

Supporting Information for

Improved Photoelectrochemical Water Oxidation Kinetics Using TiO₂ Nanorod Arrays Photoanode Decorated with Graphene Oxide in a Neutral pH Solution

Sang Youn Chae^{a,b}, Pitchaimuthu Sudhagar^c, Akira Fujishima^c, Yun Jeong Hwang^{a,d*}, Oh-Shim Joo^{a,d}

^a Clean Energy Research Center, Korea Institute of Science and Technology (KIST), 39-1 Hawolgoek Dong, Seoul 136-791, Republic of Korea

^b Department of Chemistry, Korea University, Seoul 136-713, Republic of Korea

^c Photocatalysis International Research Center (PIRC), Research Institute for Science and Technology, Tokyo University of Science, 2641 Yamazaki, Noda, Chiba 278-8510, Japan

^d Korea University of Science and Technology, 176 Gajung-dong, 217 Gajungro Yuseong-gu, Daejeon 305-350, Republic of Korea

* Corresponding author. Tel: +82-2-958-5227; Fax: 82-2-958-5809

E-mail address: yjhwang@kist.re.kr

Supplementary Figures

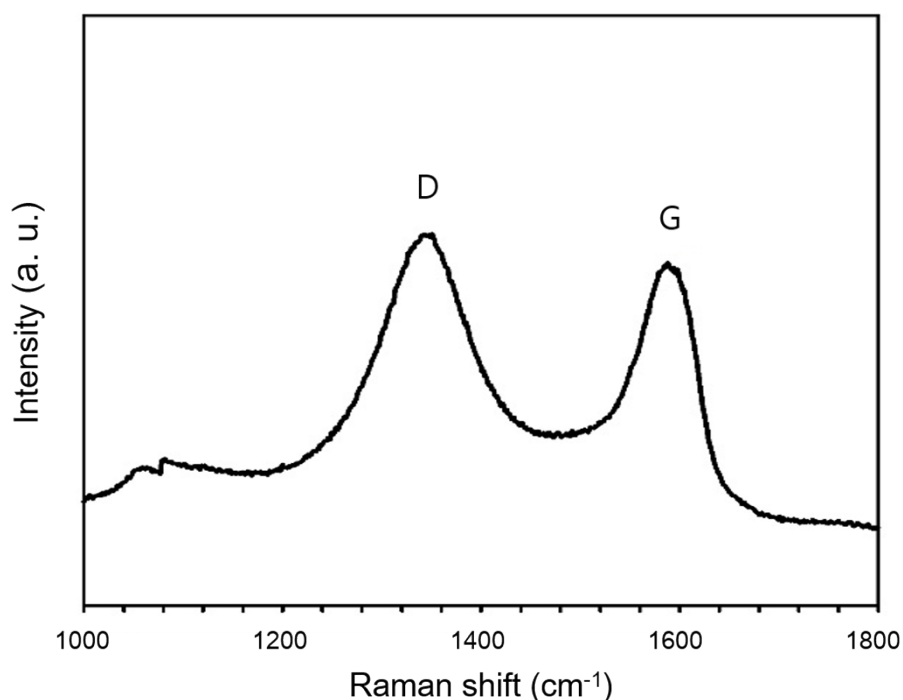


Figure S1. Raman spectrum of the prepared graphene oxide. D and G indicate D band and G band of graphene oxide.

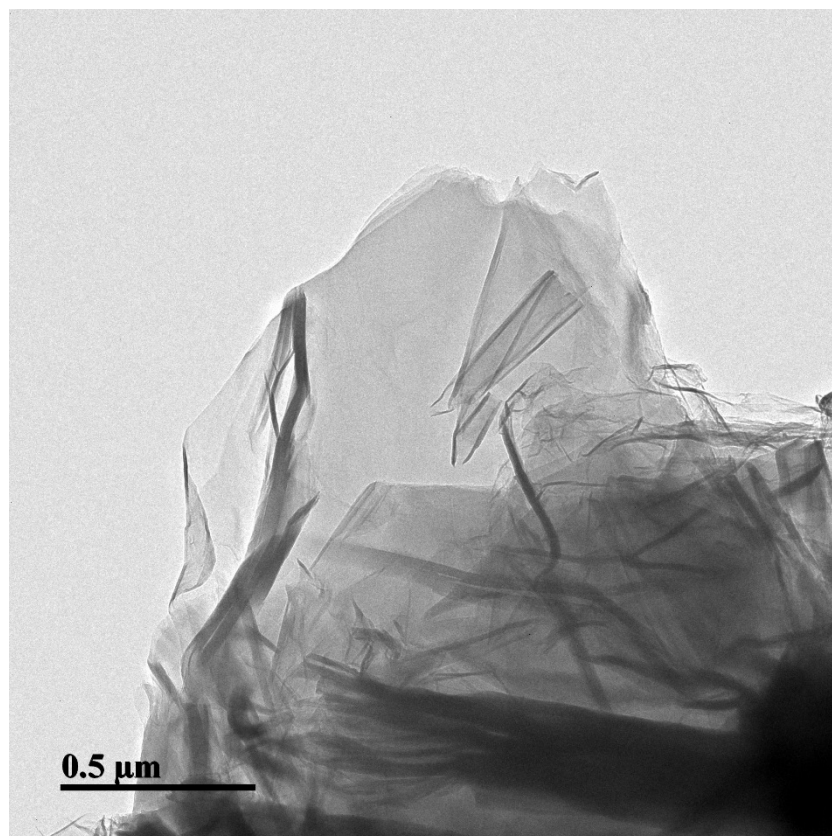


Figure S2. TEM image of the synthesized graphene oxide.

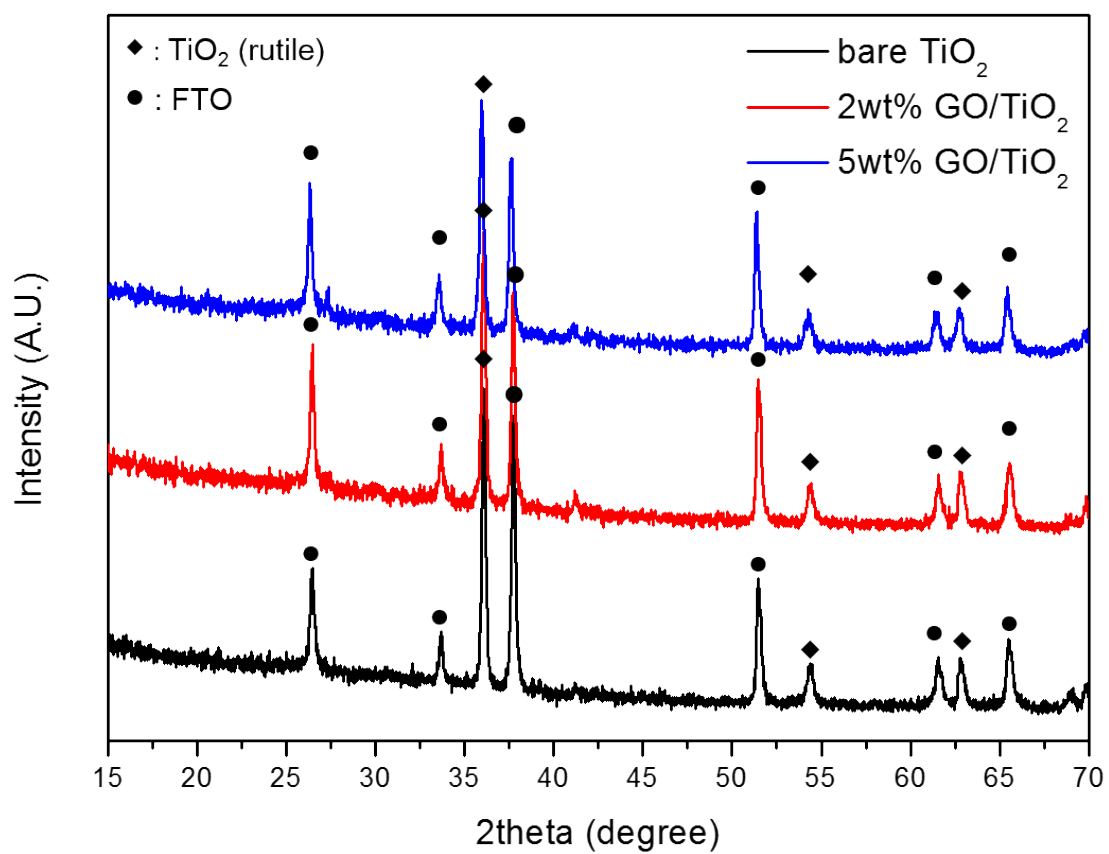
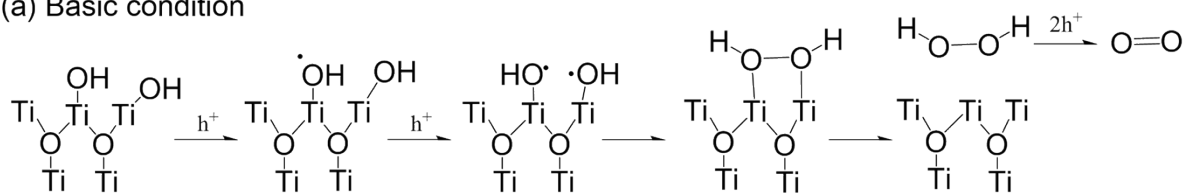
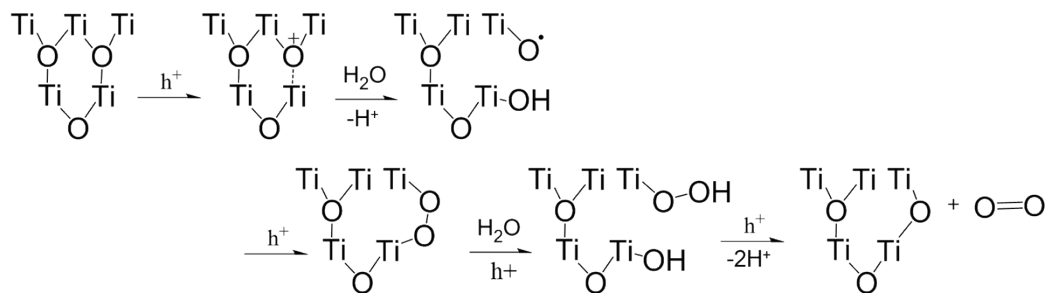


Figure S3. XRD pattern of bare TiO₂, 2 wt% GO coated TiO₂, and 5 wt% GO coated TiO₂ nanorod on FTO glass.

(a) Basic condition



(b) Neutral condition



Scheme S1. Proposed water oxidation mechanisms showing the involving chemical species (a) in a basic, and (b) in a neutral pH electrolyte adapted by the previous reports^{1,2}.

Reference

1. M. R. Hoffmann, S. T. Martin, W. Y. Choi and D. W. Bahnemann, *Chem. Rev.*, 1995, 95, 69-96.
2. R. Nakamura, T. Okamura, N. Ohashi, A. Imanishi and Y. Nakato, *J. Am. Chem. Soc.*, 2005, 127, 12975-12983.