

Revealing the Turkevich mechanism of gold nanoparticles synthesis

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Supporting Information

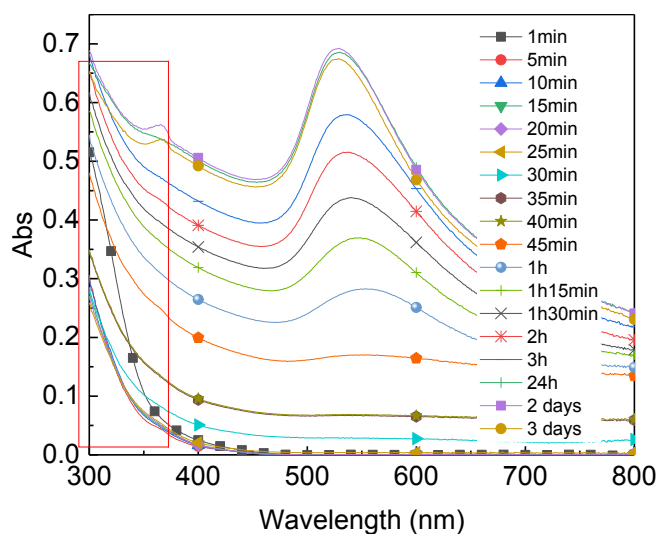


Figure S1 The time-evolution of UV-vis spectra of the experiment using sodium citrate as a reducing agent at 25 °C. Conditions: $[\text{HAuCl}_4]$: 0.25 mM, $[\text{Sodium citrate}]$: 2.5 mM, pH: 4.9.

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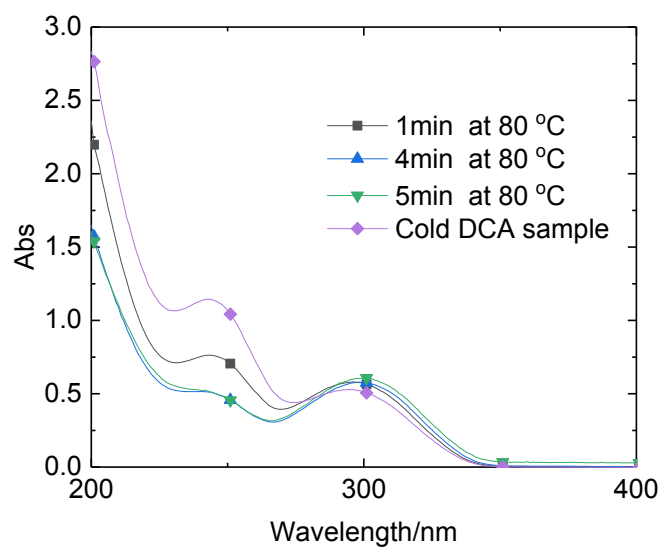


Figure S2: Thermal stability of an aqueous solution of DCA (2.5 mM) at 80°C.

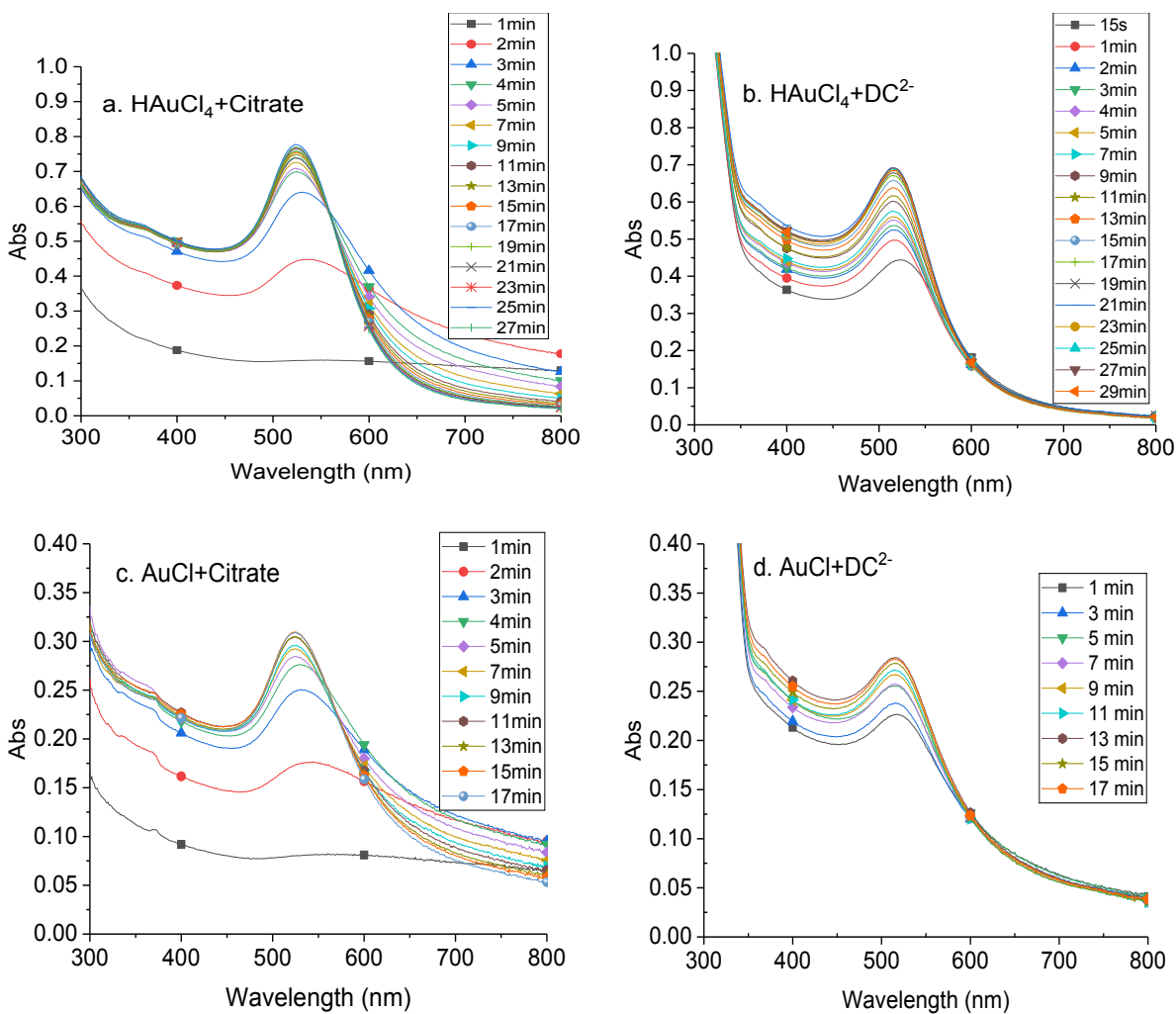


Figure S3: The time-evolution of UV-vis spectra of the experiment using different gold precursor and reducing agent. a. HAuCl_4 + citrate, b. HAuCl_4 + DC^{2-} , c. AuCl + citrate, d. AuCl + DC^{2-} . Conditions: $[\text{HAuCl}_4]$ or $[\text{AuCl}]$: 0.25 mM, [Sodium citrate] or [DCA]: 2.5 mM, pH: 4.8-4.9, 80 °C.

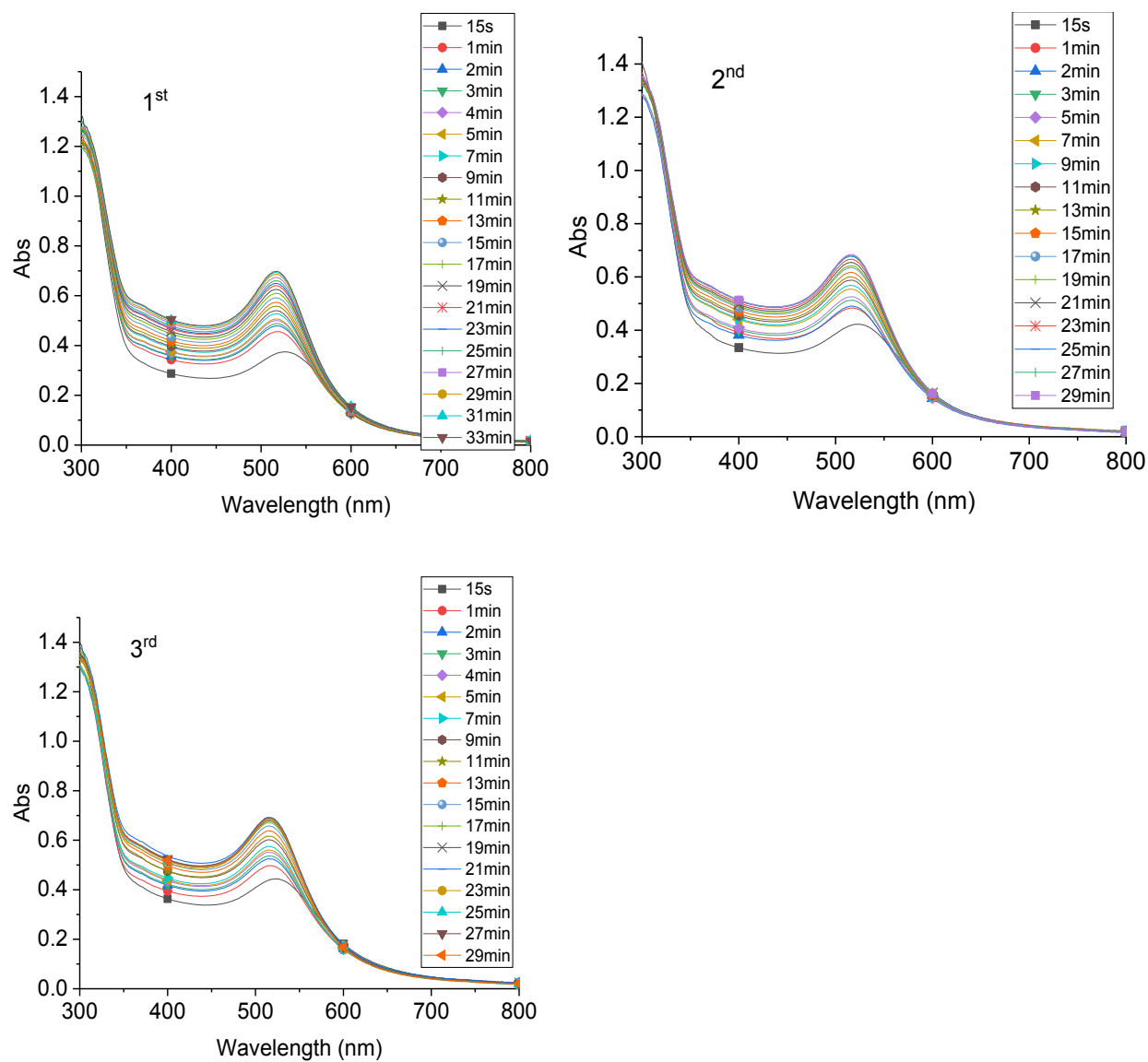


Figure S4: The reproducibility of the reduction of $\text{H[AuCl}_4\text{]}^-$ by DC^{2-} . Conditions: $[\text{H[AuCl}_4\text{]}^-]$: 0.25 mM, $[\text{DC}^{2-}]$: 2.5 mM, pH: 4.9, 80 °C.