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Mechanistic study

1. Sulfinate anion capture experiment



Typical procedures:

fac-Ir(ppy)₃ (0.001 mmol, 1 mol%), *N*-arylacrylamide **1a** (0.2 mmol, 2.0 equiv.), difluoromethyl 2-pyridyl sulfone **2** (0.1 mmol, 1.0 eq.), Na₂CO₃ (0.2 mmol, 2.0 eq.) and MeI (0.5 mmol, 5.0 eq.) were sequentially weighed into a tube. DMSO (1.0 mL) were added by using a syringe. The mixture was irradiated by a 6W blue LED at room temperature under air atmosphere for 36h. Then the volatile solvent and reagents were removed by rotary evaporation and the residue was purified by silica gel flash chromatography using petroleum ether/EtOAc to provide **3a** as colourless liquid (49% yield), **4a** as colourless liquid (26% yield) and 2-(methylsulfonyl)pyridine as white solid (64% yield). The trapping product 2-(methylsulfonyl)pyridine could also be observed with 65% yield if removing *N*-arylacrylamide **1a** from the reaction system.

2-(methylsulfonyl)pyridine:¹HNMR (600 MHz, Chloroform-*d*) δ 8.74 (d, J = 4.6 Hz, 1H), 8.10 (d, J = 7.8 Hz, 1H), 7.98 (td, J = 7.7, 1.7 Hz, 1H), 7.57 (dd, J = 6.6, 4.7 Hz, 0H), 3.24 (s, 3H).¹³C NMR (151 MHz, Chloroform-*d*) δ 158.04, 150.11, 138.31, 127.48, 121.13, 40.04. HRMS (ESI): m/z = 158.0270, calcd. For C₆H₈NO₂S ([M+H⁺]) 158.0276.

2. Radical Inhibition Experiments



Typical procedures:

fac-Ir(ppy)₃ (0.001 mmol, 1 mol%), *N*-arylacrylamide **1a** (0.2 mmol, 2.0 equiv.), difluoromethyl 2-pyridyl sulfone **2** (0.1 mmol, 1.0 eq.), Na₂CO₃ (0.2 mmol, 2.0 eq.) and TEMPO (0.2 mmol, 2.0 eq.) were sequentially weighed into a tube. DMSO (1.0 mL) were added by using a syringe. The mixture was irradiated by a 6W blue LED at room temperature under air atmosphere for 36h (monitored by ¹⁹F NMR using PhCF₃ as the internal standard).



3. Luminescence Quenching Experiments

To verify the oxidative quenching cycle, we conducted the luminescence quenching experiments.

PerkinElmer LS55 fluorescence spectrophotometer was used for the luminescence quenching experiment. Samples with *fac*-Ir(ppy)₃ were irradiated at 380 nm and emission was detected at 540 nm. Sample with Na₂CO₃ were stirred for 10 min and filtrated with a syringe filter (pore size = $0.22 \ \mu$ m) before the luminescence measurement. I₀ is the luminescence intensity without the quencher, I is the intensity with the quencher.



Fluorescence-emission of a 6.7×10^{-5} M solution of *fac*-Ir(ppy)₃ (without the quencher) in DMSO, I₀ = 985.70.



Fluorescence-emission of a 6.7×10^{-5} M solution of *fac*-Ir(ppy)₃ with *N*-arylacrylamide **1a** (6.7×10^{-3} M) in DMSO, I = 953.88.



Fluorescence-emission of a 6.7×10^{-5} M solution of *fac*-Ir(ppy)₃ with Na₂CO₃ (1.3 × 10^{-2} M) in DMSO/H₂O (4:1), I = 989.78.



Fluorescence-emission of a 6.7×10^{-5} M solution of *fac*-Ir(ppy)₃ with diffuoromethyl 2-pyridyl sulfone **2** in DMSO.



These results suggested difluoromethyl 2-pyridyl sulfone **2** could be used as an efficient quencher to $[Ir]^{III*}$, and the reaction proceeds through an oxidative quenching cycle, which is consistent with the reduction potential of difluoromethyl 2-pyridyl sulfone **2** (the first reduction potential: -1.50 V vs. SCE) and *fac*-Ir(ppy)₃ (E_{1/2}^{IV/*III}=-1.73 V Vs SCE in CH₃CN).

4. H₂¹⁸O-labeling experiment

In order to verify H₂O providing the new oxygen of quinoline-2,4-diones in this reaction, an H₂¹⁸O-labeling experiment was performed using THF (anhydrous)/H₂¹⁸O (v/v, 4:1) as the solvent. The mixture was irradiated by a 6W blue LED at room temperature under air atmosphere for 12h. Then the volatile solvent and reagents were removed by rotary evaporation and the residue was purified by silica gel flash chromatography using petroleum ether/EtOAc to afford the desired product **5a**' with 60% yield.







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