

SUPPORTING INFORMATION

Facile conversion of waste glass into Li storage materials

Seung-Su Lee^a and Cheol-Min Park^{a}*

^aSchool of Materials Science and Engineering, Kumoh National Institute of Technology, Gumi,
Gyeongbuk 39177, Republic of Korea

*Cheol-Min Park. Tel.: +82-54-478-7746; Fax: +82-54-478-7769

E-mail: cmpark@kumoh.ac.kr

Figure S1.

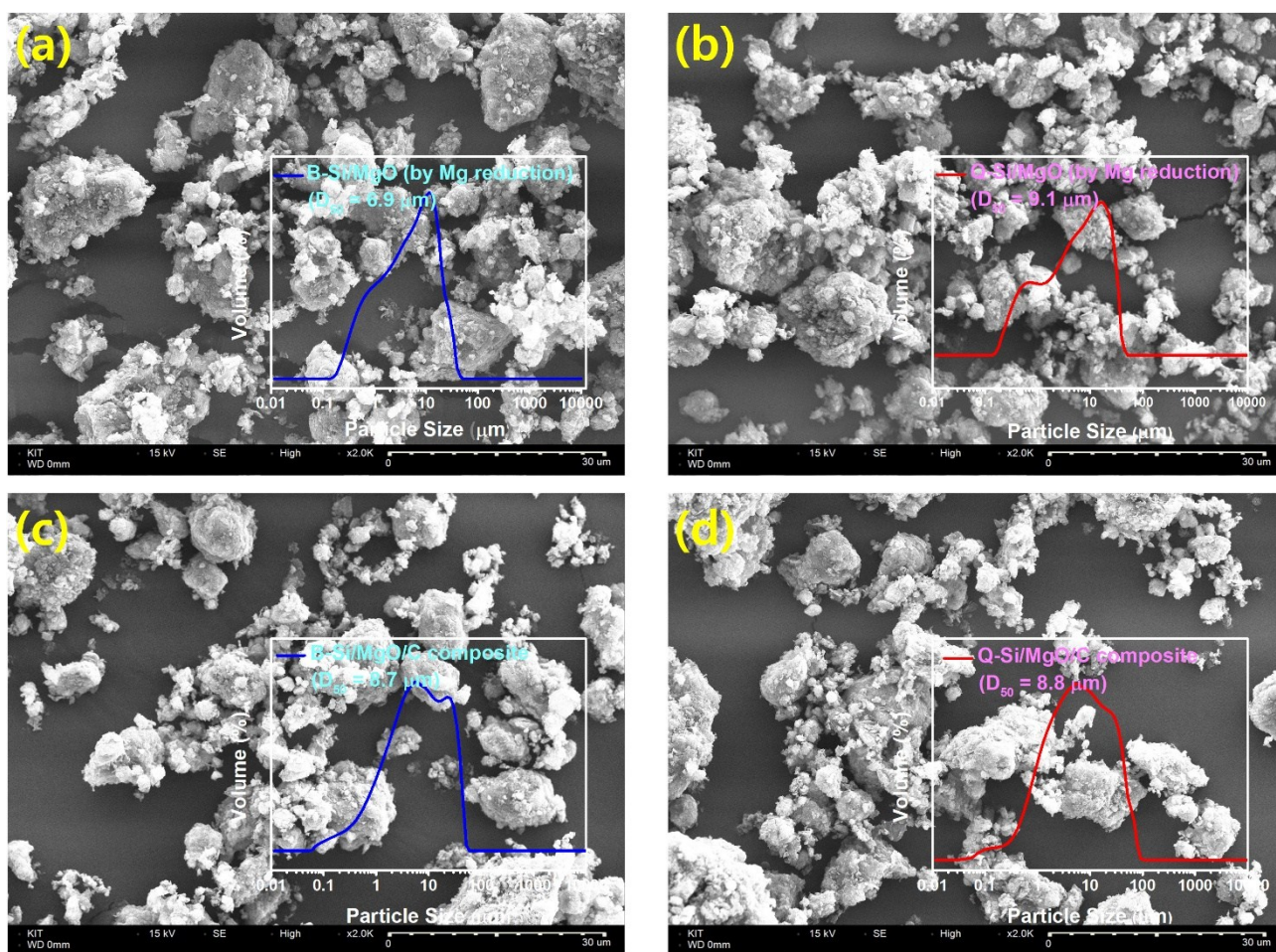


Figure S1. Particle size and morphology of Si/MgO and Si/MgO/C. SEM images and particle size distributions of (a) B-Si/MgO, (b) Q-Si/MgO, (c) B-Si/MgO/C, and (d) Q-Si/MgO/C.

Figure S2.

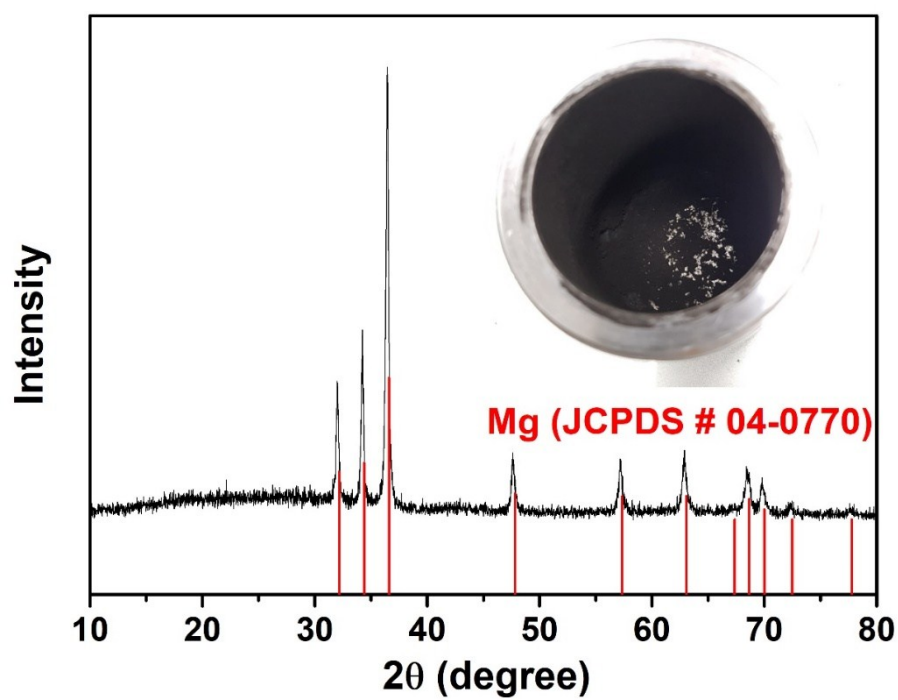


Figure S2. XRD and photography of S-Si/MgO. The S-Si/MgO was formed by the high-energy BM process using soda-lime glass and Mg powders.

Figure S3

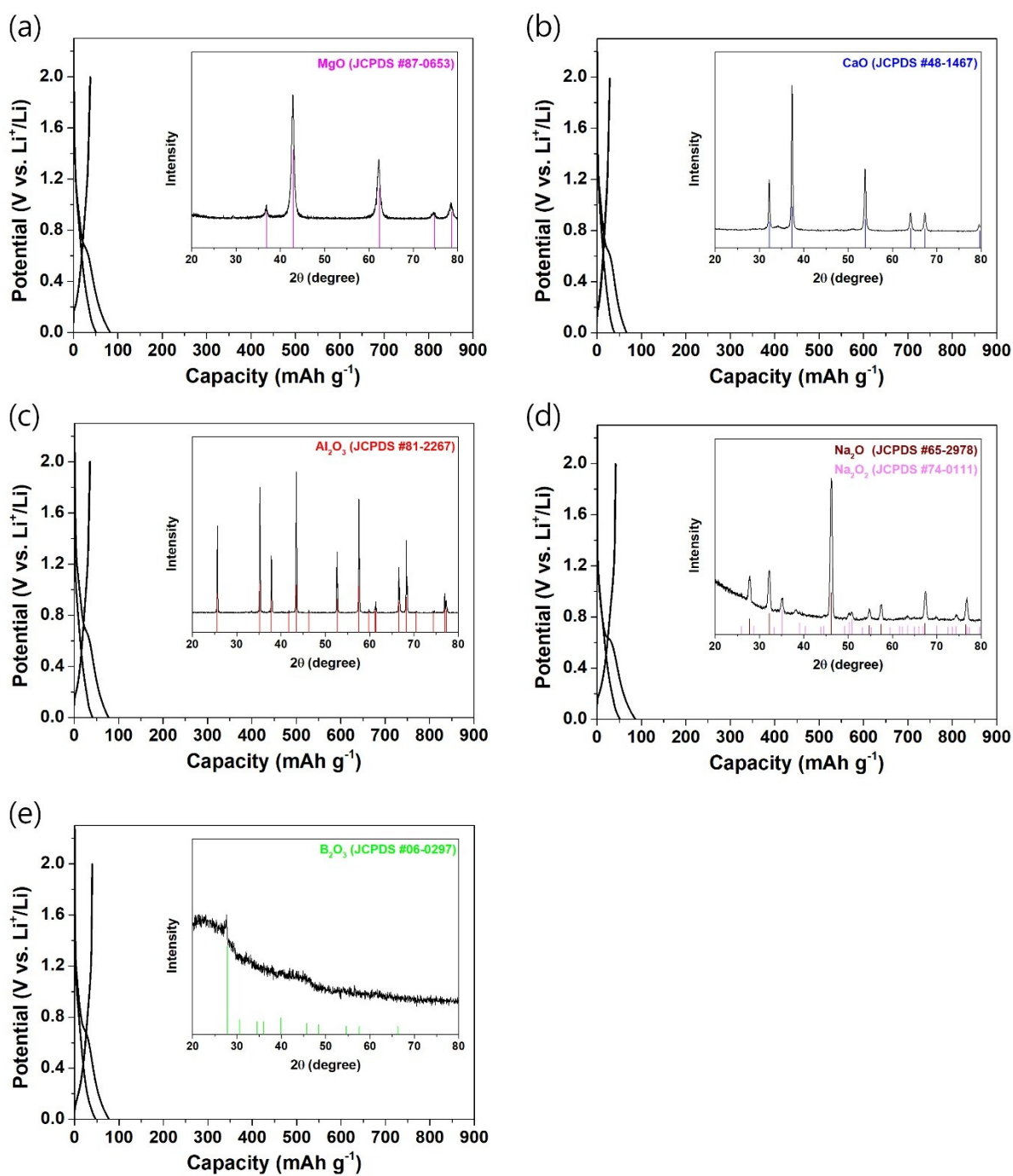


Figure S3. Electrochemical performances of various glass forming oxides. Voltage profiles and XRD results of (a) MgO, (b) CaO, (c) Al_2O_3 , (d) Na_2O , and (e) B_2O_3 .

Figure S4

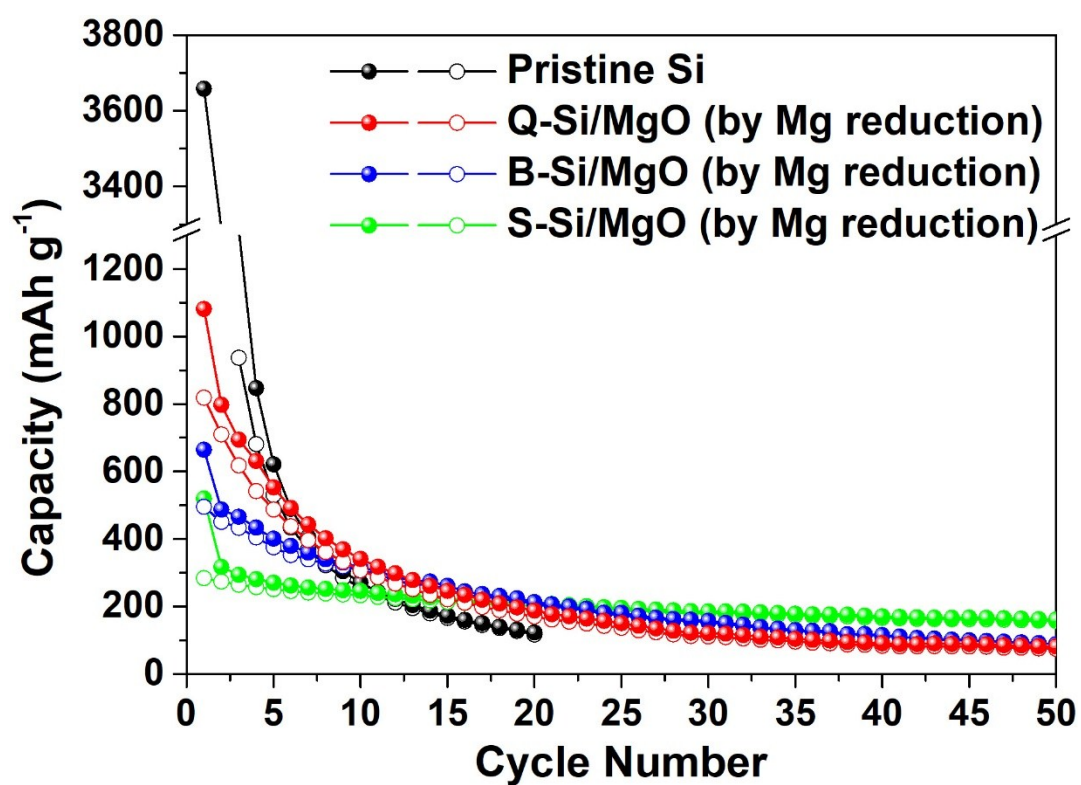


Figure S4. Cycling performance results for Si, Q-Si/MgO, B-Si/MgO, and S-Si/MgO electrodes.

Figure S5.

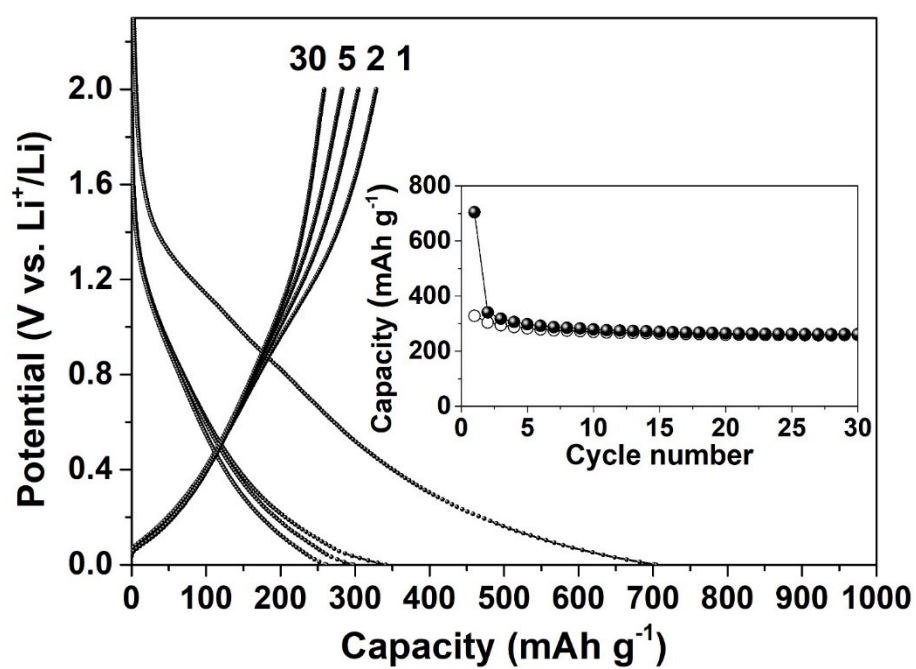


Figure S5. Electrochemical performances of the ball-milled C (Super P) electrode. Voltage profile of the ball-milled C at a current density of 100 mA g⁻¹ (Inset: Cycling behavior of the ball-milled C at a cycling rate of 100 mA g⁻¹).

Figure S6.

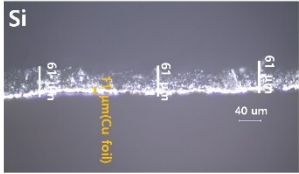



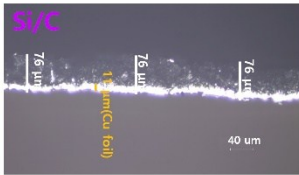



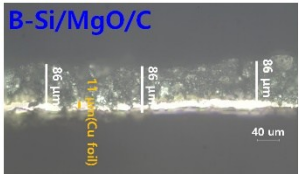



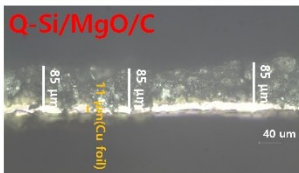



Electrodes	Electrodes before Li-reaction	Li-inserted (0 V)	Li-extracted (2 V)
(a) Si	  61 - 11 (Cu foil) = 50 μm (active material)	 126 - 11 (Cu) = 115 μm [130.0%]	 80 - 11 (Cu) = 69 μm [38.0%]
(b) Si/C composite	  76 - 11 (Cu foil) = 65 μm (active material)	 116 - 11 (Cu) = 105 μm [61.5%]	 86 - 11 (Cu) = 75 μm [15.4%]
(c) B-Si/MgO/C composite	  86 - 11 (Cu foil) = 75 μm (active material)	 108 - 11 (Cu) = 97 μm [29.3%]	 90 - 11 (Cu) = 79 μm [5.3%]
(d) Q-Si/MgO/C composite	  85 - 11 (Cu foil) = 74 μm (active material)	 112 - 11 (Cu) = 101 μm [36.5%]	 92 - 11 (Cu) = 81 μm [9.5%]

Figure S6. Electrode thickness changes in Si, Si/C, B-Si/MgO/C, and Q-Si/MgO/C electrodes according to the state-of-charge. (a) Swelling data of Si electrode (before Li-reaction, Li-inserted at 0 V, and Li-extracted at 2 V). (b) Swelling data of Si/C composite electrode (before Li-reaction, Li-inserted at 0 V, and Li-extracted at 2 V). (c) Swelling data of B-Si/MgO/C composite electrode (before Li-reaction, Li-inserted at 0 V, and Li-extracted at 2 V). (d) Swelling data of Q-Si/MgO/C composite electrode (before Li-reaction, Li-inserted at 0 V, and Li-extracted at 2 V).

Figure S7.

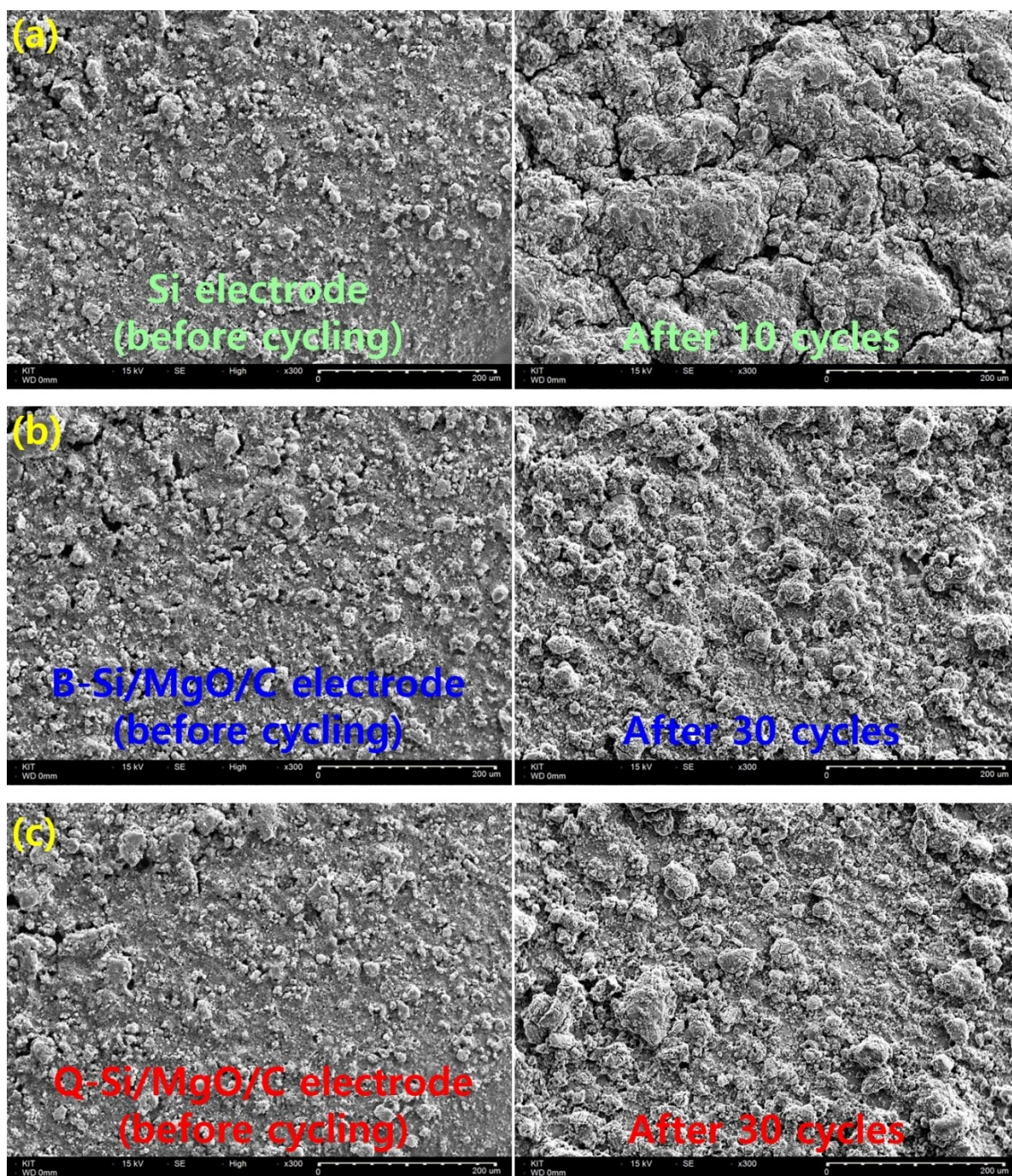


Figure S7. Morphological investigation using *ex-situ* SEM of the Si, B-Si/MgO/C, and Q-Si/MgO/C electrodes during cycling. (a) *Ex-situ* SEM results of the pristine Si electrode before cycling and after 10 cycles. (b) *Ex-situ* SEM results of the B-Si/MgO/C electrode before cycling and after 30 cycles. (c) *Ex-situ* SEM results of the Q-Si/MgO/C electrode before cycling and after 30 cycles.

Figure S8.

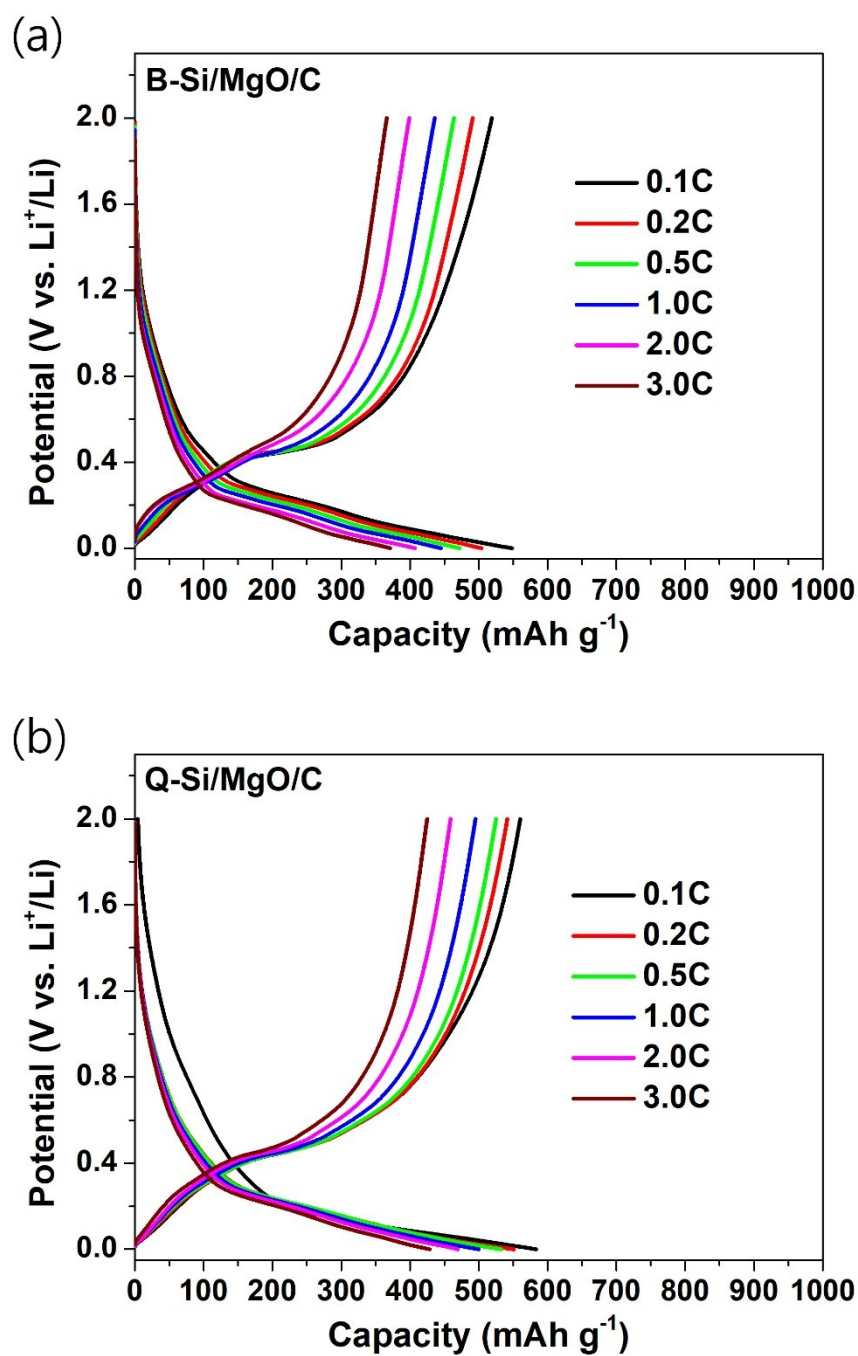


Figure S8. Voltage profiles at different C-rates of B-Si/MgO/C and Q-Si/MgO/C composite electrodes. (a) Voltage profiles of B-Si/MgO/C composite electrode. (b) Voltage profiles of Q-Si/MgO/C composite electrode.