Coupled Morphological and Structural Evolution of δ -MnO₂ to α -MnO₂ through Multistage Oriented Assembly Processes: the Role of Mn(III)

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Supporting Figures



Figure S1. (A) α -MnO₂ [MnO₆] octahedra tunnel structure model; (B) δ -MnO₂ [MnO₆] octahedra layer structure model.



Figure S2. FESEM of intermediate product of (A) NaMix, (B) Na100-2, (C) Na100-6, (D) Na100-10, (E) Na 100-14 and (F) Na80.



Figure S3. (A) The long nanosheets are cut along (001) plane, (B) The nanorods are cut along (001) plane.



Figure S4. HRTEM image of the initial NaMix indicate the size about the edge of primary nanoflakes to be 3-4 nm



Figure S5. (A) HRTEM image of the intermediate Na100-10 indicate two long nanosheets stack and thicken. (B) AFM image of the intermediate Na100-10 indicate the nanoflakes stack on the nanoribbon.

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Figure S6. Scheme of a δ -MnO₂ layer which containing 1/3 vacancy (blue octahedral) per layer octahedron, seen along c-axis. The Mn_LX (X = 1, 2 or 3) means the first, second and third Mn shells around a given Mn. The Mn_{IL}Y (Y= 1 or 2) means pairs formed by layer Mn and Mn at vacancies belonging respectively to the first and second shells.



Figure S7. XPS broad scans of (a) Mn $2p_{3/2}$ and (b) Mn 3s spectra of intermediate products at NaMix, Na100-4, Na100-10, Na100-14, and Na80.



Figure S8. Oxygen 1s spectra of intermediate products at different intervals of Na+ stabilized α -MnO₂ formation and crystal growth;



Figure S9. The TEM and HRTEM images during entire synthetic process: (a, b) NaMix, (c, d) Na100-4, (e, f) Na100-10, (g, h) Na100-14 and (i, j) Na80.



Figure S10. Schematic illustrating the ideal structure of δ -MnO₂ which can assemble to α -MnO₂.



Figure S11. XRD patterns of intermediate products at different intervals of K^+ stabilized α -MnO₂ formation and crystal growth;



Figure S12. XRD patterns of intermediate products at different intervals of Na⁺ stabilized α -MnO₂ age under 60 °C.



Figure S13 the calculated relaxed (a) Na^+ stabilized α -MnO₂ structure and (b) K^+ stabilized α -MnO₂ structure.

Supporting Tables

Table S1. K⁺ content, average oxidation states (AOS) of Mn in intermediate products at various time intervals during K⁺-stabilized α -MnO₂ formation obtained from titration and fittings of Mn (2p^{3/2})

Samples	XPS (± 0.02)	Titration	K ⁺ content (% mol)
KMix	3.80	3.95 ± 0.03	6.91 ± 0.13
K100-0	3.77	3.90 ± 0.01	4.85 ± 0.09
K100-2	-	3.85 ± 0.01	39.57 ± 0.15
K100-4	3.74	3.80 ± 0.02	5.25 ± 0.11

K100-10	3.80	3.87 ± 0.04	5.89 ± 0.04
K100-20	-	3.85 ± 0.02	8.51 ± 0.03
K80	3.80	3.82 ± 0.03	10.56 ± 0.08

Table S2. Near-surface compositions of Mn and O species derived from fittings of Mn $(2n^{3}/2)$ and O (1s)

(2p3/2) and $O(13)$.								
Sample	Mn(III)	Mn(II)	Mn(IV)	O ²⁻	OH-	H_2O		
	(±0.003)	(±0.001)	(±0.003)					
KMix	0.158	0.037	0.845	0.661	0.183	0.156		
K100-0	0.149	0.037	0.814	0.645	0.196	0.159		
K100-4	0.153	0.049	0.798	0.580	0.218	0.202		
K100-10	0.112	0.039	0.849	0.640	0.210	0.150		
K80	0.137	0.028	0.835	0.770	0.161	0.068		