

Electronic Supplementary Information

Highly selective reduction of nitroarenes by iron(0) nanoparticles in water

*Raju Dey, Nirmalya Mukherjee, Sabir Ahammed, and Brindaban C. Ranu**

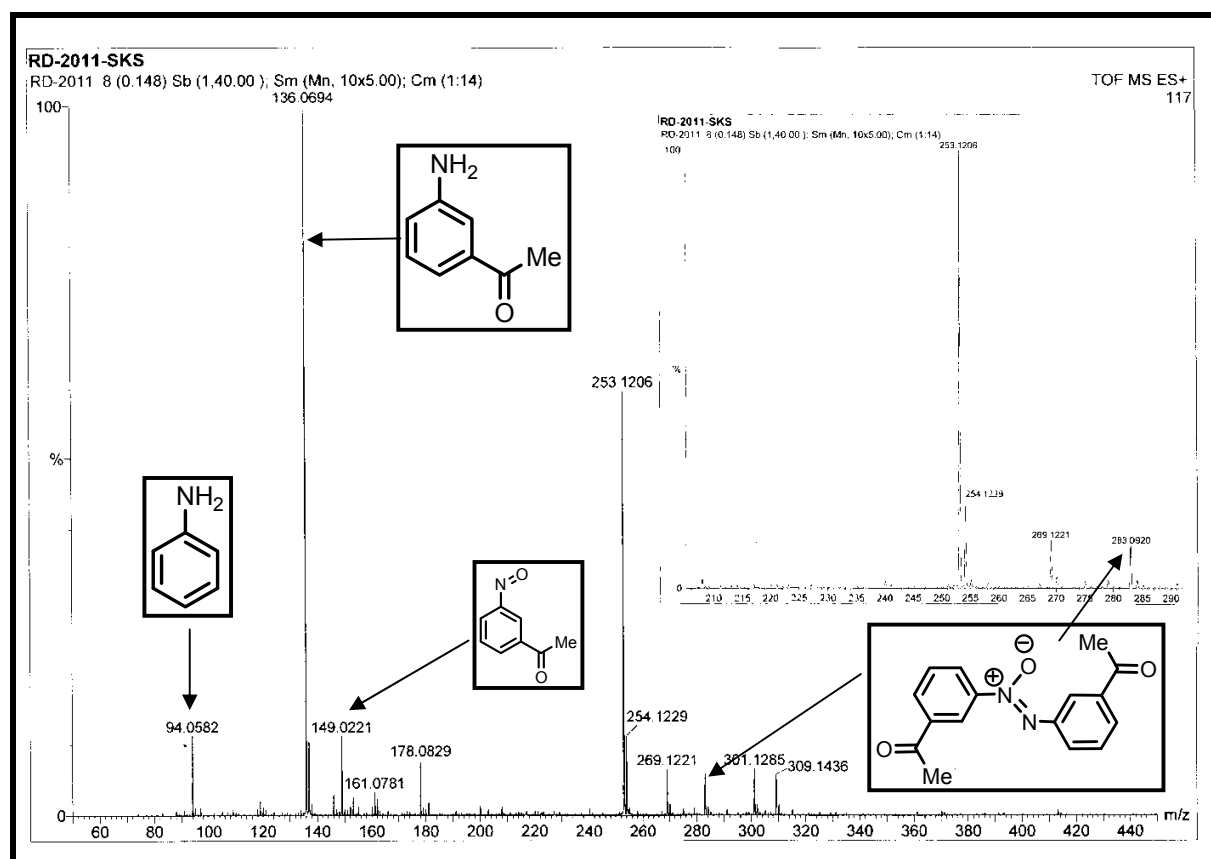
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General experimental procedure for reduction of nitroarenes

Sodium borohydride (200 mg, 5 mmol) was added to a stirring solution of ferrous sulphate-hepta-hydrate [$\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$] (834 mg, 3 mmol) and citric acid (55 mg, 0.25 mmol) in water (100 mL) and vigorous stirring was continued for 3-4 minutes. When the solution settled down the water layer was decanted. The residual black solid material of Fe nanoparticles was further washed with water two times to be ready for use in nitro reduction. The nitro compound (1 mmol) was added to these Fe nanoparticles in water (5 mL) in the same pot under constant stirring at room temperature under argon. After completion of reaction (TLC) the mixture was extracted with ethyl acetate (3 X 10 mL) (all Fe-species remained around stirring bar). Evaporation of solvent and purification by short column chromatography provided the pure product which was properly characterized by spectroscopic (IR, ^1H NMR, and ^{13}C NMR) data.

Sample preparation and recording of HRMS at intermediate stage of reaction

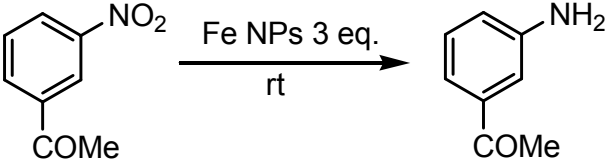
After 1 h of reaction an aliquot of reaction mixture (1 mL) was taken out and extracted with ethyl acetate (3 X 3 mL). After evaporation of solvent the crude material was subjected to HRMS and the spectra was recorded.



Heterogeneity Test

A mixture of Fe NPs (3 mmol), 3-acetylnitrobenzene (165 mg, 1 mmol) in water (5 mL) was stirred at room temperature for 1 h. At this stage (60% conversion by ^1H NMR), the catalyst was filtered off and the experiment was continued with the filtrate for another 1 h. There was no increase in the product concentration, as determined by the ^1H NMR analysis.

Table The role of solvent in the reduction process.

			
entry	solvent	time (h)	yield (%) ^a
1	Toluene	4	10
2	CH_2Cl_2	4	8
3	DMF	4	7
4	Water	2	99
5	Methanol	2	10

Yields are based on the ^1H -NMR spectra of the crude product.

Characterization of Fe(0) nanoparticles

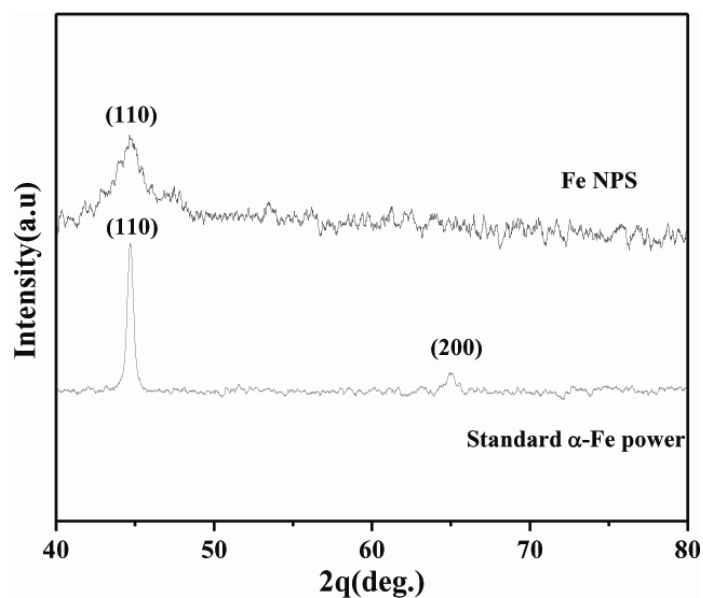


Fig. 1. XRD spectra of Fe(0) NPS

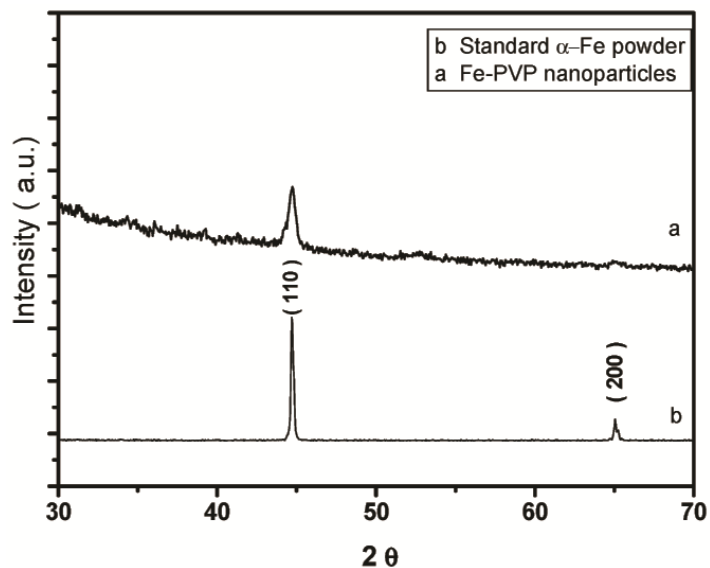


Fig. 2. XRD spectra of Fe(0) nanoparticles reported by Guo et al, *Langmuir* 2006, 22, 7867

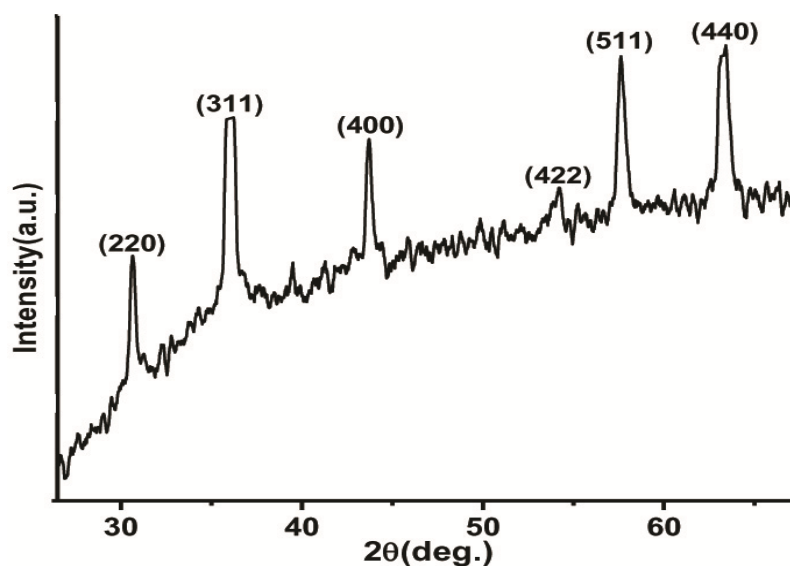


Fig. 3. XRD of end residue of Fe after reaction

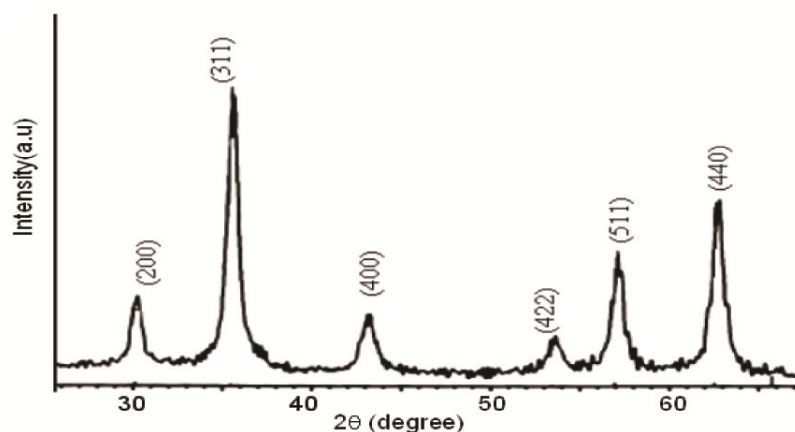


Fig. 4. XRD spectra of Fe_3O_4 nanoparticles reported by Gui-yin Li, Yu-ren Jiang, Ke-long Huang, Ping Ding, Jie Chen, *J. Alloys & Comp.* 466 (2008) 451–456

Fig. 4. shows X-ray diffraction patterns of the authentic Fe_3O_4 nanoparticles. The XRD peaks of our Fe end residue (Fig. 3.) are compared with those of standard one in JCPDS file (PDF No. 65-3107). A series of characteristic peaks at $2\theta = 30.17^\circ$, 35.46° , 43.38° , 53.69° , 57.23° , and 62.77° , which correspond to (220), (311), (400), (422), (511), and (440) Bragg reflection, respectively in Fig. 3. agree with standard magnetite (Fe_3O_4) XRD patterns in Fig. 4. The Fe_3O_4 nanoparticles are of cubic spinel structure.

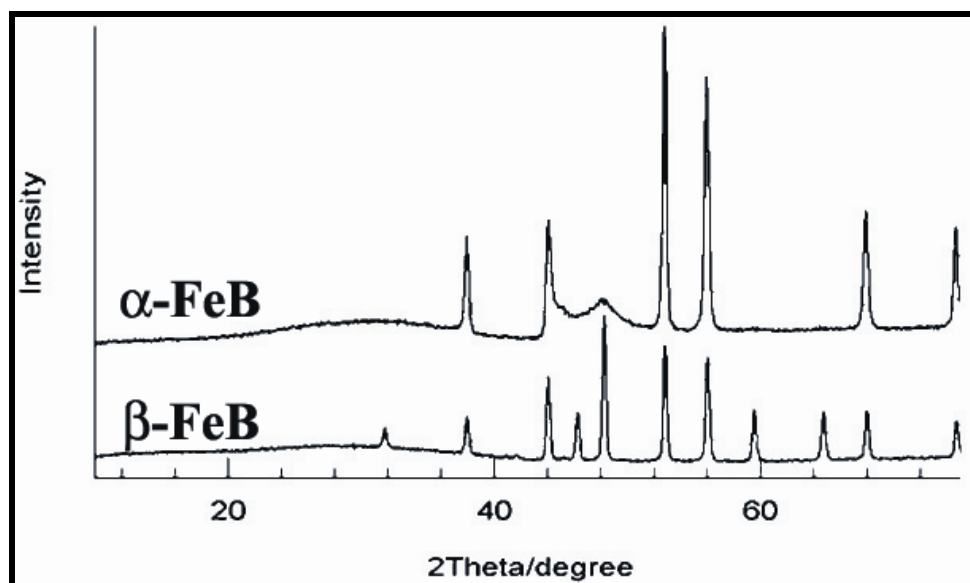


Fig. 5. X-ray diffraction patterns of the α -FeB and β -FeB

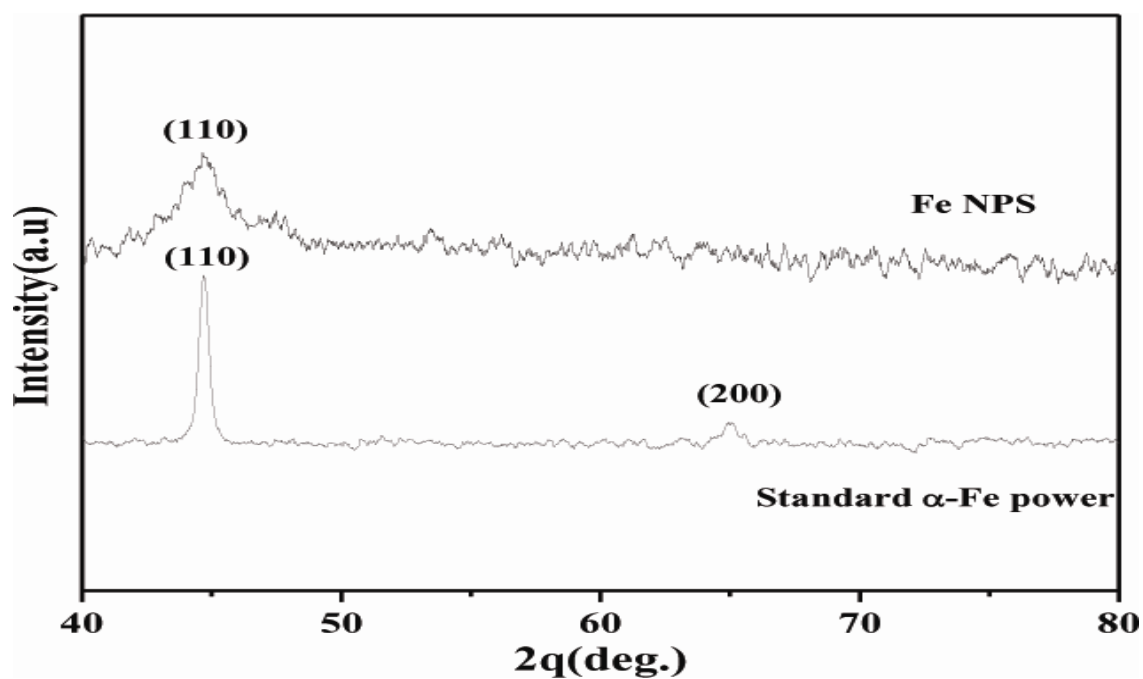


Fig. 1. XRD spectra of our Fe(0) NPS

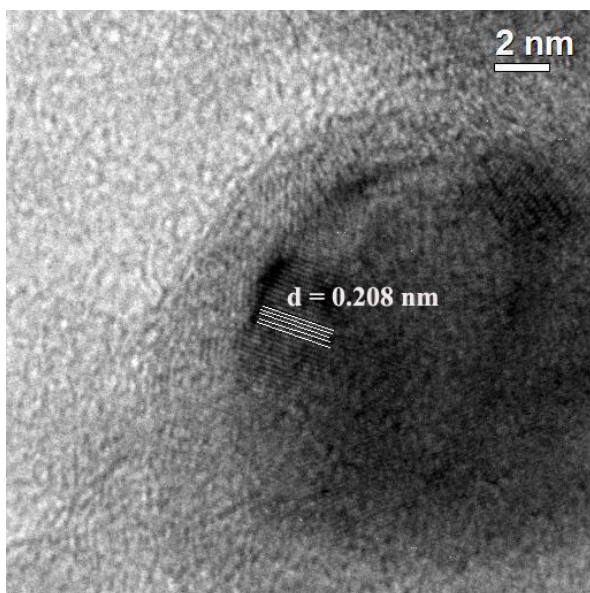


Fig. 6. HR-TEM image of Fe(0) NPs

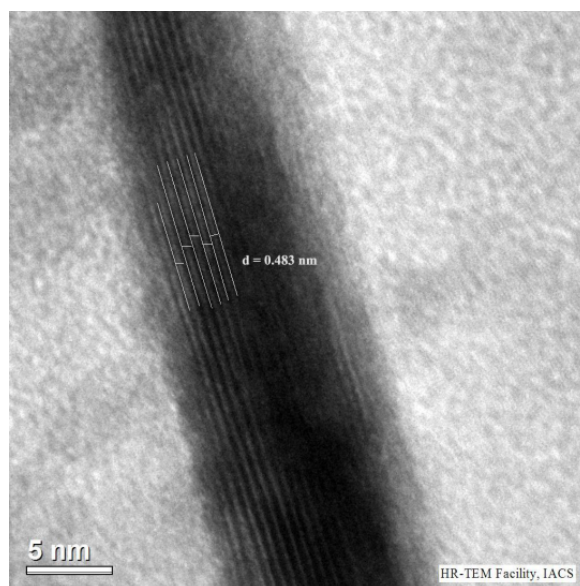


Fig. 7. HR-TEM image of of end residue of Fe (Fe_3O_4) after reaction

Fe(0) NPs: From XRD value it was found that prepared Fe(0) NPs is BBC (Body Center Cube)

Lattice d spacing value = 0.208 nm (found form HRTEM)

Standard d spacing value = 2.03 nm (JCPDS data base)

Lattice parameter = 2.94 (from calculation)

Lattice parameter = 2.87 (JCPDS data base)

Fe_3O_4 NPs: From XRD value it was found that prepared Fe(0) NPs is FCC (Body Center Cube)

d spacing value = 0.483 nm (found form HRTEM)

Standard d spacing value = 0.484 nm (JCPDS data base)

Lattice parameter = 8.36 (from calculation)

Lattice parameter = 8.38 (JCPDS data base)

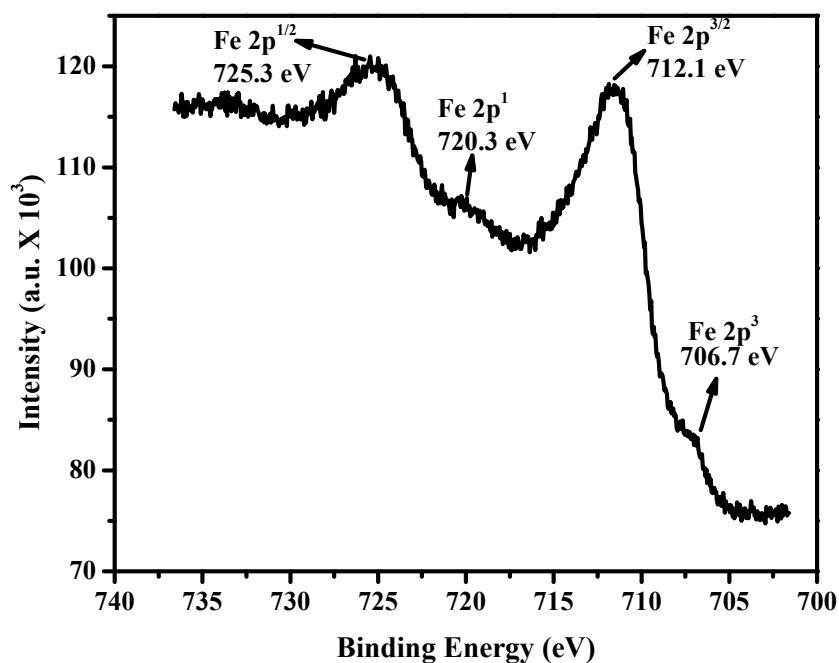


Fig. 8. XPS spectra for Fe(0) NPs

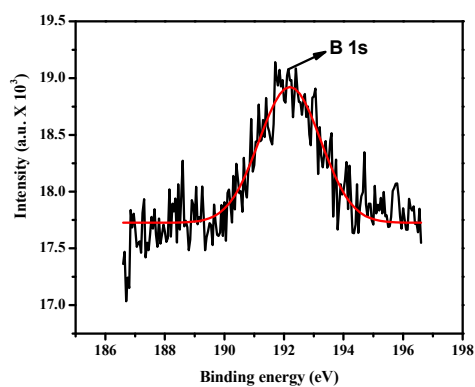


Fig. 9. The XPS spectra of our Fe(0) in the B 1s core level

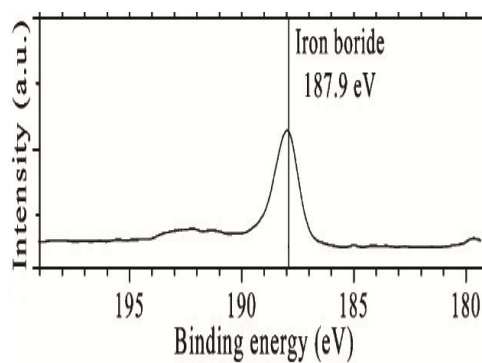
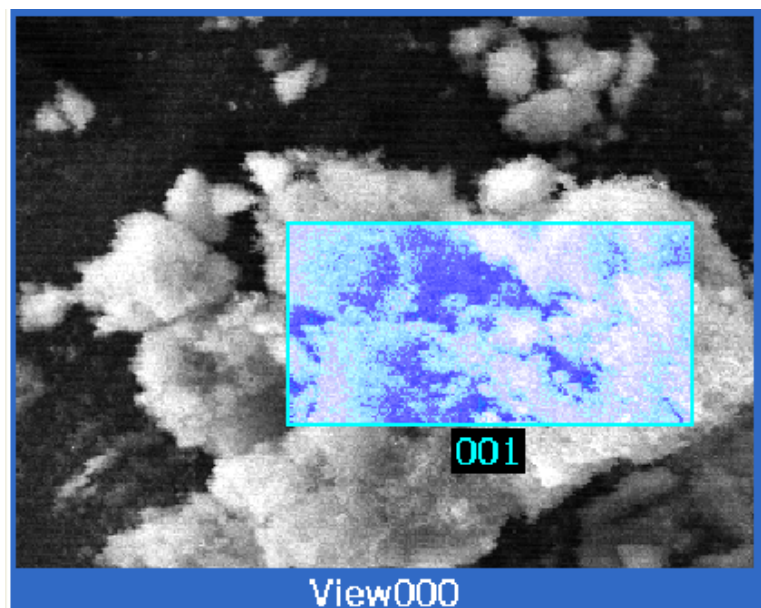
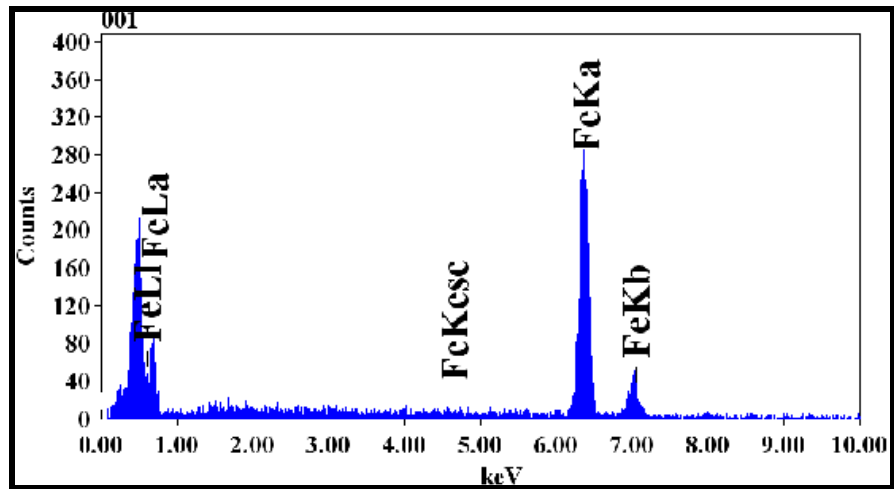


Fig. 10. The XPS spectra on the B 1s core level of FeB reported by Dehlinger et al. Surf. Coat. Technol. (2003) 331–337

EDX experiment with Fe(0) NPS



SEM Image of Fe(0) NPS

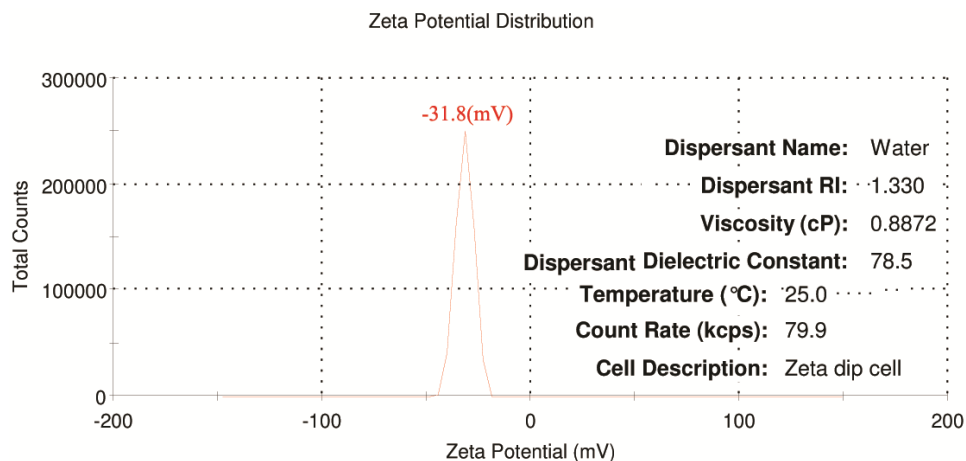


EDX spectra of selected area of Fe(0) NPS

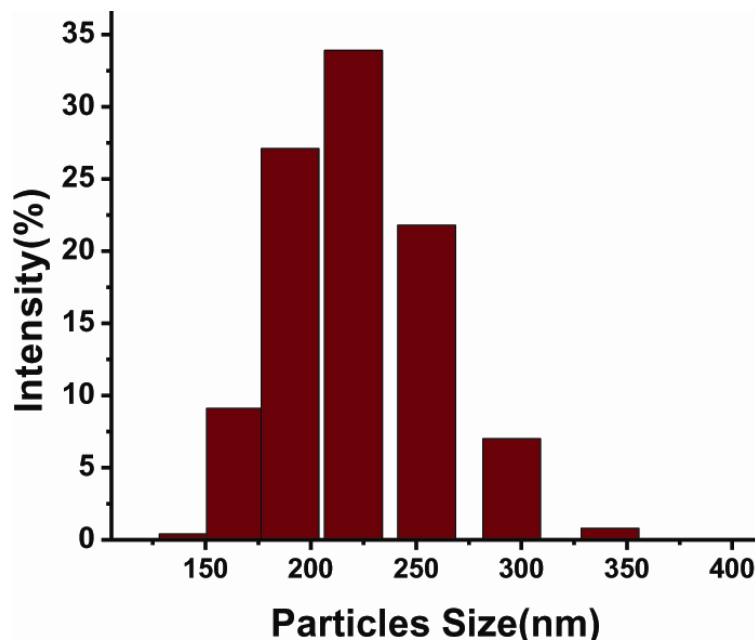
Ele...	Cl...	(keV)	mass%	Error%	At%
Fe K		6.398	100.00	1.36	100.00
Total			100.00		100.00

Composition of the material

Zeta potential value of Fe NPS:



Particles size distribution of Fe NPS (DLS exp):



Characterization data of new compounds (Table 1, entries 18 and 19)

3-Aminocinnamyl acetate (Table 1, entry 18): Yellow gummy liquid, IR (Neat): 775, 966, 1024, 1247, 1456, 1600, 1732, 2929, 3032, 3367 cm^{-1} ; ^1H NMR (CDCl_3 , 500MHz) δ 2.09 (s, 3H), 4.70 (d, $J=7$ Hz, 2H), 6.21-6.26 (m, 1H), 6.54-6.60 (m, 2H), 6.71 (s, 1H), 6.79 (d, $J=7.5$ Hz, 1H), 7.10 (t, $J=7.5$ Hz, 1H); ^{13}C NMR(CDCl_3 , 125 MHz) δ 21.1, 65.2, 113.2, 115.1, 117.4, 123.1, 129.6, 134.5, 137.4, 146.7, 170.9;
HRMS calcd for $\text{C}_{11}\text{H}_{14}\text{NO}_2[\text{M}+\text{H}]^+$: 192.1025; found: 192.1019.

3-((E)-3-(Phenylselanyl)prop-1-enyl)benzenamine (Table 1, entry 19): Yellow gummy liquid, IR (Neat): 690, 736, 777, 960, 1477, 1579, 1599, 1616, 2852, 2926, 3028, 3367 cm^{-1} ; ^1H -NMR (CDCl_3 , 500MHz) δ 3.68 (d, $J=7.5$ Hz, 2H), 6.17 (d, $J=7.5$, 2H), 6.25-6.31 (m, 1H), 6.54-6.56 (m, 1H), 6.61(s, 1H), 6.69 (d, $J=7.5$ Hz, 1H), 7.07 (t, $J=7.5$ Hz, 1H), 7.24-7.26 (m, 3H), 7.51-7.53 (m, 2H); ^{13}C -NMR (CDCl_3 , 125 MHz) δ 30.8, 112.9, 114.6, 117.2, 125.9, 127.4, 129.1, 129.5, 132.4, 133.9, 138.1, 146.6; HRMS calcd for $\text{C}_{15}\text{H}_{16}\text{NSe}[\text{M}+\text{H}]^+$: 290.0448; found: 290.0469.

Copies of ^1H NMR and ^{13}C NMR spectra of all products in Table 1

