

This contradictory information, coupled with the paucity of examples involving the addition of metal-free α -keto radicals, prompted us to study their reactivity in greater detail. At this stage it should be pointed out that the reactions of α -keto radicals promoted by, or carried out in the presence of, a Lewis acid appear to be mechanistically different from the ones reported in the manuscript. Indeed, in these cases, the Lewis acid interacts with the α -keto radical, influencing its stability and reactivity. For example, the simple coordination of the Lewis acid to the carbonyl function perturbs the redistribution of the electron density on the carbonyl group (the influence of the Lewis acids on the reactivity of α -keto radicals is now under investigation) transforming these nucleophilic species into electrophilic ones. For a similar reason, the reactivity of the $\text{PhC(O)CH}_2\cdot$ radicals cannot be equated with the examples presented in the article. In this case, the stability of the α -keto radical is also influenced by the presence of the aromatic ring conjugated to the carbonyl group (see *e.g.* Y. Koga, H. Kusama and K. Narasaka, *Bull. Chem. Soc. Jpn.*, 1998, **71**, 475). Additionally, if BEt_3 is used to generate the α -keto radicals from α -haloketones, the generated radical is captured by another molecule of BEt_3 and the corresponding boron enolate is formed (see *e.g.* K. Nozaki, K. Oshima and K. Utimoto, *Bull. Chem. Soc. Jpn.*, 1991, **64**, 403). Interestingly, α -ester or amide radicals behave as electrophilic radicals and, as expected, react with electron rich partners (see *e.g.* (a) S.-I. Usugi, H. Yorimitsu, H. Shinobo and K. Oshima, *Bull. Chem. Soc. Jpn.*, 2002, **75**, 2049; (b) D. P. Curran and C.-T. Chang, *J. Org. Chem.*, 1989, **54**, 3140; (c) B. Sire, S. Seguin and S. Z. Zard, *Angew. Chem., Int. Ed.*, 1998, **37**, 2864).