Inorganic oxide composites as high temperature CO₂ sorbents with

enhanced cycle stability

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Supporting information



Figure S1. XRD after TGA of CMC-1(black), CMC-2(blue) and CMC-3 (red) MgO (■), CaO (*)



Figure S2. The CO₂ sorption cycles of CMO-1, from 1-23rd cycle



Figure S3. The CO_2 sorption cycles of CMO-2, from 1-23rd cycle



Figure S4. The CO₂ sorption cycles of CMO-3, from 1-23rd cycle



Figure S5. The CO₂ sorption cycles of CaO-CaCO₃, from 1-23rd cycle



Figure S6. The CO_2 sorption cycles of CaO-HPACC, from 1-23rd cycle



Figure S7. First (upper) and second (lower) order kinetic plot of CO_2 uptake on CMO-1 during the first (black) and 23^{rd} cycle (red)



Figure S8. First (upper) and second (lower) order kinetic plot of CO_2 uptake on CMO-2 during the first (black) and 23^{rd} cycle (red)



Figure S9. First (upper) and second (lower) order kinetic plot of CO_2 uptake on CMO-3 during the first (black) and 23^{rd} cycle (red)



Figure S10. First (upper) and second (lower) order kinetic plot of CO_2 uptake on CaO-CaCO₃ during the first (black) and 23rd cycle (red)



Figure S11. First (upper) and second (lower) order kinetic plot of CO_2 uptake on CaO-HPACC during the first (black) and 23rd cycle (red)



Figure S12. TGA curves of ACMO-1 (magenta), ACMO-2 (cyan), ACMO-3 (blue), ACMO-4 (red) and ACMO-5 (black)



Figure S13. The CO_2 sorption cycles of ACMO-1, from 1 to 23^{rd} cycle



Figure S14. The CO₂ sorption cycles of ACMO-2, from 1 to 23rd cycle



Figure S15. The CO_2 sorption cycles of ACMO-3, from 1 to 23^{rd} cycle



Figure S16. The CO_2 sorption cycles of ACMO-4, from 1 to 23^{rd} cycle



Figure S17. The CO_2 sorption cycles of ACMO-5, from 1 to 23^{rd} cycle



Figure S18. First (upper) and second (lower) order kinetic plot of CO_2 uptake on ACMO-1 during the first (black) and 23^{rd} cycle (red)



Figure S19. First (upper) and second (lower) order kinetic plot of CO_2 uptake on ACMO-1 during the first (black) and 23^{rd} cycle (red)



Figure S20. First (upper) and second (lower) order kinetic plot of CO_2 uptake on ACMO-3 during the first (black) and 23^{rd} cycle (red)



Figure S21. First (upper) and second (lower) order kinetic plot of CO_2 uptake on ACMO-4 during the first (black) and 23^{rd} cycle (red)



Figure S22. First (upper) and second (lower) order kinetic plot of CO_2 uptake on ACMO-5 during the first (black) and 23^{rd} cycle (red)

Figure S23. EDX mapping of the Al, Ca and Mg on ACMC-4 and ACMO-4

Figure S24. The CO_2 sorption cycles of ACMO-4, from 1 to 100^{th} cycle

Figure S25. The CO₂ sorption cycles of ACMO-4, from 1 to 100^{th} cycle

Figure S26. XRD before and after 100 adsorption cycles MgO (■), CaO (*)

Figure S27. X-ray photoelectron spectra of a, b) ACMO-4 before CO₂ uptake experiments c,d) ACMO-4 after 23 cycles and e,f) ACMO-4 after 39 cycles, the Al 2p peak is shown in A,C and E, the Ca 2p peak is shown in B,D and F

The XPS spectra of ACMO-4 after 23 and 39 CO_2 cycles are shown in Figure 27a-b. The Al 2p spectra (Figure S27a) clearly demonstrated that only one type of Al existed on ACMO-4 before CO_2 sorption cycles. The changes in the Ca 2p and Al 2p peak after 23 and 39 cycles (Figure 27 c-f) indicated that the coordination around the Ca and Al had changed. New additional peaks could be fitted to the respective XPS spectra. This change corresponded to the partial formation of mayenite. The

formation of a small amount of may enite on ACMO-4 after \sim 30-40 CO₂ cycles was revealed by XRD and shown in Figure S26.

Figure S28. The CO₂ sorption cycles from a gas mixture (20% CO₂, 80% N₂) of ACMO-4 (top) and CaCO₃ (bottom), from 1 to 23^{rd} cycle, the same gas mixture was used during CO₂ uptake at 650 °C and desorption at 850 °C

Figure S29. The CO₂ sorption cycles of ACMO-4, from 1 to 7^{th} cycle where the CO₂ gas was mildly hydrated by bubbling through 5 water traps before being exposed to the ACMO-4 sorbent