Selective Hydrothermal Reductions Using Geomimicry

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Figure 2. SEM image of the iron nanopowder.



Figure 3. XRD patterns using Co-k α radiation for (top trace), the unreacted nickel powder, (middle trace) the unreacted iron powder, and (bottom trace), a mixture of the nickel and the iron powders as used in the experiments before reaction.



Figure 4. XRD patterns using Co-k α radiation for (top trace), the unreacted nickel and iron powder mixture, (middle trace) the solid products of the nickel and iron powder mixture after reaction for 1 hour at 250°C, nickel and iron, and (bottom trace), the solid products of the nickel and iron powder mixture after reaction for 68 hours at 250°C.



Figure 5. XRD patterns using Co-k α radiation for (top trace), the solid products of reaction of nickel and iron for 68 hours at 250°C, (middle trace) the solid products of reaction of nickel alone for 68 hours at 250°C, and (bottom trace), the solid products of reaction of iron alone for 68 hours at 250°C. The peaks indicated with the diamonds are associated with metallic nickel and the peaks indicated with the solid circles are associated with metallic iron.



Figure 6. XRD pattern using Co-k α radiation for the products of reaction of nickel and iron at 250°C and 40 bar for 40 hours, and peak assignments obtained using the EVA interpretative software. The products contain a mixture of unreacted nickel, some unreacted iron, corundum from the milling process, and magnetite oxidation product. The reaction products were milled to ca. 20 μ m grain size in the presence of corundum, which is the reason for the corundum peaks observed in the XRD pattern. The milling procedure was similar to that described in: Srodoi, J., Drits, V. A., McCarty, D. K., Hsieh, J. C. C., Eberl, D. D. Quantitative X-Ray Diffraction Analysis of Clay-Bearing Rocks From Random Preparations. *Clays and Clay Minerals*, **2001**, *49*, 514-528.