Electronic Supplementary Information

New Class of Rhodamine Luminophores: Design, Syntheses and Aggregation-Induced Emission Characetristics

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1. General remarks

Reagents were purchased from Wako, Nacalai Tesque, or TCI Japan. All other solvents were used without further purification. Experiments were conducted under an atmosphere of dry nitrogen or using a guard tube.

¹H-NMR and ¹³C-NMR spectra were recorded using 300 or 500 MHz spectrometers (Varian UNITY INOVA). Solvents used for NMR measurements were CDCl₃, DMSO- d_6 or TFA- d_8 , with TMS as the internal standard. Mass spectra were acquired using a JMX-700 (2) (JEOL Co., Ltd.) MS instrument. Melting points were determined on a Yanagimoto micromelting point apparatus and are uncorrected. Liquid column chromatography was conducted over silica gel (Merck Silica Gel 60 mesh 70-230). Developed TLC plates were visualized under a short-wave UV lamp, by staining with an I₂-SiO₂ mixture, and/or by heating plates that were dipped in ammonium phosphomolybdate sulfate solution. Dry THF was distilled over sodium benzophenone ketyl under argon atmosphere.

UV-vis spectra were collected on a SHIMADZU UV-1700 spectrophotometer at RT using a 1 cm or a 0.1 cm quartz cuvette. Emission spectra were collected on a HITACHI F-4500 fluorescence spectrophotometer. To obtain an accurate spectrum, spectrum correction was carried out with rhodamine B concentrated solution and a secondary-standard light source. In surface photometry, a cut filter was utilized to eliminate multiple lights, such as secondary light, due to the effects of light scattering. Fluorescence spectra of all samples were measured with excitation with 365 nm. Average particle size was measured by dynamic light scattering (ELSZ-2, Otsuka Electronics).

2. Syntheses

2-1. General Procedure for the synthesis of 2

To a solution of phthalic anhydride (1 equiv) and anhydrous $AlCl_3$ (1.1 equiv) in CH_2Cl_2 , **1** (1 equiv) was slowly added and the mixture was stirred under nitrogen at 0 °C for 4 hrs Then, the reaction mixture was poured into a mixture of water/6 M HCl, stirred for 10 min, and extracted with CH_2Cl_2 . The combined organic layers were dried over MgSO₄ and evaporated to give the crude product. Recrystallization from MeOH/H₂O solution of **2a-2c** was carried out to obtain the pure product as a yellow solid, and **2d** was purified by silica gel column chromatography to obtain the pure product as yellow viscous oil.

2-[4-(Diethylamino)-2-methoxybenzoyl]benzoic acid 2a

¹H-NMR (CDCl₃, 500 MHz): δ 8.00 (dd, 1H, J = 7.8, 0.8 Hz), 7.65 (d, 1H, J = 8.9 Hz), 7.53 (td, 1H, J = 8.9 Hz), 7.41 (td, 1H, J = 7.8, 1.1 Hz), 7.27 (d, 1H, J = 7.3 Hz), 6.24 (dd, 1H, J = 8.9, 2.3 Hz), 5.99 (sd, 1H, J = 2.1 Hz), 3.55 (s, 3H), 3.40 (q, 4H, J = 7.1 Hz), 1.20 (t, 6H, J = 7.1 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 193.9, 170.7, 162.2, 153.1, 146.4, 134.4, 132.3, 130.4, 127.9, 127.4, 126.9, 114.3, 103.8, 93.6, 55.3, 44.6, 12.6. HRMS (EI) calcd for C₁₈H₁₉NO₄ (M⁺): 327.1471, Found: 324.1471. Yield: 60 %.

2-[4-(Dipropylamino)-2-methoxybenzoyl]benzoic acid 2b

¹H-NMR (CDCl₃, 500 MHz): δ 8.00 (dd, 1H, J = 7.8, 0.92 Hz), 7.63 (d, 1H, J = 8.7 Hz), 7.52 (td, 1H, J = 7,3, 1.1 Hz), 7.40 (td, 1H, J = 7.6, 1.1 Hz), 7.26 (d, 1H, J = 6.6 Hz), 6.23 (dd, 1H, J = 9.2, 2.3 Hz), 5.97 (sd, 1H, J = 2.3 Hz), 3.54 (s, 3H), 3.28 (brs, 4H), 1.68-1.60 (m, 4H), 0.94 (t, 6H, J = 7.3 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 193.8, 170.9, 162.1, 153.5, 146.5, 134.3, 132.3, 130.3, 127.9, 127.4, 126.8, 114.3, 104.0, 93.9, 55.2, 52.8, 20.5, 11.4. HRMS (EI) calcd for C₂₁H₂₅NO₄ (M⁺): 355.1779, Found: 355.1786. Yield: 24 %.

2-[4-(Dibutylamino)-2-methoxybenzoyl]benzoic acid 2c

¹H-NMR (CDCl₃, 500 MHz): δ 8.01 (dd, 1H, J = 7.8, 0.92 Hz), 7.63 (brs, 1H), 7.53 (td, 1H, J = 7,3, 1.1 Hz), 7.41 (td, 1H, J = 7.8, 1.1 Hz), 7.26 (d, 1H, J = 7.3 Hz), 6.22 (dd, 1H, J = 9.2, 2.3 Hz), 5.98 (sd, 1H, J = 1.8 Hz), 3.55 (s, 3H), 3.31 (brs, 4H), 1.63-1.57 (m, 4H), 1.40-1.32 (m, 4H), 0.96 (t, 6H, J = 7.3 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 193.9, 170.6, 164.7, 162.2, 153.5, 146.4, 134.4, 132.3, 130.4, 127.9, 127.4, 126.9, 114.2, 103.9, 93.8, 55.2, 50.8, 29.4, 20.2, 13.9. HRMS (EI) calcd for C₂₃H₂₉NO₄ (M⁺): 383.2097, Found: 383.2093. Yield: 23 %.

2-[4-(Dihexylamino)-2-methoxybenzoyl]benzoic acid 2d

¹H-NMR (CDCl₃, 500 MHz): δ 8.02 (dd, 1H, J = 7.8, 0.92 Hz), 7.63 (brs, 1H), 7.53 (td, 1H, J = 7,6, 1.4 Hz), 7.42 (td, 1H, J = 7.8, 1.4 Hz), 7.27 (d, 1H, J = 8.7 Hz), 6.21 (dd, 1H, J = 8.9, 2.3 Hz), 5.97 (sd, 1H, J = 1.8 Hz), 3.56 (s, 3H), 3.23 (t, 4H, J = 7.8 Hz), 1.62-1.58 (m, 4H), 1.40-1.32 (m, 12H), 0.89 (t, 6H, J = 7.1 Hz). ¹³C-NMR (CDCl₃, 500 MHz): 194.0, 170.4, 162.2, 153.5, 146.3, 134.4, 132.3, 131.9, 130.8, 130.5, 129.9, 128.6, 128.0, 127.5, 127.0, 114.2, 103.95, 93.8, 55.2, 52.8, 51.1, 31.6, 27.2, 26.7, 22.6, 14.0. HRMS (EI) calcd for C₂₇H₃₇NO₄ (M⁺): 439.2723, Found: 439.2730.

2-2. General Procedure for the synthesis of **3**

To a stirred solution of **2** (1.0 equiv) in dry CH_2Cl_2 was added a solution of BBr₃ (1.9 equiv) in CH_2Cl_2 at -78 °C. After 1 hrs, the mixture was warmed to -25 °C. After completion of the reaction, the mixture was quenched with H_2O and evaporated to give the crude product. Recrystallization from MeOH/H₂O solution of **3a-3c** yielded the pure product as a yellow solid, and **3d** was purified by silica gel column chromatography to obtain the pure product as yellow viscous oil.

2-[4-(Diethylamino)-2-hydroxybenzoyl]benzoic acid 3a

¹H-NMR (CDCl₃, 500 MHz): δ 12.65 (s, 1H), 8.07 (dd, 1H, J = 7.8, 0.9 Hz), 7.58 (td, 1H, J = 7.6, 1.4 Hz), 7.51 (td, 1H, J = 7.8, 1.4 Hz), 7.34 (dd, 1H, J = 7.6, 1.1 Hz), 6.91 (d, 1H, J = 8.9 Hz), 6.11 (sd, 1H, J = 2.5 Hz), 6.03 (dd, 1H, J = 9.2, 2.5 Hz), 3.37 (q, 4H, J = 7.1 Hz), 1.18 (t, 6H, J = 7.1 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 199.0, 167.5, 165.2, 153.6, 140.6, 134.6, 131.6, 130.4, 129.5, 128.9, 127.6, 110.1, 103.5, 96.7, 44.5, 12.5. HRMS (EI) calcd for C₁₉H₂₁NO₄ (M⁺): 313.1314, Found: 313.1313. Yield: 65 %.

2-[4-(Dipropylamino)-2-hydroxybenzoyl]benzoic acid 3b

Yield: 1.0 g (52 %); ¹H-NMR (CDCl₃, 500 MHz): δ 12.53 (s, 1H), 8.10 (d, 1H, J = 7.1 Hz), 7.62 (t, 1H, J = 7.3 Hz), 7.53 (t, 1H, J = 7.3 Hz), 7.36 (dd, 1H, J = 7.8, 1.1 Hz), 6.87 (d, 1H, J = 9.2 Hz), 6.12 (s, 1H), 6.03 (d, 1H, J = 8.9 Hz), 3.26 (t, 4H, J = 7.6 Hz), 1.65-1.59 (m, 4H), 0.92 (t, 6H, J = 7.3 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ ¹H-NMR (CDCl₃, 500 MHz): δ 198.1, 169.8, 165.4, 154.4, 141.2, 134.5, 132.7, 131.1, 129.1, 128.1, 127.6, 109.9, 104.0, 97.3, 52.8, 20.5, 11.3. HRMS (EI) calcd for C₂₀H₂₃NO₄ (M⁺): 341.1630, Found: 341.1626.

2-[4-(Dibutylamino)-2-hydroxybenzoyl]benzoic acid 3c

¹H-NMR (CDCl₃, 500 MHz): δ 12.54 (s, 1H), 8.12 (d, 1H, J = 7.8 Hz), 7.63 (t, 1H, J = 7.3 Hz), 7.54 (t, 1H, J = 7.6 Hz), 7.36 (dd, 1H, J = 7.6, 0.92 Hz), 6.87 (d, 1H, J = 8.9 Hz), 6.11 (s, 1H), 6.03 (d, 1H, J = 8.5 Hz), 3.29 (t, 4H, J = 7.3 Hz), 1.58 (brs, 4H), 1.37-1.30 (m, 4H), 0.94 (t, 6H, J = 7.3 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 198.1, 169.5, 165.5, 154.4, 141.1, 134.5, 132.7, 131.1, 129.2, 128.1, 127.6, 109.8, 103.9, 97.2, 50.9, 29.4, 20.2, 13.9. HRMS (EI) calcd for C₂₂H₂₇NO₄ (M⁺): 369.1940, Found: 369.1932. Yield: 49 %.

2-[4-(Dihexylamino)-2-hydroxybenzoyl]benzoic acid 3d

¹H-NMR (CDCl₃, 500 MHz): δ 12.55 (s, 1H), 8.10 (dd, 1H, J = 8.0, 0.92 Hz), 7.62 (dd, 1H, J = 7.6, 1.4 Hz), 7.53 (td, 1H, J = 7.8, 1.4 Hz), 7.34 (dd, 1H, J = 7.1, 0.92 Hz), 6.87 (d, 1H, J = 9.2 Hz), 6.11 (sd, 1H, J = 2.5 Hz), 6.02 (dd, 1H, J = 9.2, 2.5 Hz), 3.27 (brs, 4H), 1.59 (brs, 4H), 1.33-1.30 (m, 12H), 0.90-0.88 (m 6H). ¹³C-NMR (CDCl₃, 500 MHz): δ 198.1, 169.0, 165.5, 154.4, 141.1, 134.5, 132.6, 131.1, 129.1, 128.1, 127.7, 109.8, 103.9, 97.2, 51.1, 31.59, 27.3, 26.7, 22.6, 14.0. HRMS (EI) calcd for C₂₆H₃₅NO₄ (M⁺): 425.2563, Found: 425.2569.

2-3. General Procedure for the synthesis of 4

3 (2.0 equiv) and resorcinol (1.0 equiv) were combined in methanesulfuric acid (2.0 mL) in a sealed tube and heated at 90 $^{\circ}$ C for 2 hrs. The reaction was poured into stirred ice water, its pH was adjusted to 11~12 with 1.0 M sodium hydroxide aqueous solution, and the mixture was stirred for 20 min. Then, the mixture was extracted with CH₂Cl₂ three times. The organic layers were dried over MgSO₄ and evaporated to give the crude product. This was purified by silica gel column chromatography and further recrystallized from acetonitrile solution to obtain the pure product as a white solid.

<u>3',3"-Bis(oxospriroisobenzofuran)-3,7-bis(diethylamino)benzopyranoxanthene</u> (ABPX01) 4a

 $δ^{1}$ H-NMR (THF-d₈, 500 MHz): δ 7.73-7.72 (m, 1H), 7.70 (dt, 1H, J = 7.6, 1.1 Hz), 7.61 (td, 1H, J = 7.6, 0.92 Hz), 7.52 (td, 1H, J = 7.3, 0.9 Hz), 7.47-7.43 (m, 2H), 7.16 (dt, 1H, J = 7.6, 0.9 Hz), 7.11 (s, 1H), 6.94-6.92 (m, 1H), 6.50-6.49 (m, 2H), 6.47 (d, 1H, J = 7.6 Hz), 6.45 (d, 1H, J = 7.6 Hz), 6.37-6.34 (m, 2H), 6.07 (s, 1H), 3.30-3.25 (m, 8H), 0.92 (t, 12H, J = 7.6 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 168.4, 168.3, 154.0, 153.9, 153.8, 153.7, 153.1, 153.0, 151.0, 150.9, 135.5, 134.8, 130.3, 130.0, 129.7, 129.4, 129.3, 128.8, 128.4, 128.1, 125.0, 124.9, 124.8, 124.4, 118.00, 117.8, 109.6, 106.6, 106.5, 104.4, 104.3, 98.6, 98.5, 83.0, 82.8, 53.4, 11.5. ¹³C-NMR (CDCl₃, 500 MHz): δ HRMS (FAB) calcd for C₄₂H₃₇N₂O₆ (M+H): 665.2651, Found: 665.2654. Yield: 72 %.

<u>3',3"-Bis(oxospriroisobenzofuran)-3,7-bis(dipropylamino)benzopyranoxanthene</u> (ABPX02) 4b

¹H-NMR (THF-d₈, 500 MHz): δ 7.73-7.72 (m, 1H), 7.70 (dt, 1H, J = 7.6, 1.1 Hz), 7.61 (td, 1H, J = 7.6, 0.92 Hz), 7.52 (td, 1H, J = 7.3, 0.9 Hz), 7.47-7.43 (m, 2H), 7.16 (dt, 1H, J = 7.6, 0.9 Hz), 7.11 (s, 1H), 6.94-6.92 (m, 1H), 6.50-6.49 (m, 2H), 6.47 (d, 1H, J = 7.6 Hz), 6.45 (d, 1H, J = 7.6 Hz), 6.37-6.34 (m, 2H), 6.07 (s, 1H), 3.30-3.25 (m, 8H), 1.65-1.57 (m, 8H), 0.92 (t, 12H, J = 7.6 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 168.4, 168.3, 154.0, 153.9, 153.8, 153.7, 153.2, 153.0, 151.0, 150.9, 135.5, 134.8, 130.3, 130.0, 129.7, 129.4, 129.3, 128.8, 128.4, 128.1, 125.0, 124.9, 124.8, 124.4, 118.0. 117.8, 109.6, 106.6, 106.5, 104.4, 104.3, 98.6, 98.5, 83.0, 82.8, 53.4, 21.1, 11.5. HRMS (FAB) calcd for C₄₆H₄₅N₂O₆ (M+H): 721.3278, Found: 721.3276. Yield: 56 %.

<u>3',3"-Bis(oxospriroisobenzofuran)-3,7-bis(dibuthylamino)benzopyranoxanthene</u> (ABPX03) 4c

¹H-NMR (THF-d₈, 500 MHz): δ 7.74-7.72 (m, 1H), 7.70 (dt, 1H, J = 7.6, 0.92 Hz), 7.62 (td, 1H, J = 7.6, 1.1 Hz), 7.52 (td, 1H, J = 7.3, 0.92 Hz), 7.47-7.43 (m, 2H), 7.16 (dt, 1H, J = 7.6, 0.92 Hz), 7.13 (s, 1H), 6.94-6.92 (m, 1H), 6.50-6.49 (m, 2H), 6.47 (d, 1H, J = 7.8 Hz), 6.45 (d, 1H, J = 7.8 Hz), 6.37-6.34 (m, 2H), 6.07 (s, 1H), 3.33-3.30 (m, 8H), 1.60-1.54 (m, 8H), 1.40-1.32 (m, 8H), 0.95 (t, 12H, J = 7.3 Hz). ¹³C-NMR (CDCl₃, 500 MHz): δ 168.4, 168.3, 154.0, 153.9, 153.8, 153.7, 153.2, 153.0, 151.0, 150.9, 135.4, 134.8, 130.3, 130.0, 129.7, 130.0, 129.4, 128.8, 128.5, 128.1, 125.0, 124.9, 124.8, 124.4, 118.0, 117.8, 109.6, 106.6, 106.5, 104.5, 104.4, 98.6, 98.5, 83.0, 82.8, 51.4, 30.2, 21.0, 14.3. HRMS (FAB) calcd for C₅₀H₅₃N₂O₆ (M+H): 777.3903, Found: 777.3911. Yield: 52 %.

<u>3',3"-bis(oxospriroisobenzofuran)-3,7-bis(dihexylamino)benzopyranoxanthene</u> (ABPX04) 4d

¹H-NMR (THF-d₈, 500 MHz): δ 7.74-7.72 (m, 1H), 7.70 (d, 1H, J = 7.8 Hz), 7.62 (td, 1H, J = 7.6, 0.92 Hz), 7.52 (td, 1H, J = 7.6, 0.92 Hz), 7.47-7.44 (m, 2H), 7.16 (d, 1H, J = 7.6 Hz), 7.12 (s, 1H), 6.94-6.92 (m, 1H), 6.49-6.48 (m, 2H), 6.47 (d, 1H, J = 8.7 Hz), 6.45 (d, 1H, J = 8.0 Hz), 6.36-6.34 (m, 2H), 6.06 (s, 1H), 3.31 (brs, 8H), 1.59 (brs, 8H), 1.33 (brs, 24H), 0.90 (brs, 12H). ¹³C-NMR (CDCl₃, 500 MHz): δ 168.4, 168.3, 154.0, 154.0, 153.9, 153.8, 153.7, 153.2, 153.0, 151.0, 150.9, 150.8, 135.4, 134.8, 130.3, 130.0, 129.7, 129.4, 129.3, 128.7, 128.5, 128.1, 125.0, 125.9, 124.8, 124.4, 118.0, 117.8, 109.6, 106.6, 106.5, 104.5, 104.4, 98.6, 98.5, 83.0, 82.8, 51.7, 32.7, 28.0, 27.6, 23.6, 14.4. HRMS (FAB) calcd for C₅₈H₆₉N₂O₆ (M+H): 889.5155, Found: 889.5165

3. Crystal data and experimental data of ABPX01⁰

Crystal and experimental data: C₄₂H₃₇N₂O₆, M = 665.76, orthorhombic, Fdd2, *a* = 17.5837 (11), *b* = 23.2759 (11), *c* = 18.2237 (9) Å, α , β , γ = 90°, *V* = 7458.5 (7) Å³, *Z* = 8, *D*x = 1.186 g cm⁻³, *F*(000) = 2808, μ (Mo K α) = 0.079 mm⁻¹. The data were collected on a Rigaku RAXIS-RAPID area detector using graphite-monochromated MoK α radiation at 293.1 K. A total of 17,587 reflections were measured up to θ_{max} = 27.46° (0.85 Å resolution) and merged to 14,495 reflections with *R*_{int} = 0.034. The structure was solved using CrystalStructure 3.8 and refined with CrystalStructure 3.8. Non-hydrogen atoms were refined anisotropically. Hydrogen atoms were calculated at the ideal positions and isotropically included in the calculations of the structure factors. The structure was converted at goodness of fit = 1.232, (Δ/σ)_{max} = 0.0088, $\Delta \rho_{max}$ = 2.50 eÅ⁻³, $\Delta \rho_{min}$ = -2.27 eÅ⁻³, *R* = 0.0589, and *wR* = 0.0772. Crystallographic data (excluding structure factors) for the structure of ABPX01⁰ reported in this paper are deposited with the Cambridge Crystallographic Data Centre as supplementary publication number CCDC-779022. The data can be obtained free of charge from The Cambridge Crystallographic Data Centre *via* www.ccdc.cam.ac.uk/data request/cif.

4. Scheme and figures



Scheme S1 Synthetic routes of benzophenone derivatives.



Fig. S1 Chemical structure of ABPX01⁰.



Fig. S2 ORTEP drawing of ABPX01⁰. The trans-ABPX01⁰ was confirmed in solids, of which five rings had a high degree of planarity.



Fig. S3 Effect of protolytic reaction on absorption spectra of ABPX01 in TFA/THF solutions. TFA: trifluoroacetic acid, TEA: triethylamine. Inset: photograph of the solutions under room light. In strong acid solution, the absorbance spectra having two peaks at 490 nm and 540 nm should be attributable to the cationic species $ABPX01^+$. In THF and alkaline solutions, the spectra had no peaks that were attributable to the neutral lactonoid species $ABPX01^0$.



Fig. S4 (a) Emission at various concentrations of ABPX01⁺. (b) photograph of the solutions, top: under 365 nm irradiation, bottom: under room light irradiation.



Fig. S5 (a) Emission spectra of 500 μ M of ABPX01⁺ and (b) absorption spectra of 5 μ M of ABPX01⁺ in various solvents. The emission spectrum in THF was not observed. 1.0% TFA was added to ensure protonation and prevent pseudobase formation during measurement.



Fig. S6 (a) Emission spectra of 500 μ M of ABPX02⁺ and (b) absorption spectra of 5 μ M of ABPX02⁺ in various solvents. The emission spectrum in THF was not observed. 1.0% TFA was added to ensure protonation and prevent pseudobase formation during measurement.



Fig. S7 (a) Emission spectra of 500 μ M of ABPX03⁺ and (b) absorption spectra of 5 μ M of ABPX03⁺ in various solvents. The emission spectrum in THF was not observed. 1.0% TFA was added to ensure protonation and prevent pseudobase formation during measurement.



Fig. S8 (a) Emission spectra of 500 μ M of ABPX04⁺ and (b) absorption spectra of 5 μ M of ABPX04⁺ in various solvents. The emission spectrum in THF was not observed. 1.0% TFA was added to ensure protonation and prevent pseudobase formation during measurement.



Fig. S9 (a) Emission and (b) absorption spectra of $ABPX01^+-04^+$ in chloroform. A slightly continuous red shift was observed with an increase in *N*-alkyl chain length. 1.0% TFA was added to ensure protonation and prevent pseudobase formation during measurement.



Fig. S10 Emission spectra of ABPX01 by (a) conventional photometric method and (b) surface photometric method. Shown below are the respective equipment configurations. The emission spectra by the conventional photometric method showed random shifts of the maximum wavelengths, which were caused by the internal filter effect when measurement of $ABPX01^+$ was conducted at a high concentration. The problem was solved by using the surface photometric method. The sample holder was set at an angle of 30° with respect to the excitation light. However, attention should be paid to the linearity of the emission intensity against the concentration for the accurate evaluation of high-concentration sample solutions, such as AIEE molecules. In addition, the relative quantum yields of ABPX in solution could not be determined because of the lack of a standard compound. We are now developing an apparatus for the measurement of the absolute fluorescence quantum yields of high-concentration solutions.



Fig. S11 (a) Level-off tails of 500 μ M of ABPX01⁺ were observed in chloroform, methanol, and 90% water/THF mixture solutions. In comparison, level-off tails of 500 μ M of ABPX01⁺ in chloroform could not be observed. (b) Dynamic laser scattering (time dependence of particle size distributions of 500 μ M of ABPX01⁺ in 90% water/THF mixtures). The nanoaggregate size of ABPX01⁺ in solutions was increased as a time-dependent change. 1.0% TFA was added to ensure protonation and prevent pseudobase formation during measurement.



Fig. S12 Membrane filter photographs of $ABPX01^+$ in 90% water/THF mixtures (containing 1% TFA) under (a) room light and (b) 365 nm irradiation.



Fig S13. (a) Effects of composition of glycerol-methanol mixture on emission intensity of ABPX01. In the mixtures with glycerol fractions > 80%, the emission intensity rapidly increased due to the viscosity effect (viscochromism). (b) Effects of temperature on emission intensity of ABPX01 in methanol. When a methanol solution of ABPX01 was cooled, its emission intensity was increased due to thermochromism. The above results suggested that RIR caused AIEE. 1.0% TFA was added to ensure protonation and prevent pseudobase formation during measurement.

5. NMR data of benzophenone derivatives and ABPX01-04















































