  (LM1). (b) Optimized structures for  $\text{O}_2$  and  $\text{H}_2\text{O}$  co-adsorption on  $\text{Cu}_{22}$  cluster (LM2). (c) Optimized structures for  $\text{O}_2$ ,  $\text{H}_2\text{CO}$ , and  $\text{H}_2\text{O}$  co-adsorption on  $\text{Cu}_{22}$  cluster (LM3). (d) Optimized structures for  $\text{OOH-H}_2\text{C(O)OH}$  (LM4). (e) Optimized structures for  $\text{O}_2\text{-H}_2\text{-HCOOH}$  (LM5).



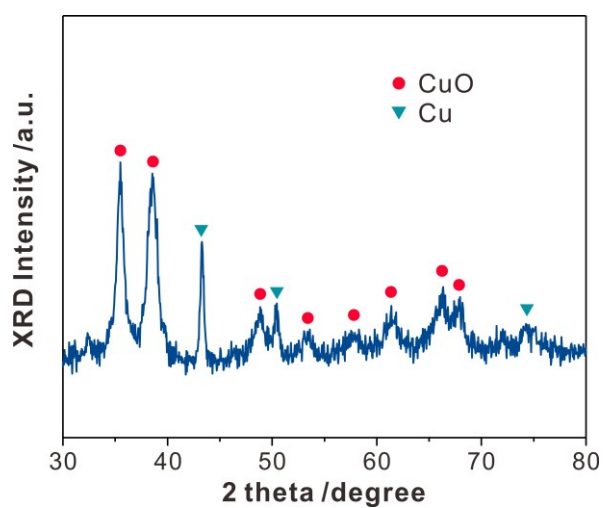
**Fig. S11** (a) Optimized structures for  $\text{Cu}_{22}\text{O}_9\text{-H}_2\text{O}$  (LM1). (b) Optimized structures for co-adsorbed states of  $\text{O}_2+\text{H}_2\text{O}$  (LM2). (c) Optimized structures for co-adsorbed states of  $\text{O}_2+\text{H}_2\text{O}+\text{H}_2\text{CO}$  (LM3). (d) Optimized product structure for  $\text{H}_2\text{O}$  dissociation. (LM4). (e) Optimized structures for  $\text{OOH-HCO}$ . (f) Optimized product structure for  $\text{HCOOH}$  formation. (g) Optimized structures for  $\text{OOH+H}$ . (h) Optimized structures for  $\text{O}_2+\text{H}_2$ .



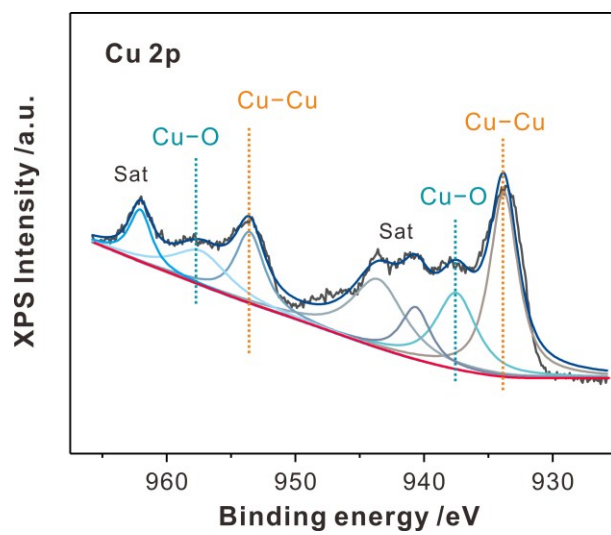
**Fig. S12** The cycle performance of the CuO and Cu NPs catalyzes the hydrogen production of alkaline formaldehyde solution.



**Fig. S13** TEM images of CuO catalyst after reaction.



**Fig. S14** XRD pattern of CuO catalyst after reaction.



**Fig. S15** Cu 2p XPS spectrum of CuO catalyst after reaction.