

*Electronic Supplementary Information (ESI)*

## **Reversible photoswitching specifically responds to mercury(II) ions: the gated photochromism of bis(dithiazole)ethene**

*Yue Wu,<sup>a,b</sup> Weihong Zhu,<sup>\*a</sup> Wei Wan,<sup>b</sup> Yongshu Xie,<sup>a</sup> He Tian,<sup>a</sup> and Alexander D. Q. Li<sup>\*b</sup>*

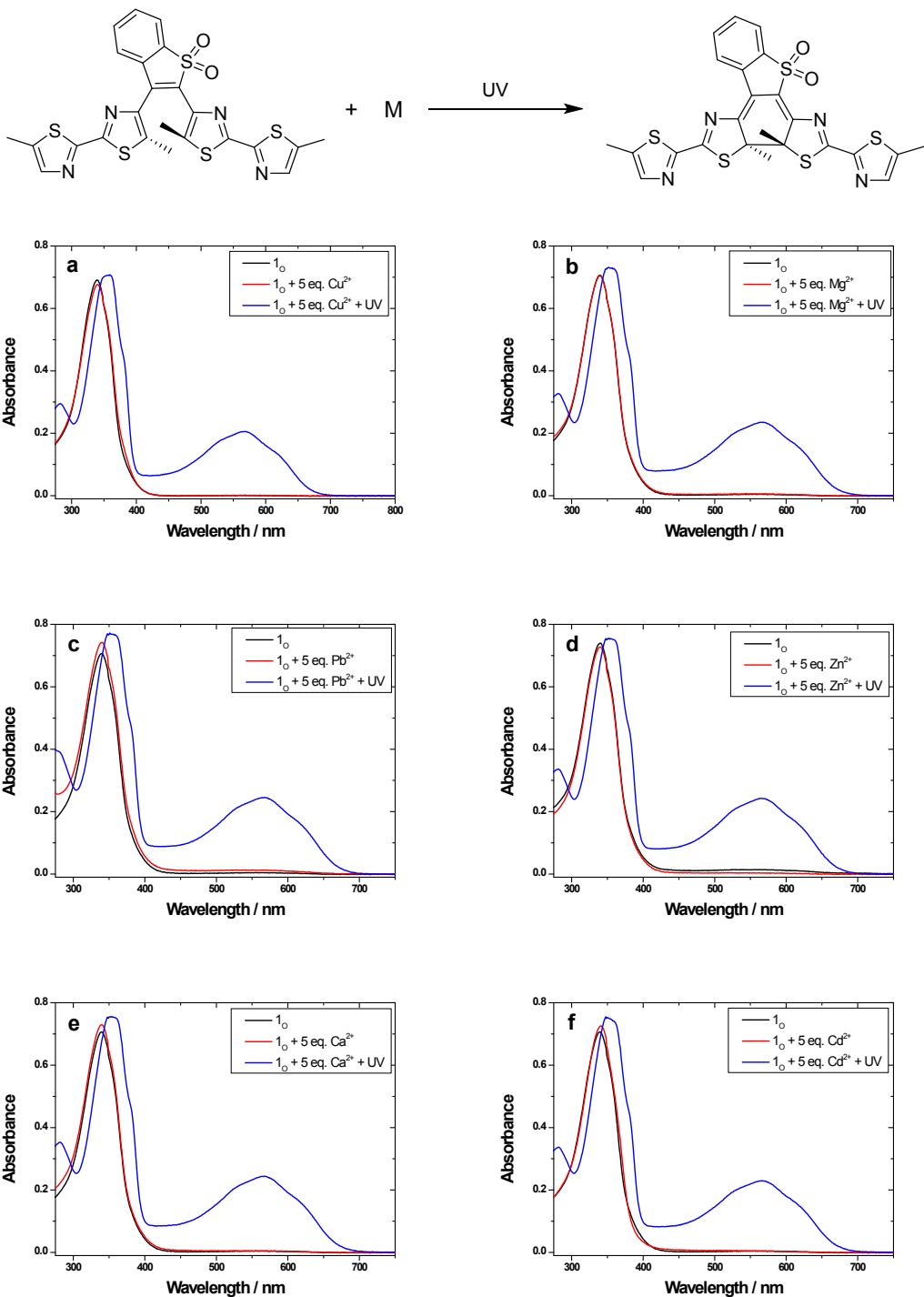
*<sup>a</sup>Key Laboratory for Advanced Materials and Institute of Fine Chemicals, Shanghai Key Laboratory of Functional Materials Chemistry, East China University of Science and Technology, Shanghai 200237, P. R. China. E-mail: whzhu@ecust.edu.cn*

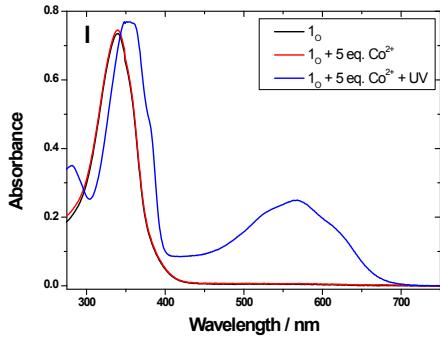
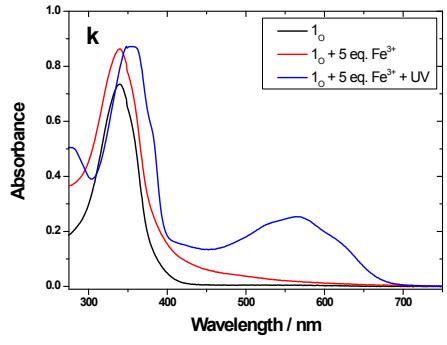
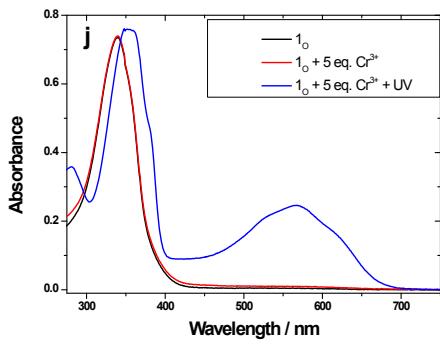
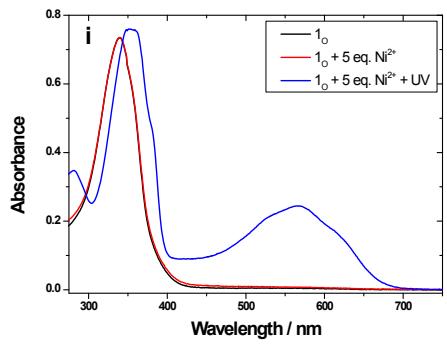
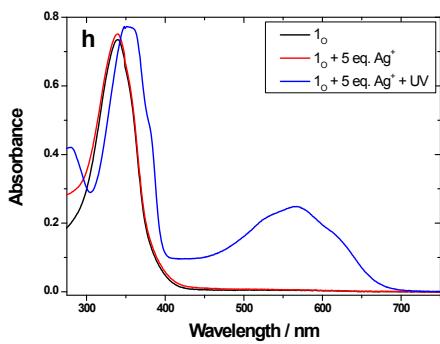
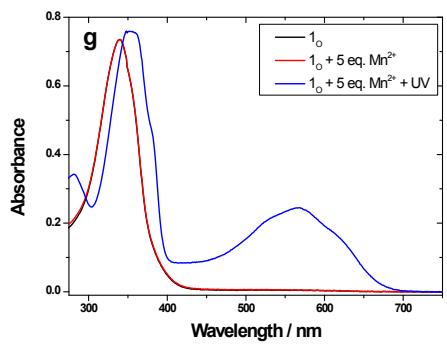
*<sup>b</sup>Department of Chemistry, Washington State University, Pullman, Washington 99164, United States. E-mail: dequan@wsu.edu*

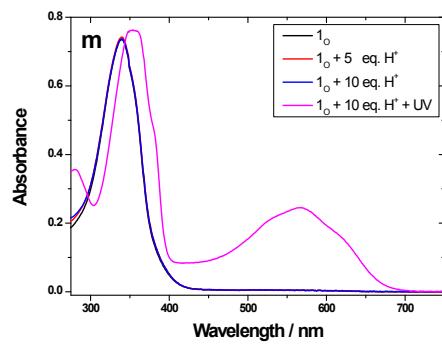
### **Contents:**

- 1. UV-Vis spectra of 1 upon adding protons and metal ions**
- 2. UV-Vis spectra of 2 upon adding Hg<sup>2+</sup>**
- 3. Titrating curve of 1 and Hg<sup>2+</sup>**
- 4. Gating of molecular Photoswitching**
- 5. Theoretical calculations**

## 1. UV-Vis spectra of **1** upon adding protons and metal ions

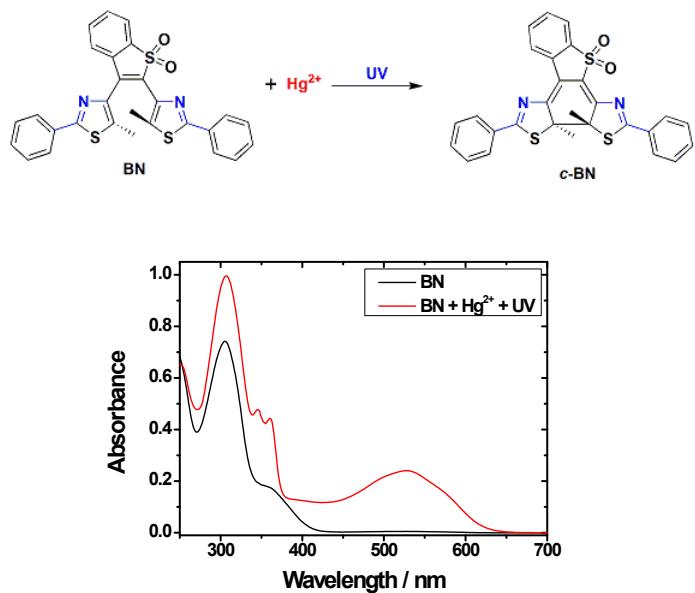






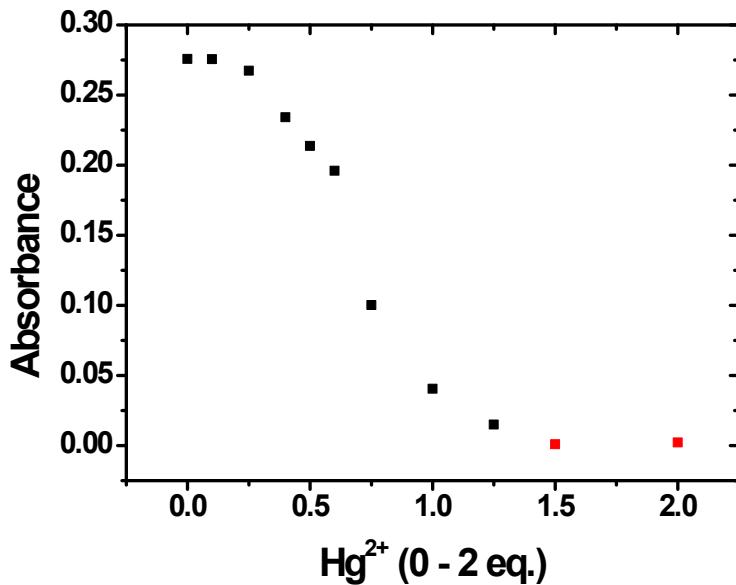
**Fig. S1** UV-Vis absorption changes in the presence of various ions (a-Cu<sup>2+</sup>, b-Mg<sup>2+</sup>, c-Pb<sup>2+</sup>, d-Zn<sup>2+</sup>, e-Ca<sup>2+</sup>, f-Cd<sup>2+</sup>, g-Mn<sup>2+</sup>, h-Ag<sup>+</sup>, i-Ni<sup>2+</sup>, j-Cr<sup>3+</sup>, k-Fe<sup>3+</sup>, l-Co<sup>2+</sup>, m-H<sup>+</sup> as perchlorates were used) of **1o** ( $2.0 \times 10^{-5}$  M) in CH<sub>3</sub>CN, and then upon irradiation with 365-nm light for 50 s.

## 2. UV-Vis spectra of **2** upon adding $\text{Hg}^{2+}$

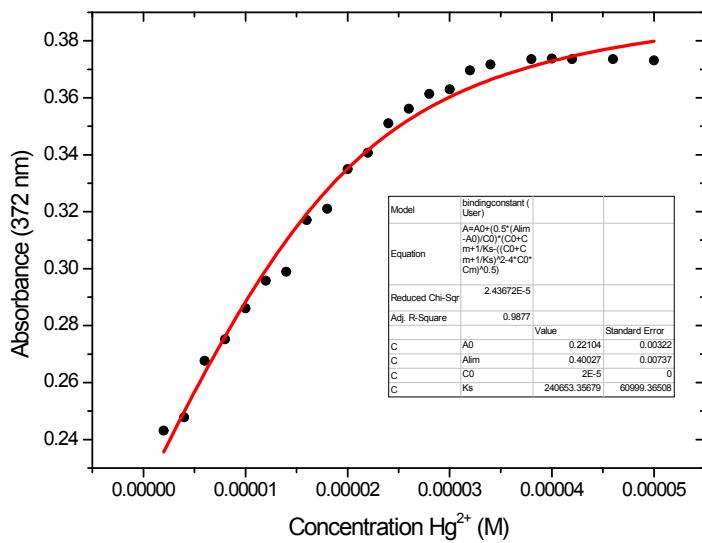


**Fig. S2** In the presence of 5 eq.  $\text{Hg}(\text{ClO}_4)_2$ , absorption spectra of **2** (**BN**) in  $\text{CH}_3\text{CN}$  (black line), and after 365-nm light irradiation (red line).

### 3. Titrating curve of **1** and $\text{Hg}^{2+}$



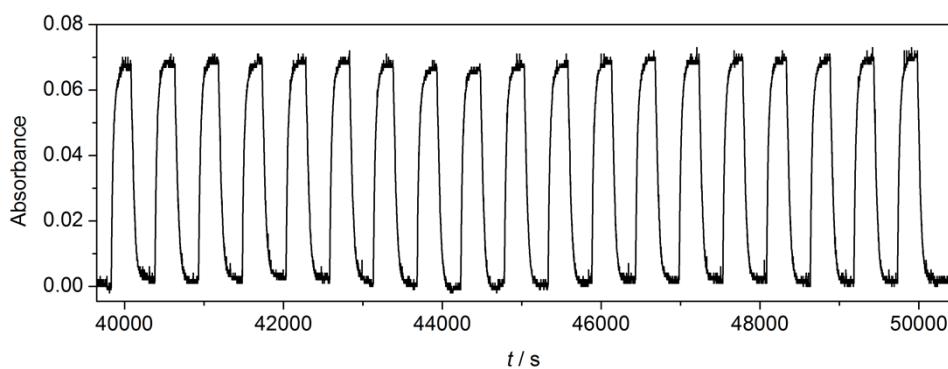
**Fig. S3** Absorption changes at 568 nm due to interconversion of **1<sub>O</sub>** to **1<sub>C</sub>** ( $2.0 \times 10^{-5}$  M) upon titration of  $\text{Hg}(\text{ClO}_4)_2$ , followed by 365-nm irradiation for 50 s, respectively.



**Fig. S4** Absorption changes at 372 nm of **1<sub>O</sub>** ( $2.0 \times 10^{-5}$  M) upon titration of  $\text{Hg}(\text{ClO}_4)_2$ .

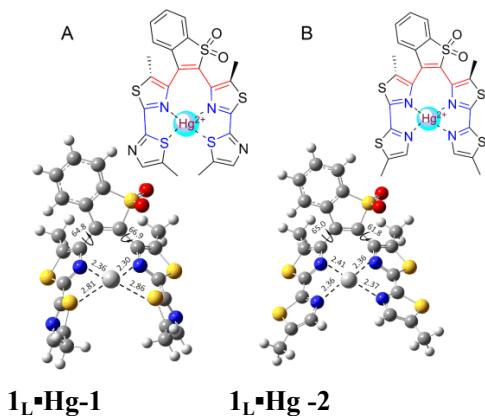
#### 4 Gating of molecular Photoswitching

In a typical experiment, the molecular photoswitch **1<sub>O</sub>** concentration was  $c = 1.0 \times 10^{-5}$  M, in 1 mL volume. The corresponding equivalence of HgClO<sub>4</sub> in CH<sub>3</sub>CN was added to the above photoswitching system, which was tested using the following conditions. The UV irradiation was carried out using a 375-nm laser operating at 6.4  $\mu\text{W/mm}^2$ . Typically UV-laser on time averaged from 200 seconds to 270 seconds. Because visible lasers have relatively low energy per photon, three visible lasers were used simultaneously to impart the ring-opening reaction for 250-350 seconds. The wavelengths are 561 nm, 532 nm, and 473 nm; together they produce 175  $\mu\text{W/mm}^2$ . These switching lasers were applied orthogonally to the incoming and out-going light used to measure absorbance values. The light source for the absorbance measurements came from an Ocean Optics DH-200-BAL module via an UV-NIR transparent optical fiber. A similar optical fiber was used to deliver the output light to a CCD detector (Ocean Optics QE65 Pro) for simultaneous measurements of absorbance at all wavelength. The absorbance data were collected using the Ocean Optics SpectraSuite software.



**Fig. S5** Alternating a 375-nm UV laser and three visible lasers (473, 532, and 561-nm) causes the interconversion between **1<sub>O</sub>** and **1<sub>C</sub>**. Partial equivalent of Hg(II) ions (**1<sub>O</sub>** + 0.5 eq. Hg<sup>2+</sup>) were added to gate the magnitude of the photoswitching. In this case, the photoswitching was reduced to 70%, but fatigue resistance remains excellent.

## 5. Theoretical calculations



**Table S1.** Structural parameters for the complexes and relative energies

Compound	Relative energy (kcal mol <sup>-1</sup> )		d <sub>f-f'</sub> (Å)	d <sub>a-b</sub> (Å)	d <sub>a-b'</sub> (Å)	d <sub>a-c</sub> (Å)	d <sub>a-c'</sub> (Å)	e'-e-d-b  (°)	e-e'-d'-b'  (°)
	1 <sub>L</sub> •Hg-1	42.3	5.16	2.36	2.30	2.81	2.86	64.8	66.9
1 <sub>L</sub> •Hg-2	0	5.30	2.41	2.36	2.36	2.37	65.0	61.8	

Quantum chemical calculations were performed within the framework of Gaussian09 program.<sup>[S1]</sup> Geometrical optimizations were carried out in gas phase at the PBE0/6-31G(d) level, except for Hg, LANL2DZ basis set was used. **1<sub>L</sub>•Hg-2** was predicted to be the most stable complex formed with Hg<sup>2+</sup>, that is, **1<sub>L</sub>•Hg-1** was 42.3 kcal mol<sup>-1</sup> less stable than **1<sub>L</sub>•Hg-2**.

[S1] M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, H. Nakatsuji, M. Caricato, X. Li, H. P. Hratchian, A. F. Izmaylov, J. Bloino, G. Zheng, J. L. Sonnenberg, M. Hada, M. Ehara, K. Toyota, R. Fukuda, J. Hasegawa, M. Ishida, T. Nakajima, Y. Honda, O. Kitao, H. Nakai, T. Vreven, J. A. Montgomery, Jr., J. E. Peralta, F. Ogliaro, M. Bearpark, J. J. Heyd, E. Brothers, K. N. Kudin, V. N. Staroverov, R. Kobayashi, J. Normand, K. Raghavachari, A. Rendell, J. C. Burant, S. S. Iyengar, J. Tomasi, M. Cossi, N. Rega, J. M. Millam, M. Klene, J. E. Knox, J. B. Cross, V. Bakken, C. Adamo, J. Jaramillo, R. Gomperts, R. E. Stratmann, O. Yazyev, A. J. Austin, R. Cammi, C. Pomelli, J. W. Ochterski, R. L. Martin, K. Morokuma, V. G. Zakrzewski, G. A. Voth, P. Salvador, J. J. Dannenberg, S. Dapprich, A. D. Daniels, O. Farkas, J. B. Foresman, J. V. Ortiz, J. Cioslowski, and D. J. Fox, Gaussian 09, Revision A.02, Gaussian, Inc., Wallingford CT, 2009