

Supporting Information

Facile synthesis of MnO₂-Ag hollow microspheres with sheet-like subunits and their catalytic properties

Dan-Ling Zhou,^a De-Jun Chen,^b Pei-Pei Zhang,^b Fang-Fang Li,^b Jian-Rong Chen,^a Ai-Jun Wang^{a*}
and Jiu-Ju Feng^{a*}

^a *College of Geography and Environmental Science, College of Chemistry and Life Science, Zhejiang Normal University, Jinhua 321004, China*

^b *College of Chemistry and Chemical Engineering, Henan Normal University, Xinxiang 453007, China*

**Corresponding author: Tel./Fax: 86-579-82282273; E-mail: ajwang@zjnu.cn (AJW) and E-mail: jjfeng@zjnu.cn (JJF)*

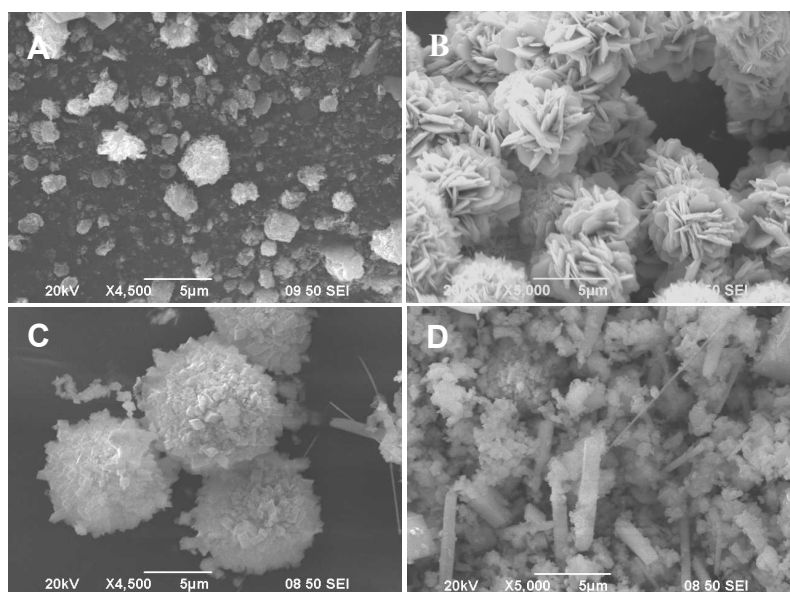


Fig. S1. SEM images of the products prepared at different reaction temperature: room temperature without hydrothermal treatment (A), 100 °C (B), 160 °C (C), and 200 °C (D).

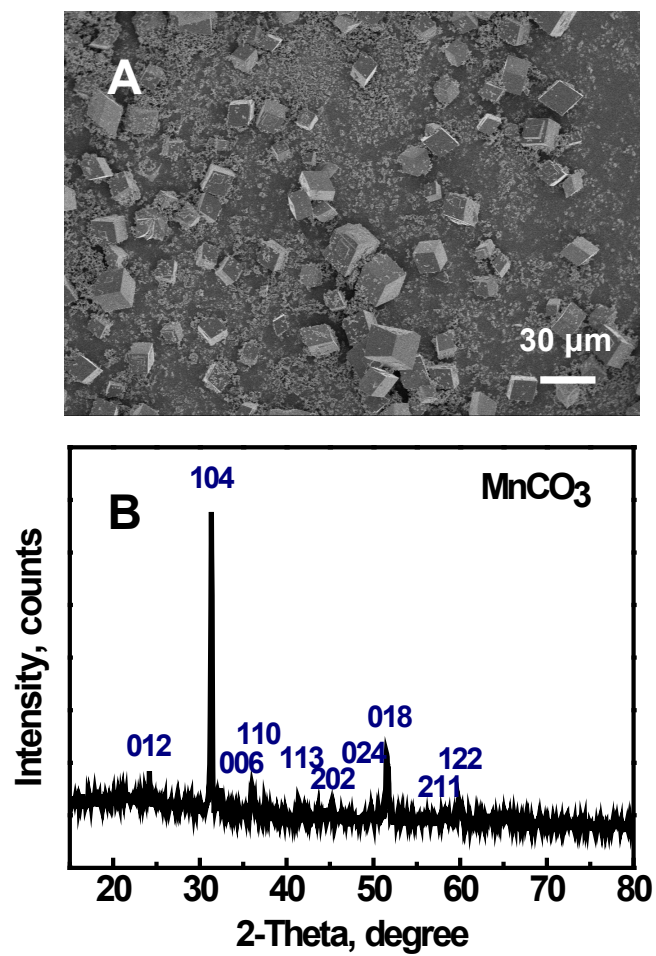


Fig. S2. (A) SEM image and XRD pattern (B) of the products prepared without AgNO₃.

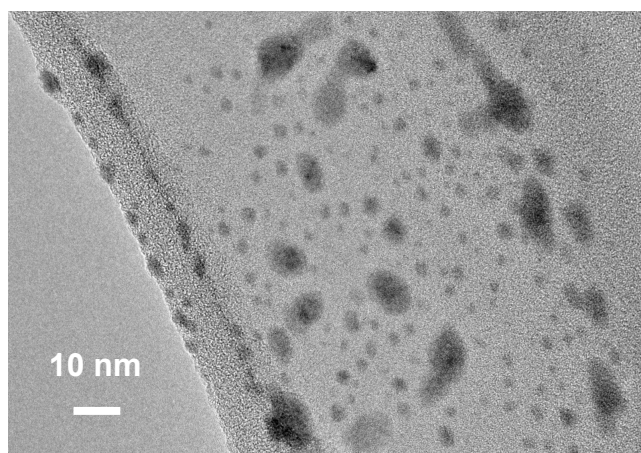


Fig. S3. TEM image of the product obtained at 130 °C for 0.5 h.

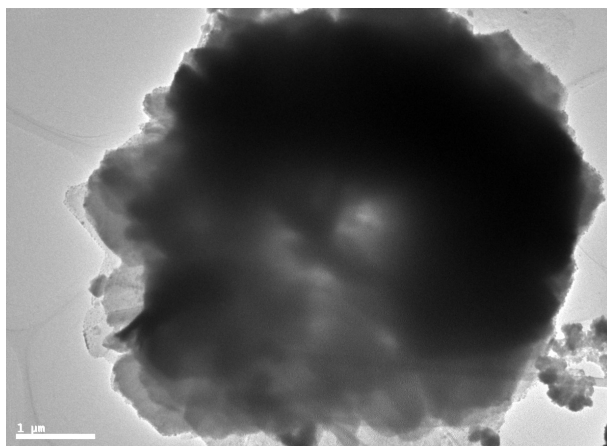


Fig. S4. High-magnification TEM image of the product obtained at 130 °C for 2 h.

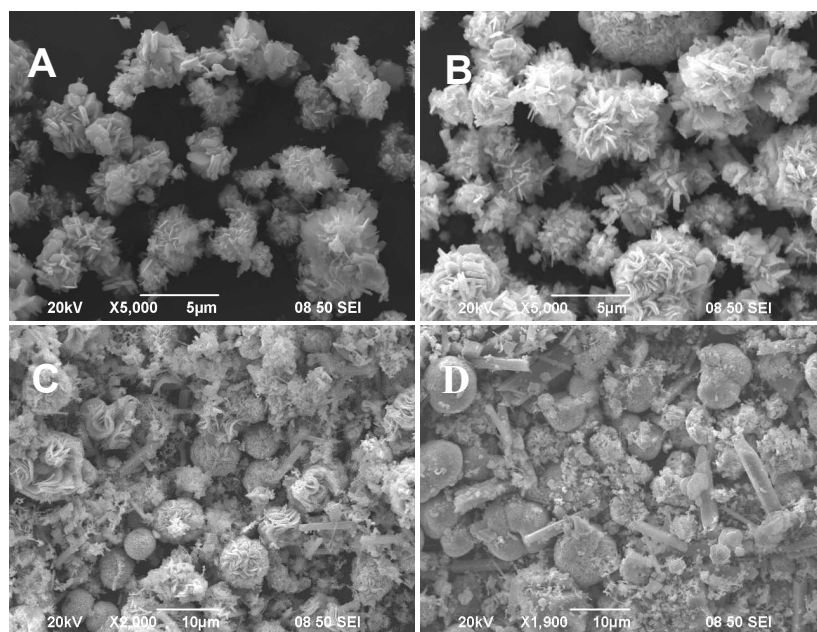


Fig. S5. SEM images of the products prepared at different reaction time: 1 h (A), 2 h (B), 5 h (C), and 10 h (D).

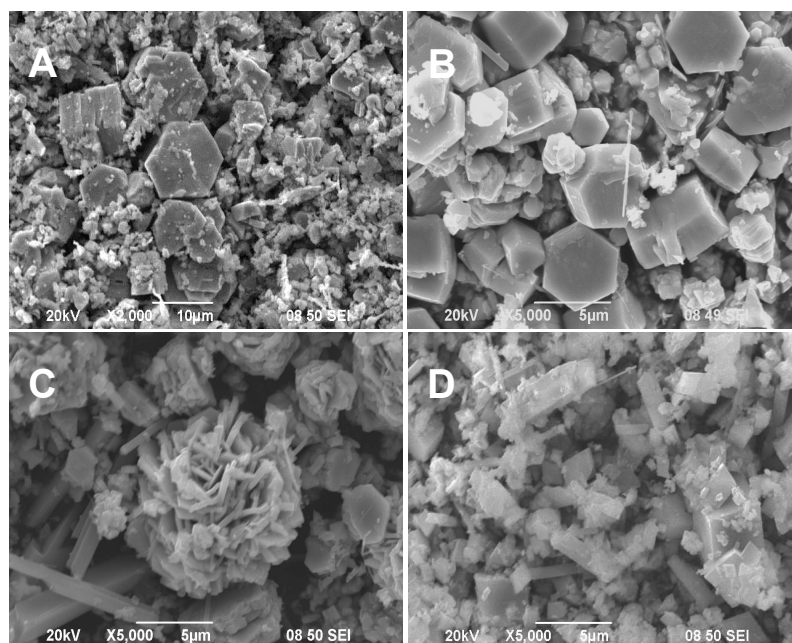


Fig. S6. SEM images of the products prepared with different amount of urea: 0.024 g (A), 0.144 g (B), 0.48 g (C), and 4.80 g (D).

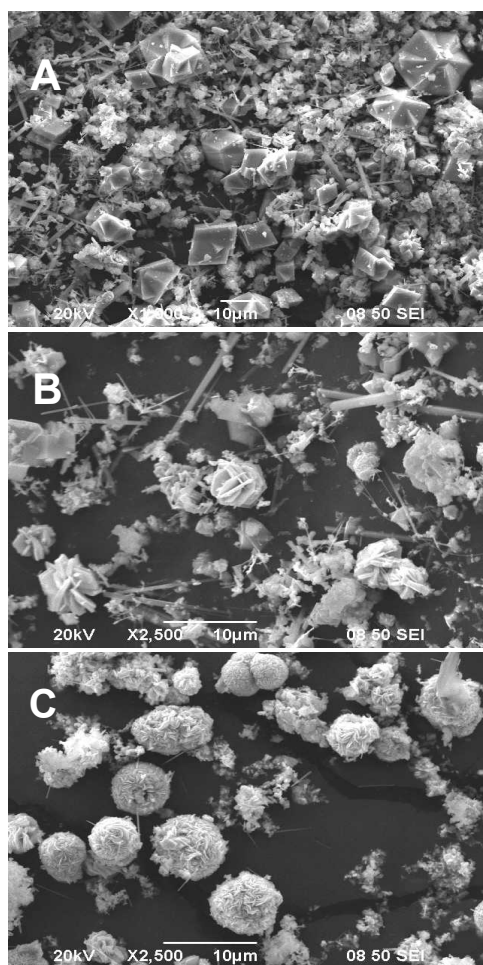


Fig. S7. SEM images of the products obtained with different molar ratios of Mn²⁺/Ag⁺:
2:1 (A), 1:1 (B), and 1:3 (C) at 130 °C.

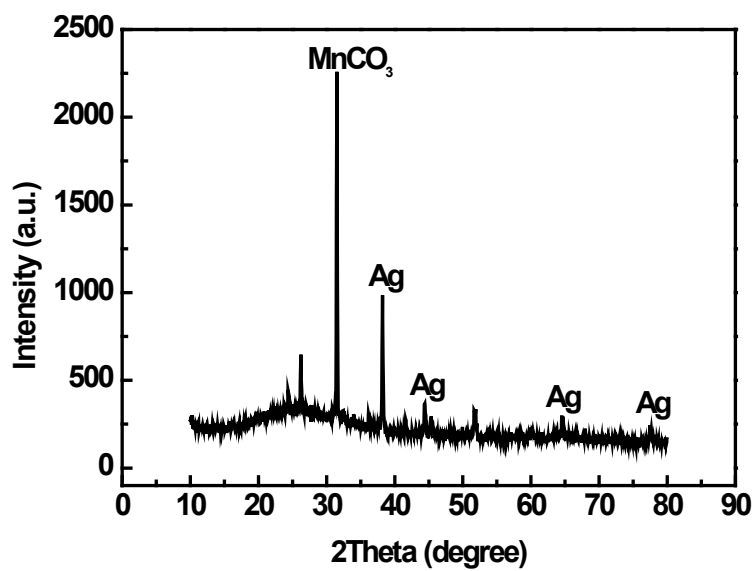


Fig. S8. X-ray diffraction pattern of the product using the molar ratio ($\text{Mn}^{2+}/\text{Ag}^{+}$) of 2:1.

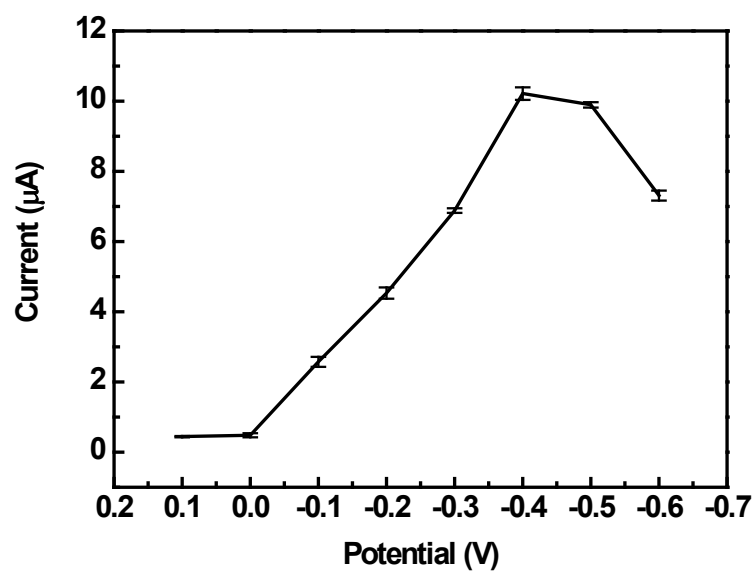


Fig. S9. Effects of the applied potentials on the steady-state responses of the MnO₂-Ag HMs modified CPE in a 25 mM phosphate solution (pH 7.0) containing 1.25 mM H₂O₂.

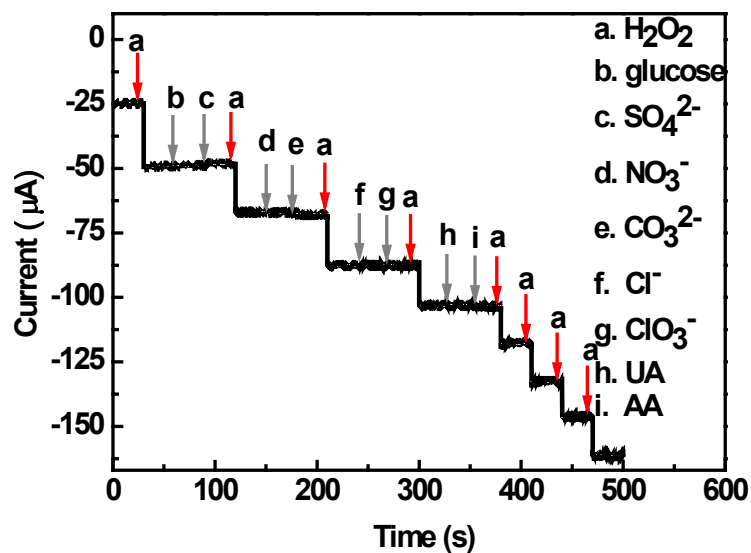


Fig. S10. Amperometric *i-t* curve of the MnO₂-Ag HMs modified CPE with successive addition of 1.0 mM H₂O₂, ClO₃⁻, UA, and AA, as well as 10 mM glucose, SO₄²⁻, NO₃⁻, CO₃²⁻, and Cl⁻ ions in a 25 mM phosphate solution (pH 7.0) at -0.4 V.

Table S1. Comparison of the MnO₂-Ag HMs sensor with other MnO₂ based H₂O₂ sensors.

Materials	Stability	Linear ranges	Detection limit/ μ M	Refs.
MnO ₂ -Ag HMs/CPE	90% (4 weeks)	1.31 μ M ~ 36.71 mM	1.31	Our work
β -MnO ₂ nanorods	90% (30 days)	2.45 μ M ~ 42.85 mM	2.45	1
Ag-MnO ₂ -MWCNTs	90% (3 days)	5.0 μ M ~ 10.4 mM	1.7	2
MnO ₂ microspheres	89.0% (4 weeks)	10.0 μ M ~ 0.15 mM	2.0	3
MnO ₂ /graphene oxide	90% (4 weeks)	5 μ M ~ 0.6 mM	0.8	4
MnO ₂ /carbon fiber	not detected	12 μ M ~ 0.26 mM	5.4	5
MnO ₂ -mesoporous carbon	92% (1 month)	0.5 μ M ~ 0.6 mM	0.07	6
MnO ₂ -VACNTs	92% (30 days)	1.2 μ M ~ 1.8 mM	0.8	7

References

- 1 A.-J. Wang, P.-P. Zhang, Y.-F. Li, J.-J. Feng, W.-J. Dong and X.-Y. Liu, *Microchim. Acta* 2011, **175**, 31-37.
- 2 Y. Han, J. Zheng and S. Dong, *Electrochim. Acta* 2013, **90**, 35-43.
- 3 L. Zhang, Z. Fang, Y. Ni and G. Zhao, *Int. J. Electrochem. Sci.* 2009, **4**, 407-413.
- 4 L. Li, Z. Du, S. Liu, Q. Hao, Y. Wang, Q. Li and T. Wang, *Talanta* 2010, **82**, 1637-1641.
- 5 S. B. Hocevar, B. Ogorevc, K. Schachl and K. Kalcher, *Electroanalysis* 2004, **16**, 1711-1716.

6 L. Luo, F. Li, L. Zhu, Z. Zhang, Y. Ding and D. Deng, *Electrochim. Acta* 2012, **77**,

179-183.

7 B. Xu, M.-L. Ye, Y.-X. Yu and W.-D. Zhang, *Anal. Chim. Acta* 2010, **674**, 20-26.