

Electrochemical Investigation of Mn₄O₄-Cubane Water Oxidizing Clusters

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Electronic Supplementary Information

Table S1. Summary of cyclic voltammetric data obtained for the oxidation of 0.4 mM **1** at a glassy carbon disc electrode in CH₂Cl₂ (0.1 M Bu₄NPF₆) at 22 °C.

| Scan rate (mVs ⁻¹) | E_p^{ox} (mV*) | i_p^{ox} (μ A) | E_p^{red} (mV*) | i_p^{red} (μ A) | ΔE_p (mV*) | E_f^o (mV*) | $[i^{red}/i^{ox}]$ |
|-----------------------------------|---------------------|--------------------------|----------------------|---------------------------|-----------------------|------------------|--------------------|
| 10 | 729 | 1.6 | 645 | 1.4 | 84 | 687 | 0.9 |
| 50 | 733 | 3.3 | 647 | 2.8 | 86 | 690 | 0.8 |
| 100 | 740 | 4.9 | 645 | 4.1 | 95 | 692 | 0.8 |
| 200 | 747 | 6.5 | 634 | 5.5 | 113 | 690 | 0.8 |
| 500 | 750 | 9.8 | 636 | 8.3 | 129 | 693 | 0.9 |

*mV vs Fc/Fc⁺.

Table S2. Summary of cyclic voltammetric data obtained for the oxidation of 0.4 mM **2** at a glassy carbon disc electrode in CH₂Cl₂ (0.1 M Bu₄NPF₆) at 22°C.

| Scan rate (mVs ⁻¹) | E_p^{ox} (mV*) | i_p^{ox} (μ A) | E_p^{red} (mV*) | i_p^{red} (μ A) | ΔE_p (mV*) | E_f^o (mV*) | $[i^{red}/i^{ox}]$ |
|-----------------------------------|---------------------|--------------------------|----------------------|---------------------------|-----------------------|------------------|--------------------|
| 10 | 578 | 1.6 | 492 | 1.3 | 86 | 535 | 0.8 |
| 50 | 592 | 3.3 | 482 | 2.6 | 110 | 537 | 0.8 |
| 100 | 606 | 4.2 | 471 | 3.3 | 135 | 538 | 0.8 |
| 200 | 633 | 5.6 | 458 | 4.0 | 175 | 545 | 0.7 |
| 500 | 656 | 9.8 | 444 | 9.0 | 212 | 550 | 0.9 |

*mV vs Fc/Fc⁺.

Table S3. Summary of cyclic voltammetric data obtained for the reduction of 0.4 mM CH₂Cl₂ solutions of **1** and **2** at a glassy carbon disc electrode at 22 °C (0.1 M Bu₄NPF₆).

| Compound 1 | | | | | Compound 2 | | | | |
|-----------------------------------|----------------------|---------------------------|---------------------|----------------------|-----------------------------------|----------------------|---------------------------|---------------------|----------------------|
| Scan rate (mVs ⁻¹) | E_p^{red} (mV*) | i_p^{red} (μ A) | E_p^{ox} (mV*) | ΔE_p (mV) | Scan rate (mVs ⁻¹) | E_p^{red} (mV*) | i_p^{red} (μ A) | E_p^{ox} (mV*) | ΔE_p (mV) |
| 10 | -591 | 1.4 | Undetected | - | 10 | -582 | 1.6 | Undetected | - |
| 50 | -591 | 2.4 | Undetected | - | 50 | -617 | 3.1 | Undetected | - |
| 100 | -608 | 3.3 | Undetected | - | 100 | -654 | 4.2 | Undetected | - |
| 200 | -646 | 5.0 | -346 | -302 | 200 | -657 | 5.9 | -371 | 286 |
| 500 | -662 | 7.3 | -365 | -299 | 500 | -716 | 8.0 | -387 | 329 |

*mV vs Fc/Fc⁺

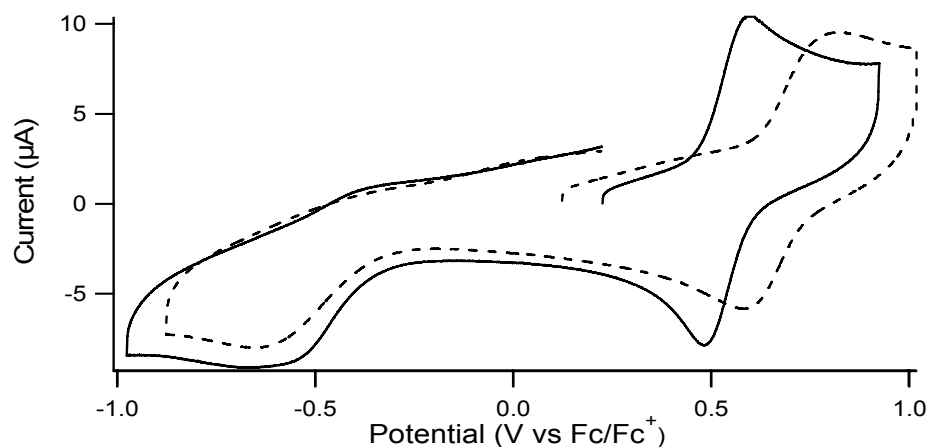


Figure S1. Cyclic voltammograms obtained at 22 °C at a glassy carbon electrode for oxidation and reduction of: A) 0.4 mM CH₂Cl₂ (0.1 M Bu₄NPF₆) solutions of **1** (- - -) and **2** (—), scan rate 250 mVs⁻¹.

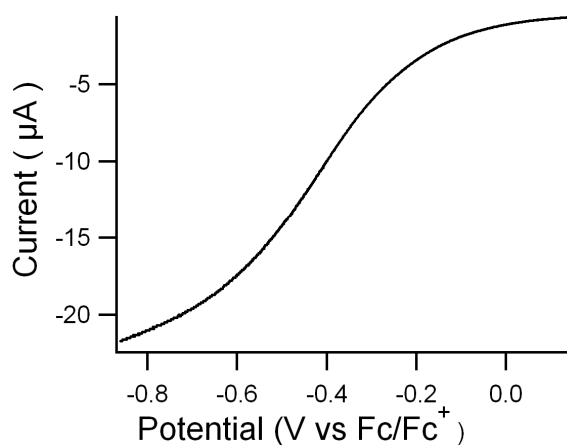


Figure S2. Glassy carbon rotated disc voltammogram obtained at 22°C for reduction of a 0.4 mM CH₂Cl₂ (0.1 M Bu₄NPF₆) solution of **1**, rotation rate = 1000 rpm, scan rate = 5 mV/s