

Supporting Information

Crystal engineering of salen type cerium complexes induced by various cerium counterions

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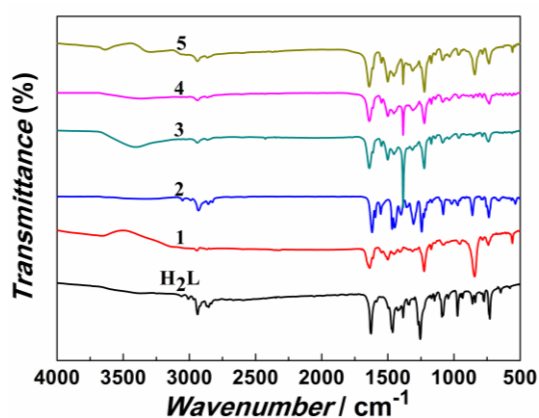


Fig. S1 IR spectra of H₂L and complexes 1–5.

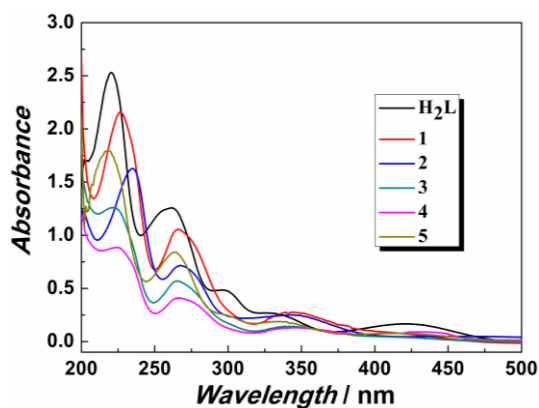


Fig. S2 UV-vis absorption spectra of H₂L and complexes 1–5 in CH₃OH.

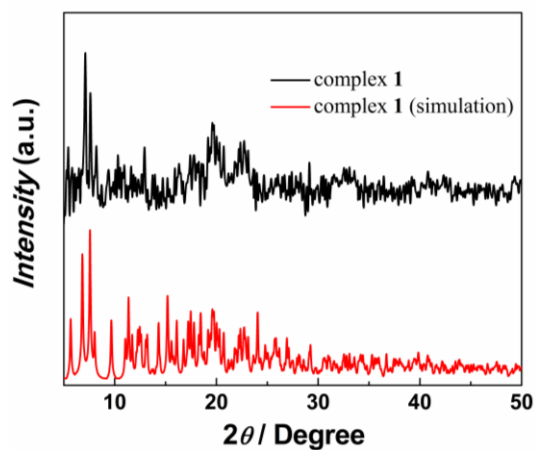


Fig. S3 The powder X-ray diffraction patterns and the simulated patterns of complex 1.

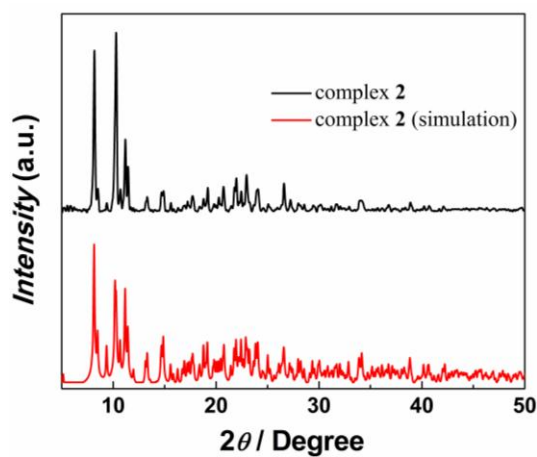


Fig. S4 The powder X-ray diffraction patterns and the simulated patterns of complex 2.

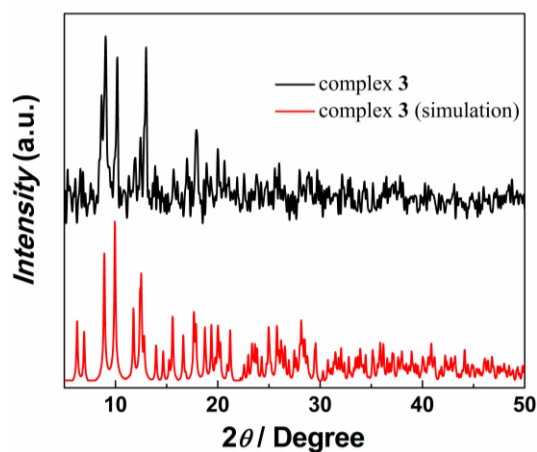


Fig. S5 The powder X-ray diffraction patterns and the simulated patterns of complex 3.

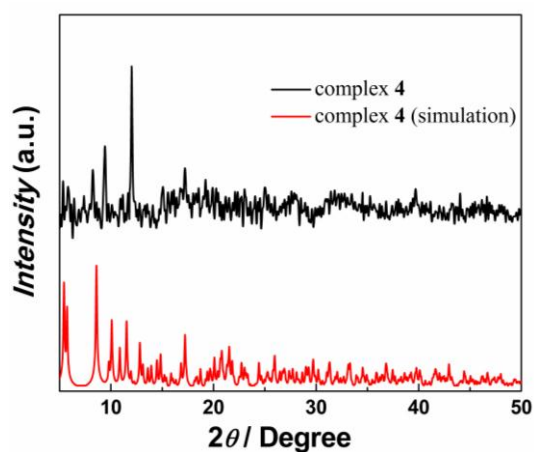


Fig. S6 The powder X-ray diffraction patterns and the simulated patterns of complex 4.

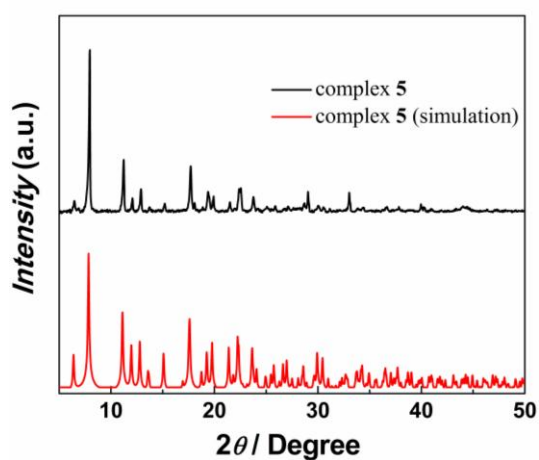


Fig. S7 The powder X-ray diffraction patterns and the simulated patterns of complex 5.

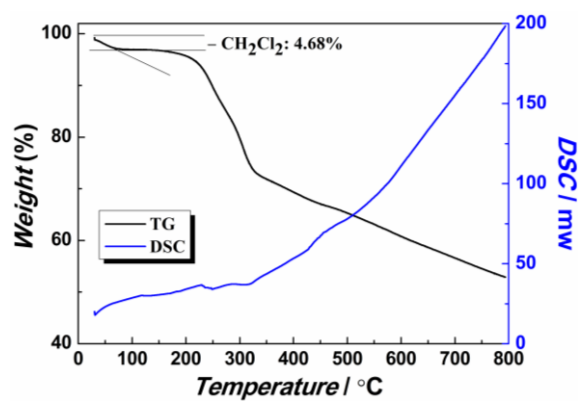


Fig. S8 TG-DSC curve for the complex 1.

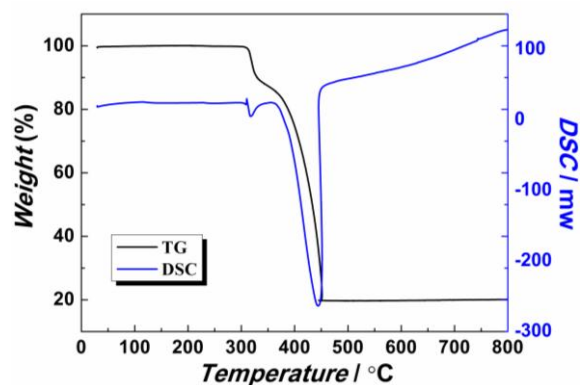


Fig. S9 TG-DSC curve for the complex 2.

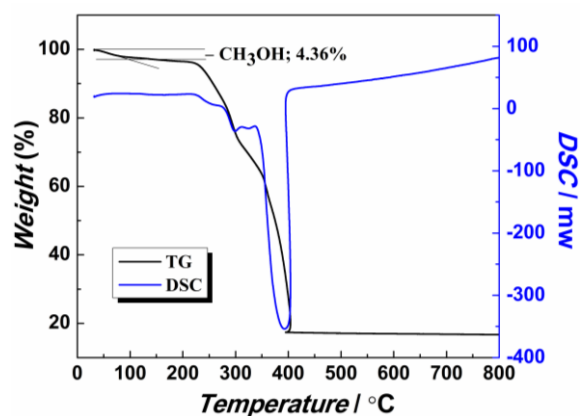


Fig. S10 TG-DSC curve for the complex 3.

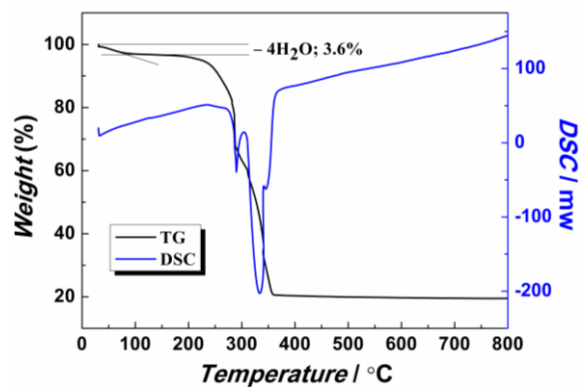


Fig. S11 TG-DSC curve for the complex 4.

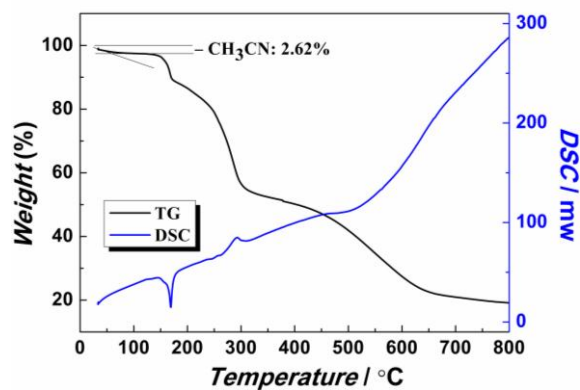


Fig. S12 TG-DSC curve for the complex 5.

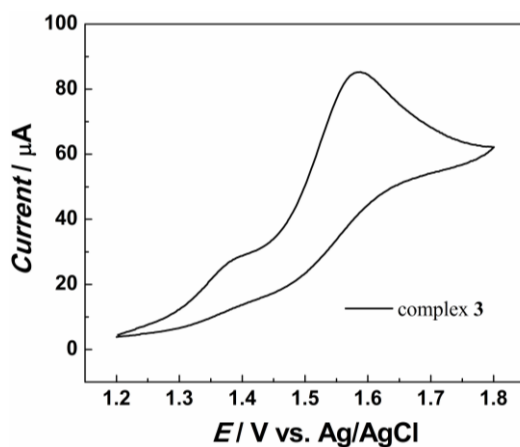


Fig. S13 CV curve of complex **3** in CH₃CN solution at 298K with 0.1 M *n*-Bu₄NPF₆ as supporting electrolyte (scan rate, 20 mV/s; working electrode, glassy carbon).

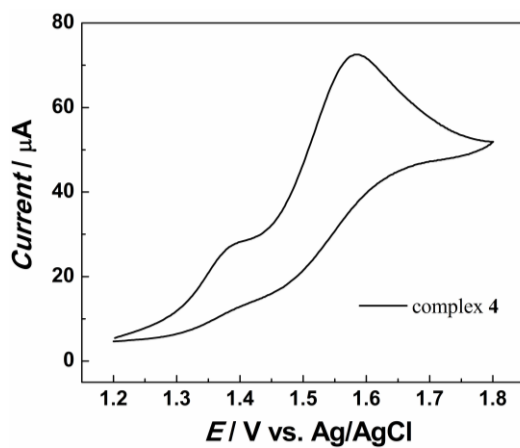


Fig. S14 CV curve of complex **4** in CH₃CN solution at 298K with 0.1 M *n*-Bu₄NPF₆ as supporting electrolyte (scan rate, 20 mV/s; working electrode, glassy carbon).

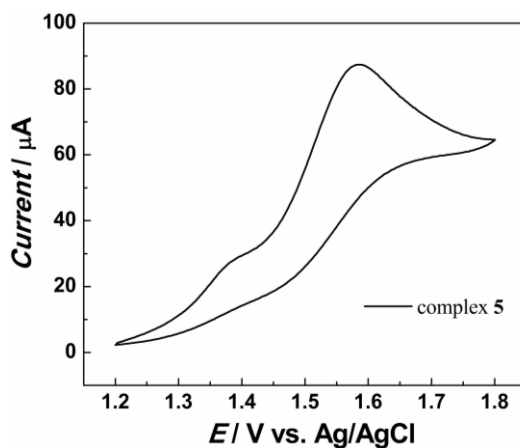


Fig. S15 CV curve of complex **5** in CH₃CN solution at 298K with 0.1 M *n*-Bu₄NPF₆ as supporting electrolyte (scan rate, 20 mV/s; working electrode, glassy carbon).