Supporting Information for

A Facile Solution-chemistry Method for $Cu(OH)_2$ Nanoribbon Arrays with Noticeable Electrochemical Hydrogen Storage Ability at Room Temperature

Peng Gao, [†,*] Milin Zhang, [†] and Quanping Xiao [†]

Key Laboratory of Superlight Materials and Surface Technology, Ministry of Education, Harbin Engineering University, Harbin, Heilongjiang 150001, P. R. China

RECEIVED DATE (automatically inserted by publisher); Email: gaopeng@hrbeu.edu.cn

Figure S1. Schematic showing the coordination assembly growth of $Cu(OH)_2$ nanoribbons.

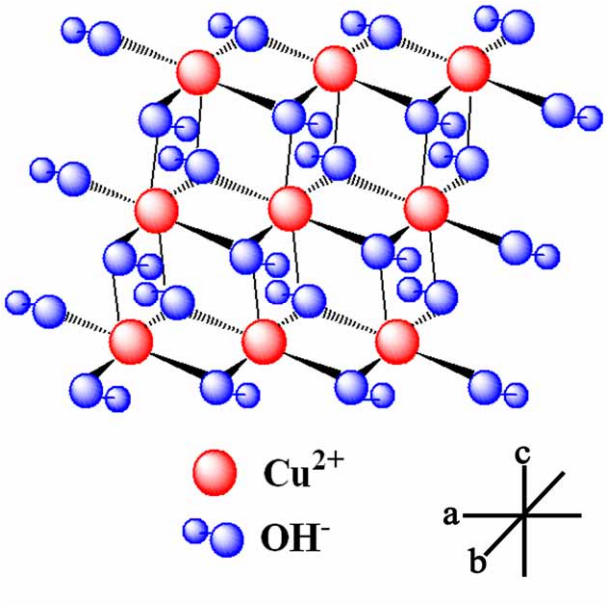

Figure S2. SEM image of the $Cu(OH)_2$ nanowires prepared only in $NH_3 \cdot H_2O$ solutions.

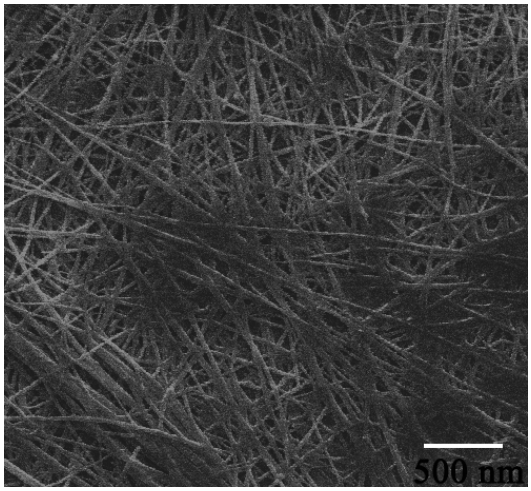

Figure S3. N_2 adsorption-desorption isotherm of the as-obtained $Cu(OH)_2$ nanoribbon arrays; inset, pore-size distribution curve obtained from the desorption data.

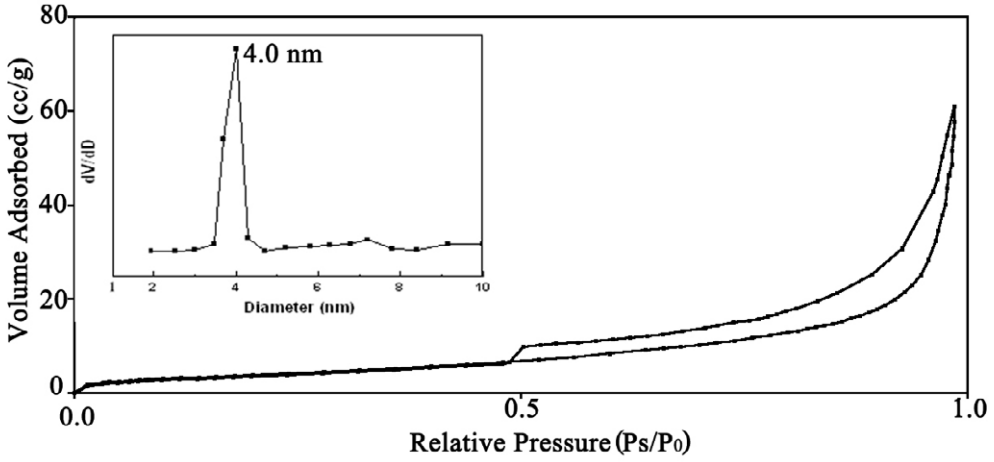

Figure S4. High magnification FE-SEM image of the boundary between the water layer and benzene layer.

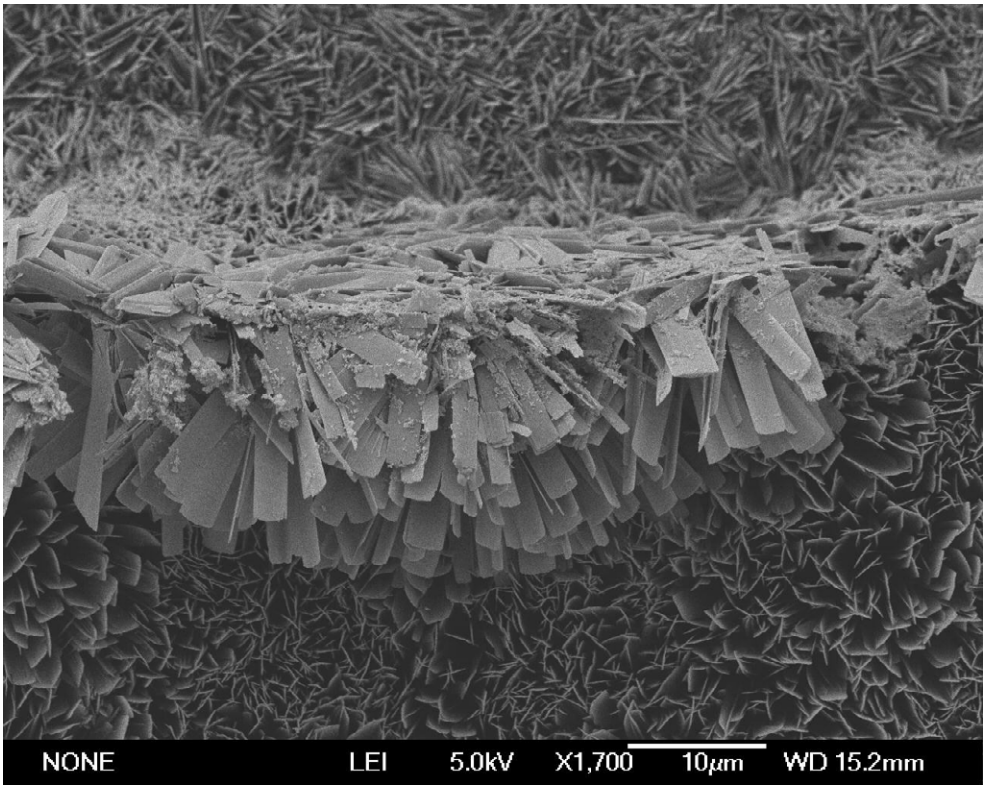

Figure S5. FE-SEM image of the $Cu(OH)_2$ nanoribbon arrays after being cycled 50 times at the charge-discharge current density of 10 mA/g.