Supporting Information

Towards a Practical Perfluoroalkylation of (Hetero)Arenes using Cobalt Nanocatalysts and Perfluoroalkyl Bromides

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1. General comments

All manipulations involving air- and moisture-sensitive organometallic compounds were carried out using standard Schlenk techniques under an argon atmosphere. All chemicals were purchased from *Aldrich*, *TCI*, *Alfa*, or *ABCR* and used without further purification. Acetone was dried over molsieves 4Å and was freezed out to remove oxygen before using. All perfluoroalkylation reactions were set up under argon in 10 mL thick-walled screw cap vial or in a 50 mL autoclave (PARR Instrument Company). Multiplets of NMR were assigned as s (singlet), d (doublet), t (triplet), q (quartet), dd (doublet of doublet), dt (doublet of triplet), td (triplet of doublet), br.s (broad singlet), and m (multiplet).

2. Analytical methods

¹H NMR spectra were recorded using Bruker AV300 (300 MHZ for 1H), Bruker AV400 (400 MHz for 1H) and Fourier 300 (300 MHZ for ¹H) spectrometers. ¹⁹F NMR spectra were obtained at 282 MHz. Conversions and yields of perfluoroalkylation reactions were determined by 19 F NMR using either 1,4-difluorobenzene or 1,2-difluorobenzene as internal standard. For quantitative ¹⁹F NMR a puls delay of 30s are required. NMR chemical shifts are reported in parts per million (ppm) downfield from TMS and were referenced to the residual proton resonance and the natural abundance. All measurements were carried out at room temperature unless otherwise stated.

High resolution mass spectra were recorded on a MAT 95XP ThermoFisher Mass Spectrometer using Electrospray Ionization mode.

XRD pattern of the materials were recorded on a PANalytical X'Pert Pro diffractometer in reflection mode with Cu K α radiation (λ =1.5406 Å) and a silicon strip detector (X'Celerator).

The XPS measurements were performed on an ESCALAB 220iXL (ThermoFischer Scientific) with monochromated Al K α radiation (E = 1486.6 eV). Samples are prepared on a stainless-steel holder with conductive double-sided adhesive carbon tape. The electron binding energies were obtained with charge compensation using a flood electron source and referenced to the C1s peak assuming sp2 carbon as main component at 284.0 eV. For quantitative analysis the peaks were deconvoluted with Gaussian-Lorentzian curves using the software Unifit 2020, the peak areas were divided by the transmission function of the spectrometer and the element specific sensitivity factor of Scofield.¹

Scanning transmission electron microscopy (STEM) measurements were performed at 80kV with a probe aberrationcorrected JEM-ARM200F (microscope: JEOL, Japan; corrector: CEOS, Germany) using annular bright field (ABF) and high angle annular dark field (HAADF) detectors. The microscope is equipped with a JED-2300 (JEOL) energy-dispersive x-rayspectrometer (EDXS) and an Enfinium ER electron energy loss spectrometer (EELS, Gatan, USA) for chemical analysis. The sample was deposed on a holey carbon supported Cu-grid (mesh 300) without any pretreatment and transferred to the microscope.

3. General procedures

3.1 Catalyst preparation

In a representative procedure for the preparation of 3% Co@N/C-800 (Cat2), Co(OAc)₂·4H₂O (0.5 mmol, 124.5 mg) and 1,10-phenanthroline (1.0 mmol, 180.2 mg) were solved in 50 mL ethanol for 30 minutes at room temperature. Next, carbon powder (689.7 mg, VULCAN® XC72R, Cabot Corporation Prod. Code XVC72R; CAS No. 1333-86-4) was added and the whole reaction mixture was refluxed for 4 hours. After cooling the reaction mixture to room temperature the ethanol was

removed in high vacuum and the remaining solid was dried at 60 °C for 12 hours. The catalyst was grinded to a fine powder and was transferred into a ceramic crucible and placed in an evacuated (5 mbar) and argon flushed oven. The oven was heated to 800 °C (25 °C per minute), and held at 800 °C for 2 hours under argon atmosphere. Next, the oven was switched off and cooled to room temperature. Argon was constantly passed through the oven during the whole process. Elemental analysis of **Cat2** (wt%): C = 92.28, H = 0.20, N = 2.70, Co = 3.05, O = 1.32

3.2 Synthesis protocol for the cobalt catalyzed perfluoroalkylation of (hetero)arenes

In a representative procedure, to a 10 mL thick-walled screw cap vial with a magnetically stirring bar Cs_2CO_3 (1 mmol, 325.8 mg) and cobalt catalayst **Cat2** (10 mol% [Co], 200.0 mg) was added. The Vial was transferred into a schlenk chamber and evacuated and refilled with argon 3 times. Under argon atmosphere (hetero)arene substrate (10 mmol) was added. Additionally perfluorohexyl bromide (1 mmol, 213 µL) was introduced and the screw cap vial was closed. The closed vial was transferred into an aluminum block and heated while stirring for 24 h at 130 °C. Thereafter, the reaction mixture was cooled down with ice and diluted with 5ml diethylether. The suspension was centrifuged for 15min at 30000 RPM. The solution was decanted and the solid was washed 3 further times with diethylether. The etheric solutions were combined, an aliquot for the quantitative ¹⁹F NMR and GCMS was taken and the solvent was evaporated under mild vacuum. The residue was purified by combi flash chromatography with pentane/ethyl acetate. For determining the yield 0.2 mmol 1,4-difluororbenzene was added as the standard and directly a ¹⁹F NMR in CDCl₃ was recorded.

3.3 Synthesis protocol for the cobalt catalyzed trifluoromethylation of (hetero)arenes

In a 50 mL autoclave, **Cat2** (10 mol% [Co], 400 mg) and Cs_2CO_3 (1 mmol, 651.6 mg) were added. Next, 2 mmol (hetero)arene and 4 mL acetone were added. The autoclave was pressurized with the full pressure (ca. 12 bar) of the CF₃Br bottle. Then the autoclave was heated to 130 °C with an aluminum block for 48 h.

3.4 Synthesis protocol for the trifluoromethylation of acetophenone and recycling of the catalyst

In a 100 mL round Schlenk, 20 mL acetone and 12 mmol acetophenone were freezed to remove oxygen. Next, in a 100 mL autoclave with mechanical stirrer from Parr Instrument **Cat2** (10 mol% [Co], 2.04 g) and Cs₂CO₃ (1.5 equiv, 18 mmol, 5.9 g) were added and under argon stream the acetone solution was transferred to the autoclave via a syringe. The autoclave is closed and spilled 5 min with argon and weighted. The CF₃Br bottle was connected to the autoclave and the full pressure was given for 3 minutes. After weighting again ca. 26 g (174 mmol) of CF₃Br was in the autoclave. The reaction was carried out at 130 °C for 48 h. After cooling to room temperature the black reaction solution was quantitatively filled in a 250 mL flask and the liquid phase was evaporated and condensed in another 250 mL flask. The black residue was suspended with water, filtrated and washed with water and CH₂Cl₂. After drying the catalyst can be reused up to 6 times. Next, the condensed colorless liquid is reduced on the rotavapor and was filled into a 20 mL volumetric flask for quantitative measurement. 0.4 ml was taken in a NMR tube, 20 μ L 1,2-difluorobenzol was used as an internal standard and the tube was filled to 0.7 mL with CDCl₃. The quantitative ¹⁹F NMR (Figure S1) was carried out with a puls delay of 30 s. The conversion was determined with GC by taking 1 mL of the stock solution and mixing it with 70.7 mg hexadecane.



Figure S1. The quantitative ¹⁹F NMR for determining the yield of trifluoromethylation of acetophenone.

3.5 An Example for the Evaluation of a typical ¹⁹F NMR

The ¹⁹F NMR spectrum gives full information with regard to yield, conversion, loss of fluorine and reduced sideproduct (R_FH). (Figure S2) Since the chemical shift of the -CF₃ group is identical for all products and educt the total amount of fluorine can be determined. Interestingly after each reaction some amount of the fluorine is missing. We envisioned that a part of the fluorine chain was decomposed into gaseous fragments, which are released during the opening of the reaction vessel. As an example of the calculations, cording to the 19F NMR, we calculated 43% yield, 88% conversion, 23% loss of fluorine and 11% reduced sideproduct for the perfluorohexylation of 4-picoline using **Cat1** as the catalyst.



Figure S2. An Example for the Evaluation of a typical ¹⁹F NMR.

4. Catalyst characterization

4.1 Powder X-ray Diffraction



Figure S3. Detailed representation of the powder diffraction pattern of the catalyst after the 6th recycling step and reference peak positions of metallic Co (pdf 00-015-0806, ICDD, 2016). The diffraction peak at 44.22° 2theta can be indexed as (111) Brag reflex of the cubic modification of metallic cobalt.

4.2 X-ray photoelectron spectroscopy



Figure S4. XPS survey spectra of the fresh and recycled catalyst after the 6th recycling of trifluoromethylation (TFM) and perfluoroalkylation (PFA). Whereas the fresh catalyst only shows signals for carbon, nitrogen, oxygen and cobalt, fluorine and cesium are clearly visible for both recycled catalysts.



Figure S5. XP spectra of the fresh and recycled catalyst after the 6th cycle of trifluoromethylation (TFM) and perfluoroalkylation (PFA). A) The C 1s spectra of the fresh catalyst shows mainly C-C and C-C/C-H bonds at 248.0 eV and 248.8 eV whereas the spectra of the recycled materials clearly depict the formation of C-O/C-OH functionalities between 286 eV and 289 eV. A broad feature at about 292 eV binding energy indicates the formation of CF₂ at the surface. B) The N 1s region can be deconvoluted with two signals at 398.6 eV and 400.6 eV which are characteristic for pyridinic and pyrrolic nitrogen.¹ D) The F 1s spectra of the recycled catalysts show one main component at 688 eV characteristic for organic fluorine groups. In the fresh catalyst no F can be found. D) Co 2p region of the fresh and recycled catalyst. Only in the fresh catalyst and after trifluoromethylation cobalt can be found with a binding energy of about 780 eV characteristic for oxidic cobalt (probably Co3O4 as main component when considering the satellite structures).²

4.3. Additional STEM data





Figure S6. Fresh catalyst: A) ADF image with regions marked from which spectra were extracted from the spectrum imaging dataset. B) EELS data shows the similarity of the C and N content in all three areas whereas in C) the EDXS data shows the different composition of the particles in area 1 (metal Co) and 2 (Co oxide). Also, in area 3 there is Co present although not in the form of a particle but probably in a finely distributed manner.



Figure S7. Fresh catalyst: A) HAADF-STEM image showing the general morphology with Co particles lighting up as bright particles compared to support. B) ABF-STEM image of a Co particle surrounded by graphitic carbon. C) HAADF-STEM images with high magnification showing dots on the surface of carbon support particle (bright onion like structure) possibly indicating a distributed Co phase.



800

Figure S8. Catalyst TFM: A) ADF-STEM image with marked areas from which EELS and EDX spectra were extracted from the spectrum imaging dataset. B) Detail of the EELS data showing the presence of Cr, Fe, Cs and F in the bright particle in area 1 while Co content is negligible. This is also verified by EDX (C) data, which in contrast show the presence of a small amount of Co in area 2, indicating a prevalence of the distributed Co phase in the recycled catalyst.



Figure S9. Catalyst TFM: A) HAADF-STEM images showing a larger Cs, F, Fe and Cr containing particle, B) ABF-STEM image of the graphitic shell of Co particles comparable to the fresh sample and C) HAADF-STEM image showing more pronounced dots in a carbon matrix probably indicating a distributed Cs containing phase.



Figure S10. Catalyst TFM: HAADF-STEM image with region marked from which the EDX spectrum stems. This spectrum indicates the element composition of the large particle, consisting of F, O, Cr, Fe, Cs.



Figure S11. Catalyst-PFA: A) ADF-STEM image with marked areas from which EELS and EDX spectra were extracted from the spectrum imaging dataset. B) Detail of the EELS data showing the presence of Cs and Co in the bright part close to the surface in area 1. In addition, the corresponding EDX (C) data shows the strong presence of Si in area 1. In contrast, area 2 contains much less Si, Co and Cs although it is still present, while F seems to negligible in both areas.



Figure S12. Catalyst-PFA: A) HAADF-STEM image of a large Cs, F and Si containing particle attached to the Vulcan support, B) HAADF-STEM image showing a close up of the area used in SZ1 with the bright parts on the surface of the support particles indicating the presence of Cs, C) ABF-STEM image corresponding to B).



Figure S13. Catalyst-PFA: HAADF-STEM image with region marked from which the EDX spectrum stems. The spectrum indicates the element composition of the large particle, mainly consisting of F, O, Si and Cs.

5. Ancillary experimental data

Table S1. Perfluorohexylation of 4-picoline with different cobalt catalysts prepared from different cobalt precursors.

$\begin{array}{c} \begin{array}{c} \begin{array}{c} catalyst (30 \text{ mg}) \\ \hline Cs_2CO_3 (1 \text{ equiv}) \\ \hline N \\ 2a \\ 1a \\ 17 \text{ equiv} \end{array} \begin{array}{c} C_6F_{13}H \\ \hline Cs_2CO_3 (1 \text{ equiv}) \\ \hline 130 \ ^\circ\text{C}, 24 \text{ h} \\ \hline N \\ 2a \\ \hline 130 \ ^\circ\text{C}, 24 \text{ h} \end{array} \begin{array}{c} \begin{array}{c} \begin{array}{c} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \begin{array}{c} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{} \end{array}{$								
entry	cobalt precursor	conversion of 2a / %	yield of 3a / %	selectivity of 3a / %	yield of 4 / %	loss of 2a / %		
1	Co (OAc)2·4H2O	82	30	37	7	38		
2	CoCl ₂ ·6H ₂ O	64	6	9	3	55		
3	CoSO4·7H ₂ O	62	5	8	8	54		
4	Co(NO ₃) ₂ ·6H ₂ O	62	4	7	4	55		

5	Co ₃ (PO ₄) ₂	60	3	5	3	55

Table S2. Perfluorohexylation of 4-picoline with different cobalt catalysts prepared using different supports.

	+ C ₆	catalyst F ₁₃ Br <u>Cs₂CO₃ (1</u> 130 °C	(30 mg) 1.5 equiv) C, 24 h	$\frac{3}{-2}C_{6}F_{13} + C_{6}F_{13}$	- ₁₃ H	
	1a 0.2 17 equiv	mmol	3	a 4	Ļ	
entry	support	conversion of 2a / %	yield of 3a /%	selectivity of 3a / %	yield of 4 / %	loss of 2a / %
1	α -Al ₂ O ₃	47	0	0	5	42
2	SnO_2	47	0	0	5	37
3	BN	48	0	0	5	41
4	B_2O_3	43	0	0	4	37
5	MgO	49	4	8	7	36
6	ZrO	43	0	0	5	37
7	CeO_2	45	0	0	5	37
8	TiO ₂ (anatase)	57	7	12	7	41
9	V_2O_5	46	0	0	4	41
10	SiO ₂	41	0	0	4	35
11	ZnO	56	8	14	9	36

Table S3. Perfluorohexylation of 4-picoline with acid-washed cobalt catalysts.

	[$+ C_6F_{13}Br - Cs_2$ $N 2a$ $1a 0.2 mmol$	alyst (30 mg) CO ₃ (1 equiv) I30 °C, 24 h	$- \underbrace{\begin{bmatrix} 3\\ -3\\ -2\\ 3a \end{bmatrix}}_{3a}^{3} C_{6} F_{13} + \underbrace{\begin{bmatrix} 3\\ -2\\ -2\\ -2\\ -2\\ -2\\ -2\\ -2\\ -2\\ -2\\ -2$	- С ₆ F ₁₃ Н 4		
entry	catalyst	cobalt precursor	conversion of 2a / %	yield of 3a /%	selectivity of 3a / %	yield of 4 / %	loss of 2a / %
1	Co@N/C	Co(OAc)2·4H2O	67	20	30	10	33
2	Co@N/C	CoCl ₂ ·6H ₂ O	64	18	28	9	34
3	Co@N/C	CoSO ₄ ·7H ₂ O	55	11	20	8	32
4	Co@N/C	Co(NO ₃) ₂ ·6H ₂ O	56	13	23	9	31
5	Co@N/C	Co ₃ (PO ₄) ₂	47	8	17	9	28
6	Co@N/MS4Å	Co(OAc)2·4H2O	41	4	10	5	33
7	Co@/TiO2 (rutile)	Co(OAc)2·4H2O	44	0	0	6	38
8	Co@N/a-Al ₂ O ₃	Co(OAc)2·4H2O	46	0	0	6	42
9	Co@N/SnO	Co(OAc)2·4H2O	46	0	0	5	41
10	Co@N/BN	Co(OAc)2·4H2O	43	0	0	5	37
11	$Co(a)N/B_2O_3$	Co(OAc)2·4H2O	46	2	4	5	39

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12	Co@/ZrO	Co(OAc) ₂ ·4H ₂ O	49	0	0	4	43		
13	Co@/CeO ₂	Co(OAc) ₂ ·4H ₂ O	54	5	10	5	42		
14	Co@/BC	Co(OAc) ₂ ·4H ₂ O	48	0	0	4	40		
15	Co@/TiO2 (anatas)	Co(OAc) ₂ ·4H ₂ O	52	3	5	5	43		
16	Co@/V2O5	Co(OAc) ₂ ·4H ₂ O	64	10	16	11	39		
17	Co@/SiO2	Co(OAc) ₂ ·4H ₂ O	42	0	0	3	37		
Reaction converse	Reaction conditions: Catalysts were washed with 0.5 M H ₂ SO ₄ for 8h at 80°C after preparation. Yields and conversions were determined by ¹⁹ F NMR with 1.4-difluorobenzene as standard.								

Table S4. Cobalt-catalyzed perfluorohexylation of 4-picoline in different solvents.

$\begin{array}{c c} & \mathbf{Cat2} (30 \text{ mg}) \\ & \mathbf{base} (1 \text{ equiv}) \\ & \mathbf{base} (1 \text{ equiv}) \\ & \mathbf{solvent} (1 \text{ mL}) \\ & 130 \ ^\circ \text{C}, 24 \text{ h} \\ & \mathbf{3a} \\ \end{array} \begin{array}{c} 3 \\ & \mathbf{C}_6 \text{F}_{13} + \text{ C}_6 \text{F}_{13} \text{H} \\ & \mathbf{N} & 2 \\ & \mathbf{3a} \\ & 4 \end{array}$								
entry	solvent	conversion of 2a / %	yield of 3a /%	selectivity of 3a / %	yield of 4 / %	loss of 2a / %		
1	heptane	97	2	2	78	8		
2	cyclohexane	100	1	1	81	10		
3	perfluorhexane	100	0	0	25	-		
4	CH ₃ CN	100	9	9	44	28		
5	acetone	95	21	22	36	26		
6	DMSO	95	0	0	29	52		
7	CH ₂ Cl ₂	19	0	0	0	17		
8	CHCl ₃	13	0	0	0	12		
9	DCE	30	0	0	0	27		
10	CH ₂ Cl ₂ /H ₂ O (3:1)	37	1	3	1	30		
11	EtOAc	79	10	13	46	15		
12	DMA	100	3	3	51	35		

Table S5.	Cobalt-catalyzed	perfluorohexylation of	4-picoline with differen	nt amount of catalyst and base.
	2	1 2	1	2

		+ C ₆ F ₁₃ Br N 2a 1a 0.2 mmol 17 equiv	Cat2 Cs₂CO₃ 130 °C, 24 h	$\rightarrow \boxed[\frac{3}{N^2}C_6F_{13}]$	+ C ₆ F ₁₃ H 4		
entry	amount of Cat2	amount of Cs ₂ CO ₃	conversion of 2a / %	yield of 3a / %	selectivity of 3a / %	yield of 4 / %	loss of 2a / %
1	40 mg	0	38	12	32	19	3
2	40 mg	1 equiv	100	47	47	41	9
3	40 mg	2 equiv	100	41	41	49	5
4	20 mg	2 equiv	93	27	29	54	8

5	20 mg	1 equiv	88	27	31	51	6
6	5 mg	1 equiv	52	<5	10	47	3

 Table S6. Cobalt-catalyzed perfluorohexylation of 4-picoline at different temperatures.

$\begin{array}{c} \begin{array}{c} & \textbf{Cat2} (30 \text{ mg}) \\ & \textbf{F}_{13}\text{Br} & \underline{\textbf{Cs}_2\text{CO}_3 (1 \text{ equiv})} \\ & \textbf{2a} & 24 \text{ h} \\ \hline \textbf{1a} & 0.2 \text{ mmol} \\ & \textbf{17 equiv} \end{array} \xrightarrow{\begin{array}{c} \textbf{Cat2} (30 \text{ mg}) \\ & \textbf{Cs}_2\text{CO}_3 (1 \text{ equiv}) \\ & \textbf{24 h} \\ & \textbf{N} & 2 \\ \hline \textbf{N} & 2 \\ & \textbf{S}_2 \\ & $									
entry	temperature	conversion of 2a / %	yield of 3a /%	selectivity of 3a / %	yield of 4 / %	loss of 2a / %			
1	90 °C	39	0	0	0	37			
2	100 °C	40	0	0	0	38			
3	110 °C	42	2	5	3	36			
4	120 °C	67	22	33	7	34			
5	130 °C	99	46	46	5	39			
6	140 °C	100	44	44	6	45			

Table S7. Cobalt-catalyzed perfluorohexylation of 4-picoline at different temperatures.

	+ N +	Cat2 C ₆ F ₁₃ Br <u>base</u> 2a 130	2 (30 mg) (1 equiv) ℃, 24 h	$\frac{3}{100}$ C ₆ F ₁₃ + C	C ₆ F ₁₃ H	
	1a 17 equiv	0.2 mmol		3a	4	
entry	base	conversion of 2a / %	yield of 3a / %	selectivity of 3a / %	yield of 4 / %	loss of 2a / %
1	No base	38	13	34	3	19
2	Cs ₂ CO ₃	100	60	60	12	22
3	K_2CO_3	87	49	56	8	22
4	KHCO ₃	78	48	61	8	11
5	KOAc	69	35	51	6	16
6	K ₃ PO ₄	83	51	61	8	14
7	K ₂ HPO ₄	63	31	49	5	19
8	КОН	87	47	54	13	22
9	Mg(OH) ₂	40	13	33	3	22
10	HNEt ₂	92	18	20	24	36
11	NEt ₃	99	10	10	42	31
12	DBU	100	11	11	42	31

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Table S8. Cobalt-catalyzed perfluorohexylation of 4-picoline with different additives.

$\begin{array}{c c} & & & & & & & \\ \hline & & & & & \\ & & & & \\ & & & &$									
entry	additive	conversion of 2a / %	yield of 3a /%	selectivity of 3a / %	yield of 4 / %	loss of 2a / %			
1	10 mol% (<i>n</i> -Bu) ₄ NBr	100	33	33	30	27			
2	15 mol% PPh ₃	100	45	45	23	25			
3	1 equiv PPh ₃	100	37	37	9	28			
4	1 equiv LiCl	92	47	51	13	27			
5	1 equiv LiBr	78	30	38	20	27			
6	1 equiv Zn (dust)	95	30	32	20	40			

Table S9. Degradation of perfluoroalkyl reagent.

	C ₆ F ₁₃ Br <u>additi</u> 2a 130 °C, 0.2 mmol	ves 24 h ► C ₆ F ₁₃ 4	зН	
entry	additives	conversion of 2a / %	yield of 4 / %	loss of 2a / %
1	$C_6F_{13}Br$	25	0	25
2	$C_6F_{13}Br + Cs_2CO_3$	27	0	26
3	$C_6F_{13}Br + Cs_2CO_3 + \textbf{Cat2}$	43	0	42
4	$C_6F_{13}Br + Cat2$	41	0	41

6. Characterization data

6.1 Perfluoroalkylated (hetero)arenes

Notes: For the ¹³C NMR (F coupled), peaks for the C of perfluoroalkyl chain and C adjacent to perfluoroalkyl chain (in most cases) are too broad to be assigned (not shown in the following data).

$$3a$$
, N 2 C_6F_{13}

2- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.62 (d, J = 5.1Hz, -Py, 1H), 7.51 (s -Py, 1H), 7.31 (d, J = 5.6 Hz, 1H, -Py), 2.46 (m, 3H, CH₃).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.37 (tt, J = 9.9 Hz, J = 2.3 Hz, 3F, CF₃), -113.60 (t, J = 14.0 Hz, 2F, ArCF2), -121.15, -121.42, -122.37, -125.73 (m, 8F, CF2, chain).

HRMS (ESI): m/z^+ calcd for $C_{12}H_6F_{13}N(M+H)^+$ 412.03654; found 412.03651.

3- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.70 (s, 1H, py), 8.62 (d, J = 4.7Hz, 1H, py), 7.32 (d, J = 4.7 Hz, 1H, py), 2.51 (m, 3H, CH3).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.35 (tt, J = 10.0 Hz and J = 2.2 Hz, 3F, CF3), -106.64 (pseudo t, J = 15.4 Hz, 2F, ArCF2), -120.64, -121.22, -122.32, -125.69 (m, 8F, CF2, chain).

HRMS (ESI): m/z^+ calcd for $C_{12}H_6F_{13}N(M+H)^+$ 412.03654; found 412.03659.

$$3b$$
, N 2 C_6F_{13}

2- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.79 (d, J = 4.4 Hz, 1H, py), 7.91 (t, J = 8.2Hz, 1H, py), 7.71 (d, J = 8.2 Hz, 1H, py), 7.52 (m, 1H, py).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (tt, J = 9.9 Hz and J = 2.3 Hz, 3F, CF3), -113.7 (pseudo t, J = 13.6 Hz, 2F, pyCF2), -121.2, -121.5, -122.4, -125.7 (m, 8F, CF2, chain).

HRMS (ESI): m/z^+ calcd for $C_{11}H_4F_{13}N(M+H)^+$ 398.02089; found 398.02073.

3- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.85 (m, 2H, py), 7.92 (d, J = 8.3Hz, 1H, py), 7.48 (m, 1H, py).

¹⁹F NMR (282 MHz, CDCl₃): δ = -80.4 (tt, J = 9.9 Hz and J = 2.3 Hz, 3F, CF3), -111.2 (t, J = 15.5 Hz, 2F, pyCF2), -121.0, -121.5, -122.4, -125.7 (m, 8F, CF2, chain).

$$3c$$
, N C_6F_{13}

¹**H NMR (300 MHz, CDCl₃):** δ = 8.97 (d, J = 41.5 Hz, 1H, pyz), 8.83 (m, 1H, pyz), 8.76 (m, 1H, pyz).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.4 (m, 3F, CF3), -114.3 (t, J = 12.6 Hz, 2F, pyzCF2), -121.1 - -121.5, -122.4, -125.8 (m, 8F, CF2 chain).

$$3d, 56 N 3C_6F_{13} C_6F_{13}$$

3- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.20 (d, J = 2.2 Hz, 1H, pyz), 8.06 (d, J = 2.2Hz, 1H, pyz), 5.15 (s,2H, NH2).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (tt, J = 9.9 Hz and J = 2.3 Hz, 3F, CF3), -122.5 (t, J = 14.2 Hz, 2F, pyzCF2), -121.1, -121.5, -122.4, -125.7 m, (8F, CF2, chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 152.77, 146.01, 133.73.$

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$:146.05, 133.76.

HRMS (ESI): m/z^+ calcd for $C_{10}H_4F_{13}N_3$ (M+H)⁺ 414.02704; found 414.02639

5- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.31 (d, J = 1.4 Hz, 1H, pyz), 8.04 (d, J = 1.4 Hz, 1H, pyz), 5.03 (s,2H, NH2).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.4 (tt, J = 9.9 Hz and J = 2.3 Hz, 3F, CF3), -113.3 (tt, J = 13.4 Hz, J = 2.9 Hz. 2F, pyzCF2), -121.2, -121.8, -122.4, -125.7 (m, 8F, CF2, chain).

6- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.25 (s, 1H, pyz), 8.16 (s, 1H, pyz), 4.90 (s,2H, NH2).

¹⁹F NMR (282 MHz, CDCl₃): δ = -80.4 (tt, J = 9.9 Hz and J = 2.3 Hz, 3F, CF3), -114.7 (t, J = 13.5 Hz, 2F, pyzCF2), -121.2, -121.5, -122.3, -125.7 (m, 8F, CF2, chain).

$$3e, \sqrt[5]{\frac{4}{N}2}C_{6}F_{13}$$

2- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 8.95 (d, J = 4.8 Hz, 2H, pyz), 7.53 (t, J = 4.8 Hz, 1H, pyz).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.4 (m, 3F, CF3), -114.5 (t, J = 12.9 Hz, 2F, pyzCF2), -121.0, -122.6 (m 6F, CF2), -125.8 (m, 2F, CF2).

¹³C NMR (75 MHz, CDCl₃): $\delta = 157.87, 123.00.$

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 157.89, 123.02.

HRMS (EI, 70 eV): m/z⁺calcd for C₁₀H₃F₁₃N (M)⁺ 398.00831; found 398.00832

4- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 9.43 (s, 1H, pyz), 9.03 (d, J = 4.9Hz, 1H, pyz), 7.70 (dd, J = 4.9 Hz, J = 1.4 Hz, 1H, pyz).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (t, J = 9.8 Hz 3F, CF3), -115.6 (t, J = 13.2 Hz, 2F, pyzCF2), -121.0, -121.4, -122.4, -125.8 (m, 8F, CF2, chain).

¹³C NMR (75 MHz, CDCl₃): δ = 159.10, 158.96, 119.08 (t, *J* = 4.3 Hz).

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 159.13, 158.99, 119.11 (t, J = 4.2 Hz).

HRMS (EI, 70 eV): m/z⁺calcd for C₁₀H₃F₁₃N (M)⁺ 398.00831; found 398.00832

5- isomer:

¹H NMR (300 MHz, CDCl₃): δ = 9.45 (s, 1H, pyz), 8.96 (s,2H, pyz).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (m, 3F, CF3), -112.1 (t, J = 12.9 Hz, 2F, pyzCF2), -120.9, -121.4, -122.4, -125.7 (m, 8F, CF2, chain).

¹³C NMR (75 MHz, CDCl₃): δ = 161.50, 155.37 (t, *J* = 6.2 Hz).

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 161.53, 155.40 (t, J = 6.3 Hz).

HRMS (EI, 70 eV): m/z⁺calcd for C₁₀H₃F₁₃N (M)⁺ 398.00831; found 398.00832

$$\overset{/}{\underset{\mathsf{Br}}{\overset{\mathsf{N}}{\longrightarrow}}} \mathbf{C}_{7}\mathbf{F}_{15}$$

3f, Br

¹**H NMR (300 MHz, CDCl₃):** δ = 3.79 (s, CH3).

¹⁹F NMR (282 MHz, CDCl₃): δ = -80.4 (tt, J = 9.1 Hz, 3F, CF3), -108.2 (t, J = 14.4 Hz, 2F, imCF2), -120.5, -120.9, -121.5, -122.3, 125.7 (m, 10F, CF2 chain).

HRMS (ESI): m/z^+ calcd for $C_{11}H_3Br_2F_{15}N_2$ (M+H)⁺ 608.84761; found 608.84737

$$5\sqrt{\frac{N^2}{5}}C_{6}F_{13}$$

3g,

2- isomer: ¹H NMR (300 MHz, CDCl₃): 7.13 (m, 1H, im), 7.00 (m, 1H, im), 3.79 (s, 3H, CH3).

¹³C NMR (75 MHz, CDCl₃): $\delta = 129.39, 125.58, 34.47.$

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 129.40, 125.61, 34.49.

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.6 (m, 3F, CF3), -107.8 (m, 2F, imCF2), -121.1, -121.3, -122.5, -125.9 (m 8F, CF2, chain).

HRMS (EI, 70 eV): C₁₀H₅F₁₃N₂ 401.03179; found 401.03168.

5- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 7.52 (s,1H, im), 7.38 (s, 1H, im), 3.70 (s, 3H, CH3).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.8 (m, 3F, CF3), -105.9 (m, 2F, imCF2), -121.0, -121.5, -122.7, -126.1 (m 8F, CF2, chain).

HRMS (EI, 70 eV): C₁₀H₅F₁₃N₂ 401.03179; found 401.03154

3h,

¹**H NMR (300 MHz, CDCl₃):** δ = 7.52 (t, J = 0.9 Hz, 1H, pz), 4.00 (t, J = 1.9 Hz, 3H, CH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (tt, J = 9.9 Hz, J = 2.3 Hz, 3F, CF3), -107.4 (t, J = 13.5 Hz, 2F, pzCF2), -120.9, -121.5, -122.3, -125.7 (m, 8F, CF2 chain).

GCMS (EI, 70 eV) m:z (relative intensity): 411 and 409 (M⁺-CF₃, 1 and 1), 211 and 209 (M⁺-C₅F₁₁, 97 and 100).

$$5\sqrt[]{N-N}{3i}C_{6}F_{13}$$

3- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 7.44 (s, 1H, pz), 6.53 (d, J = 2.1 Hz, 1H, pz), 3.99 (s, 3H, CH3).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.4 (tt, J = 9.7 Hz, J = 2.4 Hz, 3F, CF3), -109.2 (t, J = 13.7 Hz, 2F, pzCF2), -121.2, -

122.2, -122.5, -125.7 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 131.48, 106.25, 39.62.$

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 131.52, 106.28, 39.65.

HRMS (ESI): m/z^+ calcd for $C_{10}H_5F_{13}N_2$ (M+H)⁺ 401.03179; found 401.03114.

4- isomer:

¹H NMR (300 MHz, CDCl₃): δ = 7.67 (s, 1H, pz), 7.62 (s, 1H, pz), 3.96 (s, 3H, CH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (t, J = 9.1 Hz, 3F, CF3), -104.1 (t, J = 14.3 Hz, 2F, pzCF2), -121.1, -122.3, -122.5, -125.7 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 137.98$ (t, J = 4.5 Hz), 130.18 (t, J = 4.9 Hz), 39.42.

HRMS (ESI): m/z^+ calcd for $C_{10}H_5F_{13}N_2$ (M+H)⁺ 401.03179; found 401.03137.

5- isomer:

¹H NMR (300 MHz, CDCl₃): δ = 7.52 (m, 1H, pz), 6.61 (d, J = 2.0 Hz, 1H, pz), 4.0 (s, 3H, CH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (t, J = 10.0 Hz, 3F, CF3), -106.7 (t, J = 14.7 Hz, 2F, pzCF2), -120.9, -121.3, -122.4, -125.7 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 138.50, 109.91, 39.02$.

¹³C NMR Dept135 (75 MHz, CDCl₃): δ = (+): 138.53, 109.94, 39.04.

HRMS (ESI): m/z^+ calcd for $C_{10}H_5F_{13}N_2$ (M+H)⁺ 401.03179; found 401.03135.

$$5 \bigvee_{3j}^{N-N} C_{6} F_{13}$$

3- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 7.22 (s, 1H, pz), 3.89 (s, 3H, NCH3), 2.12 (s, 3H, CH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (t, J = 10.0 Hz, 3F, CF3), -108.9 (t, J = 14.1 Hz, 2F, pzCF2), -121.3, -131.8, -122.4, -125.7 (8F, CF₂ chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 131.17, 117.72, 39.41, 8.37.$

¹³C NMR Dept135 (75 MHz, CDCl₃): δ = (+): 131.20, 39.44, 8.40.

5- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 7.33 (s, 1H, pz), 3.94 (t, J = 1.7 Hz, 3H, NCH3), 2.14 (t, J = 2.3 Hz, 3H, CH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (t, J = 10.1 Hz, 3F, CF3), -107.3 (t, J = 14.2 Hz, 2F, pzCF2), -121.6 (4F, CF2 chain), -122.4, -125.8 (m, 4F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): δ = 139.84, 121.32, 39.51, 9.00.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 139.86, 39.50, 9.03.

,C₆F₁₃ 3k.

¹**H NMR (300 MHz, CDCl₃):** δ = 6.75 (pseudo t, J = 2.2 Hz, 1H, pyrrole), 6.75 (m, 1H, pyrrole), 6.16 (pseudo t, J = 3.5 Hz, 1H, pyrrole), 3.72 (s, 3H CH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (tt, J = 9.8 Hz, J = 2.3 Hz, 3F, CF3), -104.1 (t, J = 13.6 Hz, 2F, pyrroleCF2), -120.6, -121.3, -122.4, -125.8 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 128.15$, 114.54 (t, J = 5.4 Hz), 107.93, 35.83.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 128.15, 114.53 (t, J = 5.4 Hz), 107.93, 35.83 (p, J = 3.4 Hz).

GCMS (EI, 70 eV) m:z (relative intensity): 399 (M⁺, 29), 380 (11), 130 (100).

¹H NMR (300 MHz, CDCl₃): $\delta = 6.67$ (s,1H, thiophene), 2.47 (m,3H, CH3), 2.40 (s, 3H, CH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (tt, J = 10.0 Hz, J = 2.3 Hz, 3F, CF3), -105.4 (t, J = 14.3 Hz, 2F, thiopheneCF2), -121.5, -121.7, -122.4, -125.8 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 141.43$ (t, J = 4.1 Hz), 136.80, 124.59 (t, J = 5.3 Hz), 14.89, 13.82.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 124.61 (t, J = 5.5 Hz), 14.91, 13.84.

HRMS (EI, 70 eV):430.00554; found 430.00395

OMe C₆F₁₃ 3m, MeO OMe

¹**H NMR (300 MHz, CDCl₃):** $\delta = 6.14$ (s, 2H, arom), 3.84 (s, 3H, CH3), 3.80 (s, 6H, CH3).

¹⁹F NMR (282 MHz, CDCl₃): δ = -80.4 (tt, J = 9.9 Hz, J = 2.3 Hz, 3F, CF3), -102.3 (t, J = 14.3 Hz, 2F, CF2mesitylene), -121.8 (m, 4F, CF2 chain), -122.3, -125.7 (m, 4F, CF2 chain).

HRMS (ESI): [M+H]+ calcd for C₁₅H₁₁F₁₃O₃, 487.05734; found, 487.05704.



2- isomer:

¹**H** NMR (300 MHz, CDCl₃): $\delta = 6.86$ (d, J = 2.4 Hz, 1H, arene), 6.47 (m, 1H, arene), 3.80 (s, 3H, OCH3), 3.81 (s, 3H, OCH3).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (tt, J = 10.1 Hz, J = 2.7 Hz, 3F, CF3), -100.6 (t, J = 15.8 Hz, 2F, CF2aryl), -119.9, -121.7, -122.3, -125.7 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): δ = 162.67, 161.98 (t, *J* = 1.7 Hz), 123.64 (t, *J* = 3.7 Hz), 112.93, 99.40, 56.51, 55.70.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 112.93, 99.41, 56.51, 55.70.

HRMS (ESI): m/z^+ calcd for $C_{14}H_8BrF_{13}O_2$ (M+H)⁺ 534.95729; found 534.95671

4- isomer:

¹H NMR (300 MHz, CDCl₃): $\delta = 6.78$ (s, 2H, arene), 3.83 (s,6H, OCH3).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.4 (tt, J = 9.8 Hz, J = 2.4 Hz, 3F, CF3), -103.3 (t, J = 14.6 Hz, 2F, CF2aryl), -121.7, -121.8, -122.3, -125.8 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 160.83$ (t, J = 2.0 Hz), 127.62, 109.00, 56.66.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 109.03, 56.70.

HRMS (EI, 70 eV):533.94946; found 533.94946

.C₆F₁₃ 30.

¹H NMR (300 MHz, CDCl₃): δ = 7.64-754 (m, 3H, benzene), 7.54-7.46 (m, 2H, benzene).

¹⁹F NMR (282 MHz, CDCl₃): δ = -80.7 (tt, J = 9.9 Hz, J = 2.3 Hz, 3F, CF3), -110.6 (t, J = 14.9 Hz, 2F, CF2benzene), -121.3, -121.7, -122.6, -125.9 (m, 8F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): δ = 131.92, 128.60, 126.84 (t, *J* = 6.7 Hz).

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 131.88 (t, J = 1.5 Hz), 128.56, 126.80 (t, J = 6.6 Hz).



3p, ²

2- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 7.31 (m, 2H, aniline), 6.80 (m, 1H, aniline), 6.71 (d, J = 6.9 Hz, 1H, aniline), 4.22 (s, 2H, NH2).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.4 (tt, J = 9.8 Hz, J = 2.3 Hz, 3F, CF3), -108.1 (t, J = 15.1 Hz, 2F, CF2aniline), -121.3, -121.5, -122.3, -125.7 (m, 8F, CF2 chain).

GCMS (EI, 70eV) m:z (relative intensity): 411 (M⁺, 50), 392 (19), 342 (6), 173 (6), 142 (100).

2- isomer:

¹**H NMR (300 MHz, CDCl₃):** δ = 7.28 (m,1H, arene), 7.26-7.19 (m, 1H, arene), 6.65 (d, J = 8.6 Hz, 1H, arene), 4.24 (s, 2H, NH2).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -80.4 (t, J = 9.9 Hz, 3F, CF3), -108.6 (t, J = 14.8 Hz, 2F, areneCF2), -121.4 (m, 4F, CF2 chain), -122.4, -125.7, (m, 4F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 144.47, 133.06, 128.42$ (t, J = 9.1 Hz), 122.52, 119.07.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 133.08, 128.45 (t, J = 9.1 Hz), 119.09.

HRMS (ESI): m/z⁺calcd for C₁₂H₅ClF₁₃N (M+H)⁺445.99757; found 445.99763.

3r,

¹**H NMR (300 MHz, CDCl₃):** δ = 7.56 (m, 1H, aniline), 7.52 (dm, J = 8.6 Hz, 1H, aniline), 6.49 (d, J = 8.6 Hz, 1H, aniline), 4.25 (broad s, 2H, NH2).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.8 (tt, J = 10.0 Hz, 2.4 Hz, 3F, CF3), -108.9 (t, J = 14.9 Hz, 2F, CF2aniline), -121.5, -121.8, -121.9 (m, 6F, CF2 chain), -122.7, -126.1 (m, 4F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): $\delta = 145.48$, 141.50, 137.05 (t, J = 9.0 Hz), 119.70.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 141.50, 137.06 (t, J = 9.1 Hz), 119.71.

HRMS (ESI): m/z^+ calcd for $C_{13}H_5F_{15}IN (M+H)^+587.92999$; found 587.92919.



, white solid

¹**H** NMR (300 MHz, CDCl₃): δ = 7.60 (d, J = 2.1 Hz, 1H, arene), 7.51 (dd, J = 8.3 Hz, J = 1.8 Hz, 1H, arene), 6.74 (d, J = 8.3 Hz, 1H, arene), 4.79 (s, 2H, NH2).

¹⁹**F NMR (282 MHz, CDCl₃):** δ = -80.8 (t, J = 10.1 Hz, 3F, CF3), -109.4 (t, J = 13.7 Hz, 2F, arylCF2), -121.5 -121.9, -122.7, -126.1 (m, 10F, CF2 chain).

¹³C NMR (75 MHz, CDCl₃): δ = 149.13, 136.04, 134.11 (t, *J* = 9.0 Hz), 118.56, 117.82, 100.38.

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 136.05, 134.13 (t, J = 8.8 Hz), 117.84.

HRMS (EI, 70 eV): [M+H]+ calcd for C₁₄H₅F₁₅N₂ 487.0286; found 487.0282.

6.2 Trifluoromethylated heteroarenes

F₃C 5a.

2- isomer: ¹⁹F NMR (282 MHz, CDCl₃): δ = -62.1 (s, 3F, CF3) 3- isomer: ¹⁹F NMR (282 MHz, CDCl₃): δ = -63.1 (s, 3F, CF3) 4- isomer: ¹⁹F NMR (282 MHz, CDCl₃): δ = -63.3 (s, 3F, CF3)



¹H NMR (300 MHz, CDCl₃): δ = 4.14 (pseudo q, *J* = 1.3 Hz, 3H, CH3), 3.57 (s, 3H, CH3), 3.40 (s, 3H, CH3). ¹⁹F NMR (282 MHz, CDCl₃): δ = -62.3 (s, 3F, CF3).

¹³C NMR (75 MHz, CDCl₃): δ = 155.40, 151.27, 146.46, 119.94, 116.34, 33.15 (q, *J* = 2.0 Hz), 29.85, 28.15. ¹³C NMR Dept135 (75 MHz, CDCl₃): δ = (+): 33.18 (q, *J* = 1.9 Hz), 29.89, 28.19.

HRMS (EI): [M+H]+ calcd for C₉H₉F₃N₄O₂, 263.07504; found, 263.07522.

F₂C 6c H₂N NH₂

¹H NMR (300 MHz, CDCl₃): δ = 7.69 (s, arene), 5.13 (s,4H, NH2).

¹⁹**F** NMR (282 MHz, CDCl₃): δ = -62.1 (s, 3F, CF3).

¹³C NMR (75 MHz, CDCl₃): $\delta = 156.51, 135.82, 129.72 - 119.00 (q, J = 267.9 Hz), 98.51 - 97.18 (q, J = 33.4 Hz).$

¹³C NMR Dept135 (75 MHz, CDCl₃): $\delta = (+)$: 135.86 (hept, J = 4.6 Hz).

MS (EI, 70 eV): m/z (%) = 245 (100) M+, 226 (28), 198 (43).

7. References

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8. Spectra

8.1 Perfluoroalkylated (hetero)arenes
































S40



S41























161115.338.10.fid Weniger FW_5_Iso_2 Au19F-FWenig CDCl3 /opt/topspin 1611 38 -121.42 -121.91 -122.45 -4200 -108.97 -4000 -3800 -3600 3j 5 -3400 -3200 -3000 3- isomer -2800 ¹⁹F NMR 2600 -2400 2200 -2000 1800 -1600 1400 -1200 -1000 -800 -600 -400 -200 -0 F 00'E L-94-I 1.93 × 101 --200 -400 -80 f1 (ppm) 0 0 -10 -20 -30 -60 -100 -110 -120 -160 -40 -50 -70 -90 -130 -140 -150 161115.337.11.fid Weniger FW_5_Iso_1 Au1H CDCB /opt/topspin 1611 37 <7.33 77.26 cDCl3 OG (3.95 (3.94 (3.94 215 -9000 8000 C₆F₁₃ 3j 5 -7000 5-isomer 6000 ¹H NMR 5000 -4000 -3000 -2000 -1000 -0 3.18 H 1.00-1 3.15 I 4 3 2 16 15 14 13 12 11 10 9 8 7 6 f1 (ppm) 5 1 0 -1 -2 -3 -4







































8.2 Trifluoromethylated heteroarenes








