

Supporting Information

Two-step synthesis of a red-emissive warped nanographene derivative via a ten-fold C–H borylation

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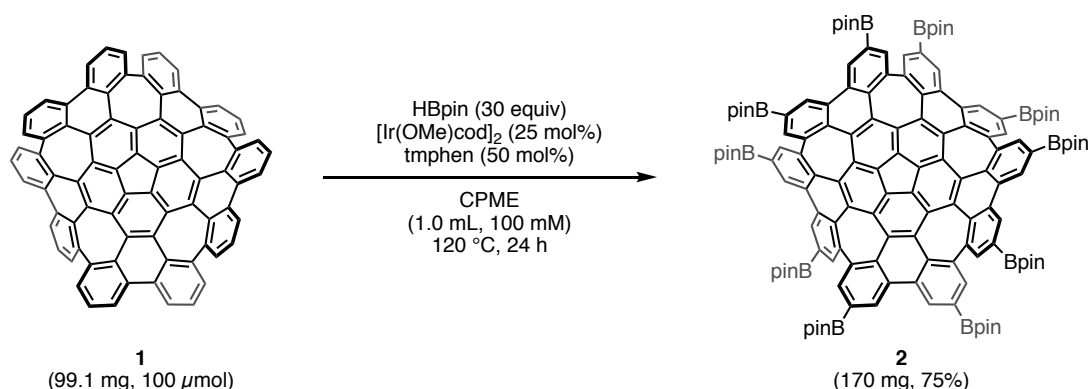
1. Experimental Section

General

Unless otherwise noted, all materials including dry solvent were obtained from commercial suppliers and used without further purification. All reactions were performed with dry solvents under an atmosphere of nitrogen in dried glassware with standard vacuum-line techniques. Work-up and purification procedures were carried out with reagent-grade solvents under air. Warped nanographene (**1**)^{S1} was synthesized according to the reported procedure.

Analytical thin-layer chromatography (TLC) was performed using E. Merck silica gel 60 F254 precoated plates (0.25 mm). The developed chromatograms were analyzed by UV lamp (254 or 365 nm). Flash column chromatography was performed with KANTO Silica Gel 60N (spherical, neutral, 40-100 μm). The high-resolution mass spectra (HRMS) were obtained from a JEOL JMS-S3000 SpiralTOF (MALDI-TOF MS). Melting points were measured on a MPA100 Optimelt automated melting point system. Preparative recycling gel permeation chromatography (GPC) was performed with a SHIMADZU Prominence high performance liquid chromatograph instrument equipped with JAIGEL-2H-40/JAIGEL-2H-40 columns using chloroform as an eluent. Nuclear magnetic resonance (NMR) spectra were recorded on a JEOL ECA600II spectrometer with UltraCoolTM probe (¹H 600 MHz, ¹³C 150 MHz) or a JEOL ECA600 (¹⁹F 565 MHz). Chemical shifts for ¹H, ¹³C, and ¹⁹F NMR are expressed in parts per million (ppm) relative to C₂H₂DCl₄ (δ 5.98 ppm), C₂D₂Cl₄ (δ 73.79 ppm), and C₆F₆ (δ –163 ppm), respectively. Data are reported as follows: chemical shift, multiplicity (s = singlet, m = multiplet, br = broad) and integration.

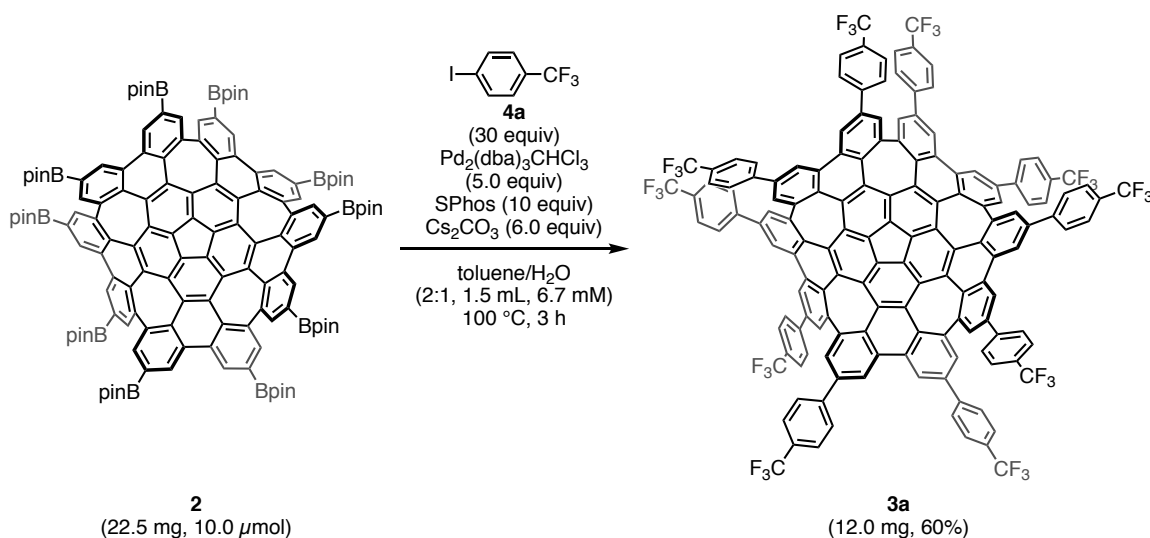
Synthesis of 2



To a 20 mL J-Young Schlenk equipped with a magnetic stirring bar were added a solution of **1** (99.1 mg, 100 μ mol, 1.0 equiv) in dry cyclopentyl methyl ether (CPME; 1.0 mL), 4,4,5,5-tetramethyl-1,3,2-dioxaborolane (HBpin; 384 mg, 3.00 mmol, 30 equiv), [IrOMe(cod)]₂ (16.6 mg, 25.0 μ mol, 25 mol%), and 3,4,7,8-tetramethyl-1,10-phenanthroline (tmphen; 11.8 mg, 50.0 μ mol, 50 mol%) under nitrogen atmosphere. The Schlenk was sealed with a J-Young screw tap and the resultant mixture was stirred at 120 °C for 24 h. After cooling the mixture to room temperature, the reaction mixture was diluted in chloroform and passed short-path silica gel chromatography with chloroform. The solution was removed under reduce pressure. The crude material was purified by GPC (eluent: chloroform) to afford **2** (170 mg, 75% yield) as an orange solid.

¹H NMR (600 MHz, C₂D₂Cl₄, 140 °C) δ 9.08 (s, 10H), 7.85 (s, 10H), 1.46 (s, 120H); ¹³C NMR (150 MHz, C₂D₂Cl₄, 140 °C) δ 138.2 (4°), 136.8 (CH), 136.0 (4°), 133.0 (4°), 130.8 (4°), 129.2 (CH), 126.9 (4°), 83.83 (4°), 24.74 (CH₃). HRMS (MALDI-TOF MS) *m/z* calcd for C₁₄₀H₁₄₀B₁₀O₂₀ [M]⁺: 2251.0970, found: 2251.0963; mp: >300 °C.

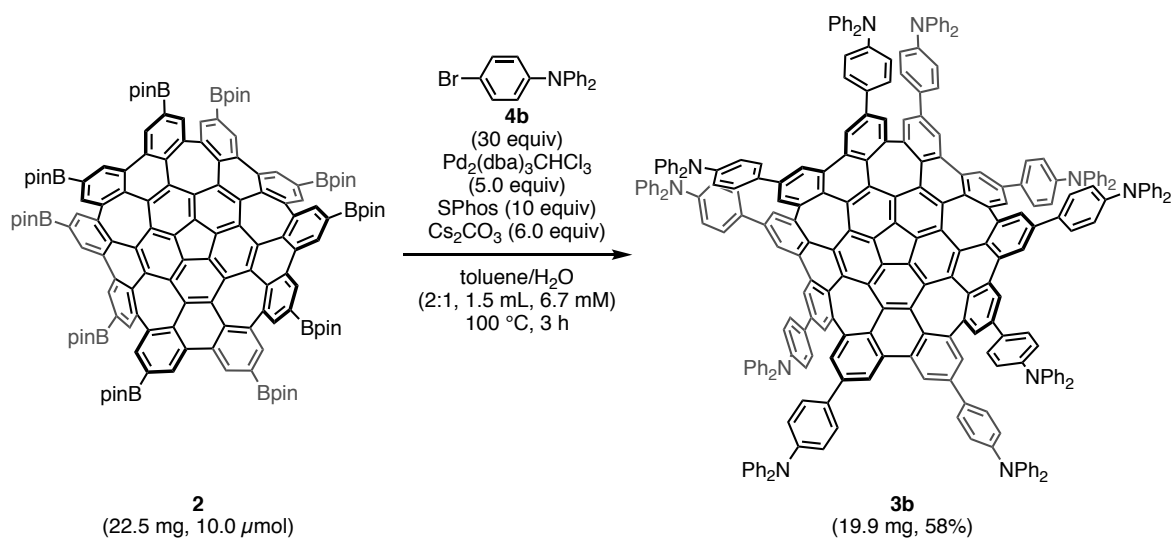
Synthesis of **3a**



To a 3 mL screw-capped pressure vessel equipped with a magnetic stirring bar were added a solution of $\text{Pd}_2(\text{dba})_3\cdot\text{CHCl}_3$ (51.8 mg, 50.0 μmol , 5.0 equiv), 2-dicyclohexylphosphino-2',6'-dimethoxybiphenyl (SPPhos; 41.1 mg, 100 μmol , 5.0 equiv) in toluene (1.0 mL), **2** (22.5 mg, 10.0 μmol , 1.0 equiv), 1-iodo-4-(trifluoromethyl)benzene (**4a**; 68.7 mg, 300 μmol , 30 equiv), and a solution of Cs_2CO_3 (19.6 mg, 60.0 μmol , 6.0 equiv) in H_2O (0.50 mL) under nitrogen atmosphere. The vessel was sealed with a PTFE-coated screw cap and the resultant mixture was stirred at 100 $^\circ\text{C}$ for 3 h under nitrogen atmosphere. After cooling the mixture to room temperature, the reaction mixture was extracted with chloroform. The combined organic layer was dried over MgSO_4 , and the solvent was removed under reduce pressure. The crude material was purified by silica-gel column chromatography (eluent: hexane/EtOAc = 6:1) and GPC (eluent: chloroform) to afford **3a** (12.0 mg, 60%) as a reddish orange solid.

^1H NMR (600 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 $^\circ\text{C}$) δ 7.80–8.00 (br m, 50H), 8.91 (br s, 10H); ^{13}C NMR (150 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 85 $^\circ\text{C}$) δ 125.9 (CH), 127.6 (CH); ^{19}F NMR (565 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 $^\circ\text{C}$) δ -63.7 (s, 30F); HRMS (MALDI-TOF MS) m/z calcd for $\text{C}_{150}\text{H}_{60}\text{F}_{30}$ [$\text{M}\cdot$] $^+$: 2430.4216, found: 2430.4290; mp: >300 $^\circ\text{C}$.

Synthesis of 3b



To a 3 mL screw-capped pressure vessel equipped with a magnetic stirring bar were added a solution of $\text{Pd}_2(\text{dba})_3\cdot\text{CHCl}_3$ (51.8 mg, 50.0 μ mol, 5.0 equiv), SPhos (41.1 mg, 100 μ mol, 5.0 equiv) in toluene (1.0 mL), **2** (22.5 mg, 10.0 μ mol, 1.0 equiv), 4-bromo-*N,N*-diphenylaniline (**4b**; 92.3 mg, 300 μ mol, 30 equiv), and a solution of Cs_2CO_3 (19.6 mg, 60.0 μ mol, 6.0 equiv) in H_2O (0.50 mL) under nitrogen atmosphere. The vessel was sealed with a PTFE-coated screw cap and the resultant mixture was stirred at 100 °C for 3 h under nitrogen. After cooling the mixture to room temperature, the reaction mixture was extracted with CH_2Cl_2 . The combined organic layer was dried over MgSO_4 , and the solvent was removed under reduce pressure. The crude material was purified by silica-gel column chromatography (eluent: hexane/EtOAc = 6:1 and chloroform) and GPC (eluent: chloroform) to afford **3b** (19.9 mg, 58%) as a red solid.

^1H NMR (600 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 °C) δ 7.08–7.30 (br m, 120H), 7.76–7.85 (br m, 30H), 8.87 (br s, 10H); ^{13}C NMR (150 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 85 °C) δ 123.1 (CH), 123.4 (CH), 124.6 (CH), 129.2 (CH), 145.1 (CH), 147.4 (CH); HRMS (MALDI-TOF MS) m/z calcd for $\text{C}_{260}\text{H}_{160}\text{N}_{10}$ [$\text{M}\cdot$] $^+$: 3423.2895, found: 3423.2880; mp: >300 °C.

2. X-ray Crystallography

Details of the crystal data and a summary of the intensity data collection parameters for **2** are listed in Table S1. A suitable crystal was mounted with mineral oil on a MiTeGen MicroMeshes and transferred to the goniometer of a Rigaku PILATUS diffractometer. Graphite-monochromated Mo K α radiation was used. The structures were solved by direct methods with (SIR-97)^{S2} and refined by full-matrix least-squares techniques against F^2 (SHELXL-2014/7)^{S3} by using Yadokari-XG software package.^{S4} The intensities were corrected for Lorentz and polarization effects. The non-hydrogen atoms were refined anisotropically. Hydrogen atoms were placed using AFIX instructions. CCDC 1893066 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif. Alert A (VERY LARGE Solvent Accessible VOID(S) in Structure) is caused by the disorder of solvent molecules. Suitable solvent molecules could not be placed because of the lack of observed electron density.

Table S1. Crystallographic data and structure refinement details of **2**

	2 ·1.5C ₄ H ₈ O
formula	C ₁₄₆ H ₁₅₅ B ₁₀ O _{21.5}
fw	2361.79
<i>T</i> (K)	123(2)
λ (Å)	0.71073
cryst syst	Triclinic
space group	<i>P</i> -1
<i>a</i> (Å)	17.1883(7)
<i>b</i> (Å)	19.8237(8)
<i>c</i> (Å)	25.0041(10)
α (deg)	112.264(4)
β (deg)	93.732(3)
γ (deg)	109.808(4)
<i>V</i> (Å ³)	7232.6(6)
<i>Z</i>	2
<i>D</i> _{calc} (g·cm ⁻³)	1.084
μ (mm ⁻¹)	0.070
F(000)	2506
cryst size (mm)	0.10 × 0.10 × 0.10
θ range (deg)	1.790–25.000
reflns collected	105367
indep reflns/ <i>R</i> _{int}	25435/ 0.2479
params	1653
GOF on <i>F</i> ²	1.064
<i>R</i> ₁ , <i>wR</i> ₂ [<i>I</i> > 2 σ (<i>I</i>)]	0.1266, 0.3072
<i>R</i> ₁ , <i>wR</i> ₂ (all data)	0.3042, 0.4142

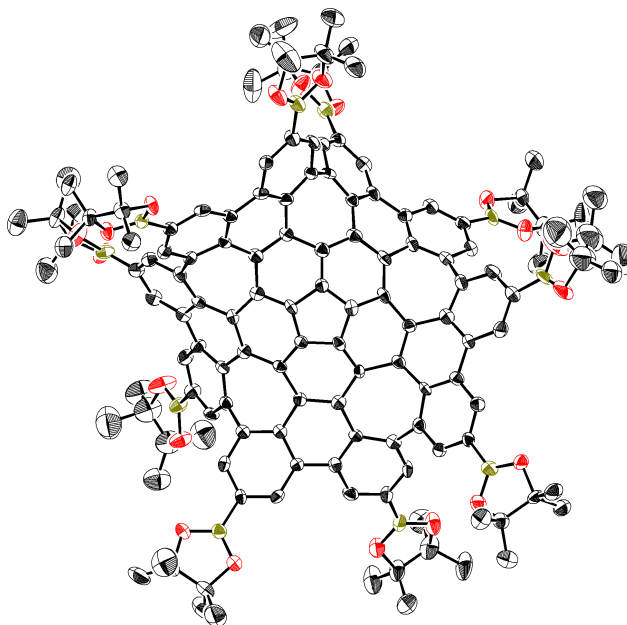


Figure S1. ORTEP drawing of **2** with 50% probability. All hydrogen atoms and solvent molecules (Et₂O) are omitted for clarity.

3. Photophysical measurement

UV–Vis absorption spectra were recorded on a Shimadzu UV-3510 spectrometer with a resolution of 0.5 nm. Emission spectra were measured on Shimadzu RF-6000 spectrometer with a resolution of 0.4 nm. Absolute fluorescence quantum yields (Φ_F) were determined on a Shimadzu RF-6000 using a calibrated integrating sphere system upon excitation at 340 nm. For FL lifetime measurements, Hamamatsu Photonics Quantaaurus-Tau® fluorescence lifetime spectrometer C11367-21 with LED as a light source was used.

Table S2. The fluorescence quantum yield (Φ_F), lifetime (τ), radiative (k_r) and nonradiative (k_{nr}) decay rate constants of **1**, **2**, **3a**, and **3b**.

	Φ_F	τ (ns)	k_r (s ⁻¹) ^a	k_{nr} (s ⁻¹) ^a
1	0.26	6.86	3.8×10^7	1.1×10^8
2	0.20	6.34	3.2×10^7	1.3×10^8
3a	0.28	6.33	4.4×10^7	1.1×10^8
3b	0.47	6.61	7.1×10^7	8.0×10^7

^a $\Phi_F = k_r \times \tau$, $k_r + k_{nr} = \tau^{-1}$

4. Computational Study

The Gaussian16 program^{S5} running on a SGI Altix4700 system was used for optimization (B3LYP/6-31G(d)).^{S6} Structures were optimized without any symmetry assumptions. Zero-point energy, enthalpy, and Gibbs free energy at 298.15 K and 1 atm were estimated from the gas-phase studies. Harmonic vibration frequency calculation at the same level was performed to verify all stationary points as local minima (with no imaginary frequency). Visualization of the results was performed by use of GaussView 5.0.9 software.

Table S3. TD-DFT vertical one-electron excitations of **2**.

Excitation [nm]	Oscillator strength (<i>f</i>)	Description
506	0.131	H→L, H-1→L+1
492	0.004	H-2, H-1→L H→L+1
479	0.005	H-3, H-2, H-1→L
467	0.048	H-3, H-1, H→L H→L

H: HOMO, L: LUMO.

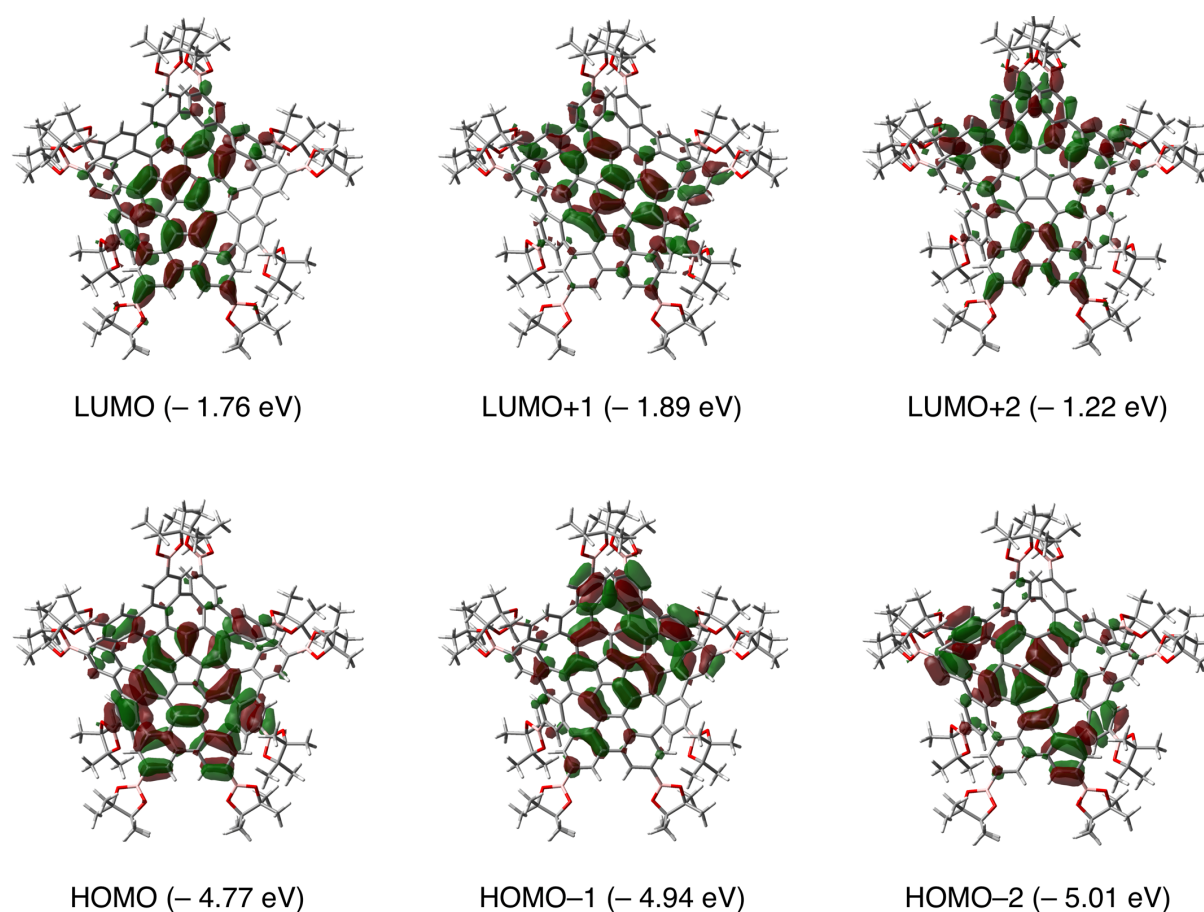


Figure S2. Distributions and energy levels of molecular orbitals of **2** (isovalue = 0.02).

Table S4. TD-DFT vertical one-electron excitations of **3a**

Excitation [nm]	Oscillator strength (<i>f</i>)	Description
515	0.180	H→L H-1→L+1
500	0.004	H-2, H-1→L H→L+1
486	0.008	H-3, H-2, H-1→L
474	0.093	H-3, H-2, H-1, H→L H-2, H-1→L+1

H: HOMO, L: LUMO.

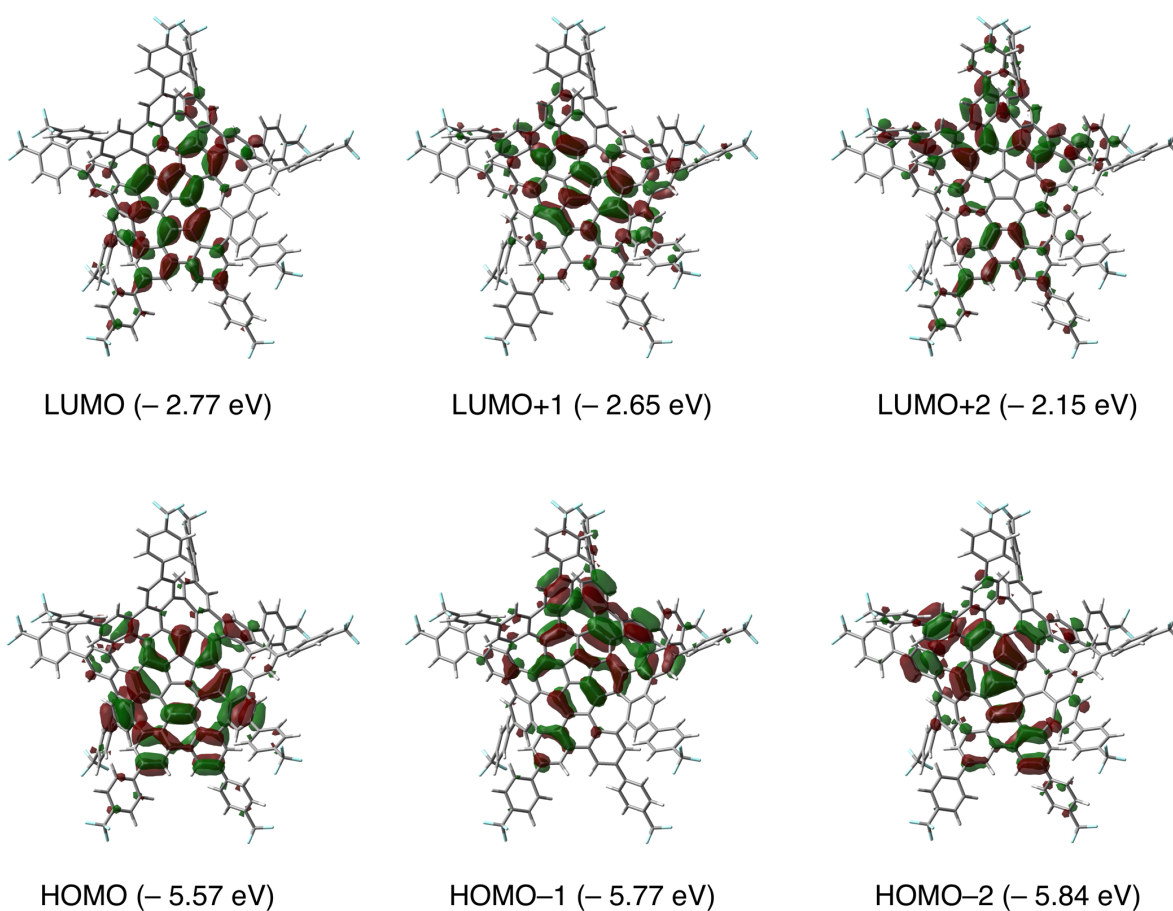


Figure S3. Distributions and energy levels of molecular orbitals of **3a** (isovalue = 0.02).

Table S5. TD-DFT vertical one-electron excitations of **3b**.

Excitation [nm]	Oscillator strength (<i>f</i>)	Description
548	0.321	H-10, H→L H-1→L+1
530	0.018	H-11, H-1→L H→L+1
524	0.011	H-12, H-3, H-2, H-1→L
515	0.033	H-13, H-3, H-2→L H-1→L+1

H: HOMO, L: LUMO.

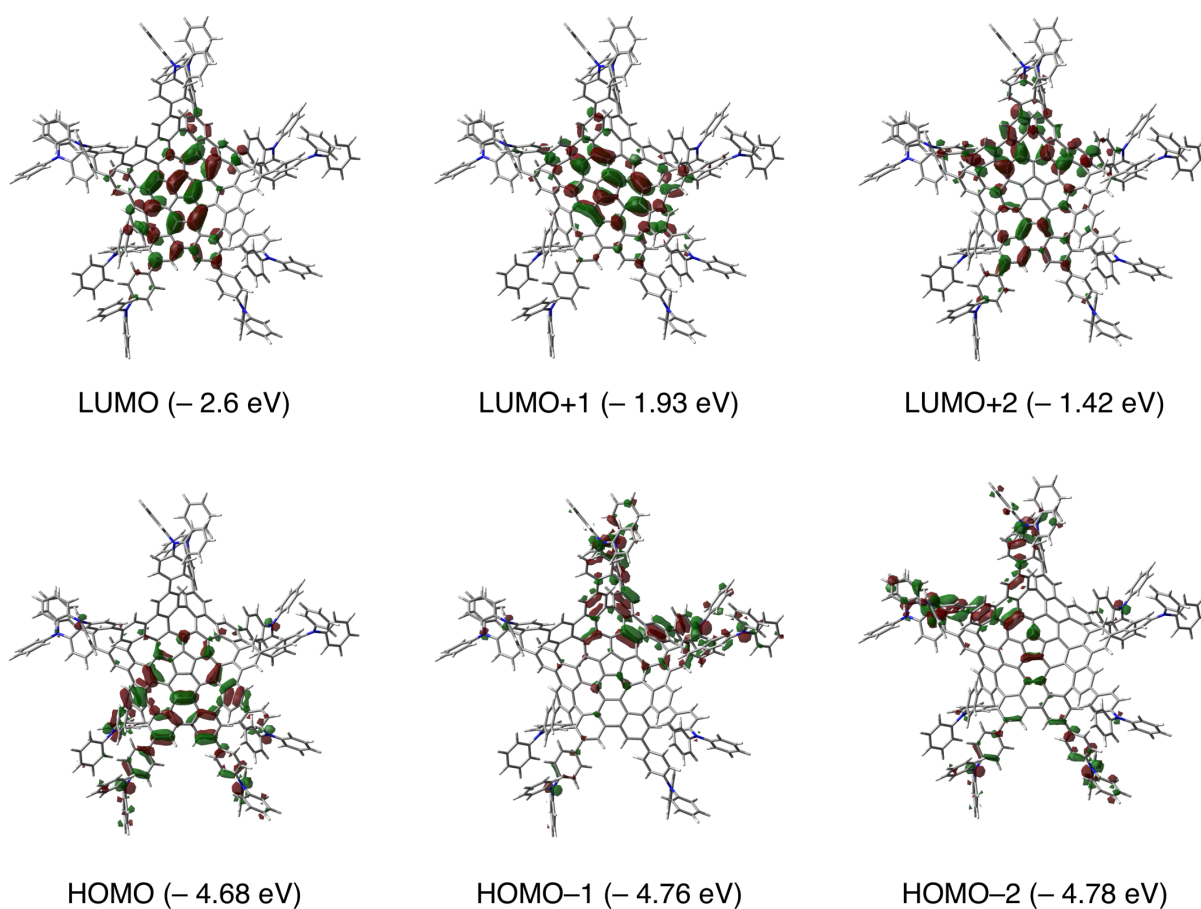


Figure S4. Distributions and energy levels of molecular orbitals of **3b** (isovalue = 0.02).

Supporting Information (Kato, Lin, Kuwayama, Nagase, Segawa, Scott, Itami)
Two-step synthesis of a red-emissive warped nanographene derivative via a ten-fold C–H borylation

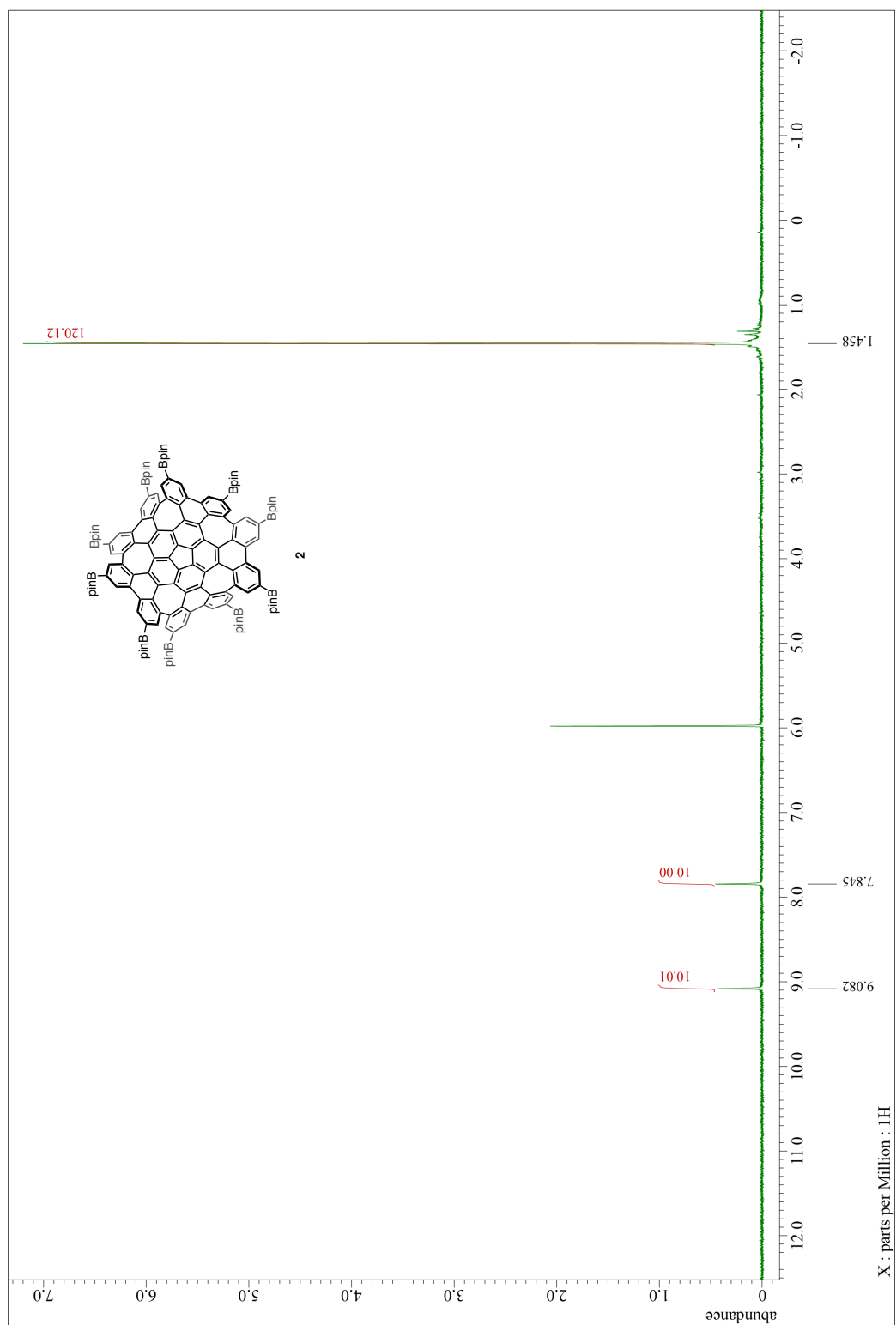
H	11.014459	2.003715	3.084224	H	4.027329	-14.524294	-3.533508	H	1.839334	11.364023	5.468395
H	11.829107	-1.372526	0.551454	C	1.419457	-14.251785	-5.709322	C	4.250490	13.532530	6.513692
C	6.903029	-3.790482	-4.014610	H	-0.289480	-12.967791	-5.416313	H	5.749803	13.073681	7.995598
C	7.443849	-2.819417	-4.875425	H	3.236585	-15.414781	-5.701319	H	2.640872	13.671601	5.083792
C	7.313708	-5.121135	-4.206619	H	1.074886	-14.637033	-6.664745	H	4.598643	14.548477	6.350511
C	8.350954	-3.154994	-5.874291								

5. References

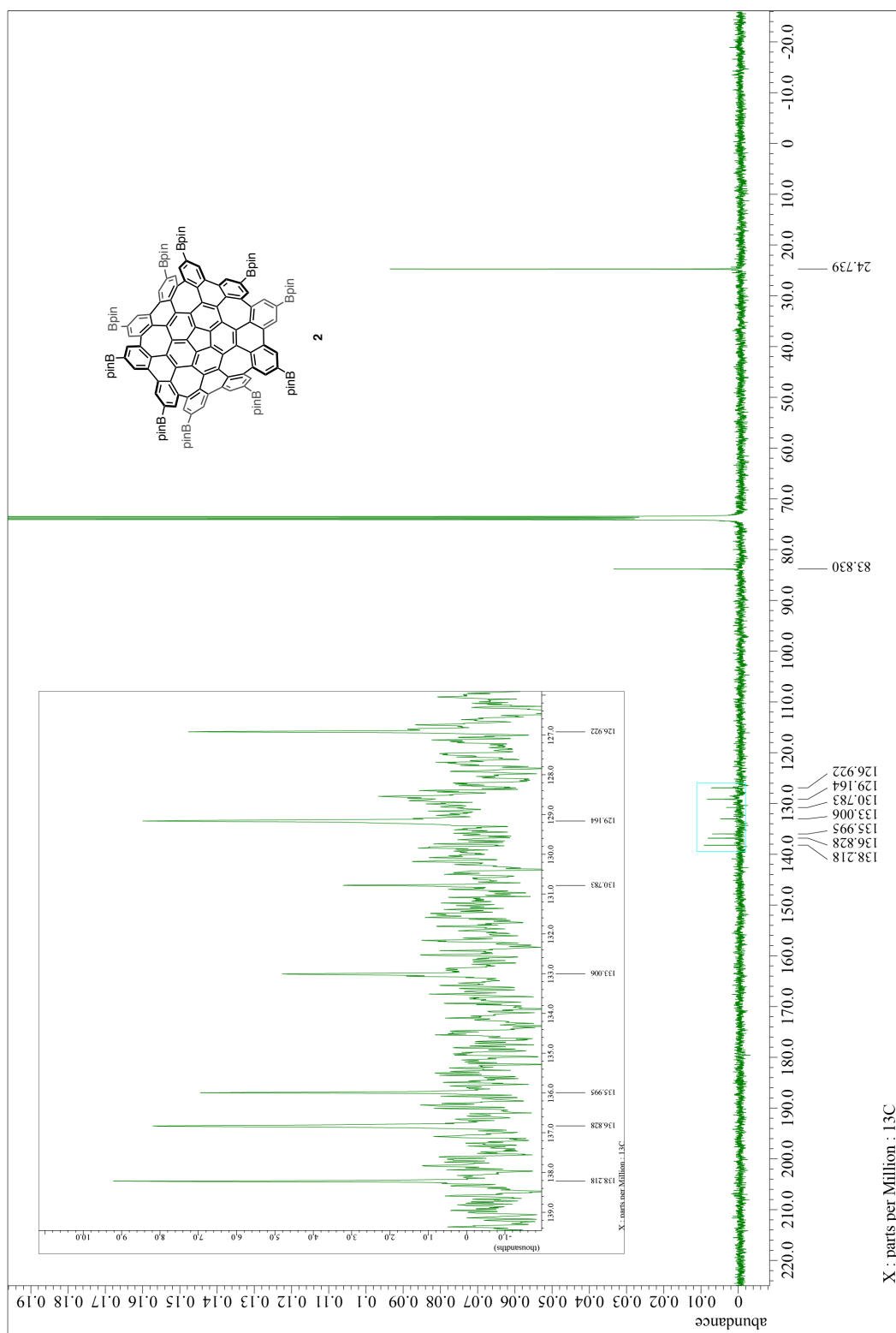
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6. ^1H NMR and ^{13}C NMR Spectra

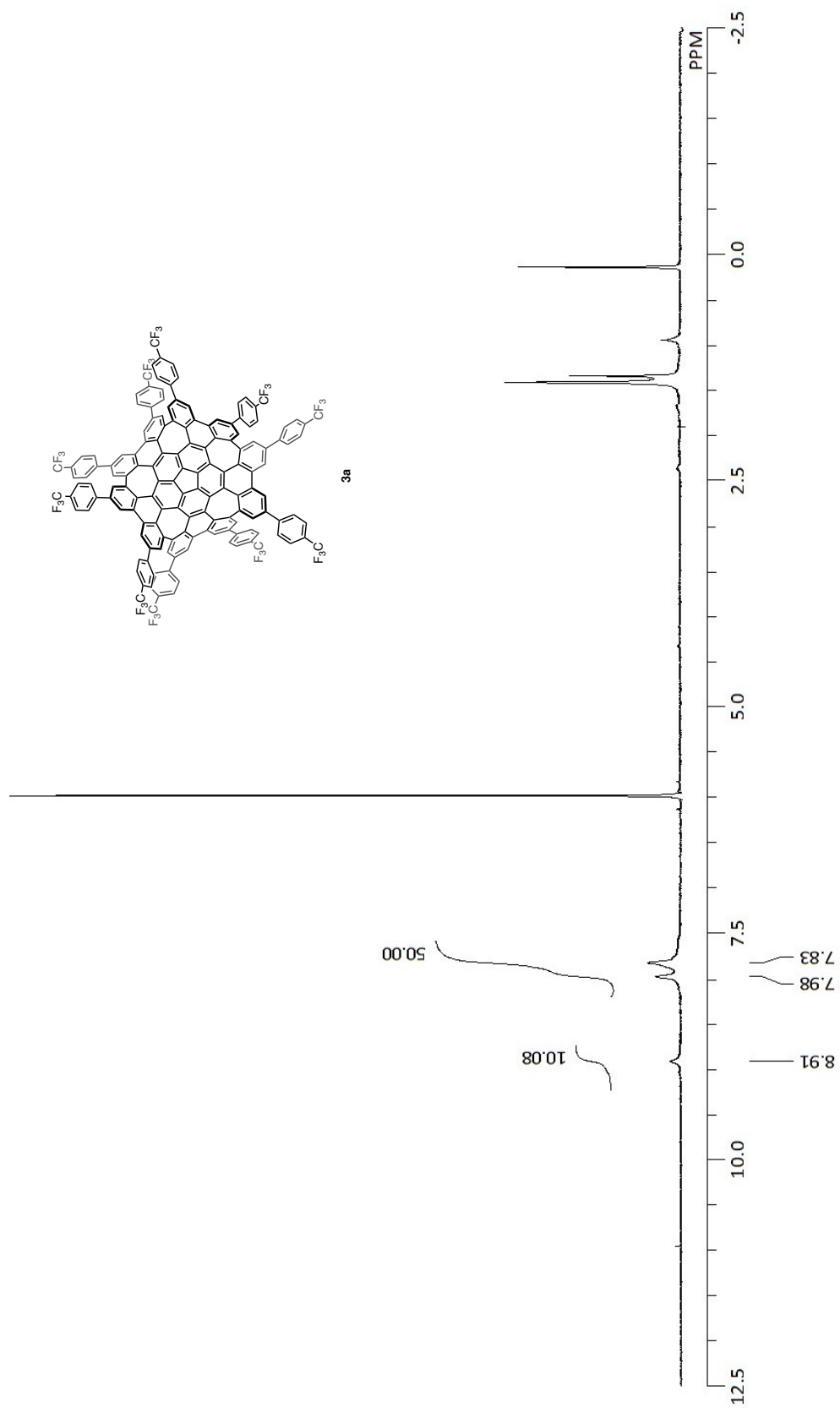
^1H NMR spectrum of **2** (600 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 °C)



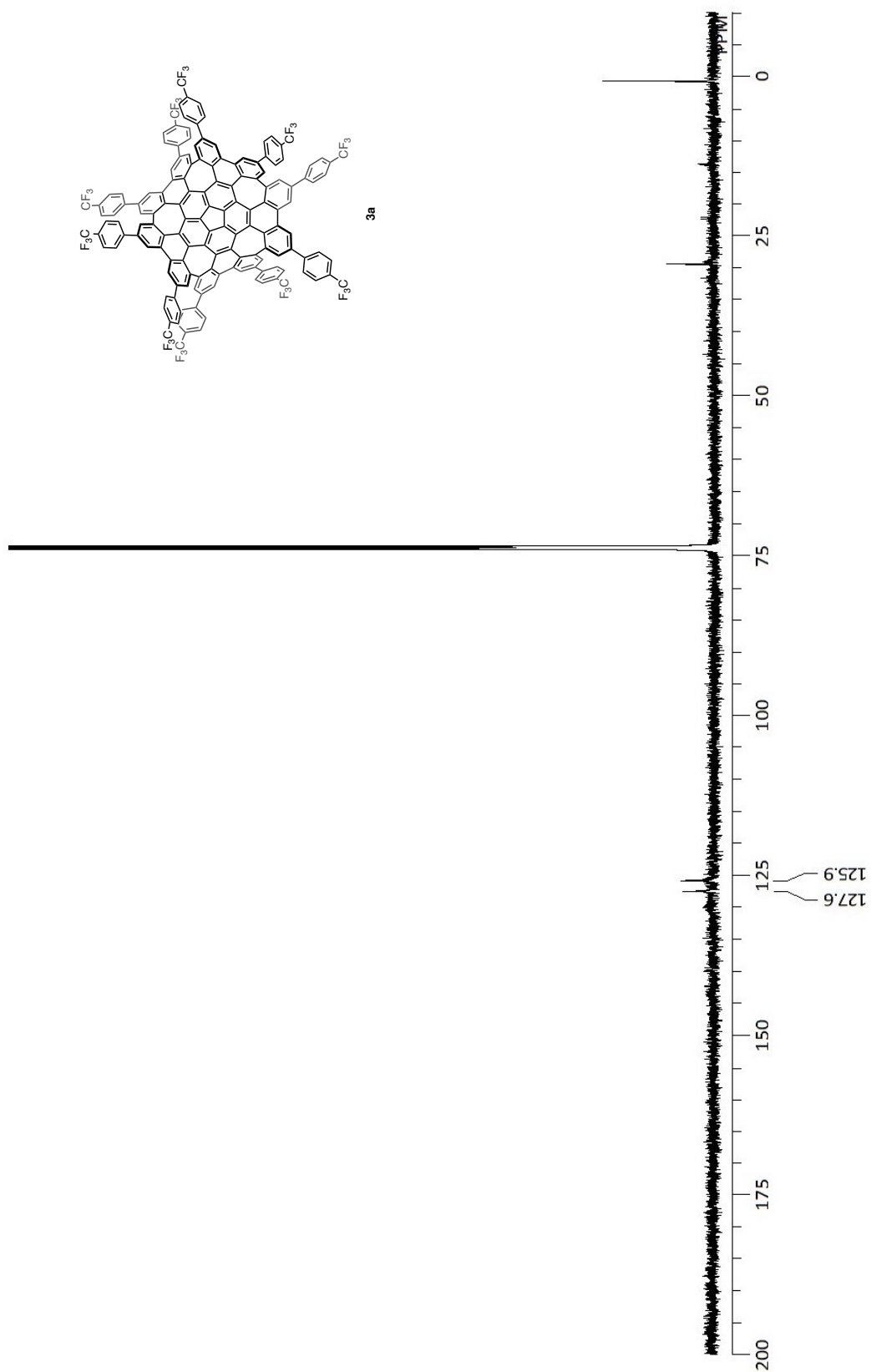
^{13}C NMR spectrum of **2** (150 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 °C)



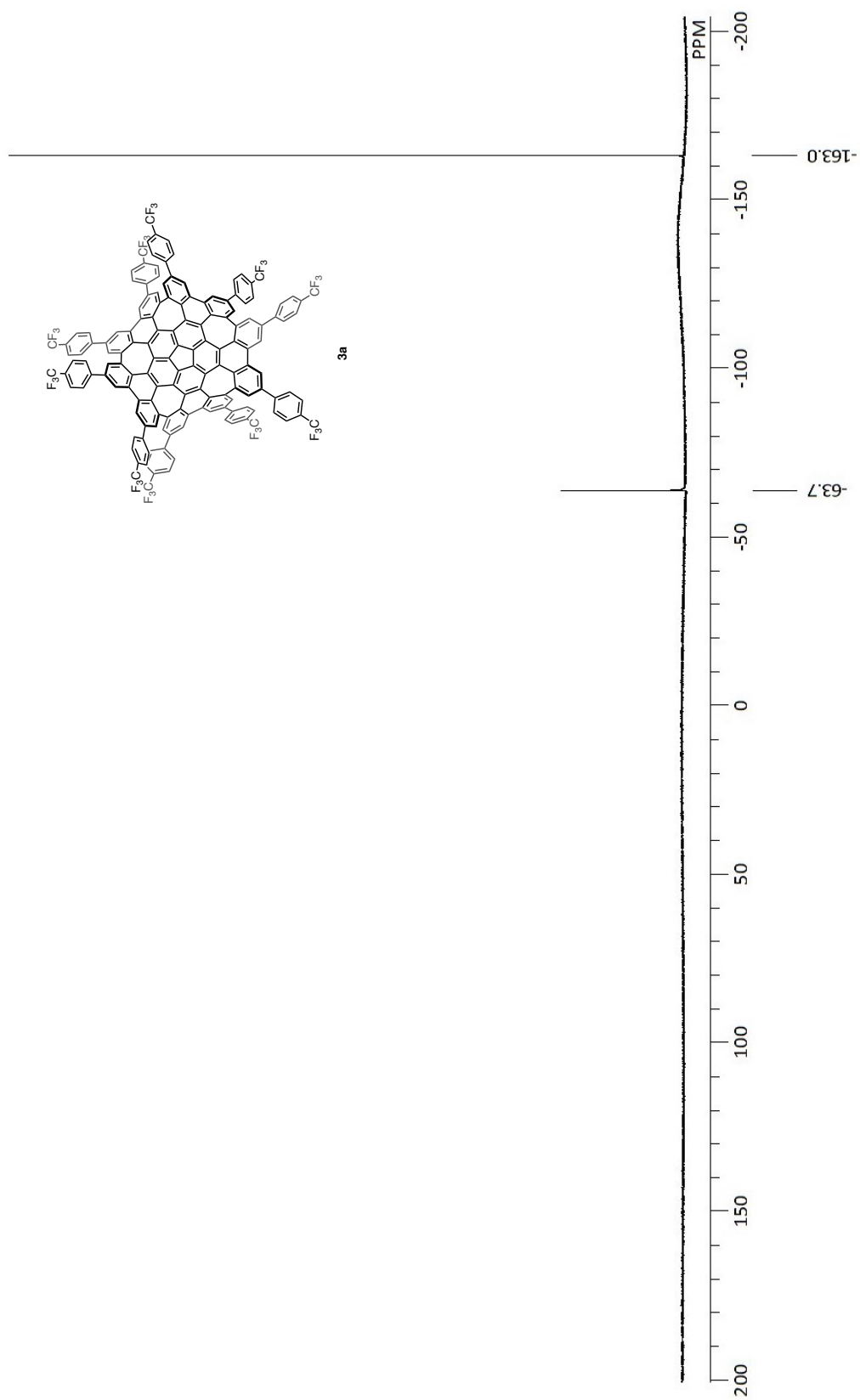
^1H NMR spectrum of **3a** (600 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 °C)



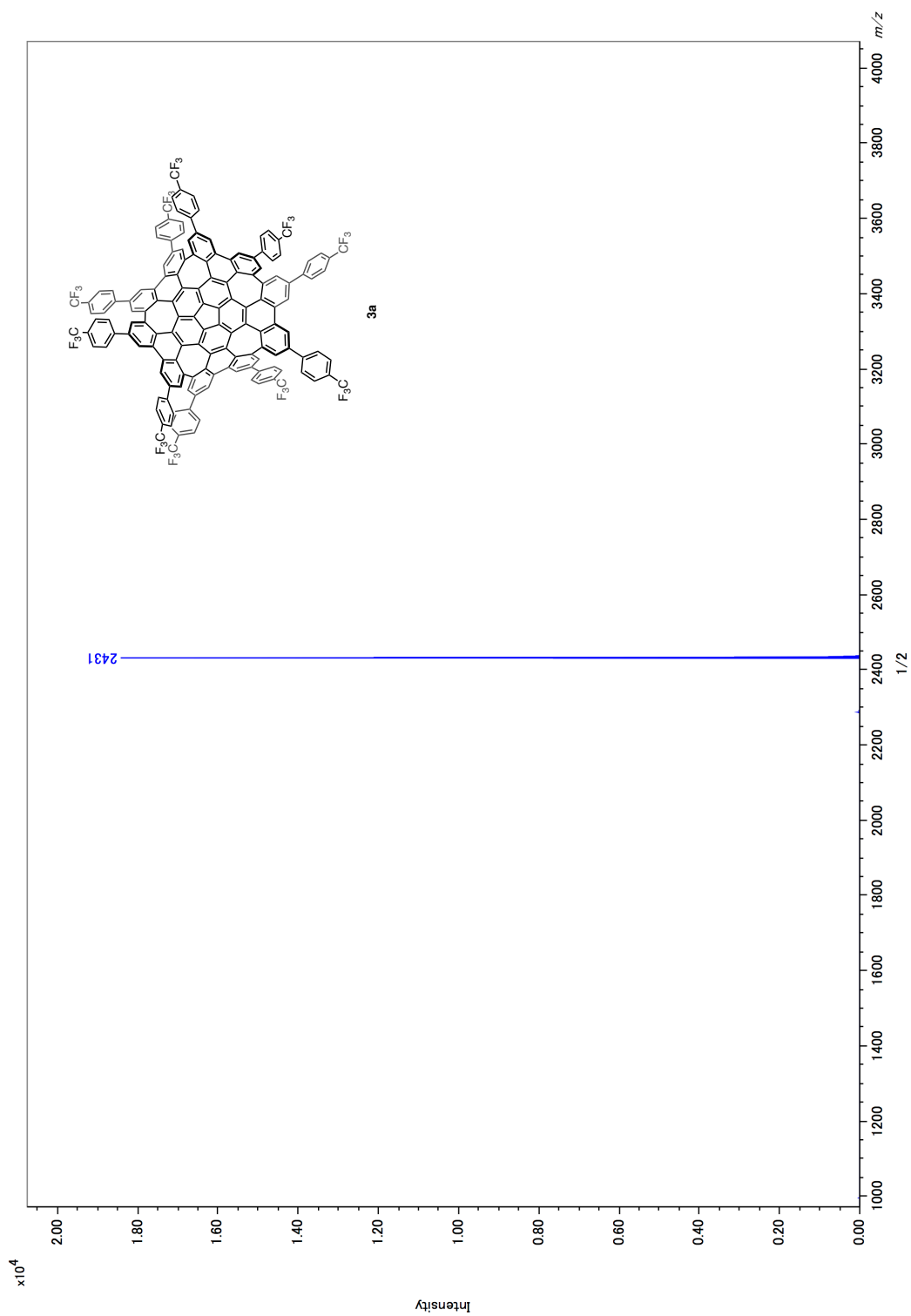
^{13}C NMR spectrum of **3a** (150 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 85 °C)



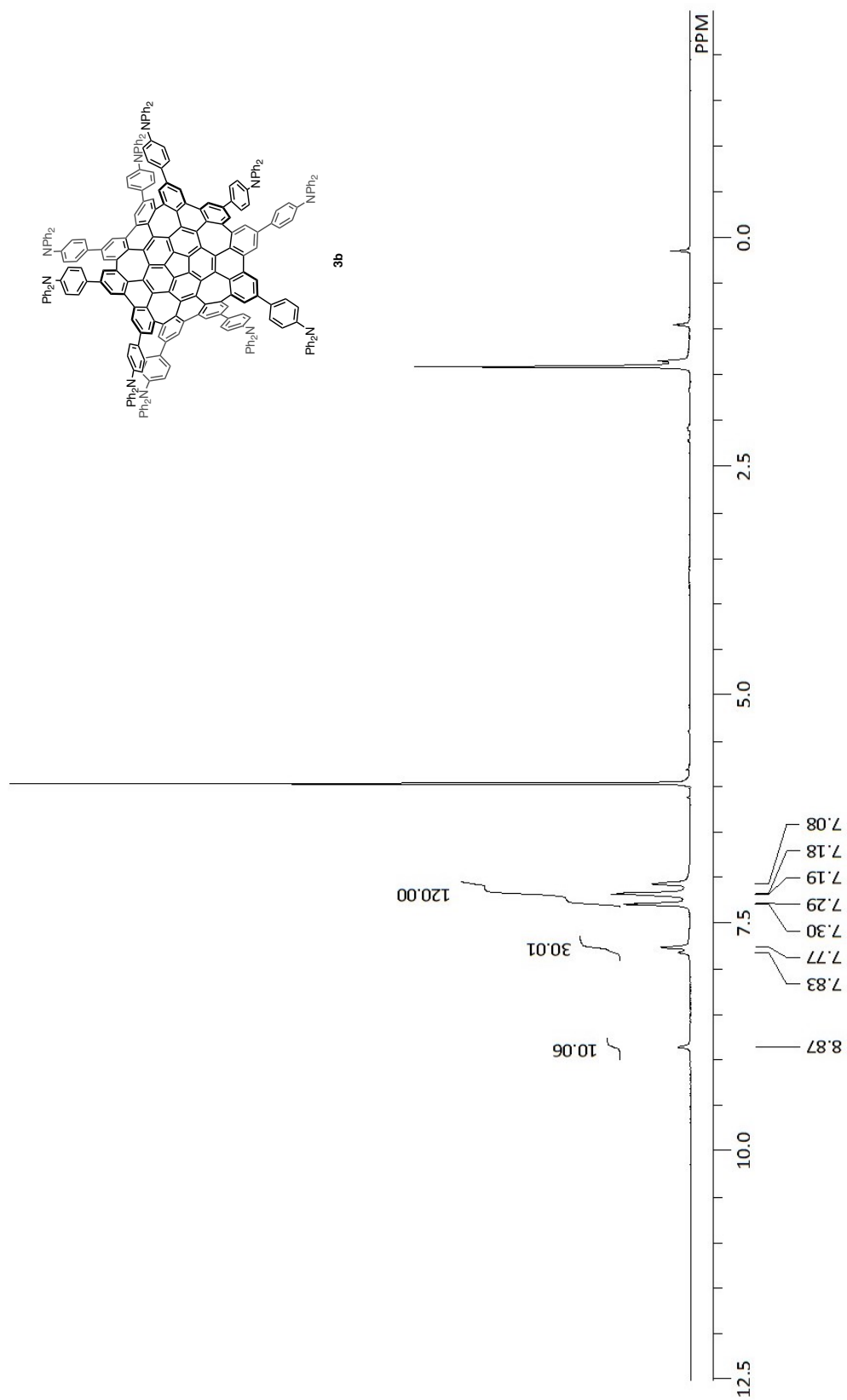
^{19}F NMR spectrum of **3a** (565 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 °C)



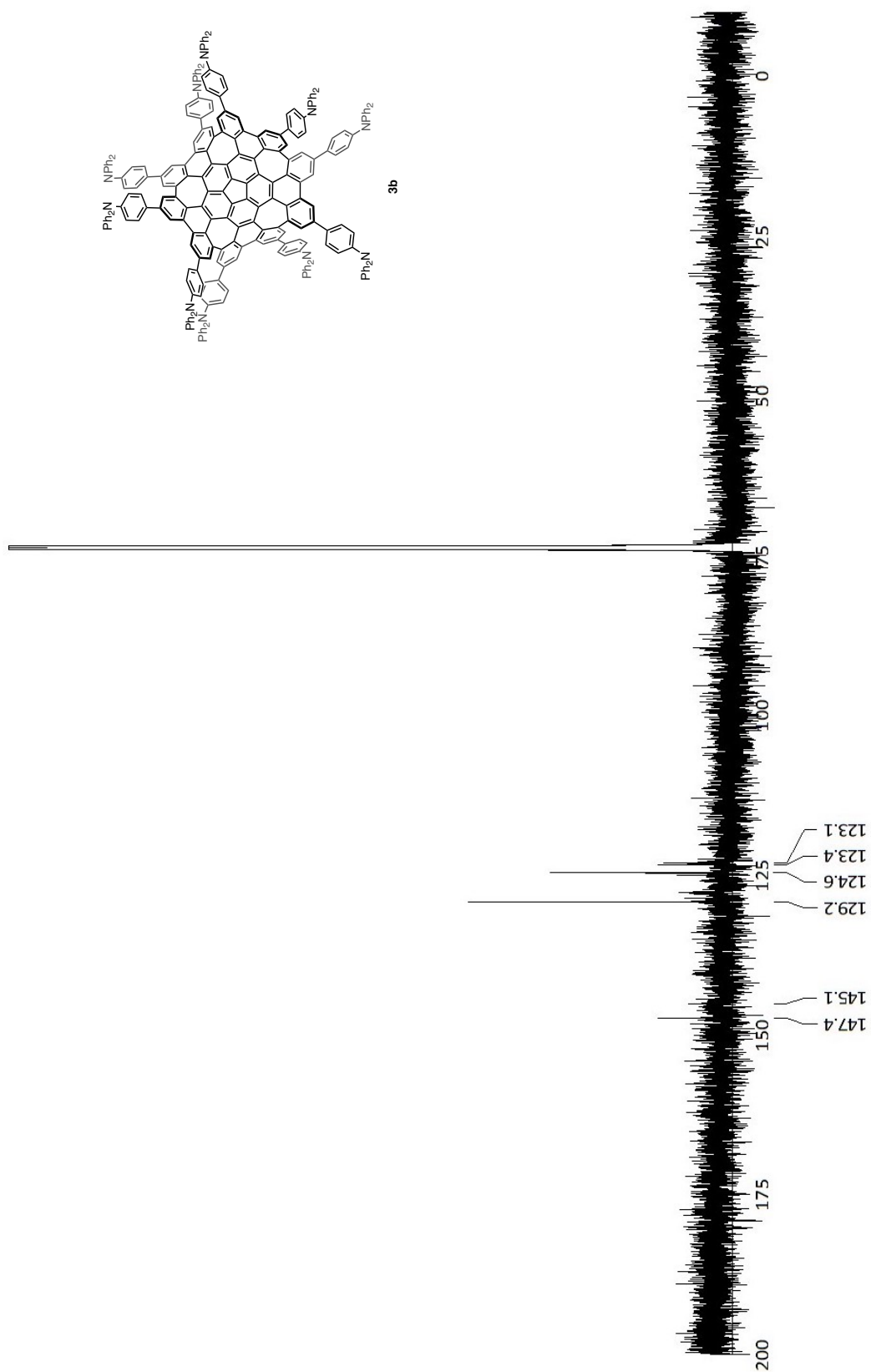
MALDI-TOF mass spectrum of **3a**



^1H NMR spectrum of **3b** (600 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 140 °C)



^{13}C NMR spectrum of **3b** (150 MHz, $\text{C}_2\text{D}_2\text{Cl}_4$, 85 °C)



MALDI-TOF mass spectrum of **3b**

