Supplementary information

## Promotion effect of Mg on post-synthesized Sn-Beta zeolite for the conversion of

## glucose to methyl lactate

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Scheme S1 Conversion of glucose to alkyl lactate in alcohol.





Fig. S1 UV-vis DR spectra of SnO<sub>2</sub>, MgO and deAl-Beta.









**Fig. S3** FT-IR spectra of CHCl<sub>3</sub> adsorbed on (a) deAl-Beta, (b) Mg-Beta, (c) Sn-Beta, (d) 0.25Mg-Sn-Beta, (e) 1Mg-Sn-Beta and (f) 4Mg-Sn-Beta at room temperature with different desorption time.





**Fig. S4** Kinetic analysis of MLA formation from fructose in methanol in the presence of 1Mg-Sn-Beta.





**Fig. S5** Kinetic analysis of MLA formation from fructose in methanol in the presence of Sn-Beta.



Fig. S6 Arrhenius plots of the retro-aldol of the fructose over 1Mg-Sn-Beta and Sn-Beta.



Fig. S7 Effect of reaction temperature on the formation of other C<sub>3</sub> products from glucose.



Fig. S8 The recyclability of 1Mg-Sn-Beta catalyst with the initial activity for MLA formation. Reaction conditions: glucose (0.37 g), 1Mg-Sn-Beta (0.20 g), methanol (15 mL),  $N_2$  (0.4 MPa), 120 °C, 0.5 h.



Fig. S9 XRD patterns of fresh and reused 1Mg-Sn-Beta.

Catalyst	S <sub>BET</sub>	S <sub>ext</sub> <sup>a</sup>	V <sub>tot</sub>	V <sub>meso</sub> <sup>a</sup>	V <sub>micro</sub>	Sn	Mg	Relative
	(m <sup>2</sup>	(m <sup>2</sup>	(mL	(mL	(mL	content	content	crystallinity
	g <sup>-1</sup> )	g <sup>-1</sup> )	g <sup>-1</sup> )	$g^{-1})^a$	g <sup>-1</sup> )	(wt%) <sup>b</sup>	(wt%) <sup>b</sup>	(%)
Fresh	515	164	0.54	0.36	0.18	1.90	0.392	100
Reused	453	157	0.49	0.34	0.15	1.85	0.384	97

Table S1 Results of N2 physisorption of fresh and reused 1Mg-Sn-Beta

<sup>a</sup>External surface area = BET surface area - micropore surface area; mesopore volume = total pore volume - micropore volume, where the micropore surface area and volume were determined by the t-plot method at a relative pressure of 0.05-0.70. <sup>b</sup>Determined by ICP.

## Reference