

**Supporting Information**

Theoretical study of electronically excited radical cations of naphthalene and anthracene as archetypical models for astrophysical observations. I. Static aspects

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**Table S1, Table S2 and FIG. S1**

TABLE S1: *Ab initio* calculated equilibrium geometry parameters of the electronic ground state of neutral naphthalene along with the experimental data. The atom numbering in the table is illustrated in the schematic drawing below the table.

Np (MP2/cc-pVDZ)	our data	Expt data <sup>a</sup>
Bond length (Å)		
C <sub>1</sub> - C <sub>2</sub>	1.4415	1.412
C <sub>2</sub> - C <sub>3</sub>	1.4270	–
C <sub>3</sub> - C <sub>4</sub>	1.3892	–
C <sub>4</sub> - C <sub>5</sub>	1.4231	1.417
C <sub>5</sub> - C <sub>6</sub>	1.3892	–
C <sub>3</sub> - H <sub>11</sub>	1.0965	1.092
C <sub>4</sub> - H <sub>12</sub>	1.0953	1.092
Bond Angle (deg)		
C <sub>1</sub> - C <sub>2</sub> - C <sub>3</sub>	118.981	–
C <sub>2</sub> - C <sub>3</sub> - C <sub>4</sub>	120.7356	–
C <sub>3</sub> - C <sub>4</sub> - C <sub>5</sub>	120.2834	–
C <sub>6</sub> - C <sub>1</sub> - C <sub>7</sub>	122.0380	–
C <sub>2</sub> - C <sub>3</sub> - H <sub>11</sub>	118.8773	–
C <sub>3</sub> - C <sub>4</sub> - H <sub>12</sub>	120.3871	–
C <sub>3</sub> - C <sub>4</sub> - H <sub>12</sub>	119.9362	–
C <sub>4</sub> - C <sub>5</sub> - H <sub>12</sub>	119.7804	–

<sup>a</sup> N. Ketkar *et al.*, *J. Mol. Struct.* 1981, **77**, 139.

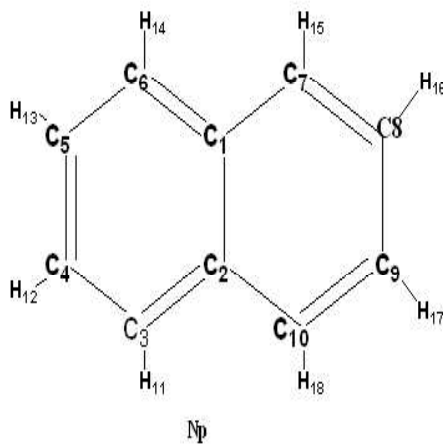
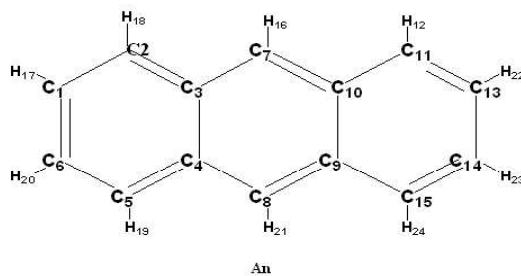


TABLE S2: Same as in Table S1 for the electronic ground state of neutral anthracene.

An (MP2/cc-pVDZ)	our data	Expt data <sup>b</sup>
Bond length (Å)		
C <sub>1</sub> - C <sub>2</sub>	1.3844	1.397
C <sub>1</sub> - C <sub>6</sub>	1.4299	1.422
C <sub>2</sub> - C <sub>3</sub>	1.4337	1.437
C <sub>3</sub> - C <sub>4</sub>	1.4511	1.437
C <sub>3</sub> - C <sub>7</sub>	1.4103	1.392
C <sub>1</sub> - H <sub>17</sub>	1.0954	–
C <sub>2</sub> - H <sub>18</sub>	1.0966	–
C <sub>7</sub> - H <sub>16</sub>	1.0979	–
Bond Angle (deg)		
C <sub>1</sub> - C <sub>2</sub> - C <sub>3</sub>	120.815	–
C <sub>2</sub> - C <sub>3</sub> - C <sub>4</sub>	118.775	–
C <sub>2</sub> - C <sub>3</sub> - C <sub>7</sub>	121.990	122.3
C <sub>2</sub> - C <sub>1</sub> - C <sub>6</sub>	120.416	120.4
C <sub>5</sub> - C <sub>6</sub> - C <sub>1</sub>	120.410	–
C <sub>3</sub> - C <sub>7</sub> - C <sub>10</sub>	121.529	121.0
C <sub>7</sub> - C <sub>10</sub> - C <sub>9</sub>	119.236	–
C <sub>7</sub> - C <sub>3</sub> - C <sub>4</sub>	119.236	118.4
H <sub>17</sub> - C <sub>1</sub> - C <sub>2</sub>	119.938	–
H <sub>17</sub> - C <sub>1</sub> - C <sub>6</sub>	119.652	–
H <sub>18</sub> - C <sub>2</sub> - C <sub>3</sub>	118.697	–
H <sub>18</sub> - C <sub>2</sub> - C <sub>1</sub>	120.488	–
H <sub>16</sub> - C <sub>7</sub> - C <sub>3</sub>	119.236	–

<sup>b</sup> N. Ketkar *et al.*, *J. Mol. Struct.* 1981, **77**, 127.



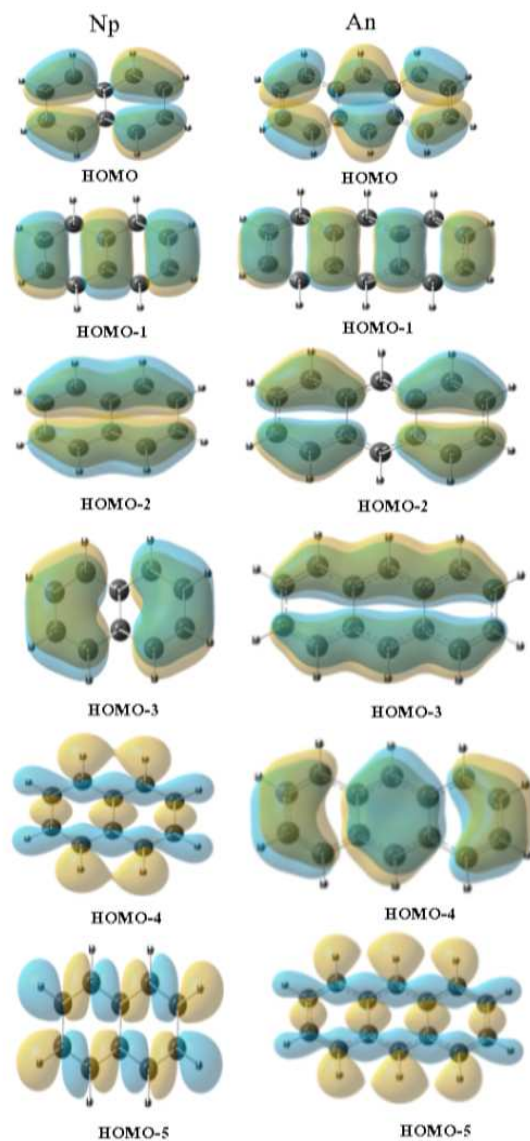


FIG. S1: The first six highest occupied molecular orbitals of the electronic ground state of neutral naphthalene and anthracene.