

Electronic Supporting Information

Synthesis of Fe₃C@porous Carbon Nanorods via carbonization for Oxygen Reduction Reaction and Zn-air Battery

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Figure S1-S4.

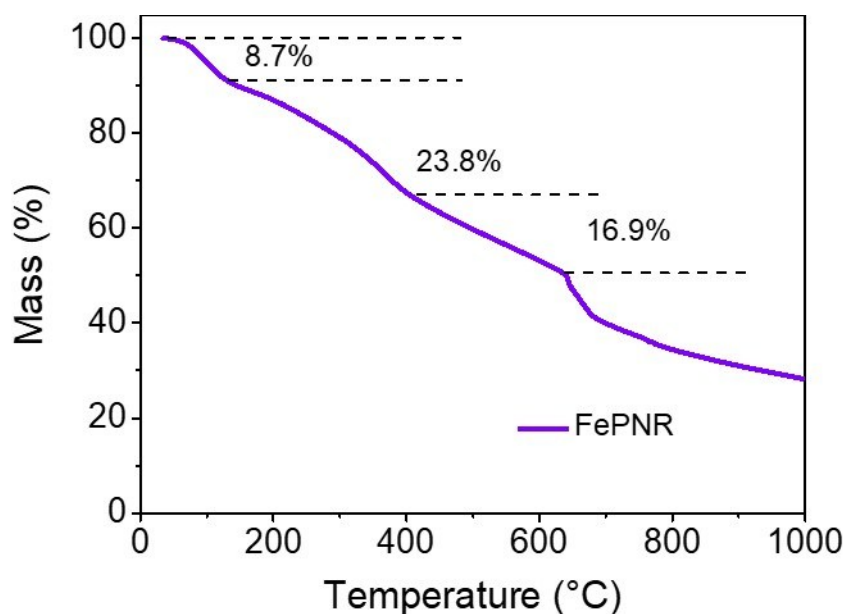


Figure S1. TG analysis of the FePNR.

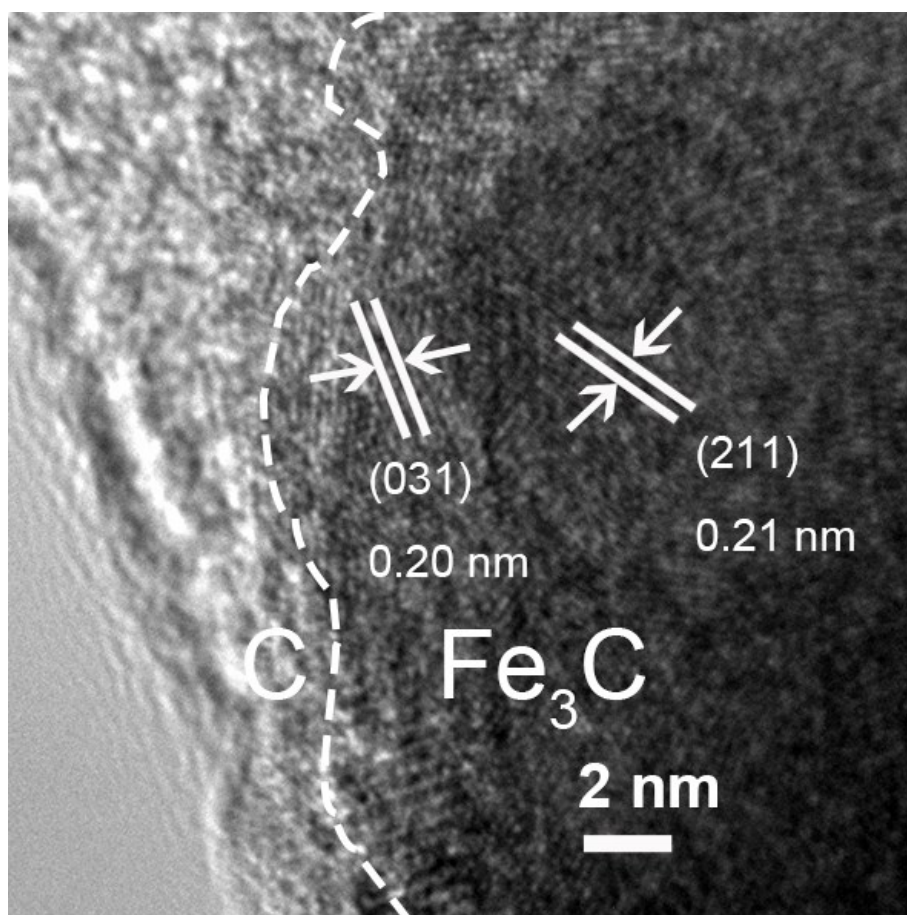


Figure S2. The HRTEM image of $\text{Fe}_3\text{C}/\text{C}$ hetero-structures.

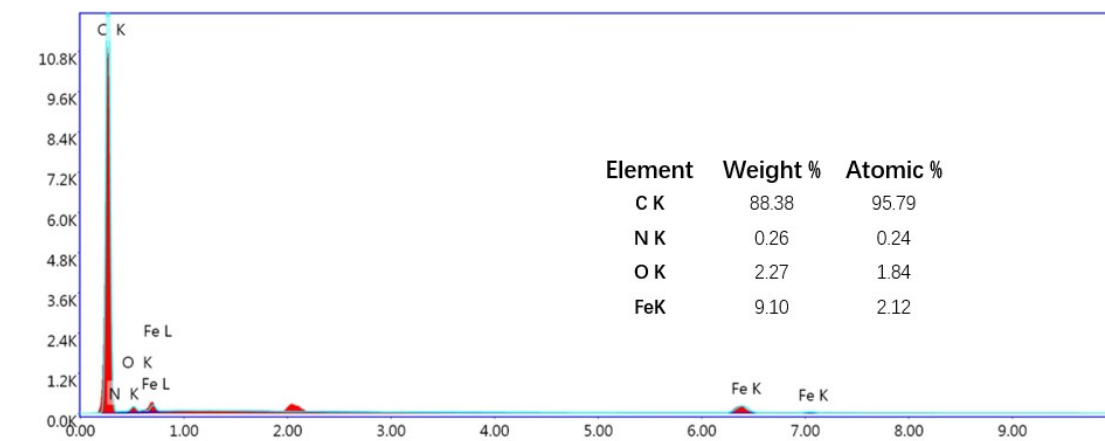


Figure S3. The EDX spectra of as-prepared of FeCNR-750

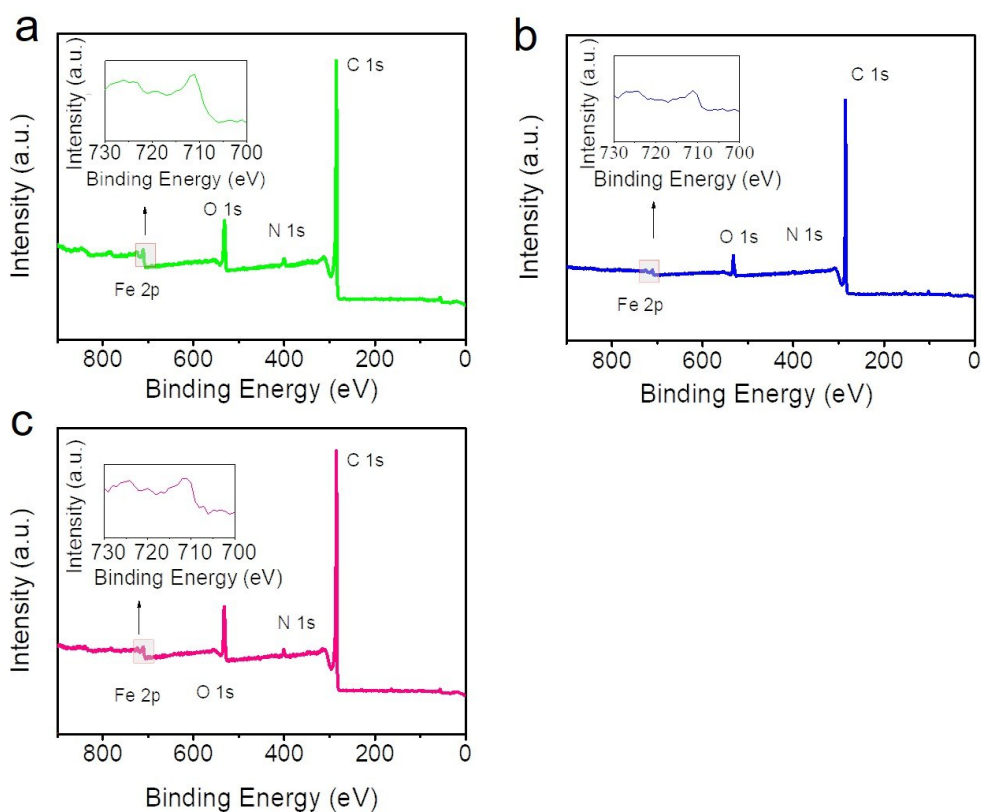


Figure S4. XPS survey spectrum of the FeCNR-700, FeCNR-750 and FeCNR-800.

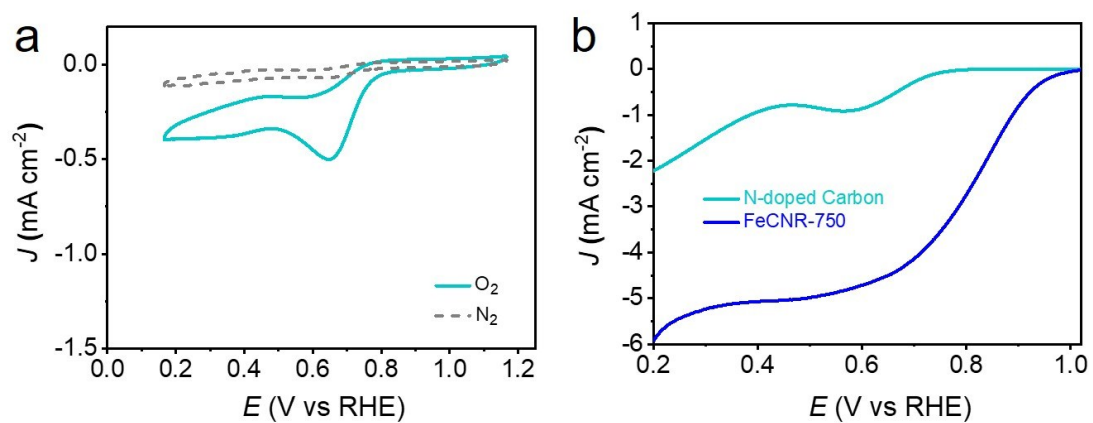


Figure S5. (a) The CV plot and (b) LCV curve of as-prepared N-doped carbon prepared by following the method of synthesizing FeCNRs without adding $\text{Fe}(\text{acac})_3$.

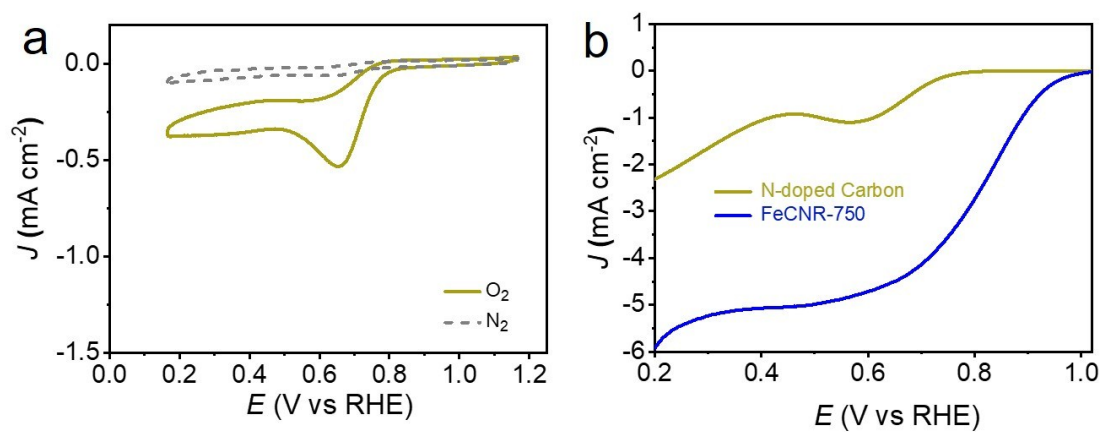


Figure S6. (a) The CV plot and (b) LCV curve of as-prepared N-doped carbon prepared by removing the Fe₃C nanopartilces.

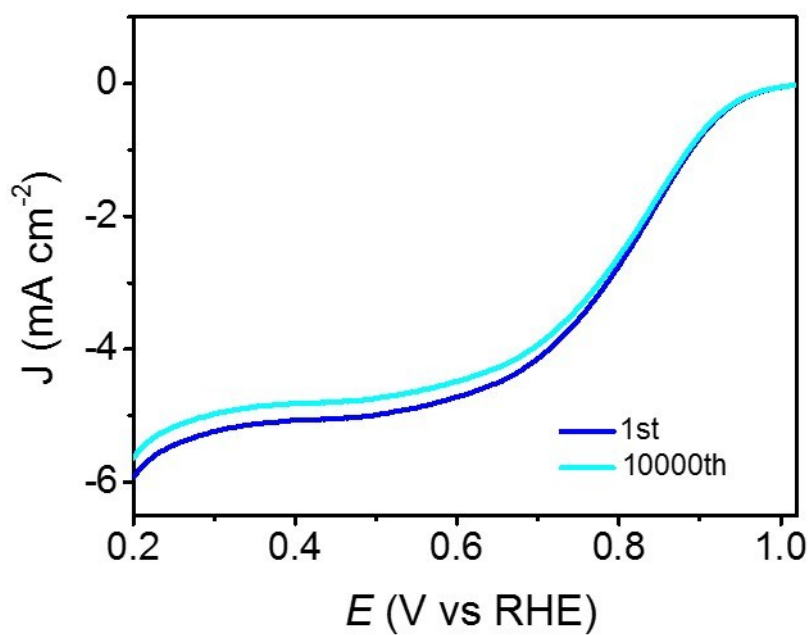


Figure S7. LSV curves of FeCNR-750 in 0.1 M KOH at 1600 rpm before and after 10000 cycles.

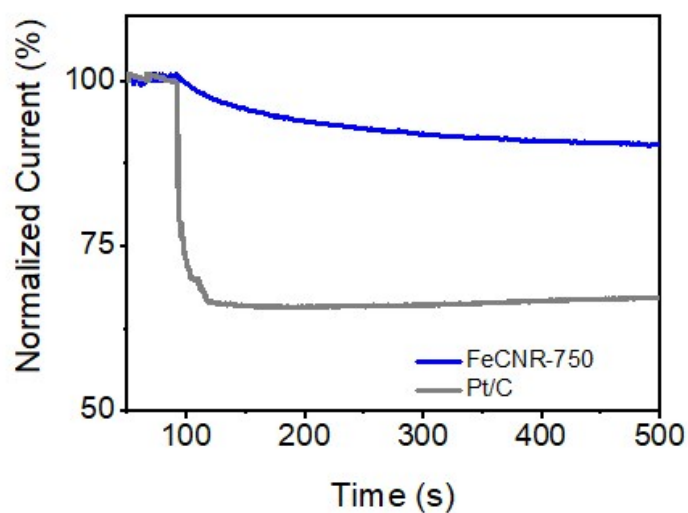


Figure S8. The methanol tolerance of FeCNR-750.

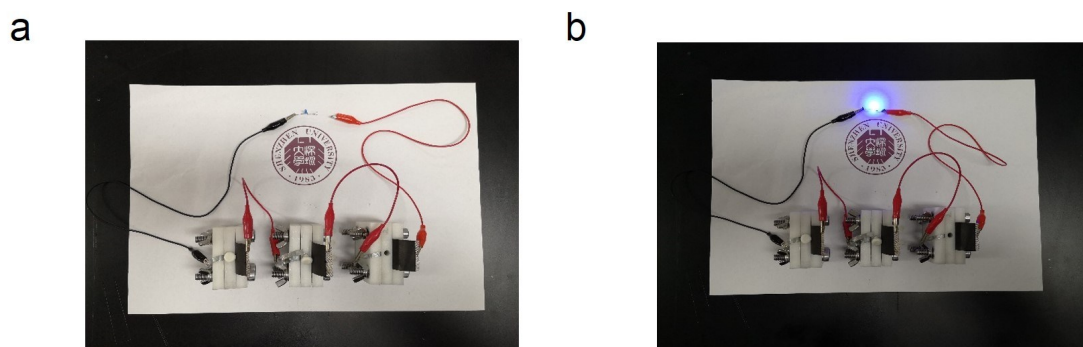


Figure S9. There is a photograph of three Zn-air batteries connected in series with FeCNR-750 as the catalysts before (a) and after (b) connecting.

Table S1. Comparison of properties of Fe derivatives for Zn-air batteries in alkaline electrolytes

Catalysts	Mass Loading mg cm^{-2}	Specific Capacity $\text{mAh} \cdot \text{g}^{-1}$ Zn	Maximum Power Density $\text{mW} \cdot \text{cm}^{-2}$	Open Circuit Voltage V	Reference
FeCNR-750	1	825.77	126.4	1.42	This work
Fe@C–NG/NCNTs	1	682.6	101.3	1.37	¹
FeNC-850		790.0	186	1.46	²

Fe–N–CNBs-600		880	257	1.53	3
F _{0.2} N _{0.2} M _{0.2} -900		637.4	95.3	1.414	4
Fe–Phen–N-800	0.5	708		1.55	5
HP-Fe-N/CNFs		701	135	1.42	6
Fe-N ₄ SAs/NPC			232		7
Fe-NC-900-M-AW			271	1.5	8
Fe-N-CNFs		614			9
N–P–Fe–C		625		1.52	10

References

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