

Electronic Supplementary Information for

**Influence of the Solvent on the Ground- and Excited-State
Buffer-Mediated Proton-Transfer Reactions of a Xanthenic Dye.**

Jose M. Paredes, Luis Crovetto, Angel Orte, Jose M. Alvarez-Pez*, Eva M. Talavera

*Department of Physical Chemistry, University of Granada, Cartuja Campus, Granada
18071, Spain.*

**Author to whom correspondence should be addressed.*

Tel.: + 34-958-243831; fax: + 34-958-244090; e-mail: jalvarez@ugr.es

	page
Figure 1S	S2
Figure 2S	S2
Table 1S	S3
Figure 3S	S3
Figure 4S	S4
Figure 5S	S4
Figure 6S	S5
Figure 7S	S5

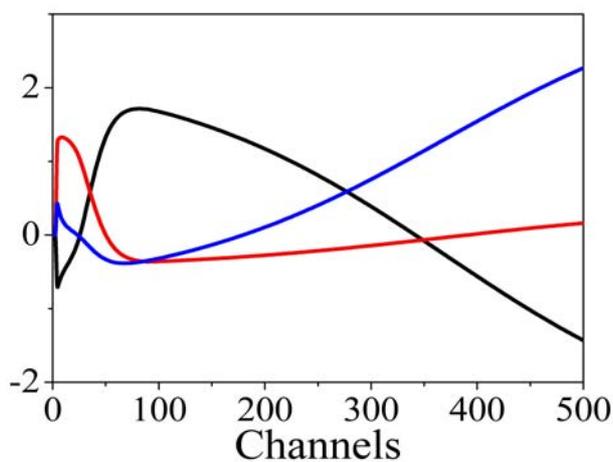


Figure 1S. Computation of time filter patterns corresponding to a long time of 3.77 ns (—) from TG_II anion fluorescence emission, a short time of 0.31 ns (—) from TG-II neutral fluorescence emission, corresponding to scattered light and background signal (—).

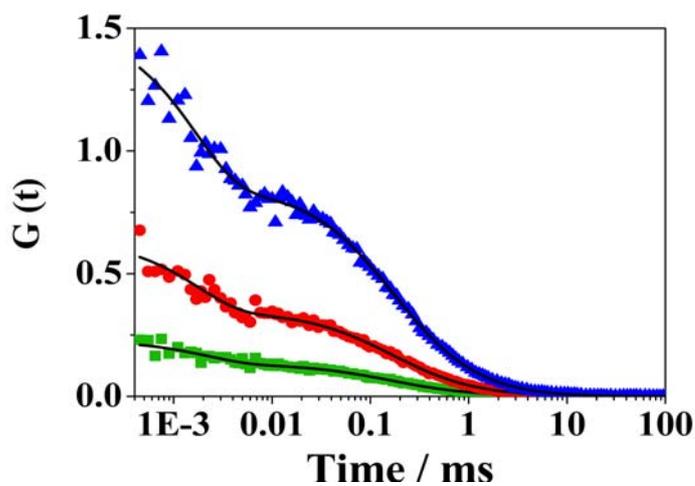


Figure 2S. Different correlation analyses of fluorescence fluctuation traces from 2 nM TG-II at pH 6.50, in the presence of 1.15 mM sodium phosphate buffer. (■) cross correlation; (●) time-gated cross correlation; and (▲) FLCS. In each case, the line represents the best fit.

Table 1S. Results obtained from the fit of the equation 2 to the experimental data shown in figure 2S, where $[C]$ is the TG-II concentration calculated from $\langle N \rangle$ in the correlation functions, f is the fraction of dark state corresponding to the neutral form of TG-II, and τ is the protonation time for the chemical equilibrium neutral-anion.

	Cross-correlation	Time-Gate	FLCS
$[C] / \text{nM}$	12.5	4.8	1.96
f	0.4208	0.4572	0.4336
τ / s	2.17×10^{-6}	1.75×10^{-6}	1.66×10^{-6}

As it can be seen, the value recovered for the fluorophore concentration by using the FLCS methodology is consistent with the nominal experimental concentration.

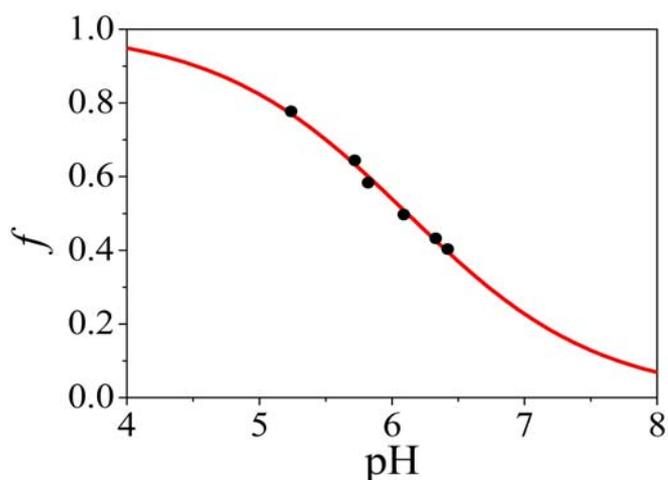


Figure 3S. Fraction of the neutral state of TG-II in water solution (f) vs. pH value, obtained from fitting eq. (2) in main text to FLCS curves of fluorescence fluctuations traces from TG-II in aqueous solution in the absence of buffer. The solid line represents the best fit of the experimental data to the equilibrium equation $f = H^+ / (K_a + H^+)$.

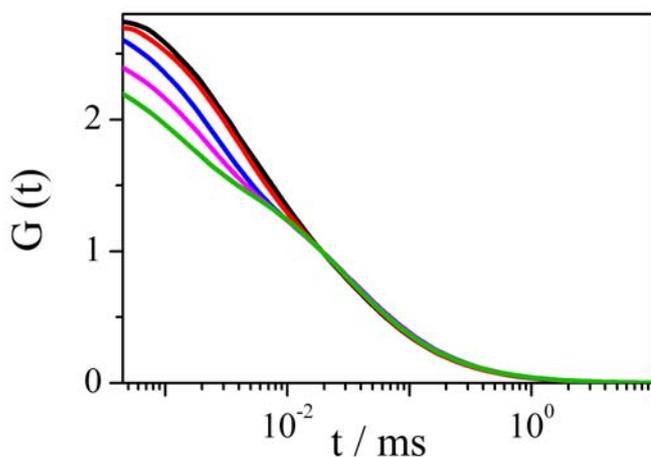


Figure 4S. Normalized lifetime filtered AC curves from TG-II solutions at different lithium phosphate buffer concentration: 0.25 mM (—), 0.50 mM (—), 1.00 mM (—), 1.50 mM (—) and 2.00 mM (—) at pH value 6.50

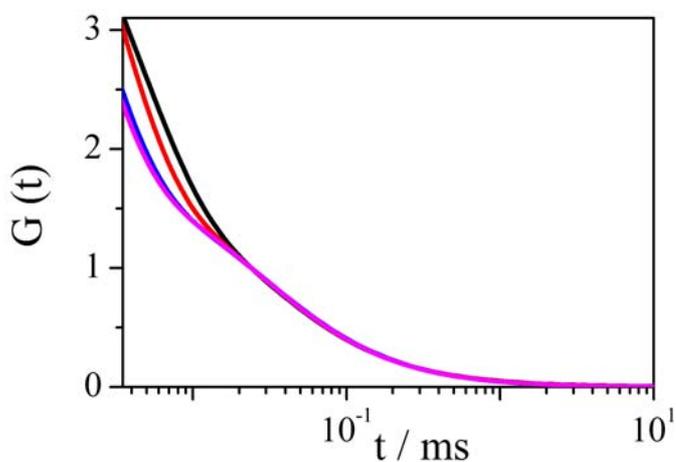


Figure 5S. Normalized lifetime filtered AC curves from TG-II solutions at different caesium phosphate buffer concentration: 0.50 mM (—), 1.00mM (—), 1.50 mM (—) and 2.00 mM (—) at pH value 5.80

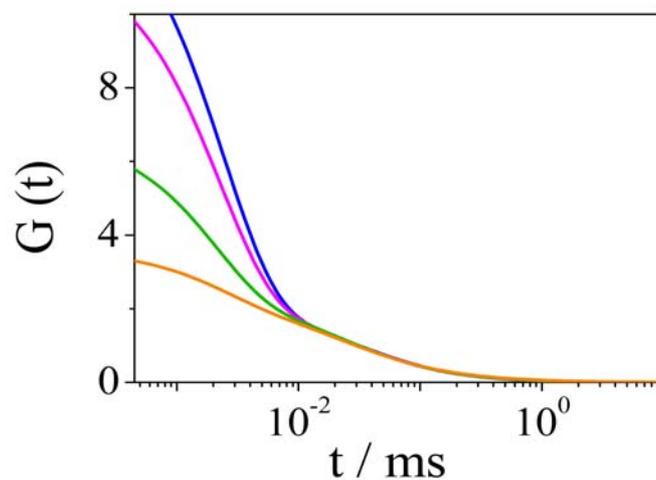


Figure 6S. Normalized lifetime filtered AC curves from TG-II solutions at different acetate buffer concentration: 0.174 mM (—), 0.390 mM (—), 0.711 mM (—) and 1.05 mM (—) at pH value 6.00

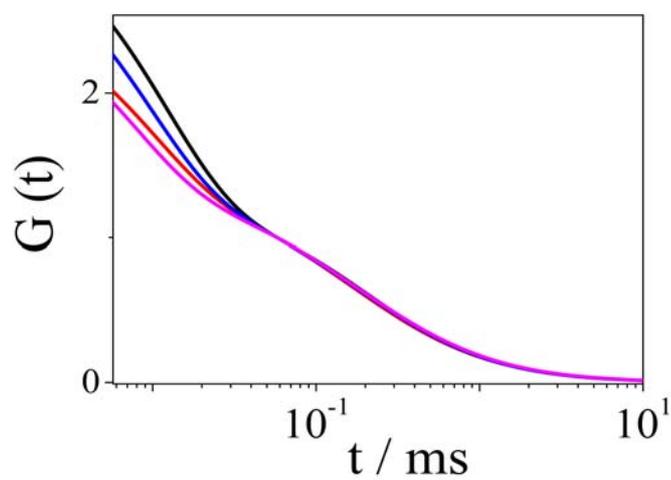


Figure 7S. Normalized lifetime filtered AC curves from TG-II solutions at different Tris buffer concentration: 0.06 mM (—), 2.38 mM (—), 6.23 mM (—) and 8.13 (—) at pH value 6.00